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# The Role of Acidic Residues and of Sodium Ion Adduction on the Gas-Phase H/D Exchange of Peptides and Peptide Dimers

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Gas-phase H/D exchange is widely used for characterizing the structure of ions. However, many structural parameters that affect the rate of H/D exchange are poorly understood, which complicates the interpretation of experimental data. Here, the effects of sodium ion adduction on the rate of H/D exchange with D<sub>2</sub>O for a series of peptides and peptide dimers with varying numbers of acidic residues are described. The maximum number of sodium ion adducts that can be accommodated by the peptides and peptide dimers in this study is  $N + 1$ , where  $N$  is the number of free carboxylic acid groups. The formation of methyl-esters at all carboxylic acid groups, or the replacement of all the acidic hydrogens with sodium ions, effectively shuts down H/D exchange with D<sub>2</sub>O. In contrast, both the rate and the extent of H/D exchange with D<sub>2</sub>O are increased for most of the peptides and peptide dimers by the adduction of an intermediate number of sodium ions. These results are consistent with the H/D exchange occurring via a salt-bridge mechanism and show that the presence of two carboxylic acid groups is much better than one. The results with peptide dimers also indicate that surface accessibility may not be a dominant factor in the extent of H/D exchange for these ions. (J Am Soc Mass Spectrom 2003, 14, 1477–1487) © 2003 American Society for Mass Spectrometry

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In solution, H/D exchange can be used to obtain information about the solvent accessibility of amino acid residues in a protein [1]. Sites of H/D exchange can be localized to the individual amino acid with NMR or to larger portions of the protein by using proteolytic enzymes (typically pepsin) and mass spectrometric analysis of the products to determine the extent of H/D exchange [2]. From these experiments, information about the tertiary structure of proteins in solution can be deduced.

In the gas phase, the rate and maximal extents of H/D exchange have been used to differentiate protein conformers. In a series of pioneering experiments of McLafferty and coworkers, as many as seven different conformers of gas-phase cytochrome *c* were separated on the basis of their H/D exchange rates and extents of exchange with D<sub>2</sub>O [3–5]. Conformers that exchanged rapidly with D<sub>2</sub>O were designated as “open” conformations due to the relatively high accessibility of exchange reagent to active hydrogen/protonation sites. Smith and coworkers showed that disulfide intact bovine proinsulin and  $\alpha$ -lactalbumin undergo more extensive H/D exchange than their disulfide reduced analogues even though the oxidized proteins are expected to have

a more compact gas-phase conformation [6]. The higher H/D exchange reactivity for proteins with intact disulfide bonds was attributed to effects of higher coulombic repulsion in the constrained proteins. Cassady and Wang studied H/D exchange of three multiply protonated peptides, each having one disulfide bond, and found that the extent of H/D exchange increased, decreased, or stayed the same upon reduction of the disulfide bond, depending on the peptide [7]. Results from molecular modeling suggested that increased H/D exchange is observed for the peptide that has a more extended gas-phase conformation upon the reduction of the disulfide bond whereas decreased exchange is observed for a peptide that has a more compact conformation after disulfide bond reduction. Many other reports of gas-phase H/D exchange of peptides and proteins have been published [8–25]. The ultimate goal of these H/D exchange experiments is to determine information about the structure and conformation of these ions in the gas phase. Other methods to determine information about gas-phase protein conformation include proton-transfer reactivity [26–29] and ion-mobility experiments [30–36].

In order to infer structural information from H/D exchange, it is important to understand the mechanism by which exchange occurs. Fundamental studies by Hunt and coworkers [37] as well as Ausloos and Lias [38] showed that for small protonated compounds, H/D exchange does not occur when the proton affinity

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of the neutral base is greater than the deuterated reagent by more than 20 kcal/mol. However, Lebrilla and coworkers clearly showed that H/D exchange can occur even when the basicity difference is greater than 20 kcal/mol for compounds that have multiple sites of similar basicity [8, 39]. They proposed a mechanism for exchange in which deuterated methanol forms two simultaneous hydrogen bonds with basic groups of the protonated compound, e.g., the carbonyl oxygen of a carboxylic acid group and the protonated amino terminus, forming a bridge between two basic sites.

Beauchamp and coworkers have investigated the H/D exchange reactivity of protonated glycine oligomers with a variety of exchange reagents [11, 12]. They found that the rate and extent of H/D exchange with D<sub>2</sub>O is a strong function of polymer length. They proposed several mechanisms by which exchange might occur. For D<sub>2</sub>O, the preferred mechanism is a "relay" mechanism similar to the mechanism proposed by Lebrilla for exchange with CH<sub>3</sub>OD [8], in which D<sub>2</sub>O is coordinated by both the protonated amino terminus and a carbonyl oxygen. The exchange occurs by shuttling a proton from the protonated amino terminus to the oxygen of D<sub>2</sub>O while simultaneously transferring a deuteron to a carbonyl oxygen of the amide backbone. Although Gly<sub>3</sub> exchanges rapidly with D<sub>2</sub>O, Gly<sub>4</sub> exchanges very slowly. For Gly<sub>4</sub>, the authors suggested that the energy released by the formation of two hydrogen bonds is insufficient to alter the gas-phase conformation of the peptide to allow binding of D<sub>2</sub>O and subsequent H/D exchange to occur.

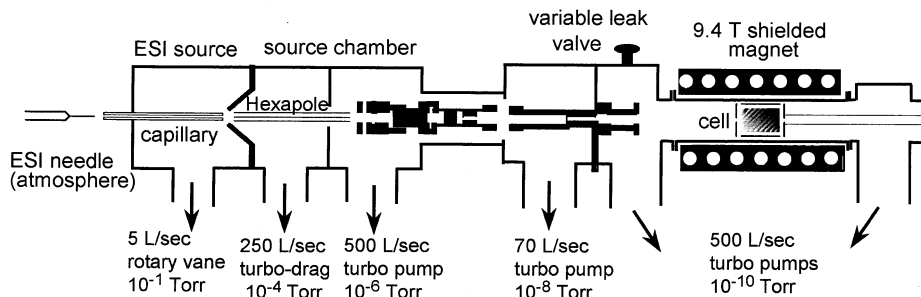
Noncovalent complexes have also been investigated using H/D exchange. Lebrilla and coworkers found that methyl-esters of several amino acids complexed with glycerol and glucose undergo no measurable gas-phase H/D exchange, indicating that a free carboxylic acid is important for H/D exchange with CH<sub>3</sub>OD for these complexes [9]. They also found that complexes of amino acids and monosaccharides undergo less H/D exchange with CH<sub>3</sub>OD than the individual amino acids themselves [9], indicating that the protonated amino acid is extensively coordinated with the monosaccharide and is not able to interact sufficiently with methanol to promote H/D exchange. Beauchamp and coworkers observed a similar result when they measured a decrease in H/D exchange with ND<sub>3</sub> for a series of protonated glycines upon complexation with crown ethers [13]. Liftshitz and coworkers found that doubly protonated dimers of leucine enkephalin exchanged more slowly with ND<sub>3</sub> than singly protonated monomers and attributed this phenomenon to the doubly protonated leucine enkephalin dimer forming a more compact structure [23]. In contrast, Green et al. [10] and Reid et al. [40] working with the exchange of homodimers of butylamine with CH<sub>3</sub>OD and the exchange of dimers of dipeptides with D<sub>2</sub>O respectively, found that dimer formation increased exchange. The authors suggested that dimer formation increases the rate of H/D exchange as the reagent molecule is able to

