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Synthesis, Crystal Structure, and Magnetic Studies of Oxo-Centered Trinuclear Chromium(III) Complexes: $[\text{Cr}_3(\mu_3\text{-O})(\mu_2\text{-PhCOO})_6(\text{H}_2\text{O})_3]\text{NO}_3 \cdot 4\text{H}_2\text{O} \cdot 2\text{CH}_3\text{OH}$, a Case of Spin-Frustrated System, and $[\text{Cr}_3(\mu_3\text{-O})(\mu_2\text{-PhCOO})_2(\mu_2\text{-OCH}_2\text{CH}_3)_2(\text{bpy})_2(\text{NCS})_3]$, a New Type of $[\text{Cr}_3\text{O}]$ Core

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The synthesis, crystal structure, and magnetic properties of two trinuclear oxo-centered carboxylate complexes are reported and discussed: $[\text{Cr}_3(\mu_3\text{-O})(\mu_2\text{-PhCOO})_6(\text{H}_2\text{O})_3]\text{NO}_3 \cdot 4\text{H}_2\text{O} \cdot 2\text{CH}_3\text{OH}$ (**1**) and $[\text{Cr}_3(\mu_3\text{-O})(\mu_2\text{-PhCOO})_2(\mu_2\text{-OCH}_2\text{CH}_3)_2(\text{bpy})_2(\text{NCS})_3]$ (**2**). For both complexes the crystal system is monoclinic, with space group $C2/c$ for **1** and $P1/n$ for **2**. The structure of complex **1** consists of discrete trinuclear cations, associated NO_3^- anions, and lattice methanol and water molecules. The structure of complex **2** is built only by neutral discrete trinuclear entities. The most important feature of **2** is the unusual skeleton of the $[\text{Cr}_3\text{O}]$ core due to the lack of peripheral bridging ligands along one side of the triangular core, which is unique among the structurally characterized $(\mu_3\text{-oxo})$ -trichromium(III) complexes. Magnetic measurements were performed in the 2–300 K temperature range. For complex **1**, in the high-temperature region ($T > 8$ K), experimental data could be satisfactorily reproduced by using an isotropic exchange model, $H = -2J_{12}S_1S_2 - 2J_{13}S_1S_3 - 2J_{23}S_2S_3$ ($J_{12} = J_{13} = J_{23}$) with $J_{ij} = -10.1 \text{ cm}^{-1}$, $g = 1.97$, and $\text{TIP} = 550 \times 10^{-6} \text{ emu mol}^{-1}$. The antisymmetric exchange interaction plays an important role in the magnetic behavior of the system, so in order to fit the experimental magnetic data at low temperature, a new magnetic model was used where this kind of interaction was also considered. The resulting fitting parameters are the following: $G_{zz} = 0.25 \text{ cm}^{-1}$, $\delta = 2.5 \text{ cm}^{-1}$, and $\text{TIP} = 550 \times 10^{-6} \text{ emu mol}^{-1}$. For complex **2**, the experimental data could be satisfactorily reproduced by using an isotropic exchange model, $H = -2J_1(S_1S_2 + S_1S_3) - 2J_2(S_2S_3)$ with $J_1 = -7.44 \text{ cm}^{-1}$, $J_2 = -51.98 \text{ cm}^{-1}$, and $g = 1.99$. The magnetization data allows us to deduce the ground term of $S = 1/2$, characteristic of equilateral triangular chromium(III) for complex **1** and $S = 3/2$ for complex **2**, which is confirmed by EPR measurements.

Introduction

Trinuclear oxo-centered metal complexes of the general type $[\text{M}_3\text{O}(\text{O}_2\text{CR})_6\text{L}_3]^{0/+}$,^{1a} where L is water or pyridine and

R is an alkyl or aryl group, have been of great interest as systems to test current theories of magnetic and electronic coupling between metal atoms. A feature that has attracted attention is that when the three metal atoms are identical, and of the same oxidation state, the complex nearly always fails to adopt the 3-fold symmetry that might be expected. One interesting phenomenon is spin frustration, which occurs when all the interactions between spin pairs cannot simultaneously have their optimal value. The simplest systems where this could be observed are an odd-member ring of antiferromagnetic coupled spins. In complexes which have the triangular $\mu_3\text{-oxo}$ bridge $[\text{M}_3\text{O}]$ core, when all three

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pairwise interactions are antiferromagnetic, there is the possibility of spin frustration if the three pairwise interactions are the same; then, only two of the three spin constraints can be satisfied simultaneously, and the system is geometrically frustrated. The ground state then is of intermediate spin, intermediate between the lowest and the highest spins possible for the three metal ions. It is well-established that the μ_3 -oxo bridge in $[M_3O(O_2CR)_6L_3]^{n+}$ complexes is dominant in propagating the exchange interaction.^{1b,c} Oxo-centered trinuclear complexes of antiferromagnetic spins are quite common for transition metal ions like Fe^{3+} ,² Mn^{3+} ,^{1c,3} and Cr^{3+} .^{2k,4} The isotropic Hamiltonian for the interaction between any two spins can be written as the scalar product of the spin operators: $H = -\sum_{ij} J_{ij} S_i S_j$. Therefore, the energy is minimized for collinear parallel or antiparallel spin alignments. The available experimental data are in good agreement with the isotropic HDVV model in the region of

