

Additions and Corrections

1989, Volume 28

Christopher King, Ju-Chun Wang, Md. Nazrul I. Khan, and John P. Fackler, Jr.*: Luminescence and Metal–Metal Interactions in Binuclear Gold(I) Compounds.

Page 2147. In Table II, the λ_{max} value of 571 nm for $[\text{Au}(\text{dppm})]_2(\text{BH}_3\text{CN})_2$ should be 490 nm.

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Runyu Han and Gerard Parkin*: $[\text{Tris}(3\text{-tert-butylpyrazolyl})\text{-hydroborato}]$ beryllium Hydride: Synthesis, Structure, and Reactivity of a Terminal Beryllium Hydride Complex.

Pages 983–988. The X-ray structure of a complex with terminal beryllium–hydride bonds was first reported for the dimer $[(\text{Me}_2\text{NCH}_2\text{CH}_2\text{NMe})\text{BeH}]_2$.¹ We thank Dr. Norman Bell for providing us with this information.

- (1) (a) Bell, N. A.; Coates, G. E.; Schneider, M. L.; Shearer, H. M. M. *J. Chem. Soc., Chem. Commun.* **1983**, 828–829. (b) Bell, N. A.; Coates, G. E.; Schneider, M. L.; Shearer, H. M. M. *Acta Crystallogr.* **1984**, C40, 608–610.

1993, Volume 32

Krishan Kumar,* C. Allen Chang, and M. F. Tweedle: Equilibrium and Kinetic Studies of Lanthanide Complexes of Macrocyclic Polyamino Carboxylates.

Page 591. Equations 8 and 9 should read as follows:

$$-d[\text{LnL}]_T/dt = k_{\text{obsd}}([\text{LnL}] + [\text{LnL}(\text{H})] + [\text{LnL}(\text{H}_2)]) \\ = [k_d + \{k_2 K_1 [\text{H}^+]/(1 + K_1 [\text{H}^+])\}] \times [\text{LnL}(\text{H})] \quad (8)$$

$$k'_{\text{obsd}} = k_{\text{obsd}}\{(1/K_1 [\text{H}^+]) + 1 + (K_1 [\text{H}^+])\} \\ = k_d + \{k_2 K_1 [\text{H}^+]/(1 + K_1 [\text{H}^+])\} \quad (9)$$

Page 592. In Table III, the values of k_d (s^{-1}), K_1 (M^{-1}), and k_2 (s^{-1}), respectively, should read as follows: $(4.0 \pm 1.5) \times 10^{-4}$, 1.3 ± 0.1 , $(1.93 \pm 0.11) \times 10^{-2}$ for $\text{Gd}(\text{DO3A})$; b , 0.35 ± 0.03 , $(1.66 \pm 0.09) \times 10^{-3}$ for $\text{Gd}(\text{HP-DO3A})$.

Tamotsu Sugimori, Kimio Shibakawa, Hideki Masuda, Akira Odani, and Osamu Yamauchi*: Ternary Metal(II) Complexes with Tyrosine-Containing Dipeptides. Structures of Copper(II) and Palladium(II) Complexes Involving L-Tyrosylglycine and Stabilization of Copper(II) Complexes Due to Intramolecular Aromatic Ring Stacking.

Page 4956. The caption for Figure 3 should read as follows: Species distributions as a function of pH in the 1:1:1 Cu(II)–bpy–L-tyr–L-phe system (1 mM). Species: a, Cu(bpy); b, Cu(bpy)(L); c, Cu(bpy)(LH₁); d, Cu(bpy)₂; e, Cu(LH₁); f, Cu(LH₂); g, Cu(bpy)(LH₂); h, Cu(LH₂)(OH); i, Cu₂(LH₁)₂.

Page 4957. The right-hand side of eq 6

$$1 + \frac{1}{K_{\text{st}} + 1}$$

should read as follows:

$$1 - \frac{1}{K_{\text{st}} + 1}$$

Hiroki Oshio,* Etsuo Ino, Iwao Mogi, and Tasuko Ito*: A Weak Antiferromagnetic Interaction between Mn^{2+} Centers through a TCNQ Column: Crystal Structures and Magnetic Properties of $[\text{Mn}^{\text{II}}(\text{tpa})(\text{TCNQ})(\text{CH}_3\text{OH})](\text{TCNQ})_2\text{CH}_3\text{CN}$, $[\text{Mn}^{\text{II}}(\text{tpa})(\mu\text{-O}_2\text{CCH}_3)]_2(\text{TCNQ})_2\cdot 2\text{CH}_3\text{CN}$, and $[\text{Mn}^{\text{II}}(\text{tpa})(\text{NCS})_2]\cdot\text{CH}_3\text{CN}$ (tpa = Tris(2-pyridylmethyl)amine).

Pages 5697–5703. In this paper, TCNQ molecules (**A**, **B**, **C**, **D**) in $[\text{Mn}^{\text{II}}(\text{tpa})(\text{TCNQ})(\text{CH}_3\text{OH})](\text{TCNQ})_2\text{CH}_3\text{CN}$ (**2**) have been assigned as being anionic. The IR of **2** showed three bands (2152 , 2179 , 2184 cm^{-1}) characteristic of $[\text{TCNQ}]^-$ and one band (2216 cm^{-1}) characteristic of $[\text{TCNQ}]^0$. These IR data lead us to conclude that **A** and **B** are monoanionic and that **C** and **D** are neutral. All other conclusions remain unaffected by this change. We thank Prof. Joel S. Miller (University of Utah) for suggesting these IR measurements.

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Younbong Park and John D. Corbett*: $\text{Pr}_{12}\text{I}_{17}\text{Fe}_2$: A Novel Hypostoichiometric Compound with Only Isolated Clusters.

Page 1706. In Table 2, some of the positional parameters and $B(\text{eq})$ values were omitted or misplaced. The corrections are as follows:

atom	x	y	z	$B(\text{eq}), \text{\AA}^2$
14			0.8390(1)	2.58(4)
17		0.82610(8)	0.2541(1)	1.75(3)
18	0.37527(8)	0.2102(1)	0.3298(1)	2.55(4)