insert itself into the hydrogen bond linking the dimer and undergo facile exchange with a bridging or relay exchange mechanism [8, 11]. Heck and coworkers also measured an increase in the extent of H/D exchange with ND<sub>3</sub> for a series of vancomycin family antibiotics upon complex formation with their bacterial cell wall mimicking peptides [20]. For example, uncomplexed ristocetin exchanges 40% of its exchangeable hydrogens, but 70% when it is complexed with ac-GA. The authors suggested that the change in H/D exchange is due to changes in the higher order structure of the antibiotic upon complexation of the peptide.

Effects of metal ion adduction on H/D exchange have also been investigated. Cotter and coworkers found that reactions of singly protonated gramicidin S, a cyclic decapeptide, with deuterated ammonia resulted in the exchange of 5–6 out of 13 labile hydrogens [17]. This could be increased to 13 by collisionally activating the ions with off-resonance excitation. However, the singly sodiated ion only exchanged two hydrogens under the same conditions, and heating the sodiated peptide reduced the exchange rate. Protonated Gly<sub>*n*</sub>, *n* = 2–5 undergoes more exchange with ND<sub>3</sub> than do their sodiated counterparts [41]. Similar results were reported by Reyzer and Brodbelt who found that H/D exchange reactions between the alkali metal cationized complexes of noncovalent polyamine complexes and ND<sub>3</sub> are nearly quenched compared to the singly and doubly protonated complexes [42]. In contrast to the results for polyamines, singly sodiated bradykinin ions exchange with D<sub>2</sub>O more than 1000 times faster than singly protonated bradykinin, rapidly exchanging all 17 labile hydrogens [18, 25]. Also, protonated bradykinin undergoes slower H/D exchange than protonated bradykinin methyl-ester [18]. These results indicate that protonated and sodiated peptides containing basic residues, and their methyl-esters, are able to undergo H/D exchange with D<sub>2</sub>O, although the rates of exchange are lower than those reported for short glycine polymers and their analogs [8, 9, 11, 12, 14, 19, 39].

Solouki et al. investigated the effects of one and two metal ion adducts on the H/D exchange rate of Arg-Gly-Asp (RGD) using ND<sub>3</sub> [24]. They found that the rate of exchange decreases in the order H<sup>+</sup> > Cs<sup>+</sup> > K<sup>+</sup> > Na<sup>+</sup>. Ab initio calculations indicate that the sodium ion forms a trivalently coordinated structure with RGD due to its relatively small size and this presumably influences the H/D exchange. However, although the doubly adducted RGD peptides continue to exchange more slowly than the singly protonated peptide, the relative rates of exchange of the doubly adducted metal ions are reversed and follow the order 2Na<sup>+</sup> > 2K<sup>+</sup> > 2Cs<sup>+</sup> [24].

Cassady and coworkers investigated the effect of the position of lysines in the exchange with CH<sub>3</sub>OD and D<sub>2</sub>O of three quadruply protonated dodecapeptides, (KGG)<sub>4</sub>, (K<sub>2</sub>G<sub>4</sub>)<sub>2</sub>, and (K<sub>4</sub>G<sub>8</sub>) [16]. They found that even though molecular modeling calculations indicate that K<sub>4</sub>G<sub>8</sub> has a more compact structure than the other



**Figure 1.** Schematic diagram of the Berkeley-Bruker FTMS instrument with a 110 mm bore magnet. Peptide ions are stored in the UHV cell and undergo exchange reactions with  $D_2O$  introduced by the variable leak valve.

peptides, it exchanges faster and to a greater extent. Molecular modeling indicates that the protonated sites of the several lysine residues of  $K_4G_8$  are in close enough proximity for a molecule of the exchange reagent to form hydrogen bonds with both a protonated lysine residue and a peptide backbone carbonyl oxygen providing a suitable orientation for exchange by a bridge or relay mechanism. This suggests that factors other than the compactness of a gas-phase peptide may play a significant role in determining the rate and extent of H/D exchange.

Here, we investigate the effects of both metal ion adduction and the number of carboxylic acid groups on H/D exchange with  $D_2O$  for a series of peptides having 5–6 amino acids and no basic groups other than the amino terminus. Their methyl-ester derivatives and peptide dimers are also studied. We demonstrate that the extent of sodium adduction is directly related to the number of acidic residues, and that a free carboxylic acid group is critical for H/D exchange with  $D_2O$  in these peptides. We show that multiple metal ion adducts can greatly increase H/D exchange, and we demonstrate that dimers of these peptide ions, which are presumably more globular or compact than the corresponding monomer, do not necessarily exchange more slowly than their corresponding monomers.

## Experimental

### Mass Spectrometry

Experiments were conducted with a 110 mm bore 9.4 T Fourier-transform mass spectrometer that was constructed in collaboration with Bruker Daltonics (Billerica, MA) (Figure 1). This instrument has been described elsewhere [43]. The ion optics and introduction system are a standard Bruker design. Ions are generated by electrospray or nanoelectrospray. This is done using either a Bruker Apollo source electrospray emitter with a syringe pump flow rate of  $1.0 \mu\text{L}/\text{min}$  and nitrogen assisted nebulization, or nanoelectrospray. Nanoelectrospray is performed using borosilicate capillaries that are pulled to a  $\sim 4 \mu\text{m}$  tip with a model P-87 capillary puller (Sutter Instruments, Novato, CA). A small volume of solution (4–10  $\mu\text{L}$ ) is injected into the borosili-

cate capillary and a platinum wire is inserted into the solution at the end of the capillary. The capillary is positioned  $\sim 2 \text{ mm}$  from the source inlet and a potential of  $\sim 900 \text{ V}$  is applied to the platinum wire.

The electrospray generated ions are accumulated in a storage hexapole prior to injection through multiple ion lenses and three stages of differential pumping into the ultra-high vacuum chamber of the instrument where the ions are trapped in a cylindrical Bruker Infinity cell. Each ion injection is accompanied by a pulse of nitrogen gas introduced into the vacuum chamber at a pressure of  $\sim 2 \times 10^{-6} \text{ torr}$  to enhance trapping and to damp the motion of the ions in the cell. All spectra shown are from single detection events. No signal averaging was done.