relatively high temperatures. However, there are a number of serious discrepancies between the predictions of the theory and the experimental data at low temperatures. In order to interpret the magnetic properties of these frustrated systems, it is necessary to go beyond the framework of the above isotropic exchange model and to introduce an antisymmetric exchange interaction, G_{AB} . This antisymmetric exchange interaction of the type

$$H_{AS} = \sum_{ij} G_{ij} [S_i \times S_j]$$

where $G_{ij} = -G_{ji}$ is the antisymmetric vector constant. As was pointed out by Tsukerblat and co-workers,⁵ this equation was introduced phenomenologically by Dzialoshinski to explain the weak ferromagnetism of α - Fe_2O_3 and manganese carbonate crystals.⁶ The microscopic meaning of the parameter G was revealed in the work of Moriya.⁷ Therefore, the operator H_{AS} is usually termed the Dzialoshinski–Moriya antisymmetric (AS) interaction. AS exchange leads to qualitative and quantitative changes in the temperature and field dependences of the magnetic susceptibility at low temperatures. These features are due to the nonlinear behavior of levels in the magnetic field when AS exchange is taken into account, and they consist of a reduction of the magnetic moments.⁵ Information about G_{ij} values for isolated complexes is very scarce. The first experimental observation of the antisymmetric exchange parameter on the electronic properties of the trinuclear complex $[Cu_3(\mu_3-OH)(pyridine-2-aldoxime)_3(SO_4)] \cdot 10.5H_2O$ was made by Tsukerblat and co-workers.⁸ Experimental applications of this exchange term have been reported for various Cu^{2+}_3 complexes^{9,10} as well as for Fe^{3+}_3 and Cr^{3+}_3 carboxylates.^{11,12}

In this work, we describe the synthesis, X-ray structure, and magnetic properties of two cyclic trimeric compounds: $[Cr_3(\mu_3-O)(\mu_2-PhCOO)_6(H_2O)_3]NO_3 \cdot 4H_2O \cdot 2CH_3OH$ (**1**) and $[Cr_3(\mu_3-O)(\mu_2-PhCOO)_2(\mu_2-OCH_2-CH_3)_2(bpy)_2(NCS)_3]$ (**2**). Complex **1** shows a similar cationic structure to those already reported,^{4k,m–4r} but the skeleton of the $[Cr_3O]$ core of complex **2** is unique among the structurally characterized $(\mu_3$ -oxo)-trichromium(III) complexes.

The magnetic behavior of complex **1**, which can be regarded as a geometrically frustrated system, is studied by

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Table 1. Crystallographic Data for Complexes **1** and **2**

	1	2
empirical formula	C ₄₄ H ₅₈ Cr ₃ NO ₂₅	C ₄₁ H ₃₆ Cr ₃ N ₇ O ₇ S ₃
formula mass	1156.87	990.95
cryst syst	monoclinic	monoclinic
space group	C2/c	P ₁ /n
Z	4	4
a (Å)	10.9596(9)	12.037(3)
b (Å)	24.9214(19)	16.864(4)
c (Å)	21.9436(17)	21.384(5)
α (deg)	90	90
β (deg)	92.268(2)	100.210(5)
γ (deg)	90	90
V (Å ³)	5988.7(8)	4272.0(18)
ρ (calcd) (g/cm ³)	1.276	1.541
μ _{calcd} (mm ⁻¹)	0.609	0.955
λ(Mo Kα) (Å)	0.71073	0.71073
T (K)	173(2)	173(2)
θ range for data collection	1.86–28.30°	1.55–20.81°
total reflns	18057	22149
reflns I > 2σ	4420	3407
reflexions collected	7436	4473
params refined	359	644
final R indices [I > 2σ(I)]	R1 = 0.0532 wR2 = 0.1622	R1 = 0.0507 wR2 = 0.1311
final R indices [for all data]	R1 = 0.0904 wR2 = 0.1809	R1 = 0.0700 wR2 = 0.1433
goodness-of-fit on F ²	0.991	1.035
largest diff peak, hole (e Å ⁻³)	1.255 –0.471	1.061 –0.472

considering that the antisymmetric exchange interaction is operating. The corresponding calculations and results are reported.

Experimental Section

Materials. All starting materials were purchased from Aldrich and were used without further purification.

Synthesis of the New Complexes. [Cr₃(μ₃-O)(μ₂-PhCOO)₆(H₂O)₃]NO₃·4H₂O·2CH₃OH (**1**). A solution of Cr(NO₃)₃·9H₂O (0.80 g, 2 mmol) in methanol (20 mL) was added to an acetonitrile (20 mL) solution of benzoic acid (1.45 g, 12 mmol) and stirred under reflux for 1 h. Green crystals, suitable for X-ray structure analysis, were grown by slow evaporation at room temperature of the green solution after 1 day (yields 90%). Anal. Calcd for **1** (C₄₄H₅₈Cr₃NO₂₅): C, 46.68; N, 1.21; H, 5.05. Found: C, 46.4; N, 1.2; H, 5.2.

