

Nitrogen–Carbon Bond Formation from N₂ and CO₂ Promoted by a Hafnocene Dinitrogen Complex Yields a Substituted Hydrazine**

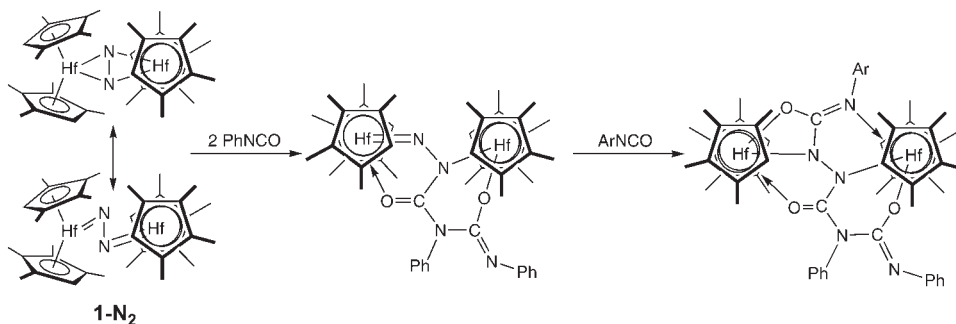
Wesley H. Bernskoetter, Emil Lobkovsky, and Paul J. Chirik*

The discovery of methods for the assembly of nitrogen–carbon bonds from abundant chemical feedstocks is a continuing chemical challenge given the importance of amines and heterocycles in pharmaceutical, agrochemical, and electronic materials applications.^[1] Synthetic ammonia currently serves as the primary nitrogen source for many of these compounds,^[2] and efforts continue to expand its role in synthesis.^[3] Because of the high energy demands associated with the Haber–Bosch process,^[4] direct elaboration of typically inert atmospheric N₂ into more value-added organic molecules is attractive for the evolution of ammonia-independent synthetic pathways.^[5]

Addition of nonpolar reagents such as dihydrogen,^[6] terminal alkynes,^[7] saturated carbon–hydrogen bonds,^[8] and proton donors^[7] to activated early-transition-metal dinitrogen compounds^[9] has emerged as a powerful approach for the assembly of N–H bonds under mild conditions. For cyclopentadienyl-supported zirconium compounds, computational^[10,11] and experimental^[10] studies have established the importance of “imido character” in the metal–nitrogen bonds to facilitate the 1,2-addition of dihydrogen and C–H bonds.

Promoting nitrogen–carbon bond formation using cycloaddition pathways has proven more challenging. Fryzuk and co-worker's report of the addition of terminal alkynes to a side-on bound zirconium dinitrogen compound to form one new N–C bond^[12] has demonstrated the validity of this approach. However in bis(cyclopentadienyl)zirconium chemistry, addition of internal alkynes or aryl isocyanates to

$[(\eta^5\text{-C}_5\text{Me}_4\text{H})_2\text{Zr}]_2(\mu_2, \eta^2, \eta^2\text{-N}_2)$ induced loss of dinitrogen.^[8,13] Because deleterious side-on, end-on isomerization of the $[\text{N}_2]^{4-}$ ligand was believed to be the origin of this unwanted reactivity, the cycloaddition chemistry of the corresponding hafnocene complex, $[(\eta^5\text{-C}_5\text{Me}_4\text{H})_2\text{Hf}]_2(\mu_2, \eta^2, \eta^2\text{-N}_2)$ (**1-N₂**), was explored. Treatment of **1-N₂** with two equivalents of PhNCO yielded a product with a new nitrogen–carbon bond where the nitrogen atom was derived from N₂. The remaining Hf=N linkage in the product is also susceptible to further functionalization as the cycloaddition of terminal alkynes and aryl isocyanates forms N–H and N–C bonds, respectively (Scheme 1).



Scheme 1. N–C bond formation by addition of phenylisocyanate to **1-N₂**.

On the basis of these initial observations, addition of various heterocumulenes to **1-N₂** seemed a viable strategy for the assembly of nitrogen–carbon bonds from coordinated N₂. An alternative bonding description of the four-electron-reduced dinitrogen ligand in **1-N₂** is a bis(amido) fragment coordinated to each hafnium center,^[14] suggesting that insertion of carbon dioxide may be feasible.^[15] Herein, we describe the formation of nitrogen–carbon bonds from coordinated dinitrogen and carbon dioxide^[16] and subsequent removal of the functionalized nitrogen core to form a dicarboxylated silyl-substituted hydrazine.

