

See discussions, stats, and author profiles for this publication at: <https://www.researchgate.net/publication/262114868>

ChemInform Abstract: Silver-Catalyzed Decarboxylative Trifluoromethylthiolation of Aliphatic Carboxylic Acids in Aqueous Emulsion.

ARTICLE *in* ANGEWANDTE CHEMIE INTERNATIONAL EDITION · DECEMBER 2014

Impact Factor: 11.26 · DOI: 10.1002/anie.201402573

CITATIONS

37

READS

93

5 AUTHORS, INCLUDING:



Feng Hu

Lamar University

13 PUBLICATIONS 123 CITATIONS

SEE PROFILE



Xinxin Shao

Chinese Academy of Sciences

11 PUBLICATIONS 261 CITATIONS

SEE PROFILE



Qilong Shen

Chinese Academy of Sciences

68 PUBLICATIONS 2,634 CITATIONS

SEE PROFILE

Silver-Catalyzed Decarboxylative Trifluoromethylthiolation of Aliphatic Carboxylic Acids in Aqueous Emulsion**

Feng Hu, Xinxin Shao, Dianhu Zhu, Long Lu, and Qilong Shen*

Abstract: A silver-catalyzed decarboxylative trifluoromethylthiolation of secondary and tertiary carboxylic acids under mild conditions tolerates a wide range of functional groups. The reaction was dramatically accelerated by its performance in an aqueous emulsion, which was formed by the addition of sodium dodecyl sulfate to water. It was proposed that the radical, which was generated from the silver-catalyzed decarboxylation in the “oil-in-water” droplets, could easily react with the trifluoromethylthiolating reagent to form the product.

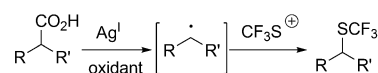
As the most electronegative element, fluorine has played a pivotal role in agrochemistry and pharmaceutical chemistry, with an increasing number of fluorinated drugs on the market and drug candidates under development at different stages.^[1] Among the different fluorinated functional groups, the trifluoromethylthio group (CF₃S) is highly valuable because of its advantageous effects, as the incorporation of the trifluoromethylthio group often improves the lipophilicity of a compound and suppresses metabolic detoxification processes to increase the in vivo lifetime of a drug.^[2] This fact has stimulated a growing interest in the development of new reagents and methods for the direct introduction of the trifluoromethylthio group under mild conditions.^[3] For instance, several electrophilic trifluoromethylthiolating reagents have been developed, which allow the trifluoromethylthiolation of common nucleophiles, such as β-ketoesters, alkynes, aryl or alkyl Grignard reagents, and lithium reagents.^[4,5a–j] In addition, significant advances have been achieved recently in the transition-metal-mediated trifluoromethylthiolation, with functional groups such as aryl boronic acids and halides undergoing efficient trifluoromethylthiolation under palladium, nickel, or copper catalysis.^[5k–p]

Despite these recent great achievements in this area, methods for the general and site-specific formation of sp³-hybridized secondary or tertiary carbon–SCF₃ bonds are

limited.^[7,8] Billard and co-workers reported the electrophilic addition of a trifluoromethylthiolated reagent to olefins in excellent yields.^[5b] Qing and co-workers reported reactions of allylsilane or tryptamine derivatives with Billard’s reagent to form secondary or tertiary alkyl trifluoromethylthioethers.^[5d,f] The groups of Shen and Shibata independently reported the trifluoromethylthiolation of β-ketoesters with two different electrophilic trifluoromethylthiolating reagents.^[6a,b,5i] More recently, the groups of Rueping and Shen reported the asymmetric trifluoromethylthiolation of β-ketoesters.^[5j,6b] Reactions using Billard’s reagent generally required the use of a strong Lewis or Brønsted acid to activate the CF₃S reagent. Other reactions only proceeded in high yields with α-substituted β-ketoesters. Thus, the development of new methods for the direct introduction of the trifluoromethylthio group at sp³-hybridized secondary or tertiary carbon centers with broad substrate scope and functional-group tolerance is highly desirable.

Aliphatic carboxylic acids are air- and water-stable, readily available, and cheap raw materials that can be easily converted into derivatives with different functional groups. One of the fundamental functional-group transformations in organic chemistry is the decarboxylative halogenation of carboxylic acids mediated by a silver salt (Hunsdiecker reaction), which has been well developed since 1930s.^[9] More recently, Hunsdiecker-type reactions using catalytic amounts of silver salts have also been developed.^[10]

Inspired by these advances and considering that the CF₃S group is generally referred to as a pseudohalide, we reasoned that a Hunsdiecker-type decarboxylative trifluoromethylthiolation (Scheme 1) may be possible if the alkyl radical generated from the well-known oxidative decarboxylation could be efficiently trapped by the electrophilic trifluoromethylthiolating reagent **1**, which we recently developed in our lab.^[6]



Scheme 1. Proposed radical trifluoromethylthiolation.

