

# Wetting Transition of Water on Graphite and Other Surfaces

S. M. Gatica,<sup>†</sup> J. K. Johnson,<sup>‡,§</sup> X. C. Zhao,<sup>‡</sup> and M. W. Cole<sup>\*,†</sup>

Department of Physics, Pennsylvania State University, University Park, Pennsylvania 16802, Department of Chemical and Petroleum Engineering, University of Pittsburgh, Pittsburgh, Pennsylvania 15261, and National Energy Technology Laboratory, U.S. Department of Energy, P O Box 10940, Pittsburgh, Pennsylvania 15236

Received: April 5, 2004; In Final Form: May 19, 2004

A wetting transition occurs when the contact angle of a liquid drop on a surface changes from a nonzero value to zero. Such a transition has never been observed for water on any solid surface. This paper discusses the value of the temperature  $T_w$  at which the transition should occur for water on graphite. A simple model, previously used for nonpolar fluids, predicts the value of  $T_w$  as a function of the well depth  $D$  of the adsorption potential. While  $D$  is not well known for the case of water/graphite, the model implies that  $T_w$  is likely to fall in the range 350–500 K. Experimental search for this transition is warranted. Water-wetting transition temperatures on other surfaces are also discussed.

## I. Introduction

Fundamental discoveries related to the thermodynamic laws describing adsorbed films on planar surfaces are still being made, despite the fact that these laws have been known for more than a century. An important example involves the contact angle  $\theta$  of a liquid drop on a surface and its relevance to wetting transitions.<sup>1–4</sup> Young's equation

$$\sigma_{gs} = \sigma_{ls} + \sigma_{lg} \cos(\theta) \quad (1)$$

implies that  $\theta$  vanishes when a particular equality is satisfied

$$\sigma_{gs} = \sigma_{ls} + \sigma_{lg} \quad (2)$$

Here,  $\sigma_{gs}$ ,  $\sigma_{ls}$ , and  $\sigma_{lg}$  are the gas–solid, liquid–solid, and liquid–gas interfacial tensions, respectively. This description in terms of macroscopic surface tensions applies only at saturated vapor pressure (svp). A wetting transition occurs when  $\theta$  changes from a nonzero value to zero.<sup>5,6</sup> Analyses describing this transition were presented some 25 years ago in complementary studies of Cahn<sup>7</sup> and Ebner and Saam.<sup>8</sup> Their research found that if a fluid does not wet a particular surface at low temperature, then a wetting transition is expected to occur at some higher temperature  $T_w$  that is less than the bulk critical temperature  $T_c$ ;  $T_w$  thus separates nonwetting from wetting regimes. A drying transition ( $\theta = \pi$ ), hypothesized in the early work, is believed now to be ruled out in the prevalent case of long-range van der Waals forces.<sup>9,10</sup>

Such wetting transitions were first observed some 10 years ago in experiments involving simple gases (He, Ne, and H<sub>2</sub>) on alkali metal surfaces;<sup>11–19</sup> soon afterward, transitions were found with Hg on the surfaces of Ta and sapphire.<sup>20–25</sup> The common origin of these phenomena is that the gas–surface interaction  $V(r)$  is only weakly attractive (implying nonwetting behavior at temperature  $T$  close to the triple temperature  $T_3$ ).<sup>13,26–28</sup> The dividing line between wetting and nonwetting

behaviors at  $T_3$  corresponds approximately to the situation when the adsorption well depth  $D = 3.5\epsilon$ ,<sup>29</sup> where  $\epsilon$  is the well depth of the adsorbate's intermolecular interaction; for a monatomic, classical fluid, this criterion is equivalent to the condition  $D/(k_B T) \approx 6$  at the triple point.<sup>30</sup> The physics of this phenomenon are qualitatively simple: at pressures close to svp, the gas is on the verge of condensing to form a bulk liquid phase. If the surface provides a very strongly attractive interaction, it nucleates the film, which grows continuously on the surface as the pressure  $P$  approaches its value  $P_0$  at svp. If, instead, the interaction is only weakly attractive, no such nucleation occurs below svp. In that case, a fluid droplet on the surface at svp will bead up instead of spreading across the surface. The difference between the two kinds of behavior is thus associated with the relative strength of cohesive and adhesive interactions.