For H/D exchange studies of peptide dimer ions, the ions were selectively accumulated in the external accumulation hexapole by increasing ion load times from 1 to 3 s, and increasing the hexapole dc offset voltage from  $\sim 2.7 \text{ V}$  to 3.5–4.5 V. Multiple hexapole accumulations were injected into the ion cell prior to H/D exchange.

### Chemicals and Chemical Modifications

$D_2O$ , leucine enkephalin (YGGFL), Ala<sub>5</sub>, Val-Glu-Pro-Ile-Pro-Tyr (VEPIPY) and Phe-Leu-Glu-Glu-Leu (FLEEL) were purchased from Sigma-Aldrich Co. (St. Louis, MO) and were used without further purification. Methanol and hydrochloric acid (37%) were purchased from EM Science (Gibbstown, NJ). Sodium chloride was purchased from Fisher Scientific (Fairlawn, NJ), and potassium chloride was purchased from Mallinckrodt (Paris, KT).

Solutions used for ESI were 1:1 water:MeOH + 2% acetic acid, and the concentration of the peptides was typically  $0.5\text{--}1 \times 10^{-3} \text{ M}$ . High peptide concentrations were used to enhance the formation of noncovalent peptide dimers. To form abundant sodium adducts, sodium chloride was added to a concentration of  $1\text{--}5 \times 10^{-3} \text{ M}$ . For some solutions of YGGFL,  $1 \times 10^{-3} \text{ M}$  potassium chloride was added to compare the effects of sodium and potassium adduction on H/D exchange with  $D_2O$ .

**Table 1.** Average fractions of labile hydrogens exchanged for a series of peptides, peptide homodimers and their sodium adducts after 50 s reaction time at  $1 \times 10^{-6}$  torr  $D_2O$ 

No. of carboxylic acid groups	Peptide	Labile hydrogens (neutral)	Weighted average fraction of hydrogens exchanged							
			Protonated	+ 1Na	+ 2Na	+ 3Na	+ 4Na	+ 5Na	+ 6Na	+ 7Na
1	YGGFL	8	0.09	0.01	0.00					
1	Ala <sub>5</sub>	7	0.14	0.15	0.01					
1	(YGGFL + YGGFL-2ME)	15	0.05	0.01	0.00 <sup>b</sup>					
1	FLEEL-2ME	7	— <sup>a</sup>	— <sup>a</sup>	0.00					
2	(YGGFL) <sub>2</sub>	16	0.01	0.02	0.49	0.04				
2	(Ala <sub>5</sub> ) <sub>2</sub>	14	0.35	0.10	0.28	0.00				
2	VEPIPY	8	0.01	0.55	0.38	0.00				
2	FLEEL-ME	8	0.10	0.18	0.43	0.01				
3	FLEEL	9	0.10	0.51	0.55	0.37	0.00			
4	(VEPIPY) <sub>2</sub>	16	0.02	0.11	0.07	0.58	0.34	0.01		
6	(FLEEL) <sub>2</sub>	18	0.35	0.43	0.55	0.44	0.45	0.55	0.05	0.00

<sup>a</sup>Not included due to overlapping isotopic distributions.<sup>b</sup>One adducted sodium and one adducted potassium ion.

Peptide methyl-esters were synthesized by dissolving the peptide in methanol until the solution was saturated or a concentration of  $2 \times 10^{-3}$  M was reached. Hydrochloric acid (37% vol/vol) was then added at 1% by volume. Complete esterification of the peptides of interest occurred overnight at room temperature. To prepare partially esterified FLEEL, the reaction was allowed to proceed at room temperature for 30 minutes. After reacting with hydrochloric acid for the desired amount of time, the esterification reaction was quenched by the addition of 1/1 vol/vol water.

### H/D Exchange

Gaseous  $D_2O$  is introduced into the vacuum chamber of the instrument through a variable leak valve (Varian, Lexington, MA) for 50 s during which time H/D exchange occurs. The pressure of the  $D_2O$  was varied between  $1 \times 10^{-8}$  and  $1 \times 10^{-6}$  torr, as measured by an uncalibrated ion gauge located  $\sim 1.1$  m from the cell. After 50 s, the leak valve was closed for 50 s to allow the pressure in the vacuum chamber to decrease prior to excitation and broad-band detection so that reasonable signal could be obtained.

## Results

The gas-phase H/D exchange characteristics of a class of peptides, their dimers, their methyl-esters, and all of the sodium adducts that they form are reported. Peptides with a varying number of acidic residues were selected to determine the effects of acidic sites on the rate and extent of H/D exchange with  $D_2O$ . All peptides in this study have only one basic site, (the amino terminus), and from zero to three carboxylic acid groups (listed in Table 1 in the order of increasing carboxylic acid groups). In Table 1, the amount of H/D exchange that each protonated or sodium adducted

species undergoes is expressed as the weighted average fraction of labile hydrogens exchanged for deuteriums after a 50 s reaction time with  $D_2O$  at  $1 \times 10^{-6}$  torr. For example, (YGGFL)<sub>2</sub> with one and two adducted sodium ions has 16 and 15 labile hydrogens, respectively. These ions exchange on average, 2% and 49% of their labile hydrogens, respectively, under these conditions. Peptide ions listed with -N ME represent ions with N carboxylic acids converted to methyl-esters. Blank entries in Table 1 are due to overlapping isotopic distributions of various ions in the mass spectra after the 50 s reaction time, which makes it difficult to determine the weighted average fraction of hydrogens exchanged. For these cases, the general exchange characteristics, i.e., slow or rapid exchange, with the same 50 s reaction time but at lower pressures of  $D_2O$ , are described in specific instances below. Gas-phase H/D exchange studies were also done on protonated and sodium adducted peptides and peptide dimers that were fully methyl-esterified (Table 2). The exchange conditions and the method of reporting the extent of H/D exchange are identical in Tables 1 and 2.

