

**Sn<sub>3</sub> and Sn<sub>10</sub> sulfonate–oxide–hydroxide clusters with two different sulfonate binding modes**

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*Dalton Trans.*, 2007, 3121 (DOI: 10.1039/b706377f). **Amendment published 21st May 2008**In Scheme 1, TSA should be changed to 2,5-Me<sub>2</sub>C<sub>6</sub>H<sub>3</sub>SO<sub>3</sub>H on the lower arrow.

The following paragraph replaces the third paragraph of page 3122:

As the cleavage of tin–aryl groups resulting in the presence of inorganic tins in **2** could originate from the prolonged treatment of **1** under acidic conditions, we undertook a study of the reaction of **1** with bulkier 2,5-dimethylbenzenesulfonic acid under milder conditions: **1** was first subjected to hydrolysis and then reacted with 2,5-dimethylbenzenesulfonic acid. So, treatment of **1** with twelve equivalents of water in THF resulted in the cleavage of the tin–alkynyl bonds. Then an equimolar reaction of 2,5-dimethylbenzenesulfonic acid at 72 °C for an additional 6 h yielded the colourless crystals of **3**  $\cdot$  2,5-Me<sub>2</sub>C<sub>6</sub>H<sub>3</sub>SO<sub>3</sub><sup>−</sup>, 2 toluene. The IR spectrum of **3** was devoid of any acetylenic absorption band and the strong bands at 3476 and 3416 cm<sup>−1</sup> indicated the presence of associated OH groups in the molecule. Compound **3** is unstable in solution and undergoes slow decomposition to yield a soluble NMR silent tin–oxo polymer.<sup>16</sup> The CPMAS <sup>119</sup>Sn investigation of **3** indicated the existence of only one tin site in the molecule with a shift of  $\delta = -600$  ppm, characteristic of a hexacoordinated monoorganotin. So, milder reaction conditions resulted in the total preservation of organic groups linked to the metal. Compound **3** consists of a six membered  $\{[2,4,6\text{-}^i\text{Pr}_3\text{C}_6\text{H}_2\text{Sn}(\mu_2\text{-OH})(\eta^2\mu_2\text{-}2,5\text{-Me}_2\text{C}_6\text{H}_3\text{SO}_3)_3(\mu_3\text{-O})]\}^+$  cation with three bridging  $\eta^2\mu_2$  sulfonate fragments and a central  $\mu_3$  oxygen atom (Fig. 2). The tin atoms are hexa-coordinated. The counteranion consists of a 2,5-dimethylbenzenesulfonate anion coordinated to the hydrogen atoms of the hydroxyl groups. Two kinds of sulfonate binding are present in **3**, as in **2**, but different from those of **2**: an electrostatic one involving three hydroxyl groups and a bridging covalent one involving two tin atoms. The coexistence of both these modes in the same cluster is, to the best of our knowledge, unprecedented.

The following paragraph replaces the second paragraph of page 3123:

**3**: H<sub>2</sub>O (0.9 g, 49.2 mmol) was added dropwise to a stirred solution of **1** (1.8 g, 4 mmol) in a mixture THF–toluene (3 : 1) (30 mL) at RT. The reaction mixture was left at RT for 9 h. 2,5-Dimethylbenzenesulfonic acid dihydrate (0.8 g, 4 mmol) was added and the solution was heated under reflux for 6 h. The solvents were removed under reduced pressure to result in a white solid. It was dissolved in toluene and the resulting solution filtered to give the colorless crystals of **3** which decompose slowly on standing. (1.2 g, 47%); mp 155–157 °C (decomp.); <sup>1</sup>H NMR (CDCl<sub>3</sub>, 250 MHz):  $\delta$  1.08–1.28 (dd, 54H, <sup>*i*</sup>Pr-CH<sub>3</sub>), 2.13–2.43 (m, 24H, Ar-CH<sub>3</sub>), 2.69–3.44 (m, 9H <sup>*i*</sup>Pr-CH), 3.26 (s, OH), 6.82–7.16, 7.48, 7.50, 7.84 (m, aryl-CH), 8.95, 9.66 (s, OH) ppm; CP MAS <sup>119</sup>Sn NMR:  $\delta = -600$  ppm; IR(KBr):  $\nu = 3476\text{s}, 3416\text{s}, 2963\text{vs}, 2913\text{m}, 2872\text{m}, 1596\text{w}, 1560\text{w}, 1492\text{w}, 1464\text{m}, 1420\text{w}, 1385\text{m}, 1364\text{m}, 1288\text{s}, 1270\text{s}, 1279\text{s}, 1148\text{m}, 1077\text{vs}, 1015\text{s}, 995\text{s}, 881\text{m}, 813\text{w}, 733\text{w}, 703\text{m}, 621\text{vs cm}^{-1}$ . Crystallographic data: **3a**: C<sub>87.5</sub>H<sub>119.5</sub>O<sub>16</sub>S<sub>4</sub>Sn<sub>3</sub>,  $M = 1911.64$ , monoclinic, space group *P*-1 (no. 2),  $a = 14.5396(1)$  Å,  $b = 17.6961(1)$  Å,  $c = 18.8745(1)$  Å,  $\alpha = 94.1887(3)^\circ$ ,  $\beta = 99.7233(3)^\circ$ ,  $\gamma = 109.3132(3)^\circ$ ,  $V = 4473.59(5)$  Å<sup>3</sup>,  $Z = 2$ ,  $\mu_{\text{calcd}} = 1.419$  g cm<sup>−3</sup>,  $T = 150(2)$  K,  $F(000) = 1973$ ,  $\rho(\text{Mo-K}\alpha) = 0.986$  mm<sup>−1</sup>, block crystal, 40 387 reflections measured, 20 465 unique ( $R_{\text{int}} = 0.0126$ ), 1000 parameters,  $R1 = 0.0329$  ( $I > 2\sigma(I)$ ),  $wR2 = 0.0829$  ( $I > 2\sigma(I)$ ), GOF = 1.033.

