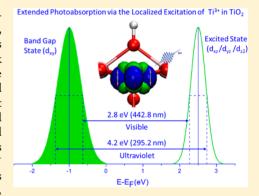


Localized Excitation of Ti³⁺ Ions in the Photoabsorption and Photocatalytic Activity of Reduced Rutile TiO₂

Zhiqiang Wang,^{†,#} Bo Wen,^{‡,||,#} Qunqing Hao,^{†,#} Li-Min Liu,*,[‡] Chuanyao Zhou,*,[†] Xinchun Mao,[†] Xiufeng Lang,[‡] Wen-Jin Yin,^{‡,⊥} Dongxu Dai,[†] Annabella Selloni,*,[§] and Xueming Yang*,[†]

Supporting Information

ABSTRACT: In reduced TiO₂, electronic transitions originating from the Ti³⁺-induced states in the band gap are known to contribute to the photoabsorption, being in fact responsible for the material's blue color, but the excited states accessed by these transitions have not been characterized in detail. In this work we investigate the excited state electronic structure of the prototypical rutile TiO₂(110) surface using two-photon photoemission spectroscopy (2PPE) and density functional theory (DFT) calculations. Using 2PPE, an excited resonant state derived from Ti³⁺ species is identified at 2.5 \pm 0.2 eV above the Fermi level (E_F) on both the reduced and hydroxylated surfaces. DFT calculations reveal that this excited state is closely related to the gap state at ~1.0 eV below E_F , as they both result from the Jahn—Teller induced splitting of the 3d orbitals of Ti³⁺ ions in reduced TiO₂. Localized excitation of Ti³⁺ ions via $3d \rightarrow 3d$ transitions from the gap state to this empty resonant state significantly increases the TiO₂



photoabsorption and extends the absorbance to the visible region, consistent with the observed enhancement of the visible light induced photocatalytic activity of TiO_2 through Ti^{3+} self-doping. Our work reveals the physical origin of the Ti^{3+} related photoabsorption and visible light photocatalytic activity in prototypical TiO_2 and also paves the way for the investigation of the electronic structure and photoabsorption of other metal oxides.

■ INTRODUCTION

Titanium dioxide (TiO₂) is well-known as one of the most widely used materials in photocatalysis and solar energy conversion. 1-9 All these applications are based on the generation of charge carriers by optical absorption, followed by their diffusion to the surface and transfer to adsorbed species. Stoichiometric rutile TiO2 is an insulator with a band gap of ~3 eV. In this material the photoexcitation involves transitions from the O 2p states in the valence band (VB) to the 3d states of the Ti (formally Ti⁴⁺) cations in the conduction band (CB). TiO₂ is very easily reduced, however. The reduction is accompanied by the formation of Ti³⁺ ions, for which a characteristic signature is the presence of a localized state in the band gap at ~0.8-1 eV below the CB edge. Many studies have focused on the Ti³⁺-induced gap states and their possible role in TiO₂ photocatalysis in recent years. ^{10–13} There is evidence that $d \rightarrow d$ transitions from the gap states contribute to the photoabsorption of reduced TiO2, as indicated by the material's blue color, 14 whose intensity increases with the degree of reduction. Because of the band gap narrowing and enhanced photoabsorption associated with the presence of the

 ${\rm Ti}^{3+}$ induced band gap state, visible light photocatalysis through ${\rm Ti}^{3+}$ self-doping has also been achieved. ^{15–19} While transitions from the band gap state are clearly important, little is known, however, about the character of the electronic excited states that are involved in these transitions, due also to the greater difficulty to experimentally access unoccupied states in comparison to occupied ones. A characterization of these states is not only scientifically interesting but also important for a better understanding and control of ${\rm TiO}_2$ photocatalysis.

In this work we utilize two-photon photoemission spectroscopy (2PPE) measurements in combination with density functional theory (DFT) calculations to investigate the excited states of the rutile ${\rm TiO_2(110)}$ surface, a well-established model system for surface science^{3,5,20–22} and theoretical^{23–27} studies of ${\rm TiO_2}$. 2PPE, which employs a 1 + 1 pump–probe scheme, is an ideal technique to investigate excited states. In the two-photon excitation process (shown schematically in Figure S1), population of an intermediate state is followed by emission to a

Received: May 7, 2015



[†]State Key Laboratory of Molecular Reaction Dynamics, Dalian Institute of Chemical Physics, Chinese Academy of Science, 457 Zhongshan Road, Dalian, 116023, Liaoning, P. R. China

[‡]Beijing Computational Science Research Center, Beijing, 100094, P. R. China

[§]Department of Chemistry, Princeton University, Princeton, New Jersey 08544, United States

International Center for Quantum Materials, Peking University, Beijing, 100871, P. R. China

 $^{^{\}perp}$ Chengdu Green Energy and Green Manufacturing Technology R&D Center, Chengdu, Sichuan 610207, P. R. China