**Table 2.** The average fraction of labile hydrogens exchanged for the peptides and peptide homodimers listed in Table 1 after methyl-esterification at all of their carboxylic acid groups

Peptide	Labile hydrogens	Weighted average fraction of hydrogens exchanged	
		Protonated	1Na
YGGFL-ME	7	0.03	0.01
(YGGFL-ME) <sub>2</sub>	14	0.02	0.00
Ala <sub>5</sub> -ME	6	0.04	0.00
(Ala <sub>5</sub> -ME) <sub>2</sub>	12	0.02	0.02
VEPIPY-2ME	6	0.02	0.01
(VEPIPY-2ME) <sub>2</sub>	12	0.01	0.00
FLEEL-3ME	6	0.04	0.02
(FLEEL-3ME) <sub>2</sub>	12	0.03	0.00

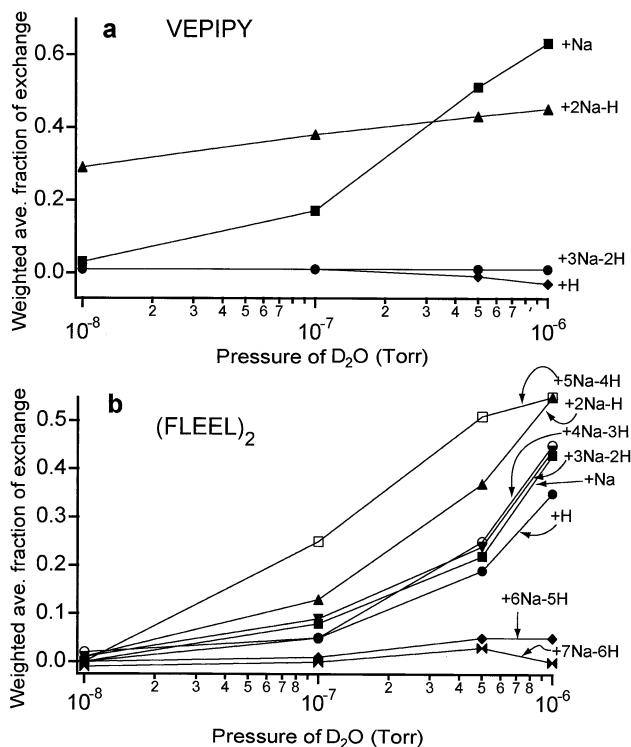
### The Maximum Number of Adducted Sodium Ions

For all of the peptides and peptide dimers in the study, the maximum number of sodium adducts observed in the ESI mass spectra is equal to  $N + 1$ , where  $N$  is the number of carboxylic acid groups in the peptide or peptide dimer. For example, the peptide FLEEL has three carboxylic acid groups and adducts up to four sodium ions. The partially esterified FLEEL peptides adduct a maximum number of sodium ions that decreases by one sodium for each methyl-ester formed at a carboxylic acid. For the peptide YGGFL, potassium adducts were formed as well as sodium adducts by adding potassium chloride to the solutions. In the case of YGGFL and (YGGFL)<sub>2</sub>, sodium and potassium are interchangeable for the purpose of determining maximal addition.

As has been observed previously, the first sodium adduct displaces the proton typically added in the ESI process to form either a charge solvated or salt-bridge form of the adducted peptide [41, 44–48]. Subsequent adducted sodium ions appear to replace carboxylic acid hydrogens and likely form ion pairs with the resulting carboxylate anions. It should be noted that the maximum number of adducted metal ions can be greater than  $N + 1$  for peptides with more basic sites than the amino terminus. For example, adducts with up to four sodium ions have been observed for doubly charged gramicidin S (cyclo[Pro-Val-Orn-Leu-D-Phe]<sub>2</sub>), a peptide that has no acidic groups [49].

### Pressure Dependence of H/D Exchange

The extent of H/D exchange after 50 s of reaction time was measured at four pressures of D<sub>2</sub>O ranging from  $1 \times 10^{-8}$  to  $1 \times 10^{-6}$  torr, for all peptides listed in Tables 1 and 2, to determine trends in the extent of H/D exchange and to study the general exchange characteristics of species that could not be distinguished at  $1 \times 10^{-6}$  torr due to overlapping isotopic distributions. Seven of the peptides and peptide dimers had multiple adducted states (protonated, singly sodiated, doubly sodiated, etc.) that exchanged more than one hydrogen during 50 s exchange at  $\sim 1 \times 10^{-6}$  torr. For all seven of these peptides, the average rate of H/D exchange for the duration of the 50 s changed with the pressure. For example, (VEIPY + 2Na – H)<sup>+</sup> has two fast exchanging hydrogens that exchange during 50 s at  $\sim 1 \times 10^{-8}$  torr, the lowest pressure investigated (Figure 2a). Little additional exchange occurs for (VEIPY + 2Na – H)<sup>+</sup> during 50 s of exchange even when the pressure of D<sub>2</sub>O is two orders of magnitude higher. In contrast, the exchange of (VEIPY + Na)<sup>+</sup> is virtually zero at low pressure, but increases with increasing pressure, exchanging an average of 4.4 hydrogens at  $1 \times 10^{-6}$  torr of D<sub>2</sub>O. The situation becomes increasingly complex with peptides and peptide dimers that have more carboxylic acid groups and thus a greater number of possible sodium adducts. For example, FLEEL dimers



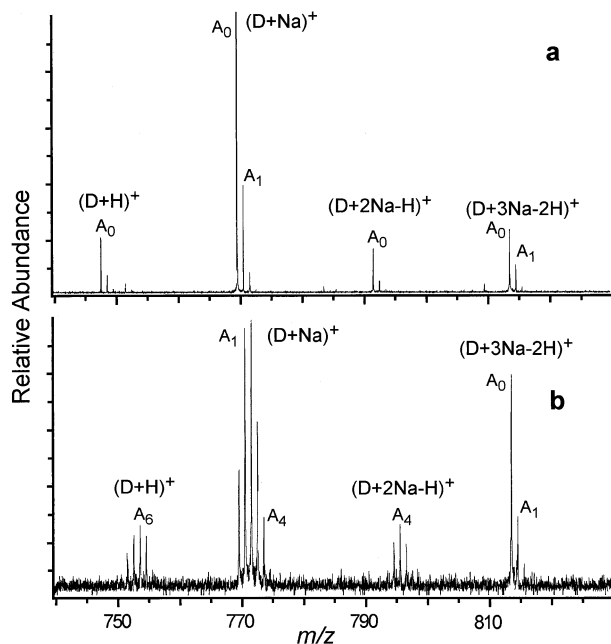
**Figure 2.** Average fraction of labile hydrogens exchanged with D<sub>2</sub>O during a 50 s reaction time at varying pressures for protonated and multiply sodiated adducts of the peptide (a) VEIPY and (b) dimers of the peptide FLEEL.

form between zero and seven sodium ion adducts (Figure 2b). The relative amount of exchange of each species changes greatly with reaction pressure. For example, (FLEEL<sub>2</sub> + 6Na – 5H)<sup>+</sup> and (FLEEL<sub>2</sub> + 4Na – 3H)<sup>+</sup> both exchange an average of  $\sim 1$  hydrogen at  $1 \times 10^{-7}$  torr. However, (FLEEL<sub>2</sub> + 4Na – 3H)<sup>+</sup> exchanges  $\sim 8$  hydrogens at  $1 \times 10^{-6}$  torr while (FLEEL<sub>2</sub> + 6Na – 5H)<sup>+</sup> exchanges an average of less than two.