Interestingly, no wetting transition involving water has ever been observed, to the best of our knowledge.<sup>31,32</sup> This situation is somewhat surprising in view of the fact that water is a much-studied fluid that does not wet many surfaces; transitions are expected to occur for each of these cases. This paper is concerned primarily with the wetting behavior of water on graphite. There exist several reasons for this choice of substrate. One is that flat surfaces of graphite, with few impurities, are readily available in a number of forms. A second is that a large database exists for other films on this surface.<sup>33–35</sup> In addition, the problem of water adsorption on graphite is related to that of water on and inside carbon nanotubes, a subject of considerable current interest.<sup>36–39</sup> Finally, and most importantly, water is known to not wet the surface of graphite at room temperature; different values of the contact angle have been reported, presumably due to surface imperfections.<sup>40–44</sup> Thus, the predicted value of  $T_w$  is an open question.

The outline of this paper is the following: The next section describes a simple model of the wetting transition, used previously with much success in predicting transition behavior of both classical and quantum fluids. Section 3 discusses calculations of the potential and the resulting values of  $T_w$  for water/graphite. Analogous values of  $T_w$  for water on other surfaces are also discussed. Section 4 summarizes our results and conclusions.

\* To whom correspondence may be addressed. E-mail: mwc@psu.edu.

<sup>†</sup> Pennsylvania State University.

<sup>‡</sup> University of Pittsburgh.

<sup>§</sup> U.S. Department of Energy.

## II. Simple Model of the Transition

In a 1991 paper, Cheng et al.<sup>45</sup> presented a so-called “simple model” that derives predictions of  $T_w$  from the adsorption potential and the  $T$ -dependent values of the adsorbate’s density  $\rho$  and  $\sigma_{lg}$ . It achieves this goal by making a drastic assumption about the difference between two of the tensions appearing in Young’s equation

$$\sigma_{ls} - \sigma_{gs} = \sigma_{lg} + \rho \int dz V(z) \quad (3)$$

The logic of this relation is that  $\sigma_{ls}$  includes two contributions: the free energy associated with terminating the liquid (roughly equal to  $\sigma_{lg}$ ) and the liquid–surface interaction energy. In eq 3, the domain of integration extends between the minimum in the adsorption potential (at  $z = z_{\min}$ ) and infinity. The qualitative justification for this relation includes some implicit assumptions: (a) that a very low density film is present on the surface just below the transition pressure, so that  $\sigma_{gs}$  is essentially the surface tension of the bare substrate, (b) that when an infinitely thick liquid is in contact with the substrate its excess free energy differs from that of the bulk liquid–vapor interface by the integrated potential energy, (c) that the function  $V(r)$  is assumed to depend on just the normal coordinate  $z$ , and (d) that the integrated potential energy is adequately described by a “sharp-kink” approximation, in which the fluid density rises from zero to its bulk value at position  $z = z_{\min}$ . If eq 3 is valid, then the wetting transition relation, eq 2, can be manipulated to yield the following equation

$$\left(\frac{2\sigma_{lg}}{\rho}\right)_T = - \int dz V(z) = I \quad (4)$$

Equations 3 and 4 provide an implicit prediction for  $T_w$ , the temperature for which the left side equals the integral  $I$ , an explicit function of the adsorption system. The accuracy of eq 4 can be tested by experiments, if the potential is accurately known, or by simulations, if it is not, by using a hypothetical potential. An extensive set of classical simulations were undertaken by Curtarolo et al., who tested the model by comparing its predictions for wetting at the triple point with simulation data for a wide variety of adsorption potentials.<sup>29</sup> The results revealed that the model works well, overall, except for situations involving the very least attractive interactions, in which case the error was not negligible. More recently, path integral Monte Carlo simulations were undertaken by Shi et al.<sup>46</sup> They found the simple model’s predictions of  $T_w$  to differ from the exact results by about 10% (near 20 K) for the isotopes of hydrogen on several alkali metal surfaces. Comparison with experimental data indicates in most cases that eq 4 predicts  $T_w$  relatively well when the potential is accurately known.<sup>11–19</sup>