Journal of the American Chemical Society

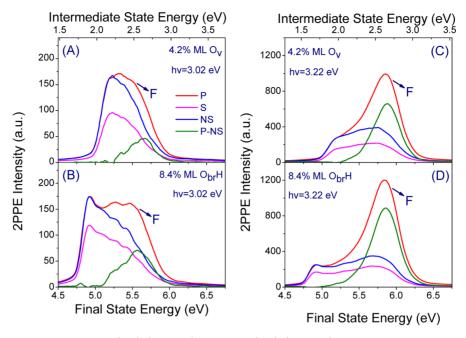


Figure 1. Typical 2PPE spectra for the r-TiO₂(110) (A and C) and h-TiO₂(110) (B and D) surfaces. The spectra were measured with both p-polarized (P) and s-polarized (S) light with a photon energy of 3.02 eV (A and B) and 3.22 eV (C and D) respectively. For comparison, S was normalized to P at the secondary electron signal edge. P-NS denotes the difference spectra, which was obtained by subtracting the normalized s-polarized data (NS) from the p-polarized data. The curves were smoothed by averaging the adjacent seven data points. The signal was integrated from -5 degrees to +5 degrees. Energies are measured with respect to E_F ; those in the top X-axis refer to the intermediate state, before absorption of the second photon. Note the different intensity scales of the spectra on the left (A and B) and right (C and D) panels.

final state above the vacuum level through absorption of a second photon. By 2PPE we observe an excited state centered at 2.5 \pm 0.2 eV above the Fermi energy ($E_{\rm F}$; hereafter, all energies are measured with respect to $E_{\rm F}$) on both the clean reduced and hydroxylated TiO₂(110) surfaces, indicating that this is an intrinsic state associated with the Ti³⁺ species in the reduced samples. DFT calculations further reveal that, similarly to the gap state at $\sim 0.8-1.0$ eV below $E_{\rm E}$, the excited resonant state originates from the Jahn-Teller induced splitting of the 3d orbitals of Ti^{3+} ions. This state can be accessed via $d \rightarrow d$ transitions from the gap state and increases significantly the photoabsorption of reduced TiO₂. Owing to the relative wide distribution of both the band gap state and excited state, the absorbance from $d \rightarrow d$ transitions ranges from the UV to the visible region, in agreement with the observed enhancement of the ${\rm TiO_2}$ photocatalytic activity under visible light irradiation through ${\rm Ti^{3+}}$ self-doping. 15–19,28 The present work not only provides a better understanding of the Ti³⁺ related visible light photocatalytic activity of the prototypical TiO₂ but also paves the way for the investigation of the electronic structure and photoabsorption of other metal oxides.

■ EXPERIMENTAL AND THEORETICAL DETAILS

All experiments were performed in a ultrahigh vacuum (UHV) apparatus (base pressure better than 5×10^{-11} mbar), which has been described in detail previously. Fig. 29 TiO₂(110) samples (Princeton Scientific Corp.) were prepared by cycles of Ar⁺ sputtering and UHV annealing at 850 K, generating a reduced substrate (r-TiO₂(110)). The cleanness and long-range order were confirmed by Auger electron spectroscopy (AES) and low energy electron diffraction (LEED), respectively. The concentration of the bridging oxygen vacancy (O_v) was determined to be 4.2% monolayer (ML; 1ML = 5.2×10^{14} cm⁻²) by temperature-programmed desorption (TPD) of water. Hydroxylated TiO₂(110) (h-TiO₂(110)) was prepared by photocatalyzed splitting of methanol on this surface, H

which was then heated to 370 K to remove species adsorbed on fivefold coordinated Ti (Ti_{5c}) sites. Methanol dissociates spontaneously at O_v sites, producing bridging methoxy (CH₃O_{br}) and hydroxyl (O_{br}H) groups.³² To minimize the amount of CH₂O_{br}, r-TiO₂(110) was first exposed to excess water followed by heating to 350 K to convert all the O_v to O_{br}H.³³ The residual water on Ti_{5c} sites after this procedure is below the detection limit of the mass spectrometer. Purified methanol was then dosed onto this surface and exposed to ultraviolet (UV) light. It should be pointed out that with this procedure 3.0% ML CH₃O_{br} still remains due to displacement, 34 but the same amount of CH₃O_{br} and O_{br}H contribute equally to the excited resonance signal (Figure S2). The second harmonic (SH, 2.95-3.59 eV) of a Ti: Sapphire oscillator (Spectra-Physics, MaiTai eHP DeepSee) was delivered in pulses of 90 fs duration and 1.5 nJ at an 80-MHz repetition rate. All photoemission measurements were performed with a TiO₂ sample temperature of 120

DFT calculations were performed with the CP2K/Quickstep package, 35 which uses a hybrid Gaussian and plane-waves approach. The hybrid HSE06 functional $^{36-38}$ was chosen for all the calculations. For comparison with the HSE06 results, selected calculations using the B3LYP hybrid functional 39 were also performed. Core electrons were described with norm-conserving Goedecker, Teter, and Hutter (GTH) pseudopotentials. 40 The wave functions of the valence electrons were expanded in terms of Gaussian functions with molecularly optimized double- ζ polarized basis sets (m-DZVP), which ensures a small basis set superposition error. 41 For the auxiliary basis set of plane waves a 280 Ry cutoff was used. Reciprocal space sampling was restricted to the Γ point.

