## Evidence of a Spiro[2.5]octa-1,4,7-trien-6-one Intermediate in the Solvolysis of $\beta$ -(p-Oxidophenylene)vinyl Bromides

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Summary The exclusive formation of 3-(p-hydroxyphenyl)-2-phenylbenzofuran from the solvolysis of both 2-(p-acetoxyphenyl)-2-(o-methoxyphenyl)-1- and 2-(p-acetoxyphenyl)-1-(o-methoxyphenyl)-2-phenyl vinyl bromide and the formation of a 1·1 mixture of 3-(p-hydroxyphenyl)-2-(o-methoxyphenyl)- and 2-(o-ethoxyphenyl)-3-(p-hydroxyphenyl)-benzofuran from 2-(p-acetoxyphenyl)-2-(o-ethoxyphenyl)-1-(o-methoxyphenyl)vinyl bromide under alkaline conditions provide firm evidence for the

presence of the title compound as the precursor of benzofurans of this type

Although spiro[2 5]octa-4,7-dien-6-one has been suggested as the solvolysis intermediate of  $\beta$ -(p-oxidophenylene)ethyl bromide, solvolysis of the corresponding vinyl derivatives has never been studied. We report evidence of a spiro[2 5]-octa-1,4,7-trien-6-one intermediate in the solvolysis of  $\beta$ -(p-oxidophenylene)vinyl bromide, namely its trapping

intramolecularly with an alkoxy-group located at a suitable position.

Heating of Z-2-(p-acetoxyphenyl)-2-(o-methoxyphenyl)-1-phenylvinyl bromide (1a) was carried out in 50% aqueous ethanol containing 10 mol. equiv. of sodium hydroxide and 5 mol. equiv. of thiophenol† at 130 °C for 13 h in a sealed tube and crystals, m.p. 151—153 °C, isolated. The <sup>1</sup>H n.m.r. spectrum of the obtained product did not show any peak in the region of methoxy-protons. Z-2-(p-Acetoxy-phenyl)-1-(o-methoxyphenyl)-2-phenylvinyl bromide (1b), the isomer of (1a), was similarly heated at 130 °C for 2 h. The <sup>1</sup>H n.m.r. spectrum of its product was identical with that of the product of (1a) and also similar crystals were obtained. The same result was attained from the solvolysis of E-(1b). The structure of the crystals could be assigned as

3-(p-hydroxyphenyl)-2-phenylbenzofuran (**5a**), confirmed by methylation to give 3-(p-methoxyphenyl)-2-phenylbenzofuran (**5b**), m.p. 87—89 °C, which was identified by a mixed melting point determination with an authentic sample. The <sup>1</sup>H n.m.r. spectra indicated that the formation of (**5a**) was quantitative in all cases. Similar solvolysis of Z-2-(p-acetoxyphenyl)-2-(p-ethoxyphenyl)-1-(p-methoxyphenyl) bromide (**1c**) gave a 1:1 mixture of 2-(p-ethoxyphenyl)- and 2-(p-methoxyphenyl)-3-(p-hydroxyphenyl)-benzofuran, (**5c**) and (**5d**).

Hence, the following reaction pathway is suggested (Scheme). First, easy removal of the acetoxy-group of (1) under the above reaction conditions gave the corresponding phenolates (2). The negatively charged vinyl bromides were solvolysed more easily than is usual for triarylvinyl bromides.‡ Exclusive formation of (5a) from (1a) and (1b) and the formation of both (5c) and (5d) from (1c) suggest that there is a common precursor in which intramolecular nucleophilic attack of the alkoxy-group gives (5). The common precursor should be the spiro[2.5]octa-1,4,7-trien-6-one (3).

Further evidence of the presence of (3) is that treatment of 5,7-di-t-butyl-1-(o-methoxyphenyl)-2-phenylspiro[2.5]-octa-1,4,7-trien-6-one (6) with 90% aqueous ethanol containing sodium hydroxide gave exclusively, 3-(3',5'-di-t-butyl-4'-hydroxyphenyl)-2-phenylbenzofuran (7).

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<sup>†</sup> Thiophenol is used as a trapping reagent for  $R^{4+}$  (T. Sonoda, S. Kobayashi, and H. Taniguchi, *Bull. Chem. Soc. Jpn.*, 1976, 49, 2560). The solvolytic reaction of (1a) without thiophenol gave the benzofuran (5a) (>90%) as a main product and small amounts of (1d) and (5b) which are the corresponding methylated compounds of (2a) and (5a).

<sup>‡ (2</sup>a),  $t_{\frac{1}{2}}$  (50% EtOH) 6·7 h at 110 °C, (2b),  $t_{\frac{1}{2}}$  (50% EtOH) 0·42 h at 80 °C, preliminary results. 2,2-Bis(p-methoxyphenyl)-1-phenylvinyl bromide,  $t_{\frac{1}{2}}$  (60% EtOH) 42 h at 140 °C (Z. Rappoport and Y. Houminer, J. Chem. Soc., Perkin Trans. 2, 1973, 1506) and 1,2-bis(p-methoxyphenyl)-1-phenylvinyl bromide (E),  $t_{\frac{1}{2}}$  (80% EtOH) 0·76 h at 120 °C (Z. Rappoport and Y. Apeloig, J. Am. Chem. Soc., 1975, 97, 836).

<sup>&</sup>lt;sup>1</sup> R. Baird and S. Winstein, J. Am. Chem. Soc., 1963, 85, 567.