One challenge associated with this approach is the possible oxidation of the product alkyltrifluoromethylthioether by the oxidant that is used to reoxidize the silver(I) salt. To facilitate the formation of the alkyltrifluoromethylthioether and minimize the side reaction, we used a biphasic system that could physically separate the alkyltrifluoromethylthioether from the oxidant, as inspired by the emulsion

[*] Dr. F. Hu,^[+] X. Shao,^[+] D.-H. Zhu, Prof. Dr. L. Lu, Prof. Dr. Q. Shen
Key Laboratory of Organofluorine Chemistry, Shanghai Institute of
Organic Chemistry, Chinese Academy of Sciences
345 Lingling Road, Shanghai 200032 (China)
E-mail: shenql@sioc.ac.cn

[+] These authors contributed equally to this work.

[**] We gratefully acknowledge financial support from the National Basic Research Program of China (2012CB821600), the Key Program of the Natural Science Foundation of China (21032006), the National Natural Science Foundation of China (21172245/2117 and 2244/21372247), and SIOC. We thank Prof. Chaozhong Li for insightful discussions.

Supporting information for this article is available on the WWW under <http://dx.doi.org/10.1002/anie.201402573>.

radical polymerization process.^[11] We herein report the use of an aqueous emulsion, which was formed by the addition of sodium dodecyl sulfate (SDS) in water and dramatically accelerated a silver-catalyzed decarboxylative trifluoromethylthiolation of secondary and tertiary alkyl carboxylic acids under mild conditions. Radical-clock and radical-cyclization experiments suggested that the reaction proceeded through a free-radical process.

In order to test if the alkyl radical could be trapped by reagent **1**, we initially studied the conversion of 1-adamantane carboxylic acid to trifluoromethylthiolated adamantane using reagent **1** in the presence of catalytic amounts of different silver salts (20 mol %) and 1.0 equivalent of K₂S₂O₈ as the oxidant in different solvents.^[12] As expected, the desired product was obtained in less than 12 % yield when silver salts such as AgNO₃, AgOTf, AgBF₄, or AgOAc were used in different solvents, such as THF, CH₃CN, DMF, CH₂Cl₂, and acetone, at room temperature or at 50 °C (results not shown in the table). Furthermore, no significant improvement was observed when the reaction was conducted in a solvent mixture such as CH₃CN/H₂O (1/1; Table 1, entry 2), acetone/H₂O (1/1), or THF/H₂O (1/1). In contrast, the reaction generated the desired product in 60 % yield when it was conducted in the presence of 1.0 equivalent of sodium dodecyl sulfate (SDS) in CH₃CN/H₂O (1/1; Table 1, entry 4). Decreasing the amount of SDS to 0.2 equivalent led to a higher yield (84 %; Table 1, entry 5). Reactions in the

presence of 0.2 equivalent of SDS conducted in acetone/H₂O (1/1) generated the product in 83 % yield, while reactions in other mixed solvents, such as DMF/H₂O, THF/H₂O, or CH₂Cl₂/H₂O, were not effective at all (Table 1, entries 7–9). Interestingly, a higher percentage of water in the mixed solvent led to an increased yield of the product (92 %; Table 1, entry 6). When the reaction was conducted in pure water with 0.2 equivalent of SDS, the desired product was generated in 98 % yield as determined by ¹⁹F NMR spectroscopy with an internal standard (Table 1, entry 10).^[13] The yield decreased to 81 % when 0.1 equivalent of SDS was used (Table 1, entry 11). The use of other surfactants, such as sodium methylsulfonate or sodium 4-dodecylbenzenesulfonate, under otherwise the same conditions resulted in much lower yields (Table 1, entries 12–14). Likewise, the use of other silver salts as the catalyst also led to lower yields (Table 1, entries 15–17). In the absence of the silver salt, no product was observed (Table 1, entry 18).