**Functionalized phosphorus analogues of the β-diketiminato ligand systems: bis(*N*-arylphosphinimino)acetonitrile-derived complexes of rhodium and iridium**

Kirsten Spannhoff, Gerald Kehr, Seda Kehr, Roland Föhlich and Gerhard Erker

*Dalton Trans.*, 2008, 3339 (DOI: 10.1039/b802032a). **Amendment published 16th July 2008**

Ref. 9 contains some errors. The correct ref. 9 should be:

Neutral derivatives: M. J. Sarsfield, M. Thornton-Pett and M. Bochmann, *J. Chem. Soc., Dalton Trans.*, 1999, 3329–3330; K. Kreischer, J. Kipke, M. Bauerfeind and J. Sundermeyer, *Z. Anorg. Allg. Chem.*, 2001, **627**, 1023–1028; P. Wei and D. W. Stephan, *Organometallics*, 2003, **22**, 601–604; V. Cadierno, J. Díez, J. García-Álvarez, J. Gimeno, M. J. Calhorda and L. F. Veiros, *Organometallics*, 2004, **23**, 2421–2433; C. Zhang, W.-H. Sun and Z.-X. Wang, *Eur. J. Inorg. Chem.*, 2006, 4895–4902.

## Effects of 5,5-substitution on dipyrrolylmethane ligand isomerization

Douglas L. Swartz II and Aaron L. Odom

*Dalton Trans.*, 2008, 4254 (DOI: 10.1039/b806947f). **Amendment published 22nd August 2008**

Unfortunately, references citing previous syntheses and use of the ligands H<sub>2</sub>cpm and H<sub>2</sub>tmcpm were excluded from the manuscript. Please see the following selected references.

H<sub>2</sub>cpm: W. H. Brown, B. J. Hutchinson and M. H. MacKinnon, *Can. J. Chem.*, 1971, **49**, 4017; T. Dube, S. Conoci, S. Gambarotta and G. P. A. Yap, *Organometallics*, 2000, **19**, 1182; T. Dube, M. Ganesan, S. Conoci, S. Gambarotta and G. P. A. Yap, *Organometallics*, 2000, **19**, 3716; M. Ganesan, S. Gambarotta and G. P. A. Yap, *Angew. Chem., Int. Ed.*, 2001, **40**, 766; D. M. M. Freckmann, T. Dube, C. D. Berube, S. Gambarotta and G. P. A. Yap, *Organometallics*, 2002, **21**, 1240; J. Ballmann, X. Sun, S. Dechert, E. Bill and F. Meyer, *J. Inorg. Biochem.*, 2007, 101, 305; S. Ramakrishnan and A. Srinivasan, *Org. Lett.*, 2007, **9**, 4769.

H<sub>2</sub>tmcpm: G. Givaja, M. Volpe, J. W. Leeland, M. A. Edwards, T. K. Young, S. B. Darby, S. D. Reid, A. J. Blake, C. Wilson, J. Wolowska, E. J. L. McInnes, M. Schöder and J. B. Love, *Chem.–Eur. J.*, 2007, **13**, 3707.

## Flexible scorpionates for transfer hydrogenation: the first example of their catalytic application

Nikolaos Tsoureas, Gareth R. Owen, Alex Hamilton and A. Guy Orpen

*Dalton Trans.*, 2008, 6039 (DOI: 10.1039/B809247H). **Amendment published 05th November 2008**

Ref. 14 contains some errors. The correct ref. 14 should be:

Tomoyuki Saito, Shigeki Kuwata and Takao Ikariya, *Chem. Lett.*, 2006, **35**, 1224.