One model potential that has been used extensively to treat the wetting problem is

$$V(z) = \frac{4C_3^3}{27D^2z^9} - \frac{C_3}{z^3} \quad (5)$$

This 3-9 potential is analogous to the Lennard-Jones 6-12 interatomic potential and is usually adopted for similar reasons; it combines a rigorously correct form of the asymptotic attraction ( $V \approx C_3/z^3$ ) and a simple power law repulsion to yield a qualitatively plausible and mathematically convenient functional form. The asymptotic van der Waals coefficient  $C_3$  is discussed in some detail in the next section. If one inserts eq 5 into eq 4,

with  $z_{\min} = [2C_3/(3D)]^{1/3}$ , one obtains

$$I = \frac{11}{24} \left(\frac{3}{2}\right)^{2/3} (C_3 D^2)^{1/3} \quad (6)$$

When combined with eq 4, this equation yields a simple relation that we employ to determine  $T_w$

$$(C_3 D^2)^{1/3} = 3.33 \left(\frac{\sigma_{lg}}{\rho}\right)_T \quad (7)$$

We emphasize that this relation is derived by combining the simple model of the wetting transition with the simple 3-9 model potential. Despite these approximations, the result appears to describe most of the wetting studies carried out so far, at least semiquantitatively. In the next section, we evaluate the predictions of the simple model for both the 3-9 potential and other water/graphite potentials.

## III. The Adsorption Potential

**A. Dispersion Interaction.** The first aspect of the adsorption potential considered here is the van der Waals dispersion coefficient  $C_3$ . While values of  $C_3$  exist for many gas/surface combinations,<sup>47</sup> no values have been calculated previously for the particular case of water/graphite. In the case of a molecule, there exist two independent contributions to this coefficient

$$C_3 = C_{\text{perm}} + C_{\text{vdw}} \quad (8)$$

The first term arises from the interaction between the molecule’s dipole moment  $\mathbf{p}$  and its image in the dielectric substrate, while the second term is due to the coupled dipolar charge fluctuations of the electrons on the molecule and those of the substrate.<sup>33</sup> In computing both of these terms, we assume a Drude model form for the surface dielectric response function

$$\frac{\epsilon(iE) - 1}{\epsilon(iE) + 1} = \frac{g}{1 + (E/E_s)^2} \quad (9)$$

Here,  $\epsilon(iE)$  is the dielectric function at imaginary energy  $iE$ , and  $E_s$  is a characteristic electronic energy of the solid. This form is exact, with  $g = 1$  and  $E_s$  equal to the surface plasmon energy, for a perfect conductor. It is also exact for a low-density solid comprised of Drude oscillators, representing the individual molecules. For many other materials, this parametrization has been shown to work well in determining the various dispersion coefficients.<sup>48</sup>

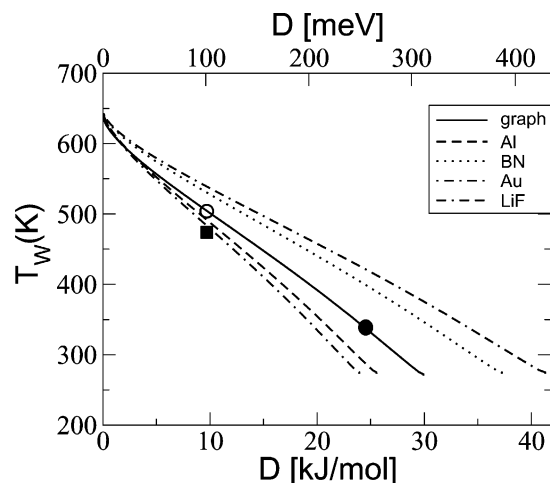
In evaluating  $C_{\text{perm}}$ , we adapt the derivation of London<sup>49,50</sup> to the case of a molecule having a permanent dipole moment  $\mathbf{p}$  oriented at an angle  $\phi$  relative to the surface normal, so that  $p_z = p \cos \phi$ , taking into account the dielectric screening at low energy (relevant to the slow rotation of  $\mathbf{p}$ )

