

Kinetic energy and angular dependence of activated dissociative adsorption of N2 on W(110): Observed insensitivity to incidence angle

D. J. Auerbach, H. E. Pfnür, C. T. Rettner, J. E. Schlaegel, J. Lee, and R. J. Madix

Citation: The Journal of Chemical Physics 81, 2515 (1984); doi: 10.1063/1.447915

View online: http://dx.doi.org/10.1063/1.447915

View Table of Contents: http://scitation.aip.org/content/aip/journal/jcp/81/5?ver=pdfcov

Published by the AIP Publishing

Articles you may be interested in

Dissociative chemisorption of oxygen on Ir(110) as a function of angle of incidence: The effects of kinetic energy and surface temperature

J. Vac. Sci. Technol. A 11, 1926 (1993); 10.1116/1.578524

Dissociative chemisorption in the N2/W(110) system: Rotational state and angular momentum polarization dependence

J. Chem. Phys. 88, 5240 (1988); 10.1063/1.454600

Potential energy surface morphology and the variation of dissociative chemisorption probabilities with kinetic energy and angle: N2/W(110)

J. Chem. Phys. 88, 2033 (1988); 10.1063/1.454078

Effect of incidence kinetic energy and surface coverage on the dissociative chemisorption of oxygen on W(110)

J. Chem. Phys. 85, 1131 (1986); 10.1063/1.451310

Summary Abstract: Kinetic energy dependence of nitrogen adsorption on W(110)

J. Vac. Sci. Technol. A 1, 1270 (1983); 10.1116/1.572086



Kinetic energy and angular dependence of activated dissociative adsorption of N_2 on W(110): Observed insensitivity to incidence angle

D. J. Auerbach, H. E. Pfnür, C. T. Rettner, and J. E. Schlaegel IBM Research Laboratory, San Jose, California 95193

J. Lee and R. J. Madix

Department of Chemical Engineering, Stanford University, Stanford, California 94305

(Received 7 June 1984; accepted 20 June 1984)

Dissociative chemisorption, the process whereby intramolecular bonds are broken and new bonds formed to surface atoms, is one of the simplest of surface chemical processes. It is a key step in a wide variety of catalytic chemistry and, when accompanied by a sizable activation barrier, is commonly the rate determining steps in such reactions. An understanding of the nature and role of such activation barriers is therefore vital to any detailed picture of surface chemical dynamics. The interaction of H₂ with Cu surfaces has served as a prototype for such studies with interest dating back to Lennard-Jones who introduced the concept of a one-dimensional barrier (1DB) to adsorption. Studies of the the dependence of the dissociative adsorption probability at zero coverage S_0 on incident kinetic energy E_i and incidence angle θ_i for H_2 /Cu have supported this picture,² as have the observation of noncosine angular distributions and non-Boltzmann kinetic energy distributions of molecules which recombine and desorb.³⁻⁵ A key feature of this 1DB model is that the dissociative adsorption probability scales with the normal component of the incident kinetic energy $E_n = E_i \cos^2(\theta_i).$

For heavier molecules, far less data is available on the dynamical aspects of activated dissociative chemisorption, despite the great chemical relevance of such processes. Recently, Cosser et al.⁶ reported a noncosine angular distribution for N_2 desorbing from W(110). This was also interpreted in terms of a 1DB model, yielding an activation barrier height of $E_a \sim 17$ kJ mol⁻¹. In recent molecular beam experiments^{7,8} performed at $\theta_i = 45^\circ$, we found that S_0 does indeed increase dramatically with E_i . However, the steepest increase occurred at $E_i \sim 80$ kJ mol⁻¹ ($E_n \sim 40$ kJ mol⁻¹). In this Communication we report measurements of the dependence of S_0 on incidence angle.⁹ Contrary to expectations, we find that $S_0(E_i,\theta_i)$ is virtually independent of θ_i for $0 < \theta_i \le 45^\circ$, with only a weak dependence at larger angles.