The rutile-TiO₂(110) surface was modeled using a repeated slab geometry with slabs of four TiO₂ trilayers and a (4 × 2) (11.836 Å × 12.994 Å) surface cell. The vacuum separation between slabs was around 15 Å. Hydroxylated TiO₂(110) surfaces were modeled by adding 1, 2, and 4 hydrogen atoms to the surface bridging oxygen sites, corresponding to $O_{br}H$ coverages of 1/8 ML, 1/4 ML, and 1/2 ML, respectively. Reduced surfaces were modeled by removing one bridging oxygen atom per unit cell. All atoms in the slab were relaxed until the maximum residual force is less than 0.02 eV/Å. We note that

the theoretical hydroxyl/defect concentration is generally larger than the experimental one. As shown below, however, the properties of the excited state do not depend on the defect concentration except for the intensity. Therefore, it is well justified to connect the experimental and theoretical results in the present study.

■ RESULTS AND DISCUSSION

Polarization dependent 2PPE spectra for different TiO₂(110) surfaces and excitation energies are compared in Figure 1. Since in our samples $E_{\rm F}$ is close to the CB edge and the work function is 4.8-5.1 eV, electrons in the VB cannot be excited into vacuum via two-photon photoemission with the photon energy used in our work. Therefore, the only possible initial states of the observed photoelectrons are the Ti³⁺ induced states in the band gap. Figure 1A shows the 2PPE spectrum for the clean reduced surface, r-TiO₂(110), acquired at an excitation energy of 3.02 eV, a value similar to that used in previous studies on water or alcohol covered TiO₂(110).⁴²⁻⁴⁵ A small additional feature (labeled as F) around 2.5 eV above $E_{\rm F}$ is observed in the p-polarized 2PPE (p-2PPE) spectrum relative to s-polarized 2PPE. When the O_v 's on r-Ti $O_2(110)$ are titrated by water, bridging hydroxyls ($O_{br}H$'s) are produced, resulting in a hydroxylated surface. ³³ The 2PPE spectra for r-TiO₂(110) (Figure 1A) and $h\text{-TiO}_2(110)$ (Figure 1B) are clearly similar, suggesting that O_v and O_{br}H contribute similarly to the unoccupied states of TiO₂(110), as found also for the occupied state in the band gap. 11,46,47

The feature F in Figure 1A and 1B is similar but much smaller than the one observed at 2.3–2.4 eV above $E_{\rm F}$ on water or alcohol covered ${\rm TiO_2(110)}.^{42-45}$ This is the reason why this feature was ignored for bare TiO2 in previous studies. Instead, the strong resonance on water or alcohol covered TiO₂(110) was ascribed to a "wet electron" state (induced by the simultaneous presence of bridging hydroxyls and a water monolayer on water/TiO₂(110)⁴² and methyl groups of methanol/methoxy and bridging hydroxyl groups on methanol/ $\text{TiO}_2(110)$, and bridging hydroxyl groups on methanol/ $\text{TiO}_2(110)$, are respectively) or to the dissociation of alcohols on the $\text{TiO}_2(110)$ surface. An important reason why the excited resonance signal is more prominent on water or alcohol covered TiO₂(110) than on the bare surface is however that water or alcohol adsorption significantly reduces the work function, 49,50 leading to the separation of the secondary electrons and the excited resonance signal in 2PPE spectra. An alternative way to increase the signal separation without adding adsorbates is to raise the 2PPE excitation photon energy. This increases the kinetic energy of the photoelectrons of interest, thus allowing their separation from the secondary electrons, and also excites electrons closer to the maxima of the band gap states which are the initial states in the present 2PPE measurements. As shown in Figure 1C and 1D, feature F in the 2PPE spectra of r-TiO₂(110) and h-TiO₂(110) becomes indeed much more pronounced with an excitation photon energy of 3.22 eV, thus confirming that this resonance is not due to adsorbates but is rather an *intrinsic* property of the $TiO_2(110)$ surface.

2PPE spectra include information on both initial and intermediate states. In order to determine the origin of the resonant feature F, both one-photon photoemission (1PPE) and 2PPE spectra with similar initial and final states were measured on an $h\text{-TiO}_2(110)$ with 3.0% ML CH₃O_{br} and 15.4% ML O_{br}H. Since equal amounts of CH₃O_{br} and O_{br}H contribute equivalently to the resonant feature F (Figure S2), CH₃O_{br} was regarded as O_{br}H in the present experiments. An

excitation energy of 5.9 eV was used for the 1PPE, and correspondingly a photon energy of 2.95 eV, half of that for the 1PPE, was used for the 2PPE measurements (see Figure S1). From the normalized photoemission spectra in Figure 2A, it

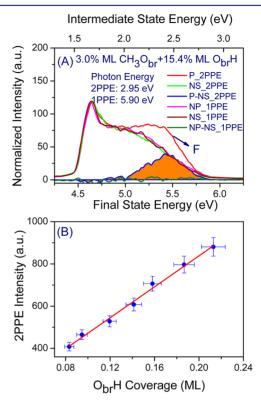


Figure 2. (A) 1PPE and 2PPE spectra of an h-TiO₂(110) surface with 3.0% ML CH₃O_{br} and 15.4% ML O_{br}H. The excitation photon energies for 1PPE and 2PPE were 5.90 and 2.95 eV, respectively. The signal in these spectra was integrated from -5 degrees to +5 degrees. Energies are measured with respect to the Fermi level. Except p-2PPE, the normalized s-2PPE, p-1PPE, and s-1PPE spectra are identical, suggesting feature F detected by p-2PPE is from an intermediate state. (B) O_{br}H coverage dependence of the excited resonance signal.