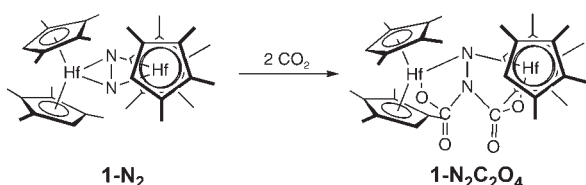
Addition of two equivalents of carbon dioxide to a solution of **1-N₂** in [D₆]benzene or toluene followed by recrystallization from a toluene solution layered with pentane at –35 °C furnished a red solid identified as $[(\eta^5\text{-C}_5\text{Me}_4\text{H})_2\text{Hf}]_2(\text{NCO}_2)_2$ (**1-N₂C₂O₄**), arising from formal insertion of two equivalents of CO₂ into the Hf–N bonds of one side of the $[\text{N}_2]^{4-}$ subunit (Scheme 2).

The ¹H NMR spectrum of **1-N₂C₂O₄** in [D₆]benzene solution exhibits the number of cyclopentadienyl resonances expected for a dimeric molecule with idealized C₂ symmetry. Treatment of the ¹⁵N-labeled dinitrogen complex, **1-¹⁵N₂**,^[14] with ¹³CO₂ gave rise to a doublet centered at $\delta = 155.77$ ppm

[*] W. H. Bernskoetter, Dr. E. Lobkovsky, Prof. P. J. Chirik
Department of Chemistry and Chemical Biology
Cornell University
Ithaca, NY 14853 (USA)
Fax: (+1) 607-255-4137
E-mail: pc92@cornell.edu

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Scheme 2.

in the ^{13}C NMR spectrum, consistent with N–C bond formation from N_2 and CO_2 . Additional support for this assignment of the solution structure was provided by ^{15}N NMR spectroscopy. Exposure of **1- $^{15}\text{N}_2$** to two equivalents of CO_2 (natural abundance) generated **1- $^{15}\text{N}_2\text{C}_2\text{O}_4$** , which exhibited two doublets ($^1J_{\text{N,N}} = 13.1$ Hz) centered at $\delta = 184.0$ and 319.6 ppm for the inequivalent nitrogen atoms. The solid-state (KBr) infrared spectrum of **1-N $_2\text{C}_2\text{O}_4$** contains two strong C=O bands centered at 1728 and 1662 cm^{-1} and a stretch at 1159 cm^{-1} assigned as a C–O band. All three bands shift appropriately (1687 , 1625 , and 1133 cm^{-1} , respectively) upon preparation of **1-N $_2^{13}\text{C}_2\text{O}_4$** from $^{13}\text{CO}_2$.

The identity of **1-N $_2\text{C}_2\text{O}_4$** was also established by single-crystal X-ray diffraction (Figure 1). In the solid state, the two hafnocene subunits are canted with respect to each other, forming a wedge dihedral angle of 87.4° . This angle is defined by the planes formed from the cyclopentadienyl centroids and the hafnium atoms. The $[\text{N}_2\text{C}_2\text{O}_4]^{4-}$ core is comprised of two fused five-membered rings that are conjoined by the N1–N2 bond. For Hf1, the N1 and O1 atoms are contained in the metallocene wedge with the functionalized nitrogen atom N2 dipped below this plane. Accordingly, the wedge for Hf2 contains N1 and O3 with the C38 carboxylate fragment distorted from the wedge.

The N1–N2 bond length of $1.471(7)\text{ \AA}$ and C37–O1 and C38–O3 bond lengths of $1.323(7)$ and $1.303(8)\text{ \AA}$, respectively, are consistent with single bonds. The contracted C37–O2 and C38–O4 bond lengths of $1.208(8)$ and $1.228(8)\text{ \AA}$ suggest localized C=O bonds. The Hf1–N1 and Hf2–N1 distances of $2.137(5)$ and $2.131(5)\text{ \AA}$ are as expected for hafnium amides.

Attempts were made to observe intermediates on the dinitrogen functionalization pathway by adding substoichiometric amounts of carbon dioxide. Treatment of **1-N $_2$** with one equivalent of CO_2 yielded two new organometallic products along with 30% of **1-N $_2\text{C}_2\text{O}_4$** and 40% of the starting dinitrogen compound. The first new product, accounting for approximately 15% of the mixture, was of C_{2v} molecular symmetry (**A**) and exhibited ^{13}C and ^{15}N NMR resonances centered at $\delta = 165.9$ and 309 ppm, respectively, upon isotopic labeling. These resonances, in combination with a coupling constant ($^1J_{\text{N,C}}$) of 12.1 Hz, are consistent with N–C bond formation.