$$C_{\text{perm}}(\phi) = \frac{gp^2(1 + \cos^2 \phi)}{16} \quad (10)$$

In the case of a randomly oriented, or freely rotating molecule, one obtains

$$C_{\text{perm}} = \langle C_{\text{perm}}(\phi) \rangle = \frac{gp^2}{12} \quad (11)$$

since the average of  $\cos^2 \phi$  for a randomly oriented dipole is  $1/3$ . In the case of water/graphite,  $g = 0.619^{33}$  and  $p = 1.85$  D,<sup>51</sup> so that (in this random case)  $C_{\text{perm}} = 104 \text{ meV} \cdot \text{\AA}^3$ .



**Figure 1.** Wetting temperature of water as a function of the well depth  $D$  of the adsorption potential on graphite (full curve), using eq 7 and  $C_3 = 1075 \text{ meV} \cdot \text{\AA}^3$ . The square is obtained from integration of the ZJ potential, eq 13, as described in the text. The open circle is based on the well depth used by Gordillo and Martí,<sup>36</sup> while the filled circle uses the calculated value of Feller and Jordan.<sup>64</sup> Other curves are dependencies computed for the indicated surfaces.

As to the van der Waals part of the interaction, an analogous modeling of the dynamical polarizability of water (with a characteristic energy  $E_{\text{water}}$ ) yields an expression

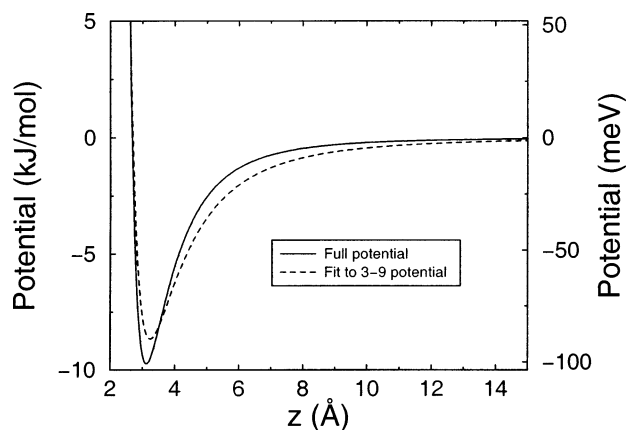
$$C_{\text{vdw}} = \frac{g\alpha E_s}{8 \left(1 + \frac{E_s}{E_{\text{water}}}\right)} \quad (12)$$

By use of data from Appendix E of Bruch et al.,<sup>33</sup> with  $E_{\text{water}} = 18.1 \text{ eV}$  and a rotationally averaged polarizability  $\alpha = 1.4 \text{ \AA}^3$ , we obtain a result  $C_{\text{vdw}} = 971 \text{ meV} \cdot \text{\AA}^3$ . From eq 8, the total  $C_3 = 1075 \text{ meV} \cdot \text{\AA}^3$ . Note that the permanent moment's contribution to this coefficient is about 10%.

By use of this value of  $C_3$  in eq 7, we can predict  $T_w$  as a function of the well depth,  $D$ . In Figure 1, we plot  $T_w$  as a function of  $D$ . The surface tension and density values for water used to construct Figure 1 were taken from ref 52.

**B. Well-Depth Calculations.** There have been numerous theoretical and semiempirical studies of the water–graphite interaction, as have recently been discussed by Werder et al. and Pertsin and Grunze.<sup>53,54</sup> The values emerging from these studies vary between 7.1 and 24.3 kJ/mol. This range reflects the wide variety of methods and assumptions employed in these calculations. Figure 1 indicates the corresponding values of the wetting temperature derived from some of these well depths.