With an efficient protocol for the decarboxylative trifluoromethylthiolation in hand, we next investigated its scope with aliphatic carboxylic acids (Table 2). The reaction conditions were quite general for tertiary and secondary alkyl carboxylic acids. A wide range of substrates with different functional groups, such as chloride, bromide, esters, and electron-rich arenes, underwent the efficient decarboxylative trifluoromethylthiolation to afford the corresponding products in good to excellent yields (Table 2, **3b–c**, **3j**, **3q–s**). A cyclopropyl carboxylic acid also underwent the decarboxylative trifluoromethylthiolation to give the corresponding product in good yields (Table 2, **3e–f**). Notably, no competing electrophilic trifluoromethylthiolation of electron-rich arenes was observed (Table 2, **3e–f**, **3l**, **3s**). In addition, the reaction was compatible with alkene and alkyne groups, thus illustrating an orthogonal reactivity of the silver-catalyzed decarboxylative trifluoromethylthiolation to classical electrophilic addition reactions (Table 2, **3k–l**). Reactions of primary alkyl carboxylic acids, however, occurred much slower and in lower yield than those of tertiary and secondary alkyl carboxylic acids. The desired products were obtained in only 20–30 % yield in a mixture of CH₃CN/H₂O (Table 2, **3t–u**). The reaction of benzoic acid with reagent **1** under the optimized conditions did not generate the trifluoromethylthiolated product at all. With regard to a possible scale-up of the reaction, the yield of **3a** dropped slightly to 71 % when the reaction was conducted on a 2.0 mmol scale (Table 2).

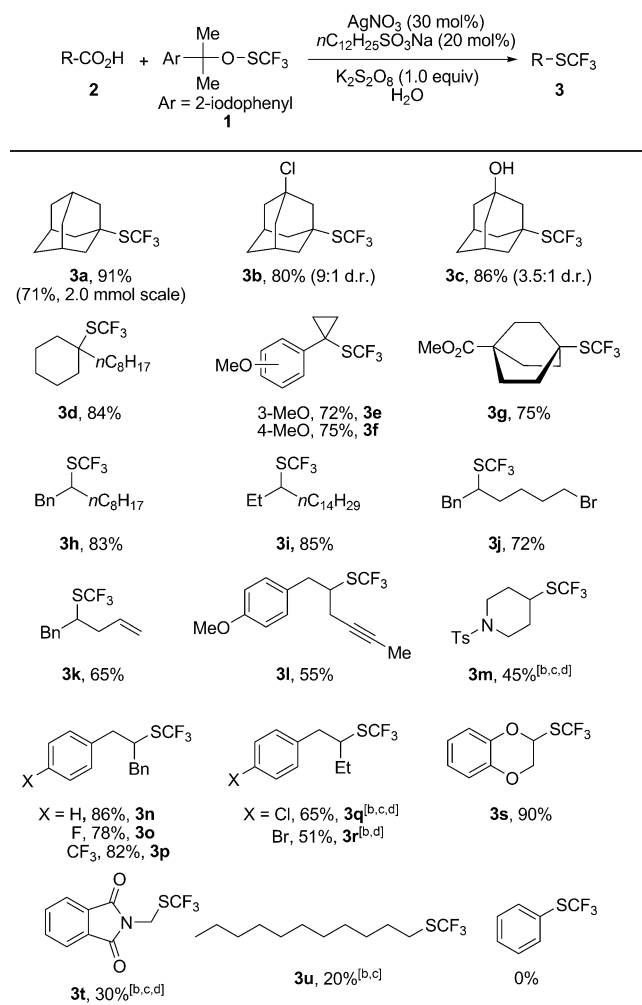
In general, silver-mediated oxidative decarboxylative functionalization was suggested to proceed through a radical pathway. The reactivity of the carboxylic acids in Table 2 decreases in the order of tertiary, secondary > primary >> aromatic, which also indicates that a radical pathway is likely. In order to gain experimental evidence to support this assertion, two sets of experiments were conducted. In the first experiment, cyclopropyl carboxylic acid **4** was prepared as a radical clock and subjected to the reaction conditions. The ring-opening products were generated in 87 % yield as a mixture of stereoisomers in a ratio of 3:1 [Eq. (1)]. No trifluoromethylthiolated cyclopropane derivative was observed by ¹H and ¹⁹F NMR spectroscopy of the crude reaction mixture.

Table 1: Optimization of reaction conditions for silver-catalyzed decarboxylative trifluoromethylthiolation.^[a]