appears that s-2PPE, p-1PPE, and s-1PPE from an h-TiO₂(110) surface are nearly identical, and do not show any feature at ~ 2.5 eV above $E_{\rm F}$. At variance with the other spectra, p-2PPE clearly displays feature F. This feature is present in 2PPE and absent in 1PPE, indicating that it originates from the intermediate state. The lifetime of the excited state was then examined by timeresolved two-photon photoemission (TR-2PPE) with a 3.14 eV and 28 fs pulse (SHG from an oscillator, Femtolasers Produktions GmbH). The two-pulse correlation from the excited state overlaps completely with the autocorrelation from a polycrystalline molybdenum surface, suggesting the lifetime of the excited state is too short to measure, as found also in a previous study. 42 Nonetheless, by comparing the two-pulse correlation from the clean reduced TiO₂(110) and those covered by water or methanol, 42,43,48 an upper limit of 20 fs for the lifetime of the intrinsic excited state could be estimated, in accord with its spectral width of 0.5-0.6 eV.

To characterize the nature of the excited state, ${\rm TiO_2(110)}$ surfaces with different amounts of point defects (for example, surface ${\rm O_{br}H}$, surface ${\rm O_{vv}}$ and subsurface Ti interstitials) are desirable. However, due to the difficulty to quantitatively assess the concentration of subsurface Ti interstitials accompanying the creation of surface ${\rm O_{v}}$'s by conventional sputtering and

Journal of the American Chemical Society

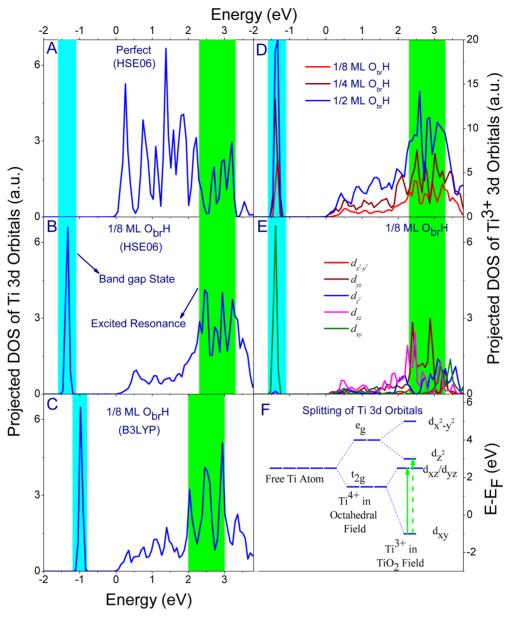


Figure 3. Projected density of states for Ti 3d orbitals: (A) Ti⁴⁺ ion in defect-free TiO₂(110), calculated using the HSE06 functional; (B) Ti³⁺ ion on h-TiO₂(110) with 1/8 ML O_b-H, calculated using the HSE06 functional; (C) same as (B), but calculated using the B3LYP functional. (D) Total projected DOS of Ti³⁺ 3d orbitals in TiO₂(110) with 1/8 ML (red), 1/4 ML (wine), and 1/2 ML (blue) hydroxyls (O_b-H). (E) Contributions of various d suborbitals to the projected DOS of Ti³⁺ 3d orbitals in TiO₂(110) with 1/8 ML O_b-H. The cyan and light green bars highlight the computed band gap state and excited resonance, respectively. (F) Crystal field splitting of Ti 3d orbitals on TiO₂(110). The positions of the energy levels have been drawn at approximately the centers of gravity of the corresponding projected DOS (see Figure S4). The experimentally observed excited resonance at 2.5 ± 0.2 eV originates from the transitions indicated by the full vertical green line, with a smaller contribution from the transitions indicated by the dashed line as well. The energy zero is at the Fermi energy (E_F) which is taken coincident with the bottom of the conduction band.

annealing, s1 we rather vary the density of surface bridging hydroxyls through photocatalyzed splitting of methanol s1,31 which is a much milder method to introduce point defects (see details in the Experimental and Theoretical Details section). Figure 2B shows 2PPE spectra of $h\text{-TiO}_2(110)$ as a function of $O_{br}H$ coverage. Feature F increases with the $O_{br}H$ density, leading to a linear relation between the integrated signal from the excited resonance and the $O_{br}H$ density. Since both oxygen vacancies and bridging hydroxyls are directly related to the reduction of the TiO_2 surface, this linear relationship suggests that the excited state is associated with Ti^{3+} species. Such a conclusion is also supported by previous inverse photoemission

spectroscopy (IPS) studies, which have shown that the density of states (DOS) around 3 eV above $E_{\rm F}$ increases dramatically with the exposure of ${\rm TiO_2(110)}$ to Ar ions. Since Ar ion sputtering preferentially removes oxygen, the changes of DOS above $E_{\rm F}$ in the IPS spectra are related to the reduction of Ti ions, consistent with our results.