**C. Results for a New Potential.** Each carbon atom in graphite has a local quadrupole moment due to the symmetry of the crystal.<sup>55–59</sup> Vernov and Steele developed an angle-explicit, position-dependent potential that accounts for the interactions of a polar molecule with the quadrupoles on the carbon atoms of graphite.<sup>57</sup> They found that the corrugation of the water–graphite potential energy surface in the  $x, y$  plane was dramatically enhanced by the inclusion of the dipole–quadrupole term. Zhao and Johnson have recently developed approximate expressions for the dipole–induced dipole, dipole–quadrupole, and quadrupole–quadrupole potential terms for polar fluid molecules interacting with graphite.<sup>60</sup> The expressions are angle averaged and integrated, so that they depend only on the normal coordinate,  $z$ . For water on graphite, only the dipole–



**Figure 2.** Potential energy function of Zhao and Johnson,<sup>60</sup> discussed in the text, evaluated at  $T = 474 \text{ K}$  (full curve). The dashed curve is a 3-9 potential fit to this potential.

induced dipole and dipole–quadrupole terms are important. These are given by

$$V_{\text{polar}} = \frac{\Delta \rho_s \pi p^2}{(4\pi\epsilon_0)^2} \left\{ \frac{\alpha_c}{2} \left[ \frac{1}{z^4} + \frac{1}{3\Delta(z + \Delta)^3} \right] + \frac{\Theta_c^2}{3k_B T} \left[ \frac{1}{z^6} + \frac{1}{5\Delta(z + \Delta)^5} \right] \right\} \quad (13)$$

where  $\Delta$  is the distance between the graphene layers,  $\rho_s$  is the density of atoms in graphite,  $\epsilon_0$  is the vacuum permittivity,  $p$  is the magnitude of the dipole moment of water,  $\alpha_c$  is the angle-averaged polarizability of a carbon atom in graphite, and  $\Theta_c$  is the quadrupole moment on each carbon atom in graphite. The polar terms in eq 13 are added to the “Steele model” 10-4-3 potential<sup>61,62</sup> to obtain the full potential.

The numerical value of the potential depends on the specific fluid potential used for water. In this work, we employ the Lennard-Jones parameters from the TIP4P potential<sup>63</sup> for calculating the 10-4-3 van der Waals solid-fluid contribution. For the polar terms, we use a point dipole having the experimental gas-phase value of 1.85 D. The potential depends on temperature, due to Boltzmann angle averaging of the dipole–quadrupole term. However, the temperature dependence is relatively weak. The water–graphite potential at 474 K and a fit of this potential to the 3-9 functional form, eq 5, are plotted in Figure 2. The main difference between the Zhao–Johnson (ZJ) potential and the 3-9 potential developed above is that the ZJ potential accounts for dipole–quadrupole interactions, but calculation of the  $C_3$  term does not. The dipole–quadrupole term accounts for 12% of the energy at the potential minimum at 474 K, whereas the induction term is 22%. Note that the 3-9 functional form does not accurately describe the ZJ potential. The 3-9 potential is not as deep as the full potential, but it has a longer range (decays more slowly).

With this new potential, one can evaluate the simple model to predict  $T_w$  in three ways. One is by directly integrating  $V(z)$ , including its weak  $T$  dependence and using eq 4. A second is to employ the simplified expression in eq 7, using a 3-9 potential fit to the theoretical potential; the fit values are  $D = 89.8 \text{ meV}$  and  $C_3 = 4751 \text{ meV} \cdot \text{\AA}^3$ . A third way to predict  $T_w$  is to use the theoretical well depth (9.74 kJ/mol or 101 meV) and the theoretical value of  $C_3$ , i.e., use the curve in Figure 1. The results in these three cases are  $T_w = 474, 416$ , and 504 K, respectively. Of these three, the value  $T_w = 474 \text{ K}$ , obtained from the direct



**TABLE 1: van der Waals Interaction Coefficient and Other Parameters Relevant to Wetting of Water on Indicated Surfaces<sup>a</sup>**

	graph	BN	Al	Au	LiF
$g$	0.619	0.38	0.98	0.84	0.31
$E_s$ (au)	0.667	0.71	0.473	0.888	0.74
$C_{\text{perm}}$ (meV·Å <sup>3</sup> )	104	64	166	142	51
$C_{\text{vdw}}$ (meV·Å <sup>3</sup> )	971	614	1278	1502	511
$C_{\text{tot}}$ (meV·Å <sup>3</sup> )	1074	678	1444	1644	562