[Cr₃(μ₃-O)(μ₂-PhCOO)₂(μ₂-OCH₂-CH₃)₂(bpy)₂(NCS)₃] (**2**). A solution of NH₄SCN (0.84 g, 10.98 mmol) in DMF (10 mL) and bpy (0.86 g, 5.49 mmol) were added to a solution of complex **1** (1.86 g, 1.83 mmol) in ethanol (15 mL). The resulting solution was allowed to stand at 138 °C for 1 day under autogenous pressure in a Teflon-lined stainless steel autoclave. Slow cooling of the resulting solution to room temperature afforded brown crystals suitable for X-ray structure analysis (yields 75%). Anal. Calcd for **2** (C₄₁H₃₆Cr₃N₇O₇S₃): C, 49.69; N, 9.89; H, 3.66; S, 9.71. Found: C, 50.0; N, 9.8; H, 3.7; S, 9.7.

Crystal Structure Determination. Crystal data and details on the data collection and refinement are summarized in Table 1. Suitable single crystals of **1** and **2** were selected and mounted on a Bruker SMART-CCD area diffractometer with graphite monochromated. Unit-cell parameters were determined from 359 frames for **1** and 244 for **2** of intensity data covering 0.3° in ω over a hemisphere of the reciprocal space by combination of three exposure sets, and refined by least-squares method. Intensities were collected using ω/2θ. Lorentz-polarization and absorption corrections were made.

Table 2. Selected Bond Lengths (Å) and Angles (deg) for **1**

Cr1–O1	1.904(3)	Cr2–O2	1.957(2)
Cr1–O3	1.976(2)	Cr2–O4	1.968(2)
Cr1–O7	1.947(2)	Cr2–O6	1.973(3)
Cr1–O5	2.031(3)	Cr2–O8	1.974(3)
Cr2–O1	1.892(3)	Cr2–O9	2.049(2)
O1–Cr1–O7	95.64(7)	O4–Cr2–O6	169.34(10)
O7–Cr1–O7#1	168.73(13)	O1–Cr2–O8	95.60(9)
O1–Cr1–O3	93.44(7)	O2–Cr2–O8	170.42(10)
O7–Cr1–O3	87.69(11)	O4–Cr2–O8	86.54(11)
O7#1–Cr1–O3	91.63(11)	O6–Cr2–O8	89.02(12)
O3–Cr1–O3#1	173.12(14)	O1–Cr2–O9	177.97(11)
O1–Cr1–O5	180.0	O2–Cr2–O9	84.80(10)
O7–Cr1–O5	84.36(7)	O4–Cr2–O9	84.05(11)
O3–Cr1–O5	86.56(7)	O6–Cr2–O9	85.97(11)
O1–Cr2–O2	93.81(8)	O8–Cr2–O9	85.84(11)
O1–Cr2–O4	94.60(8)	Cr2#1–O1–Cr2	120.01(14)
O2–Cr2–O4	94.48(10)	Cr2#1–O1–Cr1	120.00(7)
O1–Cr2–O6	95.47(9)	Cr2–O1–Cr1	120.00(7)
O2–Cr2–O6	88.31(11)		

The structures were solved by direct methods, using the SHELXS computer program,¹³ and refined by full-matrix least-squares method with the SHELXL97 computer program,¹³ using 7436 reflections for **1** and 4473 reflections for **2**. The function minimized was $\sum w|F_o|^2 - |F_c|^2|^2$, where $w = [\sigma^2(I) + (0.1067P)^2]^{-1}$ for **1** and $w = [\sigma^2(I) + (0.0927P)^2 + 4.466P]^{-1}$ for **2**, and $P = (|F_o|^2 + 2|F_c|^2)/3$, f , f' , and f'' were taken from International Tables of X-ray Crystallography.¹⁴ For **1**, the hydrogen atoms were calculated with the exception of those corresponding to the water and methanol molecules that could not be located. For **2**, 10 hydrogen atoms were calculated, and 25 hydrogen atoms were located from a difference synthesis, while the remaining hydrogen atoms were refined without constraints. The computed hydrogen atoms were refined using a riding model with the isotropic temperature factor for H atoms equal to 1.2 times the equivalent temperature factor of the atom which is linked.

Physical Measurements. Magnetic measurements were carried out in the “Unitat de Mesures Magnétiques (Universitat de Barcelona)” on polycrystalline samples (20 mg) with a Quantum Design SQUID MPMS-XL magnetometer working in the 2–300 K range. The magnetic field was 10000 G at high temperature and 200 G at low temperature. The field-dependent magnetization was measured in the applied magnetic field range 0–5 T at 2 K. The diamagnetic corrections were evaluated from Pascal’s constants. The X-band EPR spectra were recorded on powder samples with a Bruker 300E automatic spectrometer, varying the temperature between 4 and 300 K.

Results and Discussion

Description of the Structure of Complexes [Cr₃(μ₃-O)(μ₂-PhCOO)₆(H₂O)₃]NO₃·4H₂O·2CH₃OH (1**) and [Cr₃(μ₃-O)(μ₂-PhCOO)₂(μ₂-OCH₂-CH₃)₂(bpy)₂(NCS)₃] (**2**).** The structure of these complexes was determined by X-ray crystallography. For both complexes, the crystal system is monoclinic, with space group C2/c for complex **1** and P₁/n for complex **2**. Selected bond lengths and angles are listed in Tables 2 and 3 for complexes **1** and **2**, respectively.