The fraction of labile hydrogens exchanged for the peptides listed in Table 1 are for a reaction time of 50 s at a pressure of  $1 \times 10^{-6}$  torr. The individual exchange rate constants for each hydrogen exchanged could be obtained by measuring the exchange as a function of reaction time. However, this is beyond the scope of the current study.

### Average Exchange Rates

In order to quantitatively compare H/D exchange results of the peptides listed in Table 1 with previous reports of other peptides [12, 14, 18], the average rate constants of several of the peptides were approximated. This was done by modeling a series of linear, coupled differential equations with identical rate constants, graphing the solutions, and comparing the solutions with the experimental data. Some of the peptide-metal complexes had fast exchanging sites including [VEIPY + 2Na – H]<sup>+</sup> (Figure 2b) and [FLEEL + 3Na – 2H]<sup>+</sup> which each exchanged three hydrogens during 50 s



**Figure 3.** ESI mass spectra of  $(\text{Ala}_5)_2$  formed from in 0.5 mM NaCl and 1:1 water:methanol solution (a) without exchange and (b) after 50 s exchange with  $\text{D}_2\text{O}$  at a pressure of  $\sim 1 \times 10^{-6}$  torr. The notation  $A_0$  represents the monoisotopic peak,  $A_1$  the monoisotopic peak + 1 Da, and so forth.

exchange at  $\sim 1 \times 10^{-8}$  torr corresponding to an average rate constant of  $\sim 2 \times 10^{-10}$   $\text{cm}^3/\text{mol}$ . These are similar rate constants to those measured for the first three exchanges of protonated  $\text{Gly}_2$  and  $\text{Gly}_3$  polymers [12]. Most of the average rate constants approximated for the peptides in Table 1 are two or three orders of magnitude slower. These range from protonated YGGFL, ( $\sim 1 \times 10^{-13}$   $\text{cm}^3/\text{mol}$ ), to  $[\text{Ala}_5\ 2\text{Na} - \text{H}]^+$ , ( $\sim 2 \times 10^{-12}$   $\text{cm}^3/\text{mol}$ ), to  $[\text{FLEEL}_2 + 2\text{Na} - \text{H}]^+$ , ( $\sim 5 \times 10^{-12}$   $\text{cm}^3/\text{mol}$ ). Thus, the average rate constants for the peptide-metal complexes that exchanged the greatest number of hydrogens in Table 1 correspond to the slowest rate constants listed by Beauchamp and coworkers for the exchange of protonated glycine polymers with  $\text{D}_2\text{O}$ .

#### *H/D Exchange of Protonated Peptides and Peptide Dimers*

The protonated monomers of the peptides in this study undergo very little H/D exchange with  $\text{D}_2\text{O}$ , exchanging on average less than one hydrogen. Two proton-bound peptide dimers,  $[(\text{Ala}_5)_2 + \text{H}]^+$  (Figure 3a, b) and  $[(\text{FLEEL})_2 + \text{H}]^+$ , do undergo more rapid H/D exchange, whereas  $[(\text{YGGFL})_2 + \text{H}]^+$  and  $[(\text{VEPIPY})_2 + \text{H}]^+$  do not.

In contrast, none of the protonated peptides or peptide dimers with full or partial methyl-ester formation exchanges more than an average of one hydrogen after 50 s exposure to  $1 \times 10^{-6}$  torr  $\text{D}_2\text{O}$ . A similar observation was reported by Lebrilla and coworkers [9],

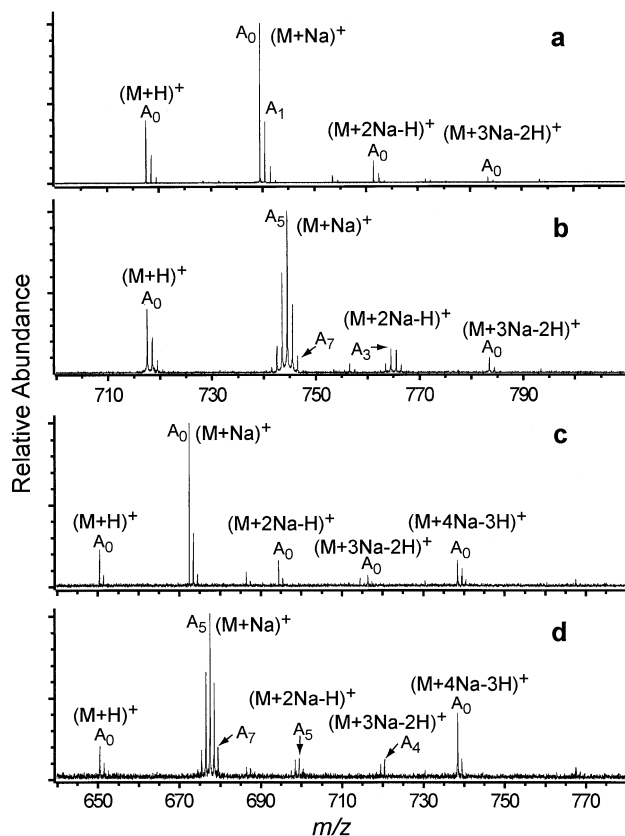
who compared H/D exchange of protonated Gly and Gly-ME in complexes with glycerol. Under exchange conditions where all seven hydrogens in the Gly complexes exchanged with deuterated methanol, no exchange was observed for the Gly-ME complexes [9]. Beauchamp and coworkers observed slow H/D exchange of one hydrogen for  $(\text{Gly}_4 + \text{H})^+$  and  $(\text{Gly}_4\text{-ME} + \text{H})^+$  for reactions up to 20 s at  $1 \times 10^{-7}$  torr  $\text{D}_2\text{O}$  [12]. Interestingly, Beauchamp and coworkers found that  $(\text{Gly}_3\text{-ME} + \text{H})^+$  exchanges all five of its labile hydrogens with  $\text{D}_2\text{O}$  within a reaction time of 20 s at  $\sim 1 \times 10^{-7}$  torr [12]. This suggests that smaller peptides than the ones in this study may exchange relatively rapidly with or without a carboxylic acid group. Methyl-esters of larger peptides with multiple basic sites can also undergo relatively rapid exchange with  $\text{D}_2\text{O}$ . Freitas and Marshall formed singly protonated bradykinin methyl-esters  $(\text{BK-ME} + \text{H})^+$  by collisionally charge stripping  $(\text{BK-ME} + 2\text{H})^{2+}$  (a necessary procedure because only the doubly protonated methyl-ester was initially present) [18]. Because ions formed by collisional charge stripping are initially 'hot', they were allowed to cool by collisions with a pulse of He buffer gas introduced into the vacuum chamber of the instrument prior to H/D exchange. After collisional cooling,  $(\text{BK-ME} + \text{H})^+$  exchanged 12–14 out of 17 exchangeable hydrogens with  $\text{D}_2\text{O}$  in 6 s at  $1 \times 10^{-5}$  torr (a comparable number of collisions as the data presented in Table 1). In contrast,  $(\text{BK-ME} + 2\text{H})^{2+}$  does not exchange any hydrogens with  $\text{D}_2\text{O}$  after 1 hr of reaction at  $1 \times 10^{-5}$  torr [18].