<sup>a</sup> Data for  $g$  and  $E_s$  are taken from Bruch et al.,<sup>33</sup> while the other quantities are computed from equations in the text.

integration, is the recommended choice since it is based on the theoretical potential, without recourse to a preconceived model potential. The 3-9 fit and resulting low value of  $T_w$  are particularly suspect because the fitted value of  $C_3$  is a factor of 4 greater than the value ( $C_3 = 1075$  meV·Å<sup>3</sup>) derived in the previous section, which has an estimated uncertainty of 20 (based on previous studies of  $C_3$  values).<sup>47</sup> The third value coincides with the one predicted from the well depth calculated by Gordillo et al.<sup>36</sup>

**D. Wetting on Other Surfaces.** Similar calculations of  $C_3$  have been carried out for water on other surfaces, with results presented in Table 1. For any one of these surfaces, say  $x$ , one can compute the dependence  $D_x(T_w)$  of the well depth on  $T_w$  from eq 7 by scaling from the functional dependence  $D_g(T_w)$  for graphite, shown in Figure 1

$$D_x(T_w) = \left( \frac{C_g}{C_x} \right)^{1/2} D_g(T_w) \quad (14)$$

Here,  $C_x/C_g$  is the ratio of van der Waals coefficients for the surface of interest to that on graphite. Figure 1 shows this behavior for these various surfaces. We have assumed only physisorption of water and that the solid retains its structure and chemical integrity in each case. As a reviewer pointed out, this last assumption is not valid for LiF, which would be partially dissolved by water (an extreme form of wetting).

#### IV. Summary and Conclusions

Although water does not wet many surfaces, no wetting transition for water has been observed. Quite general arguments predict such a transition, unless the adsorption potential is extremely weak (as occurs for Ne/Cs).<sup>10</sup> In this paper, we have predicted the wetting temperature as a function of  $D$ , using a simple model that has been found reliable for spherical adsorbates.<sup>29</sup> One reason for concern about the reliability of the prediction is that the method has not yet been tested by comparing with results for water per se. Water is greatly affected by electrostatic forces absent from the nonpolar systems. However, the simple physical picture underlying the model and its previous success both suggest that it should be quite useful in motivating and guiding experimental searches for such a transition.

Evidently, this subject can benefit from further research in three related directions. One is to carry out simulations of water/graphite based on reliable intermolecular interactions in order to determine the accuracy of the simple model employed here. A second is to further refine the water/graphite interaction potential, which is very uncertain. A third is an experimental search for the predicted wetting transition.

**Acknowledgment.** We acknowledge very helpful discussions with Ken Jordan and Erwin Vogler. We are grateful for

support from NSF (Grants 02-08520 and 03-03916 and EEC 0085480 (J.K.J.)).