The structure of complex **1** is built by trinuclear cations [Cr₃(μ₃-O)(μ₂-PhCOO)₆(H₂O)₃]⁺, with NO₃[−] ions acting as

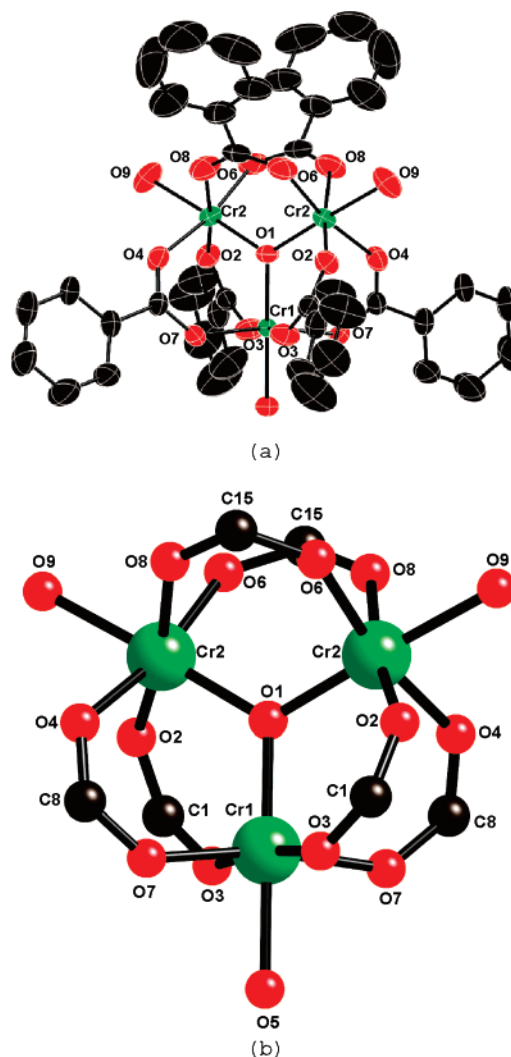
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Table 3. Selected Bond Lengths (Å) and Angles [deg] for **2**

Cr01–O1	1.922(4)	Cr02–N2	2.019(6)
Cr01–O6	1.964(4)	Cr02–N21	2.064(5)
Cr01–O7	1.968(4)	Cr02–N22	2.076(5)
Cr01–O2	2.014(4)	Cr03–O1	1.902(4)
Cr01–O5	1.992(4)	Cr03–O7	1.958(4)
Cr01–N1	2.013(6)	Cr03–O4	1.988(4)
Cr02–O1	1.895(4)	Cr03–N3	2.015(6)
Cr02–O6	1.936(4)	Cr03–N32	2.050(5)
Cr02–O3	1.982(4)	Cr03–N31	2.064(5)
O1–Cr01–O6	80.38(16)	N22–Cr02–O1	95.21(17)
O1–Cr01–O7	82.00(15)	O6–Cr02–N22	175.88(19)
O6–Cr01–O7	162.14(17)	O3–Cr02–N22	91.04(19)
O1–Cr01–O5	90.93(16)	N2–Cr02–N22	87.9(2)
O5–Cr01–O6	87.87(17)	N22–Cr02–N21	78.1(2)
O5–Cr01–O7	89.55(16)	O1–Cr02–N21	91.96(17)
O1–Cr01–N1	176.90(19)	O1–Cr03–O7	82.77(16)
O6–Cr01–N1	101.7(2)	O1–Cr03–O4	90.40(16)
O7–Cr01–N1	96.03(19)	O4–Cr03–O7	93.62(16)
O5–Cr01–N1	91.45(19)	O1–Cr03–N3	177.70(18)
O1–Cr01–O2	89.24(16)	O7–Cr03–N3	94.96(18)
O6–Cr01–O2	91.12(17)	O4–Cr03–N3	90.10(18)
O7–Cr01–O2	91.52(16)	O1–Cr03–N32	90.11(17)
O5–Cr01–O2	178.93(18)	O7–Cr03–N32	93.58(17)
N1–Cr01–O2	88.42(19)	O4–Cr03–N32	172.78(18)
O1–Cr02–O6	81.76(16)	N3–Cr03–N32	89.68(19)
O1–Cr02–O3	91.51(16)	O1–Cr03–N31	93.65(17)
O3–Cr02–O6	91.82(17)	O7–Cr03–N31	171.68(18)
N2–Cr02–O1	176.2(2)	O4–Cr03–N31	93.91(18)
O6–Cr02–N2	95.0(2)	N3–Cr03–N31	88.55(19)
O3–Cr02–N2	90.63(18)	N31–Cr03–N32	78.87(19)
O6–Cr02–N21	99.14(19)	Cr01–O1–Cr03	97.11(16)
O3–Cr02–N21	168.9(2)	Cr01–O1–Cr02	98.12(16)
N2–Cr02–N21	86.53(19)	Cr03–O1–Cr02	164.5(2)