#### *H/D Exchange of Singly Sodiated Peptides and Peptide Dimers*

Peptide monomers with a single sodium adduct can exchange slowly, as was observed for  $(\text{YGGFL} + \text{Na})^+$  and  $(\text{Ala}_5 + \text{Na})^+$ , or rapidly, as was the case for  $(\text{VEPIPY} + \text{Na})^+$  and  $(\text{FLEEL} + \text{Na})^+$  (Figure 4a, b, c, d). The methyl-esterified peptides (Table 2) undergo very little H/D exchange when adducted with a single sodium ion. There are only a few studies on the H/D exchange of noncovalent complexes of metal ions and peptides [17, 18, 24, 25], and  $\text{D}_2\text{O}$  was used as the exchange reagent in only two of these [18, 25]. Freitas and Marshall found that  $(\text{BK} + \text{Na})^+$  exchanges rapidly, incorporating all 17 exchangeable hydrogens within 10 min [18]. In contrast,  $(\text{BK} + \text{H})^+$  exchanges less than one hydrogen during an hour of reaction at  $1 \times 10^{-5}$  torr [18]. Thus, the addition of a single metal ion can significantly increase the H/D exchange of peptides with  $\text{D}_2\text{O}$ .

#### *H/D Exchange of Multiply Sodiated Peptides and Peptide Dimers*

A general trend in the extent of exchange with increasing sodium adduction is observed. The peptides and



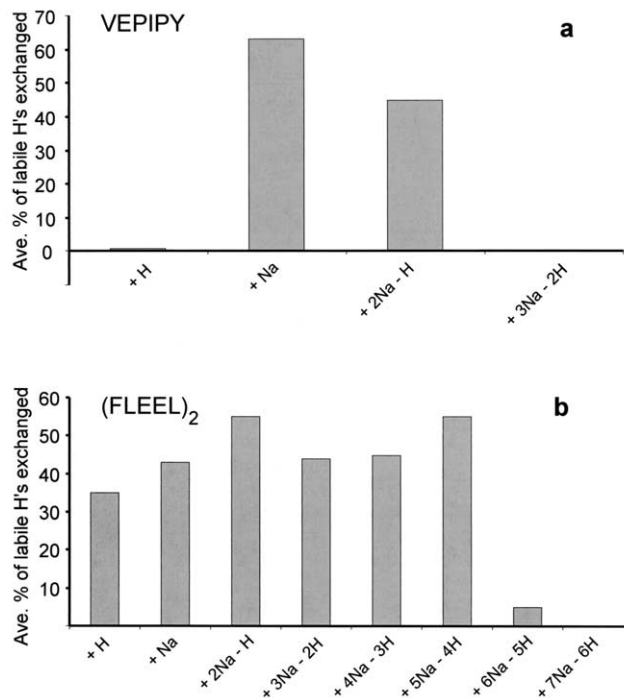
**Figure 4.** ESI mass spectra of VEIPY formed from  $\sim 1$  mM sodium chloride and 1:1 water:methanol solution (a) without exchange and (b) after 50 s exchange with  $D_2O$  at a pressure of  $1 \times 10^{-6}$  torr and FLEEL (c) without exchange and (d) after 50 s exchange with  $D_2O$  at a pressure of  $1 \times 10^{-6}$  torr.

peptide dimers in Table 1 tend to exchange slowly when protonated. With increasing numbers of sodium adducts, exchange increases until maximal exchange is reached, after which, a decrease in exchange occurs. For ions with the maximum number of sodium ions adducted, the exchange is very slow. This trend is shown graphically for VEIPY monomers and FLEEL dimers (Figure 5a, b). The exception is  $[(Ala_5)_2 + H]^+$  which exchanges slightly faster and to a greater extent than  $[(Ala_5)_2 + 2Na - H]^+$  (Figure 3a, b and Table 1).

The peptide and peptide dimers that have the maximum number of adducted sodium ions ( $N + 1$ ) are presumably lacking acidic hydrogens. In order for H/D exchange to occur with these  $N + 1$  adducted ions,  $D_2O$  would need to exchange directly with labile but non-acidic hydrogens on the peptide. Note that no H/D exchange occurs for fully methyl-esterified peptides and peptide dimers (Table 2) indicating that a carboxylic acid hydrogen plays a crucial role in the H/D exchange that occurs in these ions.

#### Sodium and Potassium Adducted $(YGGFL)_2$ Ions

To compare the effects of the adduction of a different alkali metal ion on the rate of H/D exchange,  $(YGGFL)_2$

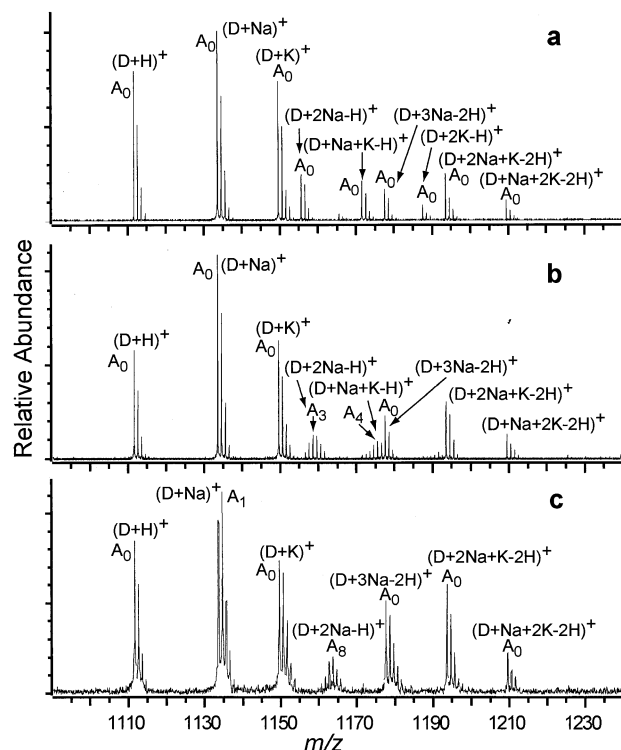


**Figure 5.** Bar graph of the average percent exchange of labile hydrogens after 50 s H/D exchange with  $D_2O$  at a pressure of  $1 \times 10^{-6}$  torr for protonated and all of the sodiated ions that form for (a) VEIPY peptides and (b) FLEEL peptide dimers.

ions were formed from an electrospray solution containing 1 mM of both sodium chloride and potassium chloride. Various peptide-metal ion complexes are readily formed, and H/D exchange was done for 50 s at several different  $D_2O$  pressures (Figure 6a, b, c). Due to overlapping isotopic distributions, it is not possible to determine the fraction of labile hydrogens exchanged for all metal adducted  $(YGGFL)_2$  ions at the highest  $D_2O$  pressure. Nonetheless, for  $(YGGFL)_2$  it is clear that dimers having one adducted metal exchange slowly, any two adducted metal ions exchange rapidly, and any three adducted metal ions exchange slowly. The dimer ion species with three adducted potassium ions was formed in low abundance using our experimental conditions.