We further explored the character of the intermediate state observed in 2PPE by DFT hybrid functional calculations, focusing on the Ti-derived states near $E_{\rm F}$ (here assumed to be at the CB minimum). Figure 3A and 3B show Ti-projected densities of states (PDOSs) for defect-free and 1/8 ML $O_{\rm br}H$ -covered $TiO_2(110)$ obtained using the HSE06 functional, ³⁷

Journal of the American Chemical Society

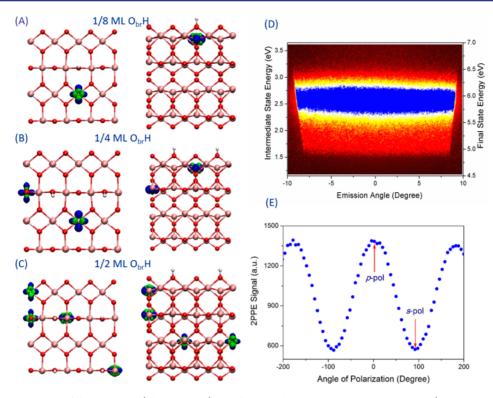


Figure 4. (A–C) Spin densities of the gap states (blue contours) and the excited resonant states at 2.5 ± 0.2 eV (green contours) on h-TiO₂(110) with different coverages of bridging hydroxyls: 1/8 (A), 1/4 (B) and 1/2 ML (C), respectively. The pink, red, and white spheres represent titanium, oxygen, and hydrogen atoms, respectively. Left panel: top-view, along [110]; right panel: side-view, along [110]. (D) Angle-resolved p-2PPE spectra for the h-TiO₂(110) surface with 8.4% ML O_b.H. The excitation photon energy is 3.35 eV. There is little angular dispersion of the excited state, suggesting the state is localized. (E) Polarization dependence of the 2PPE excited resonance signal.

which is known to provide an accurate description of the electronic structure of a wide range of materials. In the PDOS of Ti⁴⁺ cations on the stoichiometric surface (Figure 3A) we can notice two groups of electronic states, corresponding to the Ti 3d t_{2g}- and e_g-like states found in bulk TiO₂^{54,55} (see also below). A very different shape is observed instead for the PDOS of a Ti³⁺ center on the hydroxylated surface (Figure 3B). Besides the gap state at ~ 1 eV below $E_{\rm F}$, extensively discussed in previous studies, ^{13,46,56} the latter shows a reduced number of states in the range 0-2 eV and a clear enhancement above ~ 2.3 eV in comparison to the PDOS of a Ti⁴⁺ ion. Except for some small differences in the energy positions, these general features of the PDOS of Ti3+ ions are confirmed by calculations performed using B3LYP,39 a hybrid functional employed in several previous studies of defects in TiO₂⁴⁶ (Figure 3C). This indicates that the presence of Ti³⁺ derived resonant states in the range $\sim 2-3$ eV is a robust feature of the electronic structure of h-TiO₂(110) that does not depend on the hybrid functional used for the calculations. Furthermore, computed PDOS curves for Ti³⁺ ions on reduced TiO₂(110) surfaces with oxygen vacancies and Ti interstitials are qualitatively similar to those found for Ti^{3+} on h- $TiO_2(110)$ (Figure S3), in agreement with the 2PPE observation that this resonance is an inherent characteristics of Ti³⁺ species that does not depend on the presence of adsorbates. As the coverage of O_{br}H increases, both the band gap state and the excited resonance become more intense (Figure 3D), which is also in accord with the 2PPE

It is interesting to consider the contribution of the different d orbitals to the PDOS of Ti^{3+} ions (Figures 3E, S4, and S5). In bulk TiO_2 , the Ti ions occupy slightly distorted octahedral sites;

thus, their 3d orbitals split into two distinct groups, a lower group of three orbitals (d_{xy}, d_{yz}) and d_{xz} with t_{2g} -like symmetry and an upper group of two orbitals $(d_{x^2-y^2}$ and $d_{z^2})$ with e_g -like symmetry. 54 In stoichiometric TiO₂, the t_{2g}-like orbitals contribute mainly to the conduction band minimum. On hydroxylated or reduced $TiO_2(110)$, the t_{2g}^{-1} configuration of Ti³⁺ ions leads to a Jahn-Teller distortion, which further splits the three t_{2g} orbitals into a low energy state with d_{xy} character and two high energy orbitals with d_{yz} and d_{xz} character (Figure 3F). Excess electrons from reducing defects fill the d_{xy} orbitals giving rise to the occupied polaronic state in the band gap, 46 while the d_{vz} and d_{xz} orbitals move to higher energy and give rise to the excited resonance at $\sim 2-3$ eV above $E_{\rm F}$. Plots of the gap and excited states on hydroxylated TiO₂(110) surfaces with different ObrH coverages are reported in Figure 4A-C (here the states in the energy window 2.5 ± 0.2 eV were selected). The excited state appears to maintain a d-like character and remain fairly localized despite being resonant with the TiO₂ CB. The latter result is well supported by our angle-resolved 2PPE measurements (Figure 4D).