counteranion, and solvent methanol and water molecules. The ORTEP view, with atom labeling scheme and the core of the trinuclear entity are given in Figure 1a,b, respectively. Their skeletal structure is very similar to those of other trinuclear chromium(III) carboxylates.^{4c,h,j,m-r,15–19} The three metal ions lie at the corners of a nearly equilateral triangle of separations: Cr2–Cr2 = 3.276 Å and Cr1–Cr2 = 3.287 Å. The center of the triangle is occupied by triply bridging oxygen not displaced out of the plane. The μ_3 -oxide atom is the common vertex of the coordination octahedral around the three Cr³⁺ ions. The trigonal geometry at the μ_3 -O bridge is reflected by the analogous angles: Cr2–O1–Cr2 = 120.01° and Cr1–O1–Cr2 = 119.99°, very close to 120°. The benzoate ligand acts as bidentate and bridges couples of Cr³⁺ ions, so that four benzoate oxygen atoms surround each chromium(III) ion. A water molecule, in trans position to the shared oxo ligand, acts as the sixth donor. Therefore, the coordination at each Cr³⁺ ion includes six Cr–O bonds in a distorted octahedral environment. An examination of the angle values in the structure shows that all the μ_3 -O–Cr–O (benzoate) angles exceed 90°, whereas all the O(benzoate)–Cr–O(water) ones are smaller (Table 2). The trinuclear cations are self-assembled through hydrogen bonds

**Figure 1.** (a) ORTEP view of complex [Cr₃(μ_3 -O)(μ_2 -PhCOO)₆(H₂O)₃]·NO₃·4H₂O·2CH₃OH (**1**) with atom labeling scheme. (b) Core of the trinuclear entity of **1**.

giving rise to a 2-D supramolecular structure in the *xy* plane (Figure S1, Supporting Information). The O9, O5 atoms of the water molecules coordinated to the chromium ions, and the O13/O13' and O12 oxygen atoms of the NO₃[−] counteranion and the O18 of the methanol molecules of crystallization participated in this supramolecular structure. The shortest inter-trinuclear distance Cu–Cu is 8.15 Å. Short contact distances are shown in Table S1 (Supporting Information).

The structure of complex **2** is built only by neutral trinuclear [Cr₃(μ_3 -O)(μ_2 -PhCOO)₂(μ_2 -OCH₂-CH₃)₂(bpy)₂-(NCS)₃] entities. The ORTEP view, with atom labeling scheme, and the core of the trinuclear entity are given in Figure 2a,b, respectively. The three metal atoms lie at the corners of a nearly isosceles triangle of separations Cr(01)···Cr(02) = 2.883 Å and Cr(01)···Cr(03) = 2.866 Å along equivalent sides of the isosceles triangle, and Cr(02)···Cr(03) = 3.768 Å along the inequivalent one. The center of the triangle is occupied by a triple-bridging oxygen atom. The equivalent Cr(01)–Cr(02) and Cr(01)–Cr(03) sides of the triangle are each bridged by the μ_2 -oxygen atom of the ethanolate ligand and by a bidentate benzoate group; no bridges were present along the inequivalent Cr(02)–Cr(03)

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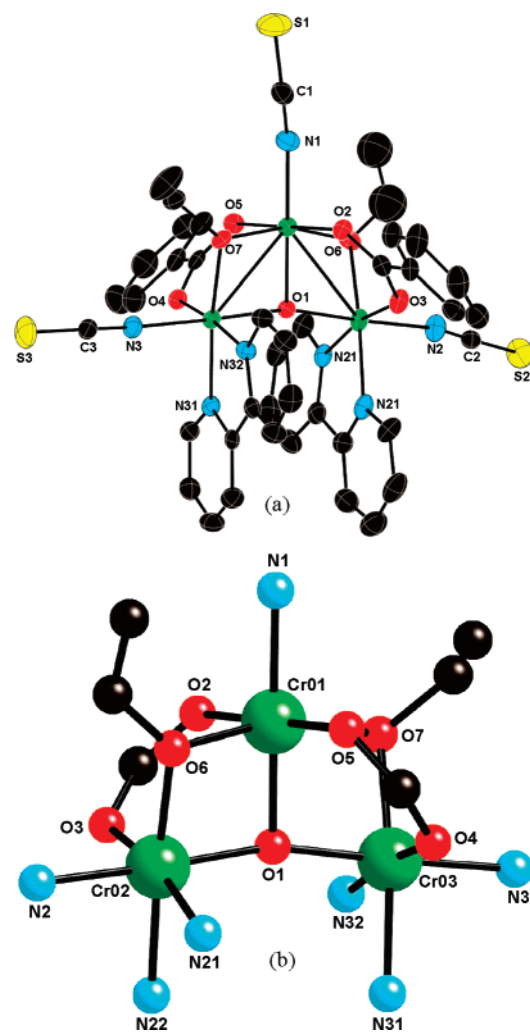


Figure 2. (a) ORTEP view of complex [Cr₃(μ₃-O)(μ₂-PhCOO)₂(μ₂-OCH₂-CH₃)₂(bpy)₂(NCS)₃ (2) with atom labeling scheme. (b) Core of the trinuclear entity of 2.

side. This lack of peripheral bridging ligands between Cr2–Cr3 is the most important structural feature which is unique among the structurally characterized (μ₃-oxo)trichromium(III) complexes, and this characteristic is responsible for the anomalous Cr–O1–Cr: 98.12° (Cr(01)–O1–Cr(02)), 97.11° (Cr(01)–O1–Cr(03)), and 164.52° (Cr(02)–O–Cr(03)). This indicates that the hybrids of the oxygen central atoms are farther than regular sp² hybridization. Chromium atoms Cr(02) and Cr(03) are also coordinated to bipyridine nitrogen atoms, and all three chromium atoms have a terminal thiocyanate ligand linked by the nitrogen atom. The Cr–O–Cr angles from the μ₂-ethanolate are 95.30° (Cr(02)–O6–Cr(01)) and 93.79° (Cr(01)–O7–Cr(03)). The neutral trinuclear entities are isolated; no hydrogen bonds or π–π stacking are operating in the structure, so we can conclude that only van der Waals forces are responsible for the packing in the global structure.