Entry	AgX	Additive	Solvent	Yield [%] ^[a]
1	AgNO ₃	–	CH ₃ CN	11
2	AgNO ₃	–	CH ₃ CN/H ₂ O	12
3	AgNO ₃	–	H ₂ O	< 2
4	AgNO ₃	<i>n</i> C ₁₂ H ₂₅ SO ₃ Na ^[c]	CH ₃ CN/H ₂ O	60
5	AgNO ₃	<i>n</i> C ₁₂ H ₂₅ SO ₃ Na	CH ₃ CN/H ₂ O	84
6	AgNO ₃	<i>n</i> C ₁₂ H ₂₅ SO ₃ Na	CH ₃ CN/H ₂ O ^[d]	92
7	AgNO ₃	<i>n</i> C ₁₂ H ₂₅ SO ₃ Na	acetone/H ₂ O	66
8	AgNO ₃	<i>n</i> C ₁₂ H ₂₅ SO ₃ Na	DMF/H ₂ O	< 2
9	AgNO ₃	<i>n</i> C ₁₂ H ₂₅ SO ₃ Na	CH ₂ Cl ₂ /H ₂ O	< 2
10	AgNO ₃	<i>n</i> C ₁₂ H ₂₅ SO ₃ Na	H ₂ O	98 (91) ^[e]
11	AgNO ₃	<i>n</i> C ₁₂ H ₂₅ SO ₃ Na ^[f]	H ₂ O	81
12	AgNO ₃	CH ₃ SO ₃ Na	H ₂ O	–
13	AgNO ₃	4-(<i>n</i> C ₁₂ H ₂₅)C ₆ H ₄ SO ₃ Na	H ₂ O	70
14	AgNO ₃	<i>n</i> Bu ₄ NHSO ₄	H ₂ O	3
15	AgSbF ₆	<i>n</i> C ₁₂ H ₂₅ SO ₃ Na	H ₂ O	11
16	AgOTf	<i>n</i> C ₁₂ H ₂₅ SO ₃ Na	H ₂ O	71
17	AgOAc	<i>n</i> C ₁₂ H ₂₅ SO ₃ Na	H ₂ O	59
18	–	<i>n</i> C ₁₂ H ₂₅ SO ₃ Na	H ₂ O	–

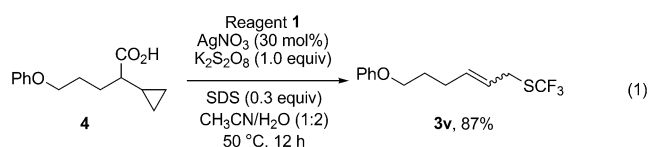
[a] Reaction conditions: 1-adamantane carboxylic acid (0.3 mmol), reagent **1** (0.45 mmol), AgX (30 mol %), additive (20 mol %), K₂S₂O₈ (0.3 mmol), solvent (2.0 mL, 1:1 (v/v) for mixed solvent) at 50 °C for 12 h. [b] Yields were determined by ¹⁹F NMR spectroscopy with 1-fluoronaphthalene as the internal standard. [c] 1.0 equiv of SDS was used. [d] 1:2 (v/v). [e] Yield of isolated product. [f] 10 mol % of SDS was used.

Table 2: Scope of silver-catalyzed decarboxylative trifluoromethylthiolation.^[a]

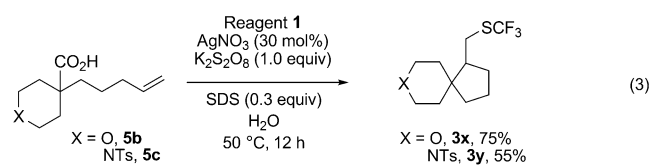
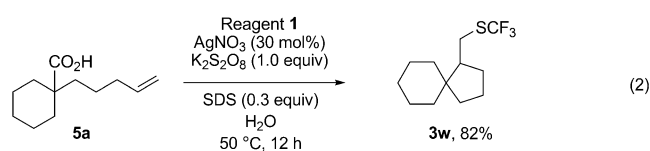


[a] Reaction conditions: alkyl carboxylic acid (0.3 mmol), reagent **1** (0.45 mmol), AgNO₃ (30 mol%), K₂S₂O₈ (0.3 mmol), SDS (20 mol%), H₂O (2.0 mL) at 50 °C for 12 h, yields of isolated products are given. [b] The reaction was conducted in CH₃CN/H₂O (1:2). [c] 24 h. [d] CH₃-(CH₂)₁₁SO₃Na (0.3 mmol) was used.

In the second set of experiments, compounds **5a–c** were synthesized and subjected to the silver-catalyzed reactions with reagent **1** in water to generate the ring-closed trifluoro-



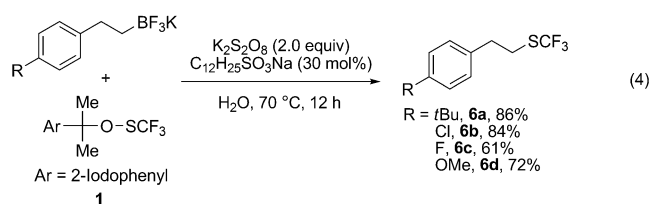
methylthiolated spiro products **3w–y** in 82%, 75%, and 55% yield, respectively [Eq. (2),(3)]. We reasoned that the radicals generated from the silver-mediated oxidative decarboxylation undergo irreversible intramolecular 5-*exo*-cyclization at a much faster rate than those of the bimolecular radical quenching process. Both sets of experiments strongly sug-



gested the involvement of a free-radical mechanism in the silver-catalyzed decarboxylative trifluoromethylthiolation. The experiments further suggested that the formation of primary alkyl radicals from primary alkyl carboxylic acids was difficult under the current conditions.^[10] Once the primary alkyl radical was formed, it could easily react with reagent **1** to form the trifluoromethylthioether.