#### References and Notes

- (1) Rowlinson, S.; Widom, B. *Molecular theory of capillarity*; Clarendon: Oxford, 1989.
- (2) Davis, H. T. *Statistical mechanics of phases, interfaces and thin films*; VCH: New York, 1996.
- (3) Dietrich, S. In *Phase Transition and Critical Phenomena*; Domb, C., Lebowitz, J. L., Eds.; Academic: London, 1988; Vol. 12, p 1.
- (4) Ross, D.; Bonn, D. *Rep. Prog. Phys.* **2001**, *64*, 1085.
- (5) Thus far, only first-order transitions have been observed on solid surfaces, but higher order transitions have been observed on liquid surfaces; see ref 6.
- (6) Ragil, K.; Meunier, J.; Broseta, D.; Indekeu, J.; Bonn, D. *Phys. Rev. Lett.* **1996**, *77*, 1532.
- (7) Cahn, J. W. *J. Chem. Phys.* **1977**, *66*, 3667.
- (8) Ebner, C.; Saam, W. F. *Phys. Rev. Lett.* **1977**, *38*, 1486.
- (9) Ebner, C.; Saam, W. F. *Phys. Rev. B* **1987**, *35*, 1822.
- (10) Ancilotto, F.; Curtarolo, S.; Toigo, F.; Cole, M. W. *Phys. Rev. Lett.* **2001**, *87*, 206103.
- (11) Ross, D.; Taborek, P.; Rutledge, J. E. *Phys. Rev. Lett.* **1995**, *74*, 4483–4486.
- (12) Ross, D.; Phillips, J. A.; Rutledge, J. E.; Taborek, P. *J. Low Temp. Phys.* **1997**, *106*, 81–92.
- (13) Bojan, M. J.; Stan, G.; Curtarolo, S.; Steele, W. A.; Cole, M. W. *Phys. Rev. E* **1999**, *59*, 864–873.
- (14) Mistura, G.; Lee, H. C.; Chan, M. H. W. *J. Low Temp. Phys.* **1994**, *96*, 221.
- (15) Ross, D.; Taborek, P.; Rutledge, J. E. *Phys. Rev. B* **1998**, *58*, R4274.
- (16) Hallock, R. B. *J. Low Temp. Phys.* **1995**, *101*, 31.
- (17) Hess, G. B.; Sabatini, M. J.; Chan, M. H. W. *Phys. Rev. Lett.* **1997**, *78*, 1739.
- (18) Demolder, B.; Bigelow, N.; Nacher, P.; Dupont-Roc, J. *J. Low Temp. Phys.* **1995**, *98*, 91.
- (19) Wyatt, A. F. G.; Klier, J.; Stefany, P. *Phys. Rev. Lett.* **1995**, *74*, 1151.
- (20) Yao, M.; Hensel, F. *J. Phys.: Condens. Matter* **1996**, *8*, 9547.
- (21) Hensel, F.; Yao, M. *Ber. Bunsen-Ges. Phys. Chem.* **1998**, *102*, 1798.
- (22) Ohmasa, Y.; Kajihara, Y.; Yao, M. *J. Phys.: Condens. Matter* **1998**, *10*, 11589.
- (23) Ohmasa, Y.; Kajihara, Y.; Yao, M. *Phys. Rev. E* **2001**, *63*, 051601.
- (24) Kozhevnikov, V. F.; Arnold, D. I.; Naurzakov, S. P.; Fisher, M. E. *Phys. Rev. Lett.* **1997**, *78*, 1735.
- (25) Kozhevnikov, V.; Arnold, D.; Naurzakov, S.; Fisher, M. *Fluid Phase Equilib.* **1998**, *150*, 625.
- (26) Finn, J. E.; Monson, P. A. *Phys. Rev. A* **1989**, *39*, 6402–6408.
- (27) Chizmeshya, A.; Cole, M. W.; Zaremba, E. *J. Low Temp. Phys.* **1998**, *110*, 677.
- (28) Ancilotto, F.; Faccin, F.; Toigo, F. *Phys. Rev. B* **2000**, *62*, 17035.
- (29) Curtarolo, S.; Stan, G.; Bojan, M. J.; Cole, M. W.; Steele, W. A. *Phys. Rev. E* **2000**, *61*, 1670.
- (30) We note that the gases Ne, Ar, Kr, and Xe have the ratio  $\epsilon/k_B T_3 = 1.7$  approximately where  $\epsilon$  values are well depths from experimental data. (See Table A2 in Bruch et al.)<sup>33</sup>
- (31) A different kind of surface-wetting transition has been attributed to the presence of water on the surface of caprolactam. See ref 32.
- (32) Wong, G. C. L.; de Jeu, W. H.; Bögels, G.; van den Berg, E. P. G.; Bennema, P. *Phys. Rev. B* **1997**, *55*, 1388–1390.
- (33) Bruch, L. W.; Cole, M. W.; Zaremba, E. *Physical Adsorption: Forces and Phenomena*; Clarendon Press: Oxford, 1997.
- (34) Shrimpton, N. D.; Cole, M. W.; Steele, W. A.; Chan, M. H. W. *Surface Properties of Layered Structures*; Kluwer Academic Publishers: Dordrecht, 1992.
- (35) Zeppenfeld, P. *Physics of Covered Surfaces, New Series, Group III*; Springer: Berlin, 2001; Vol. 42.
- (36) Gordillo, M. C.; Martí, J. *Phys. Rev. B* **2003**, *67*, 205425.
- (37) Maibaum, L.; Chandler, D. *J. Phys. Chem. B* **2003**, *107*, 1189.
- (38) Rivera, J. L.; McCabe, C.; Cummings, P. T. *Nano Lett.* **2002**, *2*, 1427.
- (39) Werder, T.; Walther, J. H.; Jaffe, R. L.; Halicioglu, T.; Noca, F.; Koumoutsakos, P. *Nano Lett.* **2001**, *1*, 697.
- (40) Luna, M.; Colchero, J.; Baro, A. M. *J. Phys. Chem. B* **1999**, *103*, 9576.
- (41) Miura, K.; Morimoto, T. *Langmuir* **1991**, *7*, 374.
- (42) Schrader, M. E. *J. Phys. Chem.* **1980**, *84*, 2774.
- (43) Morcos, I. *J. Chem. Phys.* **1972**, *57*, 1801.
- (44) Fowkes, F. M.; Harkins, W. D. *J. Am. Chem. Soc.* **1940**, *62*, 3377.