A second new organometallic product (**B**) was also formed in 15% yield and exhibited overall C_2 or C_s molecular symmetry as judged by ^1H NMR spectroscopy. A similar ^{13}C NMR shift of $\delta = 167.3$ ppm and coupling constant $^1J_{\text{C,N}} = 9.9$ Hz were observed upon isotopic labeling and are also consistent with N–C bond formation. Importantly, neither **A** nor **B** converts into **1-N $_2\text{C}_2\text{O}_4$** with addition of CO_2 . Tentative

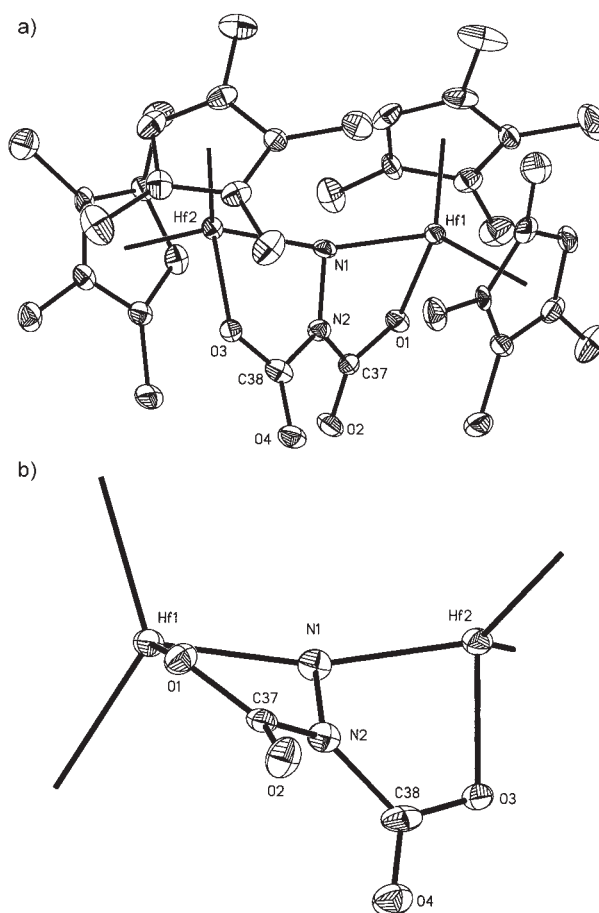


Figure 1. a) ORTEP drawing of the molecular structure of **1-N $_2\text{C}_2\text{O}_4$** with 30% probability ellipsoids. b) A view of the core of the molecule. Hydrogen atoms are omitted for clarity.

structures based on spectroscopic and reactivity data and literature precedent^[15] are presented in Figure 2. Unfortunately, definitive characterization of these compounds was hampered by their instability ($t_{1/2} = 6\text{--}10$ h in $[\text{D}_6]\text{benzene}$) and difficulty in separating them from **1-N $_2\text{C}_2\text{O}_4$** and **1-N $_2$** . By way of comparison, the half-life for **1-N $_2\text{C}_2\text{O}_4$** is approximately 4 days in $[\text{D}_6]\text{benzene}$ at 23°C . Because both **A** and **B** only appear at lower carbon dioxide concentrations, they most likely arise from a unimolecular isomerization process that competes with insertion of a second equivalent of CO_2 during the formation of **1-N $_2\text{C}_2\text{O}_4$** . The analogous reaction, whereby the zirconium dinitrogen complex, $[(\eta^5\text{-C}_5\text{Me}_4\text{H})_2\text{Zr}]_2(\mu_2, \eta^2, \eta^2\text{-N}_2)$, was treated with two equivalents of CO_2 has also been conducted and yields a product distinct from those isolated from **1-N $_2$** . The chemistry of zirconium remains an active area of investigation.

With a hafnocene complex containing a functionalized dinitrogen fragment in hand, attempts were made to remove the $[\text{N}_2\text{C}_2\text{O}_4]^{4-}$ core as a stable organic molecule. Because **1-N $_2\text{C}_2\text{O}_4$** contains potentially robust hafnium–oxygen bonds, strong electrophiles were initially explored. Addition of excess Me_3SiI to solutions of **1-N $_2\text{C}_2\text{O}_4$** in $[\text{D}_6]\text{benzene}$ or toluene produced an immediate change in color from orange

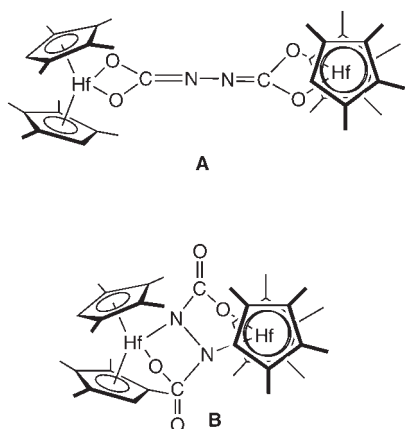


Figure 2. Proposed structures for **A** and **B** obtained from treatment of **1-N₂** with less than two equivalents of CO₂.

