

Additions and Corrections

Extreme Resistance of Superhydrophobic Surfaces to Impalement: Reversible Electrowetting Related to the Impacting/Bouncing Drop Test

P. Brunet,* F. Lapierre, V. Thomy, Y. Coffinier, and R. Boukherroub *Langmuir* **2008**, *24*, 11203–11208.

The Acknowledgment published in the original Article is updated as follows:

Acknowledgment. The authors thank S. Coudert for his kind support during the visualizations, P. Lefebvre-Legry for her technical support, and R. Blossey for fruitful discussions. The Centre National de la Recherche Scientifique (CNRS) is gratefully acknowledged for financial support.

LA900144D

10.1021/la900144d

Published on Web 02/03/2009

Particulation of Hyperbranched Aromatic Biopolyesters Self-Organized by Solvent Transformation in Ionic Liquids

Dongjian Shi, Tatsuo Kaneko, and Mitsuru Akashi* *Langmuir* **2007**, *23*, 3485–3488.

There is a mistake in the assignment of a carboxyl proton in trifluoroacetic acid (TFA) in the mixture of TFA with DMF (1/1) from Figure 3. In the original report, the frequency range was set from 12 to -3 ppm. This inadequate frequency allowed the peak to appear incorrectly, because a folded signal will appear in the transformed spectrum when a signal exists outside the frequency range owing to a limited number of data points.¹

Therefore, the frequency range was reset from 20 to -1 ppm, and the ^1H NMR spectra of common DMF, TFA, and their mixture (not a deuterate solvent) were measured again using chloroform-*d* (with 1 wt % TMS) as the reference solvent. The correct spectra are given in Figure 1A. A downfield shifted peak then appeared at 17 ppm and was assigned to the carboxylate ions in TFA after mixing with DMF. Moreover, the DMF protons of the aldehyde and methyl groups showed a slight shift (about 0.1 ppm) toward the low magnetic field upon mixing with TFA. Therefore, the interaction between TFA and DMF was proposed as shown in Figure 1B.²

Although the chemical shift of the carboxyl proton had an incorrect value in the original paper, the particles are surely prepared by mixing two homogeneous P(4HCA-*co*-DHCA) copolymer solutions with DMF and TFA, since the solubility of the copolymer was changed upon mixing with DMF and TFA.

We thank Professor Yoshihisa Inoue from Osaka University for pointing out this problem.

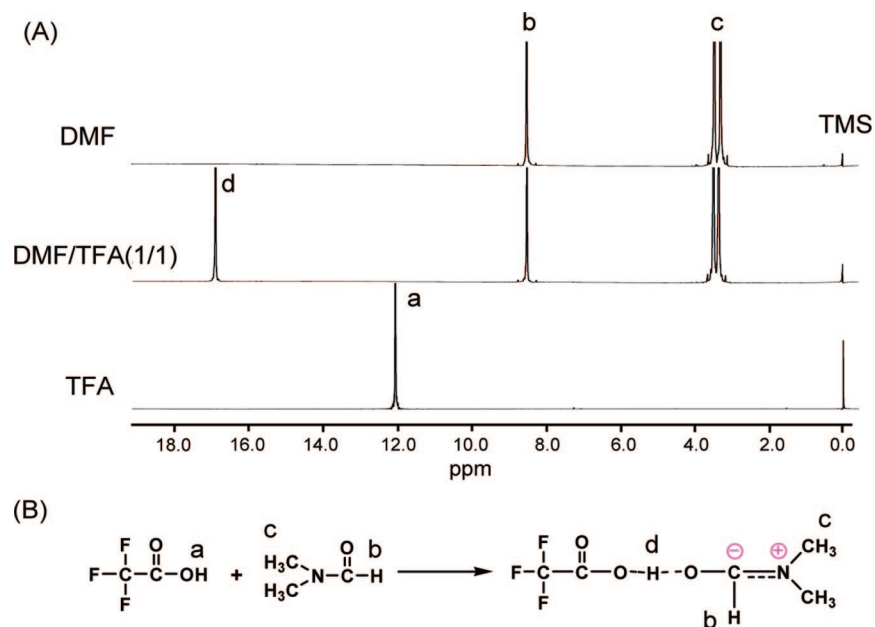


Figure 1. (A) ^1H NMR spectra of DMF, TFA, and their mixture. (B) Proposed reaction scheme occurring in the mixture of DMF and TFA solvents.

(1) Hatada, K.; Kitayama, T. *NMR Spectroscopy of Polymers*; Springer-Verlag: Berlin, Heidelberg, New York, **2004**; Chapter 1, pp 8–13.
 (2) Pang, T. S.; Ng, S. *Spectrosc. Lett.* **1973**, 6, 771.

LA804266D

10.1021/la804266d
 Published on Web 01/26/2009