Erratum: Nuclear Magnetic Resonance of the Aquated Proton. II. Chloroauric Acid Tetrahydrate. Phase Transitions and Molecular Motion

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Citation: The Journal of Chemical Physics **56**, 3734 (1972); doi: 10.1063/1.1677768
View online: http://dx.doi.org/10.1063/1.1677768
View Table of Contents: http://scitation.aip.org/content/aip/journal/jcp/56/7?ver=pdfcov
Published by the AIP Publishing

Our model calculations indicated that, for normal-direction kinetic isotope effects ($k_{light} > k_{heavy}$), significant deviations of r from the 1.8–2.0 range may occur only when the individual ^{14}C and ^{13}C kinetic isotope effects are of unusually low magnitude and/or are associated with easily detectable temperature-dependence anomalies.[4] Although kinetic-isotope-effect temperature-dependence determinations have not been carried out for the oxaloacetic acid decarboxylation, the individual effects are probably too large to be associated with temperature-dependence anomalies. Thus, in this case also, either experimental error or factors not taken into account in the basic statistical-thermodynamic theory must be operative.

* Supported in part by the U.S. Atomic Energy Commission under Contract AT(30-1)-3663.

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and B. E. Tabor for the communication which led to the discovery of this error.

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[J. Chem. Phys. **55**, 5629 (1971)]

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The figures in this paper are given in the correct order but the figure captions are not in the correct order. The following table gives the correspondence between the correct figure numbers and the figure captions as they appeared in print.

Correct numbering	1	2	3	4	5	6	7	8
Printed caption	3	4	8	2	5	1	7	6

Erratum: Corrections to the Fuoss–Onsager Theory of Electrolytes

[J. Chem. Phys. **53**, 2173 (1970)]

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In the first line of the expression for L_2 in Eq. (20), the coefficient of the curly brackets should read $-(1+q^2)/8q$ rather than $-(1+q^2)/2q$; the quantity added to L_2 in making this correction should also be added to the expression for Λ_2 in Eq. (23). The comparison with experiment (Tables I and II and Figs. 3–10) is based on the correct formulas and hence is unaffected. The authors would like to thank E. Pitts

Erratum: Flash Photolytic Production, Reactive Lifetime, and Collisional Quenching of $O_2(b\ ^1\Sigma_g^+,\ v'=0)$

[J. Chem. Phys. **52**, 5502 (1970)]

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In Fig. 7 on p. 5506, the correct decay rates are exactly half those indicated on the ordinate. In Fig. 11 on p. 5508, the open circles represent the NH_3 data while the closed circles represent the H_2O data. In Fig. 12 on p. 5509, the units of pressure should be millitorr. The deactivation rate coefficients which are recorded in Table I are correct as shown except for He and Ne which should read $\sim 1.0\times10^{-17}$ cm^3 molecule$^{-1}\cdot$sec^{-1}.