Magnetic Studies. [Cr₃(μ₃-O)(μ₂-PhCOO)₆(H₂O)₃]NO₃·4H₂O·2CH₃OH (1). The temperature dependence of the magnetic susceptibility data in χ_MT form (χ_M being the magnetic susceptibility per [Cr₃] unit) is shown in Figure 3. The χ_MT data decrease smoothly from a value around 4.73 emu mol^{−1} K at room temperature, smaller than the value

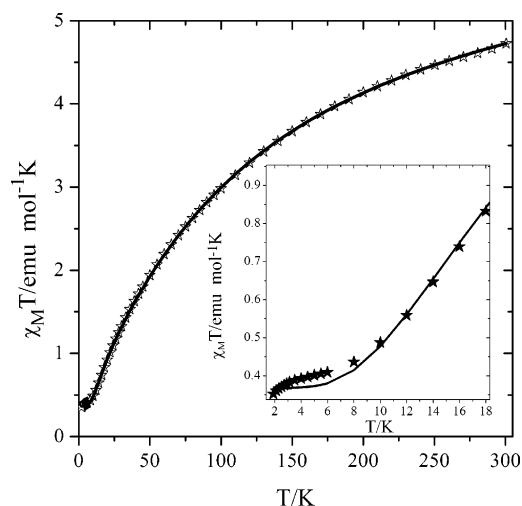


Figure 3. Magnetic susceptibility data in χ_MT form of 1. Solid lines above the experimental data are theoretical curves derived from Hamiltonian 1. In the inset is shown the low-temperature susceptibility data along with the theoretical curve.

5.628 emu mol^{−1} K expected for three independent Cr³⁺ ions, to the value of 0.43 emu mol^{−1} K at 10 K. Below that temperature, a plateaulike feature appears until 2 K (0.35 emu mol^{−1} K) which is clear evidence that the total spin (S_T) for the ground state is S_T = 1/2.

The isotropic Heisenberg Dirac van Vleck (HDvV) Hamiltonian (eq 1) is the main term used to interpret the magnetic data.

$$H = -2J_{12}S_1S_2 - 2J_{13}S_1S_3 - 2J_{23}S_2S_3 \quad (1)$$

Superexchange has been studied in a large number of oxo-centered trinuclear complexes, and it is known that the coupling is antiferromagnetic ($J_{ij} < 0$). Since the three chromium atoms of the [Cr₃O] unit define a quasiequilateral triangle, the three Cr³⁺ ions can be considered equivalent, and so the three J_{ij} constants are equal. If we assume three equal J_{ij} 's, the data are reasonably reproduced with $J_{ij} = -10.1$ cm^{−1}, $g = 1.97(1)$, and TIP = 550 × 10^{−6} emu mol^{−1} (TIP = temperature-independent paramagnetism). The quality of the magnetic susceptibility data does not vary significantly if we adopt an isosceles triangle ($J_{12} = J_{13} \neq J_{23}$). The resolution of our data in the high-temperature range, however, does not allow us to make an accurate determination of the difference $|J_{12} - J_{23}|$. In the inset of Figure 3 is shown the low-temperature magnetic data where an obvious discrepancy exists between the magnetic model and the experimental points. In order to clarify the low-temperature behavior, a systematic measurement of the susceptibility data was carried out at different magnetic fields and is shown in Figure 4. It is well documented that the antisymmetric exchange interaction plays an important role in the magnetic behavior of the system, and so a new magnetic model was used where this kind of interaction was also considered:

$$H = -2J_0(S_1 \cdot S_2 + S_2 \cdot S_3) - 2J_1(S_1 \cdot S_3) + G([S_1 \times S_2] + [S_2 \times S_3] + [S_3 \times S_1]) + (g_z \cos \theta + g_{xy} \sin \theta)\beta H(S_1 + S_2 + S_3) \quad (2)$$

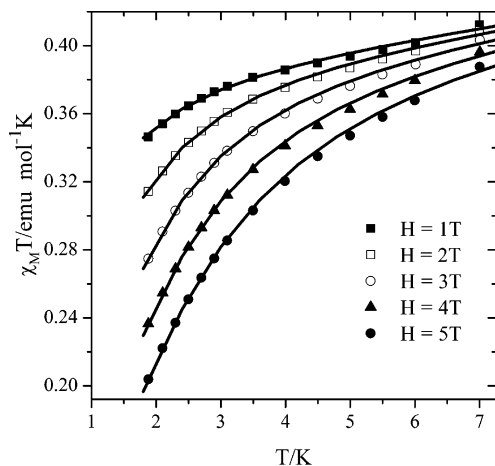


Figure 4. Low-temperature susceptibility data, at different magnetic fields, in the form of $\chi_M T$ vs T of **1**. The solid lines represent the theoretical curves according to Hamiltonian 2. See text for details.