To probe if a free alkyl radical can easily react with reagent **1** to form alkyl trifluoromethylthioether, we studied the reaction of alkyl trifluoroborate potassium salts with reagent **1** in the absence of a silver catalyst, as it is well-known that alkyl radicals can be easily formed from alkyl trifluoroborates in the presence of an oxidant.^[14] Reactions of four different primary alkyl trifluoroborate salts with reagent **1** formed the corresponding alkyl trifluoromethylthioether in good yields [Eq. (4)]. The reactions occurred in much lower yields in the absence of SDS. These results supported our assertion that free primary alkyl radicals can easily react with reagent **1** to form the products.

To gain some understanding about the effect of the surfactant on the silver-catalyzed decarboxylative trifluoromethylthiolation, we studied the kinetics of the reactions in



the presence of different amount of SDS by monitoring the reaction with ¹⁹F NMR spectroscopy. A comparison of the silver-catalyzed reactions of 1-adamantane carboxylic acid with reagent **1** in the presence of 0, 6, 10 and 20 mol % SDS is shown in Figure 1. The reaction in the absence of SDS occurred sluggishly with less than 15% conversion even after 6 hours. In contrast, the reaction occurred much faster in the presence of SDS. Interestingly, reactions in the presence of 10 mol % SDS or above were significantly faster and occurred in much higher yields than those using 6 mol % SDS. Notably, an emulsion was formed for reactions using 10 or 20 mol % SDS, which indicated the formation of “oil-in-water” droplets^[15] (Figure 1, right). The size of the droplets was deter-

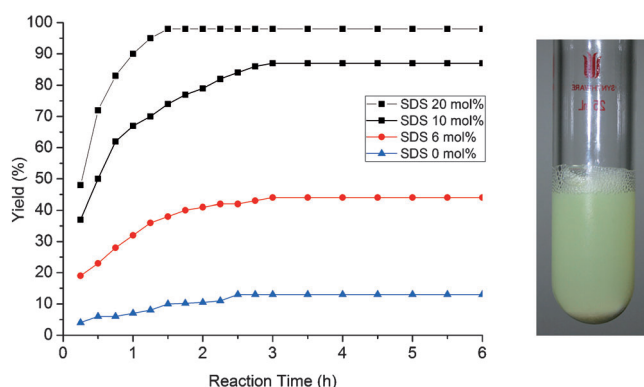
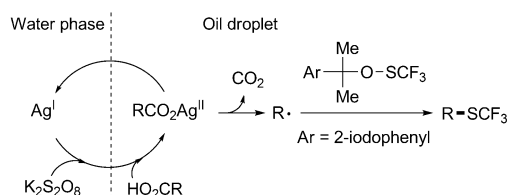


Figure 1. Left: A comparison of the silver-catalyzed decarboxylative trifluoromethylthiolation in the presence of 0, 6, 10, and 20 mol% of SDS. Right: Image of the reaction mixture.

mined to be 230–280 nm by dynamic light scattering (DLS) measurements.

Based on these experimental results, we proposed a working mechanism for the micelle-accelerated, silver-catalyzed decarboxylative trifluoromethylthiolation (Scheme 2). Initially, the organic compounds were emulsified in a solution of SDS and water to form “oil-in-water” droplets. Silver(II) carboxylates, which were generated from the reactions of



Scheme 2. Proposed mechanism.

silver(I) salts with the alkyl carboxylic acids in the presence of $K_2S_2O_8$, could penetrate into the droplets. The alkyl radical, generated from the silver(II) carboxylate through a concerted decarboxylation pathway,^[16] then reacted with reagent **1** to form the corresponding product and a tertiary alkoxy radical. The alkoxy radical decomposed to form 1-(2-iodophenyl)-ethanone. This ketone was isolated in every silver-catalyzed decarboxylative trifluoromethylthiolation reaction, in agreement with the current mechanistic assumption. On the other hand, in the water phase, the silver(I) intermediate was easily reoxidized by $K_2S_2O_8$ to regenerate the active silver(II) catalyst.