#### Partially Esterified FLEEL Exchange

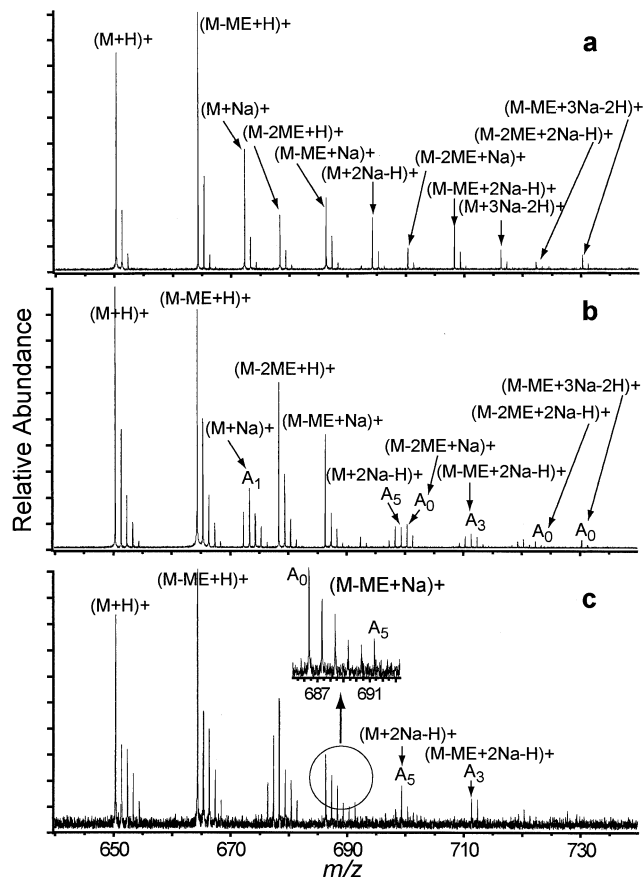
In addition to fully methyl-esterified FLEEL, which has methyl-esters formed at all three carboxylic acid groups, FLEEL was also partially methyl-esterified by allowing the peptide esterification reaction to proceed for 30 min. prior to quenching. Under these conditions, methyl-esters form at zero to two carboxylic acid groups. The partially esterified FLEEL monomers were exchanged with  $D_2O$  in the gas phase to determine the effects of removing some of the carboxylic acid groups on the H/D exchange rates of the peptide-sodium ion complexes. FLEEL, with one of the three carboxylic acid groups methyl-esterified, exchanges similarly to VE-



**Figure 6.** ESI mass spectra of (YGGFL)<sub>2</sub> in 1 mM NaCl and KCl and 1:1 water:methanol (a) without exchange, and after 50 s H/D exchange with D<sub>2</sub>O at a pressure of (b)  $\sim 1 \times 10^{-7}$  torr, and (c)  $\sim 1 \times 10^{-6}$  torr.

PIPY (two carboxylic acid groups). Both exchange slowly in the protonated form, rapidly when one or two sodium ions are added (Figure 7a, b, c), and do not exchange when three sodium ions are added. FLEEL, with methyl-esters at two of the three carboxylic acid positions exchanges similarly to YGGFL (one carboxylic acid group). Monomers exchange slowly in both the protonated form and with one or two adducted sodium ions.

A noticeable feature in the exchange spectrum of partially esterified FLEEL is the appearance of a bimodal isotopic distribution for (FLEEL-ME + Na)<sup>+</sup> (Figure 7c insert). These results clearly indicate the presence of two or more isomers or conformers. The methyl-esterification reaction is relatively non-specific, so that methyl-esterification can occur at any one of the three carboxylic acids, resulting in potentially three isomers. The positional isomers for methyl-esterification of FLEEL-ME do not appear to be resolvable by H/D exchange. Only a single isotopic pattern is measured for (FLEEL-ME + H)<sup>+</sup> and for (FLEEL-ME + 2Na - H)<sup>+</sup>. Each of these ions also has the potential for three isomers. For the latter ion, each isomer has only two free carboxylic acids available to form ion pairs with the two sodium ions. In contrast, each positional isomer for (FLEEL-ME + Na)<sup>+</sup> has two carboxylic acids available to form ion pairs with the adducted sodium ion and thus six different isomers of this ion can be produced. Thus, it seems reasonable that the bimodal isotopic



**Figure 7.** ESI mass spectra of partially esterified FLEEL in 60 mM NaCl and 1:1 water:methanol (a) without exchange and (b) after 50 s exchange with D<sub>2</sub>O at a pressure of (b)  $\sim 1 \times 10^{-7}$  torr and (c)  $\sim 1 \times 10^{-6}$  torr. The inset shows isomeric separation for FLEEL with one methyl-ester and one noncovalently adducted sodium ion.

distribution observed for (FLEEL-ME + Na)<sup>+</sup> results from different positions of sodium adduction.

## Discussion

### H/D Exchange Mechanism

On the basis of experimental results with glycine homopolymers and semi-empirical calculations, Beauchamp and coworkers proposed several mechanisms for gas-phase H/D exchange with D<sub>2</sub>O [12]. The "relay" mechanism [11, 12], which is very similar to a bridging mechanism proposed independently [8], has become a popular way to mechanistically describe and explain H/D exchange data [25, 50]. In the "relay" mechanism, D<sub>2</sub>O coordinates between a protonated site and a remote basic site, forming two hydrogen bonds. H/D exchange occurs by the transfer of a deuteron from the water to the basic site coincident with the transfer of a proton from the protonated site to the water. Intramolecular deuterium-hydrogen rearrangement subsequently distributes deuterium to other sites on the molecule. Alternatively, a "salt-bridge" mecha-



nism was proposed involving proton transfer from a carboxylic acid to D<sub>2</sub>O and the stabilization of the resulting carboxylate anion by a nearby charge site [12]. While not conclusive, our results are most consistent with H/D exchange occurring primarily by a “salt-bridge” mechanism for the small peptide-metal complexes listed in Table 1.