where J_0, J_1 are the isotropic exchange interactions (due to the symmetry lowering) and \mathbf{G} is the antisymmetric exchange vector parameter for two-center interaction. In trigonal systems, the components of \mathbf{G} are assumed to follow the relation $G_z \gg (G_x, G_y) \approx 0$ (z being the 3-fold axis), which can be derived from the properties of the \mathbf{G} vector.⁷ This kind of interaction perturbs only the two lowest Kramers doublets of $S_T = 1/2$, and the diagonalization yields the following energy levels:

$$E = \pm \frac{[\Delta E^2 + |M|^2 \pm 2\sqrt{\delta^2|M|^2 + |M \cdot \mathbf{G}|^2}]^{1/2}}{2} \quad (3)$$

where $|M| = \sqrt{(g_{xy}^2 \sin^2 \theta + g_z^2 \cos^2 \theta)} \beta H$ is the Zeeman term, $|G| = 4\sqrt{3}|G_{zz}|$, assuming $|G_{xx}| = |G_{yy}| \sim 0$ which is valid for a weakly axial system with $J_0 \approx J_1$, $|M \cdot \mathbf{G}| = 4\sqrt{3}G_{zz}g_z\beta H \cos \theta$, and $\Delta E = \sqrt{\delta^2 + |G|^2}$ while the δ factor is equal to $4|J_0 - J_1|$. It must be pointed out that the sign of the \mathbf{G} is also undefined under these conditions. It has been shown that $S_T = 3/2$ is not perturbed by the antisymmetric exchange term, \mathbf{G} , and does not affect the description of the $S_T = 1/2$ states.

The components of the susceptibility tensor were calculated according to a generalized susceptibility equation,⁸ and it was necessary to consider the parallel and perpendicular susceptibility components separately, since they are affected differently by the \mathbf{G} vector. Using this approach, it was possible to fit the low-temperature data, 1.8–8 K, and the fitting results are shown in Figure 4 as solid lines while the resulting fitting parameters are, assuming $g_{xy} = g_z = 2.0$ as fixed, $|G_{zz}| = 0.25 \text{ cm}^{-1}$, $\delta = 2.5 \text{ cm}^{-1}$, $\text{TIP} = 550 \times 10^{-6} \text{ emu mol}^{-1}$. In order to investigate the influence of \mathbf{G} and δ parameters in the low-temperature behavior of the susceptibility, different simulations were carried out and are shown in Figure 5. While each of the two parameters varied, the other one was fixed to the value obtained from the fitting procedure. When the δ factor increases, the susceptibility plateaus are shifted toward higher values (closer to $0.375 \text{ emu mol}^{-1} \text{ K}$) while an increase in the \mathbf{G} parameter causes a drastic divergence from the plateau to appear. In Figure

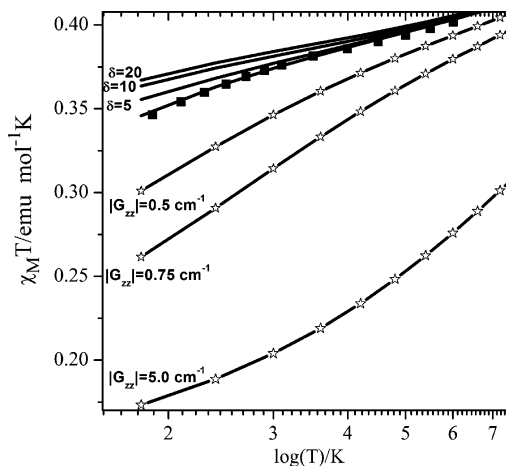


Figure 5. Simulation of the susceptibility equation in the form $\chi_M T$ vs $\log T$ for different values of the \mathbf{G} and δ parameters at external field $H = 1\text{ T}$ according to eq 2. Solid squares are the experimental points of the susceptibility data at $H = 1\text{ T}$ of **1**.

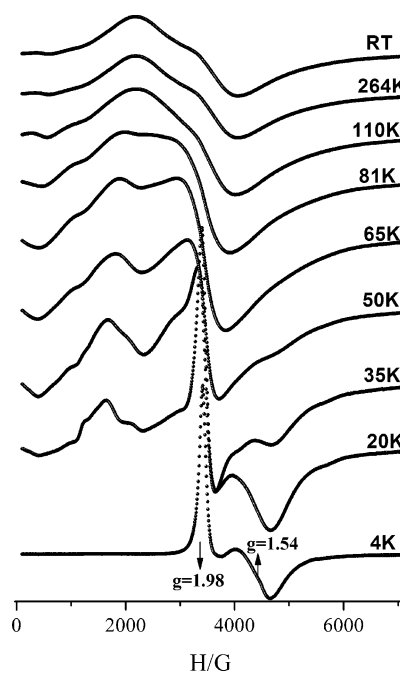


Figure 6. EPR spectra from a polycrystalline sample of **1** recorded at different temperatures from 4 K to room temperature. Other EPR conditions: modulation amplitude, 10 G_{pp}; microwave frequency, 9.40 GHz.

S2 (Supporting Information) is shown a simulation where only the Zeeman part of eq 1 was considered fixing $(\mathbf{G}, \delta) = (0, 0)$. As it is clearly shown, a more anisotropic Hamiltonian is necessary.