In summary, we have developed a silver-catalyzed decarboxylative trifluoromethylthiolation of secondary and tertiary alkyl carboxylic acids under mild conditions. The reaction tolerates a wide range of functional groups. In addition, the reactions were dramatically accelerated by their performance in an aqueous emulsion, which was formed by the addition of sodium dodecyl sulfate (SDS) to water. The use of water as the solvent for organic reactions is advantageous, as it is cheap, nontoxic, and nonflammable, and has extraordinary

solvent effects.^[13,17] Radical-clock and radical-cyclization experiments suggested that the reaction proceeded through a free-radical process. The effect of aqueous emulsions on other reactions is currently investigated in our laboratory.

Received: February 18, 2014

Published online: ■ ■ ■ ■ ■, ■ ■ ■ ■ ■

Keywords: decarboxylation · fluorine · radicals · silver · trifluoromethylthiolation

- [1] a) P. Kirsch, *Modern Fluoroorganic Chemistry: Synthesis, Reactivity, Applications*, Wiley-VCH, Weinheim, **2004**; b) K. Müller, C. Faeh, F. Diederich, *Science* **2007**, *317*, 1881; c) J.-P. Bégue, D. Bonnet-Delpon, *Bioorganic and Medicinal Chemistry of Fluorine*, Wiley, Hoboken, **2008**; d) I. Ojima, *Fluorine in Medicinal Chemistry and Chemical Biology*, Wiley-Blackwell, Chichester, **2009**; e) V. A. Petrov, *Fluorinated Heterocyclic Compounds: Synthesis, Chemistry, and Applications*, Wiley, Hoboken, NJ, **2009**.
- [2] a) A. Leo, C. Hansch, D. Elkins, *Chem. Rev.* **1971**, *71*, 525; b) R. Filler, *Biomedical Aspects of Fluorine Chemistry*, Kodansha, Tokyo, **1982**; c) L. M. Yagupolskii, A. Y. Ilchenko, N. V. Kondratenko, *Russ. Chem. Rev.* **1974**, *43*, 32; d) C. Hansch, A. Leo, R. W. Taft, *Chem. Rev.* **1991**, *91*, 165.
- [3] a) F. Leroux, P. Jeschke, M. Schlosser, *Chem. Rev.* **2005**, *105*, 827; b) B. Manteau, S. Pazenok, J.-P. Vors, F. R. Leroux, *J. Fluorine Chem.* **2010**, *131*, 140; c) V. N. Boiko, *Beilstein J. Org. Chem.* **2010**, *6*, 880; d) A. Tlili, T. Billard, *Angew. Chem.* **2013**, *125*, 6952; *Angew. Chem. Int. Ed.* **2013**, *52*, 6818; e) T. Liang, C. N. Neumann, T. Ritter, *Angew. Chem.* **2013**, *125*, 8372; *Angew. Chem. Int. Ed.* **2013**, *52*, 8214; f) I. Kieltch, P. Eisenberger, A. Togni, *Angew. Chem.* **2007**, *119*, 768; *Angew. Chem. Int. Ed.* **2007**, *46*, 754; g) G. Danoun, B. Bayarmagnai, M. F. Grünberg, L. J. Gooßen, *Chem. Sci.* **2014**, *5*, 1312; h) F. Toulgoat, S. Alazet, T. Billard, *Eur. J. Org. Chem.* **2014**, *12*, 2415.
- [4] Selected examples for electrophilic trifluoromethylthiolation using CF_3SCl : a) L. M. Yagupolskii, N. V. Kondratenko, G. N. Timofeeva, *J. Org. Chem. USSR* **1984**, *20*, 103; b) A. Haas, G. Möller, *Chem. Ber.* **1996**, *129*, 1383; c) S. Munavalli, D. K. Rohrbaugh, F. R. Longo, H. D. Durst, *Phosphorus Sulfur Silicon Relat. Elem.* **2002**, *177*, 1073; d) S. Munavalli, D. K. Rohrbaugh, D. I. Rossman, W. G. Wagner, H. D. Durst, *Phosphorus Sulfur Silicon Relat. Elem.* **2002**, *177*, 1021; e) A. Haas, D.-Y. Oh, *Chem. Ber.* **1969**, *102*, 77.
- [5] a) A. L. Ferry, T. Billard, B. R. Langlois, E. Bacque, *J. Org. Chem.* **2008**, *73*, 9362; b) A. Ferry, T. Billard, B. R. Langlois, E. Bacque, *Angew. Chem.* **2009**, *121*, 8703; *Angew. Chem. Int. Ed.* **2009**, *48*, 8551; c) A. Ferry, T. Billard, E. Bacque, B. R. Langlois, *J. Fluorine Chem.* **2012**, *134*, 160; d) Y. Yang, X.-L. Jiang, F.-L. Qing, *J. Org. Chem.* **2012**, *77*, 7538; e) F. Baert, J. Colomb, T. Billard, *Angew. Chem.* **2012**, *124*, 10528; *Angew. Chem. Int. Ed.* **2012**, *51*, 10382; f) J. Liu, L. L. Chu, F.-L. Qing, *Org. Lett.* **2013**, *15*, 894; g) S. Alazet, L. Zimmer, T. Billard, *Angew. Chem.* **2013**, *125*, 11014; *Angew. Chem. Int. Ed.* **2013**, *52*, 10814; h) Q. Xiao, J. Sheng, Z. Chen, J. Wu, *Chem. Commun.* **2013**, *49*, 8647; i) Y.-D. Yang, A. Azuma, E. Tokunaga, M. Yamasaki, M. Shiro, N. Shibata, *J. Am. Chem. Soc.* **2013**, *135*, 8782; j) T. Bootwicha, X. Liu, R. Pluta, I. Atodiresi, M. Rueping, *Angew. Chem.* **2013**, *125*, 13093; *Angew. Chem. Int. Ed.* **2013**, *52*, 12856; k) G. Teverovskiy, D. S. Surry, S. L. Buchwald, *Angew. Chem.* **2011**, *123*, 7450; *Angew. Chem. Int. Ed.* **2011**, *50*, 7312; l) C.-P. Zhang, D. A. Vici, *J. Am. Chem. Soc.* **2012**, *134*, 183; m) C. Chen, L.-L. Chu, F.-L. Qing, *J. Am. Chem. Soc.* **2012**, *134*, 12454; n) C. Chen, Y. Xie, L.-L. Chu, R.-W. Wang, X.-G. Zhang, F.-L. Qing, *Angew.*