Requisite for exchange by the “salt-bridge” mechanism is the presence of charge sites. For these peptides, we observe the complexation of a maximum of  $N + 1$  metal ions, where  $N$  is the number of carboxylic acid groups in the peptide or peptide dimer. The maximal value of  $N$  is consistent with the first metal ion displacing the proton normally added by the ESI process and then subsequent metal ions displacing carboxylic acid hydrogens and forming ion pairs. Complexation of metal ions provides multiple charge sites that are available to stabilize carboxylate anions formed by proton transfer to D<sub>2</sub>O. For peptides and peptide dimers in Table 1 that have multiple carboxylic acid groups, the stepwise addition of metal ions typically increases the rate of exchange until a maximum is reached and additional metal ions begin to decrease the rate of exchange, presumably due to a scarcity or absence of acidic protons.

While the complexation of metal ions is expected to facilitate salt-bridge formation and tends to increase the rate of H/D exchange, methyl-esterification of the carboxylic acid positions blocks salt-bridge formation and virtually eliminates H/D exchange for the protonated and sodiated peptides studied (Table 2). This result is consistent with H/D exchange by the “salt-bridge” mechanism as there are no carboxylic acids available for proton transfer in the methyl-esterified peptides. The absence of H/D exchange of the peptide methyl-esters with one adducted sodium is also consistent with the “relay” mechanism and other mechanisms that have been proposed [12] as no acidic protons remain for proton transfer to the exchange reagent.

Several distinct groups of H/D exchange rates are measured for these peptide-metal ion complexes, consistent with exchange by multiple mechanisms. Three of the peptides exchange an average of 2–3 hydrogens after 50 s exchange at  $\sim 1 \times 10^{-8}$  torr which corresponds to average exchange rates of  $1\text{--}2 \times 10^{-10}$  cm<sup>3</sup>/mol. These exchange rates are very similar to three fast-exchanging sites measured for protonated Gly<sub>2</sub>, Gly<sub>3</sub>, and Gly<sub>3</sub>-ME that were attributed to exchange by the “relay” mechanism [12]. The rest of the protonated peptide and peptide-metal ion complexes in Table 1 that exchange more than one hydrogen during 50 s at  $\sim 1 \times 10^{-6}$  torr have average rate constants of  $2\text{--}5 \times 10^{-12}$  cm<sup>3</sup>/molecule. These are approximately two orders of magnitude smaller than the rate constants for exchange processes attributed to the “relay” mechanism. However, rate constants on the order of  $2\text{--}5 \times 10^{-12}$  cm<sup>3</sup>/molecule correspond quite well to rate constants measured for single exchanges of betaine, and protonated Gly<sub>1</sub>, Gly<sub>4</sub>, and Gly<sub>5</sub> that were proposed to

exchange by the “salt-bridge” or “flip-flop” mechanisms [12]. Thus, the measured rate constants, the general increase in exchange rate with the adduction of metal ions and the dramatic decrease in exchange rate upon complete methyl-esterification are all consistent with exchange by the “salt-bridge” mechanism. It should be noted that these results do not rule out exchange by other mechanisms or a combination of mechanisms.

### *Cooperative Effects of Multiple Carboxylic Acids*

The observed results in Tables 1 and 2 are well described by placing all of the protonated peptides and peptide dimers and their metal ion complexes into two categories, those having two or more carboxylic acid groups, and those having less than two carboxylic acid groups. All of the peptides and peptide dimers having two or more carboxylic acid groups undergo rapid H/D exchange when they are either protonated or complexed with one or more metal ions. All of the peptides and peptide dimers that have less than two carboxylic acids exchange no more than one hydrogen with D<sub>2</sub>O. This classification holds true for all of the species studied including FLEEL that has been partially methyl-esterified. Protonated or metal ion complexed FLEEL with only one remaining carboxylic acid does not exchange with or without adducted sodium ions, while metal ion complexes of FLEEL, with two remaining carboxylic acids, exchanges multiple hydrogens. More data are necessary to determine if this “two are much better than one” mechanism holds true for H/D exchange with D<sub>2</sub>O for this class of peptides.

### **Conclusions**

The rate and extent of gas-phase H/D exchange with D<sub>2</sub>O have been measured for a series of peptides with 5–6 amino acids residues, their noncovalent homodimers and all of the peptide-sodium ion complexes that they form. All of the peptides studied have a single basic site at the amino terminus and 1–3 acidic sites. A maximum of  $N + 1$  sodium ions form complexes with these peptides and peptide homodimers, where  $N$  is the total number of carboxylic acids. This maximal adduction is consistent with sodium ion displacement of the proton normally added in the electrospray process and all of the acidic carboxylic acid hydrogens. Presumably, the adducted sodium ions form ion pairs with the carboxyl anions. Protonated and sodium adducted peptides having only one carboxylic acid exchange less than one hydrogen during 50 s of exchange at  $1 \times 10^{-6}$  torr. Peptides and peptide dimers with multiple carboxylic acid groups tend to exchange more rapidly when complexed with one or more sodium ions than when protonated. Peptides and peptide dimers with  $N + 1$  adducted sodium ions or with all carboxylic acid groups converted to methyl-esters exchanged on average less than one hydrogen in 50 s at  $1 \times 10^{-6}$  torr. With

the same exposure time and pressure, multiple hydrogens are exchanged for sodium complexes of FLEEL peptides with only one esterified carboxylic acid.

Results of these experiments are consistent with H/D exchange of these peptides occurring, at least in part, by a "salt-bridge" mechanism [12] as sodium ion adduction is seen to generally increase the rate of H/D exchange. The presence of metal ions could be expected to assist in the stabilization of carboxylate anions formed by proton transfer from a carboxylic acid to D<sub>2</sub>O. The average rate constants for the exchange of peptide and peptide-metal ion complexes that undergo multiple exchanges are  $1\text{--}5 \times 10^{-12} \text{ cm}^3/\text{molecule}$  which are very similar to the rate constants for slow exchange of betaine and Gly<sub>4</sub> and Gly<sub>5</sub> that were suggested to exchange by a "salt-bridge" mechanism [12]. Three of the peptides exchanged 2–3 hydrogens with average rate constants two orders of magnitude larger, suggesting that H/D exchange may be occurring by multiple mechanisms.

A general goal of gas-phase H/D exchange is to reveal aspects of gas-phase conformation. The experiments presented here suggest that properties not necessarily related to conformation, such as the inclusion of acidic groups in the primary structure, are very significant in determining the rates of H/D exchange with D<sub>2</sub>O. The adduction of one or more sodium ions also significantly affects the rate of H/D exchange, but the impact of sodium adduction on the conformation of these peptides is not known. Peptide dimers showed exchange that could be either faster or slower than the corresponding monomer despite the fact that the dimers would be expected to have lower surface-accessibility normalized to their mass. Information about the gas-phase conformations of these peptides and their various dimers and adducts could be obtained from ion mobility cross-sectional measurements and molecular modeling [30–32, 41]. Such studies may help to determine the role of conformation on the rate of H/D exchange measured for these peptides and peptide dimers.

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