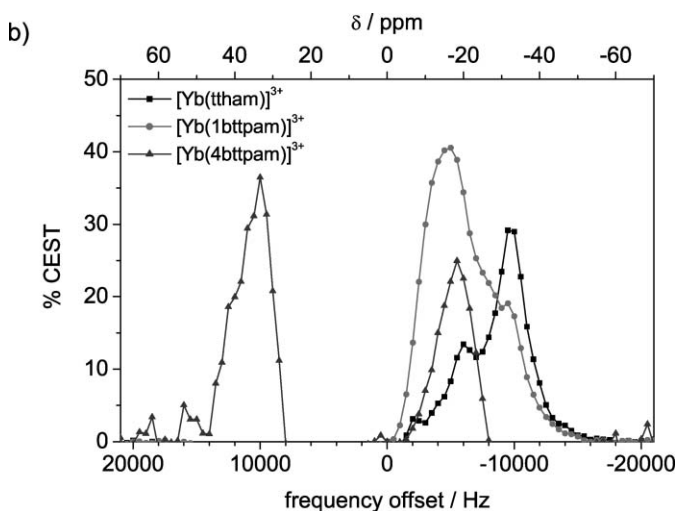
## Triethylenetetramine penta- and hexa-acetamide ligands and their ytterbium complexes as paraCEST contrast agents for MRI

Dirk Burdinski, Johan Lub, Jeroen A. Pikkemaat, Diana Moreno Jalón, Sophie Martial and Carolina Del Pozo Ochoa

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An incorrect earlier version of Fig. 4b had erroneously been submitted with the manuscript. In the correct Fig. 4b the red and blue curves remain unaltered and the corrected black curve shows lower intensities compared to the published data. This correction of Fig. 4b does not impact the discussion and conclusions of the published work, since only the peak positions are relevant in this context. The peak positions have been published and discussed correctly.

The published version of Fig. 4b should be amended as shown below:



### Continuous synthesis of organic–inorganic hybridized cubic nanoassemblies of octahedral cerium oxide nanocrystals and hexanedioic acid

Seiichi Takami, Satoshi Ohara, Tadafumi Adschiri, Yutaka Wakayama and Toyohiro Chikyow

*Dalton Trans.*, 2008, 5442 (DOI: 10.1039/b801099d). *Amendment published 24th July 2008*

The scale bar in Fig. 4g is incorrect and should read 200 nm.

### Stable $\text{Cl}_3^+$ salts and attempts to prepare $\text{CHI}_2^+$ and $\text{CH}_2\text{I}^+$

Ines Raabe, Daniel Himmel, Sonja Müller, Nils Trapp, Martin Kaupp and Ingo Krossing

*Dalton Trans.*, 2008, 946 (DOI: 10.1039/b714271d). *Amendment published 8th September 2008*

This article reported the unsuccessful attempts to make stable salts of  $\text{CH}_2\text{I}^+$ . As noted in the abstract: “All particles were also calculated at the MP2/TZVPP level to predict the vibrational and electronic spectra”, calculations of  $\text{CH}_2\text{I}^+$  and related molecules were reported. However recent studies which reported the first gas phase spectra of this ion [*J. Am. Chem. Soc.*, 2006, **128**, 9320; *J. Phys. Chem. A*, 2007, **111**, 10562] and gave detailed information on the electronic and vibrational spectra were omitted.

### Characterization and alkane oxidation activity of a diastereopure seven-coordinate iron(III) alkylperoxo complex

Silvia Gosiewska, Hjalmar P. Permentier, Andries P. Bruins, Gerard van Koten and Robertus J. M. Klein Gebbink

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Ernst E. van Faassen (Debye Institute, Surfaces, Interfaces and Devices, Utrecht University) appears in the acknowledgements section, whereas he should have appeared in the list of authors. The correct list of authors for this communication is as shown below.

Silvia Gosiewska, Ernst E. van Faassen, Hjalmar P. Permentier, Andries P. Bruins, Gerard van Koten and Robertus J. M. Klein Gebbink

### Coordination properties of $\text{Cu(II)}$ and $\text{Ni(II)}$ ions towards the C-terminal peptide fragment -TYTEHA- of histone H4

T. Karavelas, G. Malandrinos, N. Hadjiliadis, Piotr Mlynarz, Henryk Kozłowski, M. Barsam and I. Butler

*Dalton Trans.*, 2008, 1215 (DOI: 10.1039/b716863b). *Amendment published 27th February 2008*

One of the author's surname has been misspelled and should appear as Barsan rather than Barsam.

The Royal Society of Chemistry apologises for these errors and any consequent inconvenience to authors and readers.  
**Additions and corrections can be viewed online by accessing the original article to which they apply.**