The experiments are performed with a supersonic molecular beam of N_2 incident on a W(110) surface mounted in a UHV chamber (base pressure $<2\times10^{-11}$ Torr). E_i was varied from 9 to 200 kJ mol⁻¹ by seeding in H_2 or He and by changing the nozzle temperature. (Energies are determined by time-of-flight techniques.) The crystal used for the data points reported here was 1.39° off from the W(110) direction. Towards the end of this study, the crystal was repolished to within 0.5° and the data was found to be insensitive to this change. The sample was mounted on a manipulator which permitted accurate control of the incidence angle and provided e-beam heating and liquid nitrogen cooling. Temperature, monitored with a thermocouple spot welded to the back of the sample, was controlled by a feedback system

interfaced to a minicomputer. The computer was used to control all the sequencing of the experiment: cleaning, dosing with N_2 , temperature ramping, and moving the sample in and out of the beam.

Two methods were used to measure S_0 : determination of the initial slope of a surface coverage vs exposure curve and direct measurement of the fraction of the beam which reflects from the surface (following the approach of King and Wells¹⁰). For the former, coverage measurements were made by recording the area under temperature programmed desorption spectra; exposure was determined by measuring the flux of the molecular beam ($\sim 10^{15}-10^{16}$ molecules/s/cm²) and controlling the exposure time (down to $\sim 10^{-2}$ s). All the data points presented below were obtained by this method. The latter method, appropriate for $S_0 \gtrsim 0.05$, was used as an independent check. Agreement between these methods was always within a factor of 2, and usually much better. Further details of the apparatus and measurement procedures are given elsewhere. 8,11-13

Figure 1 displays S_0 values obtained at various incidence angles plotted against E_i . This data, obtained with a sample temperature of 800 K, clearly indicates the activated nature of the adsorption process. However, the sharp increase in adsorption probability is seen to occur at $E_i \sim 80$

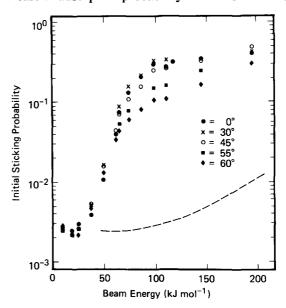


FIG. 1. Initial sticking probability for N_2 on W(110) as a function of beam energy at various angles of incidence for a sample temperature of 800 K. It is seen that the data are insensitive to incidence angle for $0^{\circ} \lesssim \theta \lesssim 45^{\circ}$. The dashed line indicates the sticking probabilities predicted for $\theta = 60^{\circ}$ obtained by assuming normal energy scaling for the $\theta = 0^{\circ}$ data. Clearly the actual 60° results fall much closer to the uncorrected 0° data. Error bars ($\sim \pm 25\%$) have been omitted for clarity of presentation.

kJ mol $^{-1}$ (point of maximum slope on a linear plot), which is much higher than the prediction of Cosser et al. For low beam energies, we find $S_0 \approx 2.5 \times 10^{-3}$, in agreement with the value of $3 \pm 1 \times 10^{-3}$ for N_2 gas at 300 K reported in the literature. He literature. We also note that the increase of S_0 with E_i is far from a step function, contrary to the behavior expected for a single barrier. In this respect the data are similar to that reported for H_2 on H_2 on H_3 on H_4 on H_5 observed here is considerably larger than the factor of H_5 observed for that system.

It is clear from Fig. 1 that for $\theta_i \leq 45^\circ$, S_0 is virtually independent of θ_i . Thus we find a complete breakdown of the normal energy scaling behavior observed for H₂ on Cu. Normal energy scaling must result from any 1DB model or indeed from any model in which the molecule-surface potential depends on z, the normal coordinate, but not on x or v. A much stronger dependence of potential on z than on x or v(weak corrugation) is quite characteristic of low energy atomic and molecular interactions with low index planes of metal surfaces. 17 Weak corrugation is observed 18 in the Ar-Pt(111) system even at 200 kJ mol⁻¹, the highest energy studied. Rotational excitation of NO in collisions with Ag(111) has been shown to scale with $E_n^{13,19}$ up to at least 180 kJ mol⁻¹. The breakdown of normal energy scaling reported here is all the more surprising since we have recently observed almost perfect normal energy scaling for CH4 dissociative chemisorption on this same crystal surface.²⁰

Possible explanations of the observed insensitivity of S_0 to θ_i fall into two classes: those which consider the surface to be strongly corrugated or rough on an atomic scale, and those which postulate a long-lived surface intermediate or precursor state which mediates the randomization of the ini-

tial energy modes. He scattering (30% at 500 K) indicates that the surface is smooth as seen by thermal energy atoms.