- Chem.* **2012**, *124*, 2542; *Angew. Chem. Int. Ed.* **2012**, *51*, 2492; o) C.-P. Zhang, D. A. Vicic, *Chem. Asian J.* **2012**, *7*, 1756; p) Z. Weng, W. He, C. Chen, R. Lee, D. Dan, Z. Lai, D. Kong, Y. Yuan, K.-W. Huang, *Angew. Chem.* **2013**, *125*, 1588; *Angew. Chem. Int. Ed.* **2013**, *52*, 1548.
- [6] The structure of reagent **1** was originally proposed to be trifluoromethylthio-substituted hypervalent iodine. However, it was revised by Buchwald and co-workers to be trifluoromethyl-substituted thioperoxide, based on a combination of spectroscopic techniques, derivatization experiments, and the crystalline sponge method; a) X.-X. Shao, X.-Q. Wang, T. Yang, L. Lu, Q. Shen, *Angew. Chem.* **2013**, *125*, 3541; *Angew. Chem. Int. Ed.* **2013**, *52*, 3457; b) X.-Q. Wang, T. Yang, X. Cheng, Q. Shen, *Angew. Chem.* **2013**, *125*, 13098; *Angew. Chem. Int. Ed.* **2013**, *52*, 12860; c) E. V. Vinogradova, P. Müller, S. L. Buchwald, *Angew. Chem.* **2014**, *126*, 3189.
- [7] Selected examples of nucleophilic trifluoromethylthiolation of activated primary alkyl halides: a) E. H. Man, D. C. Coffman, E. L. Muetterties, *J. Am. Chem. Soc.* **1959**, *81*, 3575; b) J. F. Harris, Jr., *J. Org. Chem.* **1966**, *31*, 931; c) W. A. Sheppard, *Tetrahedron* **1971**, *27*, 945; d) H. E. Borowski, A. Haas, *Chem. Ber.* **1982**, *115*, 533; e) S. Munavalli, D. I. Rossman, D. K. Rohrbaugh, C. P. Ferguson, H. D. Durst, *J. Fluorine Chem.* **1996**, *76*, 7; f) W. Tyrre, D. Naumann, B. Hoge, Y. L. Yagupolskii, *J. Fluorine Chem.* **2003**, *119*, 101; g) D. Kong, Z. Jiang, S. Xin, Z. Bai, Y. Yuan, Z. Weng, *Tetrahedron* **2013**, *69*, 6046.
- [8] Radical trifluoromethylations of dialkyldisulfide or other substrates were reported previously; a) T. Billard, N. Roques, B. R. Langlois, *J. Org. Chem.* **1999**, *64*, 3813; b) C. Pooput, M. Medebielle, Jr., W. R. Dolbier, *Org. Lett.* **2004**, *6*, 301; c) C. Pooput, Jr., W. R. Dolbier, M. Medebielle, *J. Org. Chem.* **2006**, *71*, 3564; d) A. Studer, *Angew. Chem.* **2012**, *124*, 9082; *Angew. Chem. Int. Ed.* **2012**, *51*, 8950; e) Y. Li, A. Studer, *Angew. Chem.* **2012**, *124*, 8345; *Angew. Chem. Int. Ed.* **2012**, *51*, 8221; f) B. Zhang, C. Mück-Lichtenfeld, C. G. Daniliuc, A. Studer, *Angew. Chem.* **2013**, *125*, 10992; *Angew. Chem. Int. Ed.* **2013**, *52*, 10792.
- [9] a) A. Borodine, *Justus Liebigs Ann. Chem.* **1861**, 119, 121; b) H. Hunsdiecker, C. Hunsdiecker, *Ber. Dtsch. Chem. Ges. B* **1942**, *75*, 291; c) R. G. Johnson, R. K. Ingham, *Chem. Rev.* **1956**, *56*, 219; d) C. V. Wilson, *Org. React.* **1957**, *332*; e) D. Crich, *Comp. Org. Synth.* **1991**, *7*, 717.
