

Letters

Let's All Visit Mole City!

As a college professor who has just recently finished teaching stoichiometry in my General Chemistry class, I very much appreciated Addison Ault's commentary, *Mole City: A Stoichiometric Analogy* (1). In fact, I plan to encourage my students to prepare maps on posters for extra credit using his plan.

I do find one small quibble. A stoichiometry problem need not start at the house of a Starting Material and end at the house of a Product. For example: How many grams (moles, liters, etc.) of oxygen are required to completely oxidize a given amount of iron? Both oxygen and iron reside in the houses of the Starting Materials (on the same side of Mole River), and yet one must cross the Mole Street Bridge (use the stoichiometric ratio) to solve the problem. I plan to ask my students to critique their maps and think of ways to include such problems.

The author replies:

I enjoyed reading Mark Armstrong's letter about Mole City. I can now imagine six moles (grams; liters;...) of Oxygen looking for action. On the way to Mole Street they stop by the House of Hydrogen. How many moles (grams; liters;...) of Hydrogen should they pick up to ensure that everyone has a good time at the reactor? The possibilities are endless....

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Literature Cited

1. Ault, A. Mole City: A Stoichiometric Analogy. *J.Chem. Educ.* 2006, 83, 1587–1588.

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