Two of us (RJM and JL) gratefully acknowledge the support of the DOE Office of Basic Energy Sciences, Grant DE-AT03-79ER10490.

- ¹J. E. Lennard-Jones, Trans. Faraday Soc. 28, 333 (1932).
- ²M. Balooch, M. J. Cardillo, D. R. Miller, and R. E. Stickney, Surf. Sci. 46, 358 (1974).
- ³W. van Willigen, Phys. Lett. A 28, 80 (1968).
- ⁴M. J. Cardillo, M. Balooch, and R. E. Stickney, Surf. Sci. 44, 310 (1974); 50, 263 (1975).
- ⁵G. Comsa, R. David, and B. J. Schumacher, Surf. Sci. 85, 45 (1979).
- ⁶R. C. Cosser, S. R. Bare, S. M. Francis, and D. A. King, Vacuum 31, 503 (1981).
- ⁷D. J. Auerbach, J. E. Schlaegel, J. Lee, and R. J. Madix, J. Vac. Sci. Technol. A 1, 1271 (1983).
- ⁸J. Lee, R. J. Madix, J. E. Schlaegel, and D. J. Auerbach, Surf. Sci. (in press).
- ⁹This paper is concerned only with the initial sticking or dissociative adsorption probability, i.e., the value obtained in the limit of zero surface coverage.
- ¹⁰D. A. King and M. G. Wells, Surf. Sci. 29, 454 (1971).
- ¹¹H. E. Pfnuer, C. T. Rettner, and D. J. Auerbach (to be published).
- ¹²A. W. Kleyn, A. C. Luntz, and D. J. Auerbach, Phys. Rev. Lett. 47, 1169 (1981).
- ¹³A. W. Kleyn, A. C. Luntz, and D. J. Auerbach, Surf. Sci. 117, 33 (1982).
- ¹⁴T. E. Madey and J. T. Yates, Nuovo Cimento Suppl. 5, 486 (1967).
- ¹⁵P. W. Tamm and L. D. Schmidt, Surf. Sci. 26, 286 (1971).
- ¹⁶C. Somerton and D. A. King, Surf. Sci. 89, 391 (1979).
- ¹⁷F. O. Goodmann and H. Y. Wachman, *Dynamics of Gas Surface Scattering* (Academic, New York, 1976).
- ¹⁸J. E. Hurst, L. Wharton, K. C. Janda, and D. J. Auerbach, J. Chem. Phys. 78, 1559 (1983).
- ¹⁹G. D. Kubiak, J. E. Hurst, H. G. Rennagel, G. M. McClelland, and R. N. Zare, J. Chem. Phys. **79**, 5163 (1983).
- ²⁰C. T. Rettner, H. E. Pfnür, and D. J. Auerbach (to be published).

COMMENTS

Remarks on the magnitude of van der Waals induced dipole moments in physisorption

Robert A. Kromhout and Bruno Linder Chemical Physics Program, The Florida State University, Tallahassee, Florida 32306

(Received 15 March 1984; accepted 17 May 1984)

In a recent paper, Wandelt and Hulse¹ presented a detailed analysis of the observed local work function of palladium with less than a monolayer coverage of xenon. It is generally assumed that the change in the work function of the metal is associated with the dipole moment induced in the adsorbed atom. Wandelt and Hulse mentioned three possibilities which could account for the sign and magnitude of the adatom dipole moment and its face specificity. They were: (1) charge transfer, (2) "static" (van der Waals) induced moments, and (3) change in the Xe "dynamic polarization" due to the virtual 5p to 6s excitation.

The paper by Wandelt and Hulse is an important experimental contribution to the subject of physisorption and has far reaching theoretical implications. In particular, it points out that neither nonbonded charge transfer nor van der Waals induced dipole moment theories can account for face specificity. However, the magnitude of the induced dipole moment, the authors noted, can be reproduced approximately by the model calculations of Antoniewicz^{2,3} and of Sunnen and Schmeits.⁴

The purpose of this note is to point out that, together with Galatry and Girard,⁵ we believe that the use of the