Based on the above analysis, the feature F observed in 2PPE should be attributed to localized excitation of Ti^{3+} ions via $d \to d$ transitions from the gap state to the unoccupied excited state. On $\mathrm{TiO_2}(110)$ the three main symmetry directions are [110], [1 $\overline{10}$], and [001]. The latter corresponds to the direction of the strands of edge-sharing (slightly distorted) $\mathrm{TiO_6}$ octahedra in the rutile structure, while [110] and [1 $\overline{10}$] correspond to the directions of the main axes of the $\mathrm{TiO_6}$ octahedra along the strands. This suggests that the intra-atomic $\mathrm{Ti}^{3+} d \to d$ transitions should be easily excited by polarized light with the electric field parallel to [110] or [1 $\overline{10}$], whereas these

transitions should be forbidden or much weaker when the field is parallel to [001]. Our polarization dependent 2PPE results agree well with this prediction. In our measurements, the $[1\overline{1}0]$ direction is in the incidence plane; thus, the components of the electric field of p-polarized light are along [110] and [$1\overline{10}$], while those of s-polarized light are along [001] (Figure S6). From Figure 4E one can see that as the polarization is varied gradually from s to p, the signal of peak F increases monotonically. In fact, the signal changes periodically with the polarization, with maxima and minima at p- and spolarization, respectively. Since the profiles of s-2PPE and 1PPE are nearly identical (Figure 2A), the signal in s-2PPE comes from the band gap states at $\sim 0.8-1.0$ eV below $E_{\rm F}$ without populating the excited state at \sim 2.5 eV above $E_{\rm F}$; in other words, s-polarized light is "blind" to the excited state. Only by using polarized light with the electric field along [110] and/or [110] can we detect the excited state, consistent with our symmetry analysis.

SUMMARY

In summary, in this work we have identified an excited resonant state centered at 2.5 \pm 0.2 eV above $E_{\rm F}$ associated with Ti³⁺ species in reduced TiO2. By DFT calculations, we have shown that this excited state is closely related to the band gap state, as they both result from the Jahn-Teller-induced splitting of the 3d orbitals of Ti³⁺ ions. The $d \rightarrow d$ transitions from the Ti³⁺ related occupied band gap state at $\sim -(0.8-1.0)$ eV to the empty excited state at ~2.5 eV increase the photoabsorption in the ultraviolet (UV) region. At the same time, the relative wide distribution of these two states also extends the absorbance to the visible range (Figure 5). This extended photoabsorption together with the near-infrared absorption produce an absorption minimum in the blue, which is the reason why reduced TiO₂ samples appear as blue. ¹⁴ In addition, the $d \rightarrow d$ transitions that we have identified explain the underlying mechanism of the enhanced and extended absorption of reduced TiO2 in the visible region, leading to a better



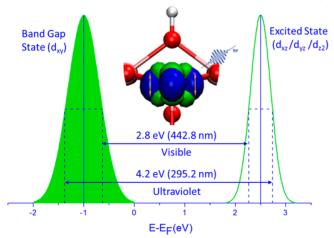


Figure 5. Schematic of the enhanced and extended photoabsorption via $3d \rightarrow 3d$ transitions from the band gap state to the excited resonant state at 2.5 ± 0.2 eV in reduced ${\rm TiO}_2$ (see Figure 3F for details). Blue and green charge density isosurfaces in the middle panel represent the band gap and excited states, respectively. The experimentally measured fwhm (full width at half maxima) of the band gap state and the excited state are $\sim 0.8 \, {\rm eV}^{12,47}$ and $\sim 0.5 \, {\rm eV}_1$, respectively.

understanding of the enhanced visible light photocatalytic activity through ${\rm Ti}^{3+}$ self-doping. ^{15–19,28} Our work also paves the way for the investigation of the electronic structure and photoabsorption of other metal oxides.

ASSOCIATED CONTENT

Supporting Information

Schematic of one-photon photoemission (1PPE) and two-photon photoemission (2PPE) from TiO_2 . Contribution of $\text{O}_{\text{br}}\text{H}$ and $\text{CH}_3\text{O}_{\text{br}}$ to the excited state. Projected DOS of Ti^{3+} 3d orbitals in $\text{TiO}_2(110)$ slabs containing Ti interstitials or oxygen vacancies calculated with the HSE06 functional. Contributions of different d orbitals to the projected DOS of Ti^{3+} 3d states in 1/8 ML $\text{O}_{\text{br}}\text{H}$ covered $\text{TiO}_2(110)$ and Ti^{4+} 3d states in stoichiometric $\text{TiO}_2(110)$. Contributions of different d orbitals to the projected DOS of Ti^{3+} 3d states on $\text{TiO}_2(110)$ covered with 1/4 ML and 1/2 ML bridging hydroxyls. Schematic overview of the experimental geometry. The Supporting Information is available free of charge on the ACS Publications website at DOI: 10.1021/jacs.5b04483.

AUTHOR INFORMATION

Corresponding Authors

*limin.liu@csrc.ac.cn

*chuanyaozhou@dicp.ac.cn

*aselloni@princeton.edu

*xmyang@dicp.ac.cn

Author Contributions

[#]Z.W., B.W., and Q.H. contributed equally.