- [10] a) Z. Wang, L. Zhu, F. Yin, Z. Su, Z. Li, C. Li, *J. Am. Chem. Soc.* **2012**, *134*, 4258; b) F. Yin, Z. Wang, Z. Li, C. Li, *J. Am. Chem. Soc.* **2012**, *134*, 10401; c) Z. Li, L. Song, C. Li, *J. Am. Chem. Soc.* **2013**, *135*, 4640; d) C. Zhang, Z. Li, L. Zhu, L. Yu, Z. Wang, C. Li, *J. Am. Chem. Soc.* **2013**, *135*, 14082; e) M. P. Sibi, Y. Landais, *Angew. Chem.* **2013**, *125*, 3654; *Angew. Chem. Int. Ed.* **2013**, *52*, 3570; f) S. Bhadra, W. I. Dzik, L. Goossen, *J. Am. Chem. Soc.* **2012**, *134*, 9938.
- [11] a) P. A. Lovell, M. S. El-Aasser, *Emulsion Polymerization and Emulsion Polymers*, Wiley, Chichester, **1997**; b) C.-S. Chern, *Principle and Applications of Emulsion Polymerization*, Wiley, Hoboken, **2008**.
- [12] Reviews on silver-mediated oxidative radical reactions: a) M. A. Duncton, *J. Med. Chem. Commun.* **2011**, *2*, 1135; b) C. Punta, F. Minisci, *Trends Heterocycl. Chem.* **2008**, *13*, 1; c) F. Minisci, F. Fontana, E. Vismara, *J. Heterocycl. Chem.* **1990**, *27*, 79; d) F. Minisci, A. Citterio, C. Giordano, *Acc. Chem. Res.* **1983**, *16*, 27; e) F. Minisci, *Synthesis* **1973**, 1.
- [13] a) T. Perchyonok, *Radical Reactions in Aqueous Media*, RSC, Cambridge, **2010**; b) C.-J. Li, *Chem. Rev.* **2005**, *105*, 3095; c) C.-J. Li, T.-H. Chan, *Comprehensive Organic Reactions in Aqueous Media*, Wiley, Hoboken, **2007**.
- [14] C. Ollivier, P. Renaud, *Chem. Rev.* **2001**, *101*, 3415.
- [15] T. F. Tadros, *Applied Surfactants: Principles and Applications*, Wiley-VCH, Weinheim, **2005**.
- [16] J. M. Anderson, J. K. Kochi, *J. Org. Chem.* **1970**, *35*, 986.
- [17] a) A. Lubineau, J. Auge in *Modern Solvents in Organic Synthesis* (Ed.: P. Knochel), Springer, Berlin, **1999**; b) C.-J. Li, *Chem. Rev.* **1993**, *93*, 2023.

Communications

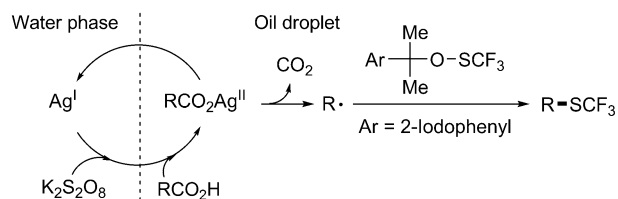


Trifluoromethylthiolation

F. Hu, X. Shao, D.-H. Zhu, L. Lu,
Q. Shen*



Silver-Catalyzed Decarboxylative
Trifluoromethylthiolation of Aliphatic
Carboxylic Acids in Aqueous Emulsion



Oil-in-water: Secondary and tertiary carboxylic acids were successfully transformed to their trifluoromethylthiolated derivatives using the title reaction, which tolerates a wide range of functional

groups. The reaction was dramatically accelerated by the use of an aqueous emulsion as the reaction medium, which was formed by the addition of sodium dodecyl sulfate (SDS) in water.