- (45) Cheng, E.; Cole, M. W.; Saam, W. F.; Treiner, J. *Phys. Rev. Lett.* **1991**, 67, 1007–1010.
- (46) Shi, W.; Johnson, J. K.; Cole, M. *Phys. Rev. B* **2003**, 68, 125401.
- (47) Vidali, G.; Ihm, G.; Kim, H. Y.; Cole, M. W. *Surf. Sci. Rep.* **1991**, 12, 133–181.
- (48) Rauber, S.; Klein, J. R.; Cole, M. W.; Bruch, L. W. *Surf. Sci.* **1982**, 123, 173.
- (49) London, F. Z. *F. Phys.* **1930**, 63, 245.
- (50) London, F. Z. *F. Phys. Chem.* **1930**, B11, 222.
- (51) Clough, S. A.; Beers, Y.; Klein, G. P.; Rothman, L. S. *J. Chem. Phys.* **1973**, 59, 2254–2259.
- (52) Data from NIST Standard Reference Database 69. March 2003. Release: *NIST Chemistry WebBook*. <http://webbook.nist.gov/>.
- (53) Werder, T.; Walther, J. H.; Jaffe, R. L.; Halicioglu, T.; Koumoutsakos, P. *J. Phys. Chem. B* **2003**, 107, 1345.
- (54) Pertsin, A.; Grunze, M. *J. Phys. Chem. B* **2004**, 108, 1357.
- (55) Nicholson, D.; Cracknell, R. F.; Parsonage, N. G. *Mol. Sim.* **1990**, 5, 307–314.
- (56) Cracknell, R. F. *Monte Carlo and Theoretical Studies of Aqueous Lamellar Systems*. Ph.D. Thesis, Imperial College, London, 1990.
- (57) Vernov, A.; Steele, W. A. *Langmuir* **1992**, 8, 155–159.
- (58) Whitehouse, D. B.; Buckingham, A. D. *J. Chem. Soc., Faraday Trans.* **1993**, 89, 1909–1913.
- (59) Hansen, F. Y.; Bruch, L. W.; Roosevelt, S. E. *Phys. Rev. B* **1992**, 45, 11238–11248.
- (60) Zhao, X. C.; Johnson, J. K. *Mol. Sim.*, submitted.
- (61) Steele, W. A. *Surf. Sci.* **1973**, 36, 317–352.
- (62) Steele, W. A. *Chem. Rev.* **1993**, 93, 2356–2378.
- (63) Jorgensen, W. L.; Chandrasekhar, J.; Madura, J. D.; Impey, R. W.; Klein, M. L. *J. Chem. Phys.* **1983**, 79.
- (64) Feller, D.; Jordan, K. D. *J. Phys. Chem. A* **2000**, 104, 9971–9975.