Notes

The authors declare no competing financial interest.

ACKNOWLEDGMENTS

We thank Ulrike Diebold and Andrea Vittadini for a critical reading of our manuscript. C.Z., L.L., and X.Y. acknowledge the financial support from the National Natural Science Foundation of China (Grant Nos. 21203189, 21321091, 51222212), the Ministry of Science and Technology of China (Grant Nos. 2013CB834605, 2011CB922200), the Key Research Program of the Chinese Academy of Science (Grant No. KGZD-EW-T05), and the State Key Laboratory of Molecular Reaction Dynamics (Grant No. ZZ-2014-02). A.S. acknowledges the support of DoE-BES, Division of Chemical Sciences, Geosciences and Biosciences under Award DE-FG02-12ER16286.

REFERENCES

- (1) Linsebigler, A. L.; Lu, G. Q.; Yates, J. T. Chem. Rev. 1995, 95, 735.
- (2) Gratzel, M. Nature 2001, 414, 338.
- (3) Diebold, U. Surf. Sci. Rep. 2003, 48, 53.
- (4) Fujishima, A.; Zhang, X.; Tryk, D. A. Surf. Sci. Rep. 2008, 63, 515.
- (5) Henderson, M. A. Surf. Sci. Rep. 2011, 66, 185.
- (6) Schneider, J.; Matsuoka, M.; Takeuchi, M.; Zhang, J.; Horiuchi, Y.; Anpo, M.; Bahnemann, D. W. Chem. Rev. 2014, 114, 9919.
- (7) Ma, Y.; Wang, X.; Jia, Y.; Chen, X.; Han, H.; Li, C. Chem. Rev. 2014, 114, 9987.
- (8) Bai, Y.; Mora-Seró, I.; De Angelis, F.; Bisquert, J.; Wang, P. Chem. Rev. 2014, 114, 10095.
- (9) Thompson, T. L.; Yates, J. T. Chem. Rev. 2006, 106, 4428.
- (10) Henrich, V. E.; Dresselhaus, G.; Zeiger, H. J. Phys. Rev. Lett. 1976, 36, 1335.
- (11) Yim, C. M.; Pang, C. L.; Thornton, G. Phys. Rev. Lett. 2010, 104, 036806.