to pale yellow. The ¹H NMR spectrum of the kinetic product in [D₆]benzene exhibited C_s molecular symmetry with a new resonance for SiMe₃ centered at δ = 0.44 ppm. Accordingly, a new ²⁹Si NMR peak was observed at δ = 24.41 ppm. Based on these data and an X-ray diffraction study (see Supporting Information), the kinetic product of Me₃SiI addition was identified as the dihafnocene complex bearing a bridging iodide and a silylated carbonyl (Scheme 3). The higher symmetry observed by solution NMR spectroscopy suggests a rapid silicon migration between the carbonyl ligands.^[17]

Further monitoring of the addition of Me₃SiI to **1-I(N₂C₂O₄SiMe₃)** by NMR spectroscopy established near-quantitative formation of (η⁵-C₅Me₄H)₂HfI₂ along with a new organonitrogen species identified as the dicarboxyl silyl-substituted hydrazine, (Me₃Si)₂N-N(CO₂SiMe₃)₂ (**2**; Scheme 4). The ¹H NMR spectrum of **2** exhibits two singlets at δ = 0.24 ppm and 0.39 ppm and the ²⁹Si NMR spectrum reveals resonances at δ = 6.16 and 24.41 ppm. The ¹³C-labeled isotopologue, prepared from **1-N₂**¹³C₂O₄, exhibits a singlet at δ = 157.70 ppm in the ¹³C NMR spectrum. This peak splits into a doublet (¹J_{C,N} = 20.6 Hz) upon additional labeling with

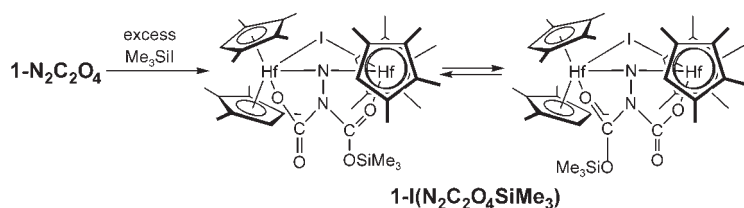
¹⁵N₂. In addition, GC-MS analyses of **2** prepared with both natural abundance and ¹³C-labeled carbon dioxide produced molecular ions with *m/z* 408 and 410, respectively, further confirming its synthesis from N₂ and CO₂. The formation of (η⁵-C₅Me₄H)₂HfI₂ with concomitant release of **2** offers the potential for recycling the hafnium^[18,19] if more compatible routes for the synthesis of **1-N₂** are discovered. Note that optimal procedures for the separation of **2** from (η⁵-C₅Me₄H)₂HfI₂ have not been developed.

In summary, functionalization of a side-on bound hafnocene dinitrogen complex with carbon dioxide has been accomplished. Both multinuclear NMR spectroscopy and X-ray diffraction studies established the identity of the product as arising from formal insertion of two equivalents of CO₂ into an [N₂]⁴⁻ ligand with significant hafnium diamide character. Elaboration of the core of the molecule was accomplished by treatment with Me₃SiI, resulting in the synthesis of a dicarboxylated silylated hydrazine directly from N₂ and CO₂.

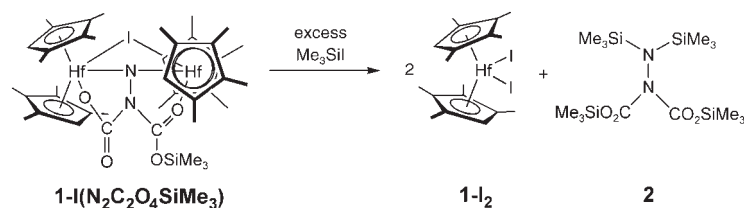
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Scheme 3.



Scheme 4.

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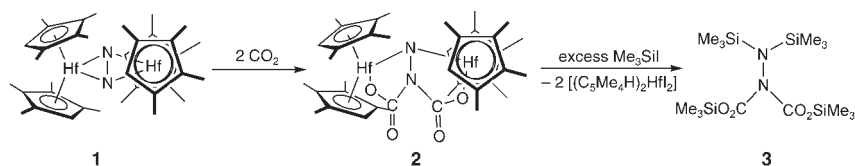
Communications



N–C Bond Formation

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P. J. Chirik*    

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Hydrazine



Hafnocene and hydrazine: The addition of two equivalents of carbon dioxide to the hafnocene dinitrogen complex **1** proved effective for the formation of

nitrogen–carbon bonds directly from N_2 and CO_2 . Subsequent treatment of **2** with Me_3SiH liberated the dicarboxylated silyl-substituted hydrazine **3**.