- (12) Wendt, S.; Sprunger, P. T.; Lira, E.; Madsen, G. K. H.; Li, Z. S.; Hansen, J. O.; Matthiesen, J.; Blekinge-Rasmussen, A.; Laegsgaard, E.; Hammer, B.; Besenbacher, F. *Science* **2008**, 320, 1755.
- (13) Mao, X. C.; Lang, X. F.; Wang, Z. Q.; Hao, Q. Q.; Wen, B.; Ren, Z. F.; Dai, D. X.; Zhou, C. Y.; Liu, L. M.; Yang, X. M. J. Phys. Chem. Lett. 2013, 4, 3839.
- (14) Khomenko, V. M.; Langer, K.; Rager, H.; Fett, A. Phys. Chem. Miner. 1998, 25, 338.
- (15) Justicia, I.; Ordejon, P.; Canto, G.; Mozos, J. L.; Fraxedas, J.; Battiston, G. A.; Gerbasi, R.; Figueras, A. Adv. Mater. 2002, 14, 1399.
- (16) Zuo, F.; Wang, L.; Wu, T.; Zhang, Z. Y.; Borchardt, D.; Feng, P. Y. J. Am. Chem. Soc. 2010, 132, 11856.
- (17) Zuo, F.; Bozhilov, K.; Dillon, R. J.; Wang, L.; Smith, P.; Zhao, X.; Bardeen, C.; Feng, P. *Angew. Chem., Int. Ed.* **2012**, *51*, 6223.
- (18) Mao, C. Y.; Zuo, F.; Hou, Y.; Bu, X. H.; Feng, P. Y. Angew. Chem., Int. Ed. 2014, 53, 10485.
- (19) Zuo, F.; Wang, L.; Feng, P. Y. Int. J. Hydrogen Energy 2014, 39, 711.
- (20) Dohnalek, Z.; Lyubinetsky, I.; Rousseau, R. Prog. Surf. Sci. 2010, 85, 161.
- (21) Henderson, M. A.; Lyubinetsky, I. Chem. Rev. 2013, 113, 4428.
- (22) Pang, C. L.; Lindsay, R.; Thornton, G. Chem. Rev. 2013, 113, 3887.
- (23) Rossmeisl, J.; Qu, Z. W.; Zhu, H.; Kroes, G. J.; Norskov, J. K. J. Electroanal. Chem. **2007**, 607, 83.
- (24) Kowalski, P. M.; Camellone, M. F.; Nair, N. N.; Meyer, B.; Marx, D. *Phys. Rev. Lett.* **2010**, *105*, 146405.
- (25) Deskins, N. A.; Rousseau, R.; Dupuis, M. J. Phys. Chem. C 2011, 115, 7562.
- (26) Cheng, J.; Liu, X.; Kattirtzi, J. A.; VandeVondele, J.; Sprik, M. Angew. Chem., Int. Ed. 2014, 53, 12046.
- (27) Cheng, J.; VandeVondele, J.; Sprik, M. J. Phys. Chem. C 2014, 118, 5437.
- (28) Zhou, C.; Ma, Z.; Ren, Z.; Mao, X.; Dai, D.; Yang, X. Chemical Science 2011, 2, 1980.
- (29) Ren, Z. F.; Zhou, C. Y.; Ma, Z. B.; Xiao, C. L.; Mao, X. C.; Dai, D. X.; LaRue, J.; Cooper, R.; Wodtke, A. M.; Yang, X. M. Chin. J. Chem. Phys. **2010**, 23, 255.
- (30) Henderson, M. A. Surf. Sci. Rep. 2002, 46, 1.
- (31) Guo, Q.; Xu, C.; Ren, Z.; Yang, W.; Ma, Z.; Dai, D.; Fan, H.; Minton, T. K.; Yang, X. J. Am. Chem. Soc. 2012, 134, 13366.
- (32) Zhang, Z. R.; Bondarchuk, O.; White, J. M.; Kay, B. D.; Dohnalek, Z. J. Am. Chem. Soc. 2006, 128, 4198.
- (33) Bikondoa, O.; Pang, C. L.; Ithnin, R.; Muryn, C. A.; Onishi, H.; Thornton, G. *Nat. Mater.* **2006**, *5*, 189.
- (34) Henderson, M. A.; Otero-Tapia, S.; Castro, M. E. Faraday Discuss. 1999, 114, 313.
- (35) VandeVondele, J.; Krack, M.; Mohamed, F.; Parrinello, M.; Chassaing, T.; Hutter, J. Comput. Phys. Commun. 2005, 167, 103.
- (36) Heyd, J.; Scuseria, G. E.; Ernzerhof, M. *J. Chem. Phys.* **2003**, *118*, 8207.
- (37) Krukau, A. V.; Vydrov, O. A.; Izmaylov, A. F.; Scuseria, G. E. J. Chem. Phys. **2006**, 125, 224106.
- (38) Guidon, M.; Hutter, J.; VandeVondele, J. J. Chem. Theory Comput. 2009, 5, 3010.
- (39) Becke, A. D. J. Chem. Phys. 1993, 98, 1372.
- (40) Goedecker, S.; Teter, M.; Hutter, J. Phys. Rev. B: Condens. Matter Mater. Phys. 1996, 54, 1703.
- (41) VandeVondele, J.; Hutter, J. J. Chem. Phys. 2007, 127, 114105.
- (42) Onda, K.; Li, B.; Zhao, J.; Jordan, K. D.; Yang, J. L.; Petek, H. Science 2005, 308, 1154.
- (43) Li, B.; Zhao, J.; Onda, K.; Jordan, K. D.; Yang, J. L.; Petek, H. Science 2006, 311, 1436.
- (44) Zhou, C. Y.; Ren, Z. F.; Tan, S. J.; Ma, Z. B.; Mao, X. C.; Dai, D. X.; Fan, H. J.; Yang, X. M.; LaRue, J.; Cooper, R.; Wodtke, A. M.; Wang, Z.; Li, Z. Y.; Wang, B.; Yang, J. L.; Hou, J. G. *Chemical Science* **2010**, *1*, 575.
- (45) Ma, Z.; Guo, Q.; Mao, X.; Ren, Z.; Wang, X.; Xu, C.; Yang, W.; Dai, D.; Zhou, C.; Fan, H.; Yang, X. J. Phys. Chem. C 2013, 117, 10336.

- (46) Di Valentin, C.; Pacchioni, G.; Selloni, A. Phys. Rev. Lett. 2006, 97, 166803.
- (47) Kurtz, R. L.; Stockbauer, R.; Msdey, T. E.; Roman, E.; Desegovia, J. L. Surf. Sci. 1989, 218, 178.
- (48) Zhou, C.; Ma, Z.; Ren, Z.; Wodtke, A. M.; Yang, X. Energy Environ. Sci. 2012, 5, 6833.
- (49) Hugenschmidt, M. B.; Gamble, L.; Campbell, C. T. Surf. Sci. 1994, 302, 329.
- (50) Zhao, J.; Li, B.; Onda, K.; Feng, M.; Petek, H. Chem. Rev. 2006, 106, 4402.
- (51) Henderson, M. A. Surf. Sci. 1995, 343, L1156.
- (52) See, A. K.; Bartynski, R. A. J. Vac. Sci. Technol., A 1992, 10, 2591.
- (53) See, A. K.; Thayer, M.; Bartynski, R. A. Phys. Rev. B: Condens. Matter Mater. Phys. 1993, 47, 13722.
- (54) Glassford, K. M.; Chelikowsky, J. R. Phys. Rev. B: Condens. Matter Mater. Phys. 1992, 46, 1284.
- (55) Labat, F.; Baranek, P.; Domain, C.; Minot, C.; Adamo, C. J. Chem. Phys. 2007, 126, 154703.
- (56) Islam, M. M.; Calatayud, M.; Pacchioni, G. J. Phys. Chem. C 2011, 115, 6809.