Magnetization data in the field range 0–6.5 T at 2 K are shown in Figure S3 (Supporting Information) along with the Brillouin function of a $S_T = 1/2$ system with $g = 2$. This behavior is in accordance with the susceptibility data and reveals that the ground state is indeed an $S_T = 1/2$ one. It was very difficult to extract information about the previous parameters from the magnetization data.

EPR Spectroscopy. EPR spectra from a polycrystalline sample of complex **1** at different temperatures from 4 K to room temperature are shown in Figure 6. At 4.2 K a strong absorption peak is observed at $g_{\text{eff}} = 1.97$ accompanied by a derivative feature at $g_{\text{eff}} = 1.70$. We attribute these signals

to an $S = 1/2$ with axial g -tensors (g_{\parallel}, g_{\perp}) = (1.98, 1.54). The temperature dependence of these signals indicates that their origin is from the ground state since after 20 K new signals appear at lower magnetic fields revealing the population of excited states.

For exchange coupled trimers, the large anisotropy in the g -tensor for the $S = 1/2$ ground state with a rather small g_{\perp} value is usually attributed to the effects of an antisymmetric exchange interaction according to eq 2.^{20–22} The effect of this antisymmetric interaction is to induce an axial anisotropy in the g -tensor. More explicitly, the g_{\parallel} component is not affected whereas a shift toward lower g -values is observed for the g_{\perp} component. From the EPR findings, compound **1** has isosceles magnetic symmetry or lower ($\delta \neq 0$); the antisymmetric exchange is important ($\mathbf{G} \neq 0$), and $\Delta E > h\nu$. The resulting relationship is¹¹

$$g_{\perp} = g_0 \left[\frac{\delta^2 - (h\nu)^2}{\Delta E^2 - (h\nu)^2} \right]^{1/2} \quad (4)$$

where $h\nu$ is the microwave energy ($\approx 0.31 \text{ cm}^{-1}$), g_0 is the true g calculated from the magnetic studies ($g_{xy} = g_z = 2.0$), and all the other symbols have the meaning defined in eq 2. According to the fitting results of the susceptibility data, $g_{\perp}/g_0 = 0.78$, which is close to the value obtained from the EPR spectrum (0.77). This is clear evidence that the susceptibility results are in accordance with the EPR findings.

[Cr₃(μ_3 -O)(μ_2 -PhCOO)₂(μ_2 -OCH₂-CH₃)₂(bpy)₂(NCS)₃] (**2**). The temperature dependence of the magnetic susceptibility data in the $\chi_M T$ form (χ_M being the magnetic susceptibility per [Cr₃] entity) is shown in Figure 7, top. At room temperature, the $\chi_M T$ value is $3.29 \text{ emu mol}^{-1} \text{ K}$, smaller than the value of $5.628 \text{ emu mol}^{-1} \text{ K}$ expected for three independent Cr³⁺ ions, and decreases to the value of $1.88 \text{ emu mol}^{-1} \text{ K}$ at 50 K. After that temperature, a plateau appears until 2 K, which is clear evidence that the total spin (S_T) for the ground state is equal to $3/2$. The fit of the experimental data, corrected for diamagnetic contributions and TIP ($240 \times 10^{-6}/\text{Cr}^{3+}$),²³ was performed according with the isotropic HDvV Hamiltonian (eq 5)

$$H = -2J_1(S_1S_2 + S_1S_3) - 2J_2(S_2S_3) \quad (5)$$

where J_1 corresponds to the coupling through the equivalent sides of the triangle and J_2 corresponds to the nonequivalent one (Figure 7, top). The free parameters were J_1 , J_2 , and g . The fit made by the irreducible tensor operator formalism (ITO), using CLUMAG program,²⁴ gave the following results: $J_1 = -7.44 \text{ cm}^{-1}$, $J_2 = -51.98 \text{ cm}^{-1}$, $g = 1.99$ and $R = 2.67 \times 10^{-4}$.

Magnetization data in the field range 0–5 T at 2 K is shown in Figure 7 (bottom) along with the Brillouin function

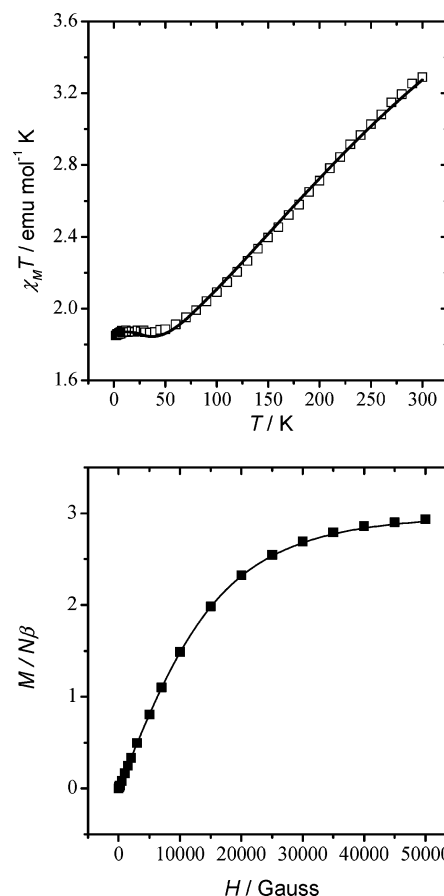


Figure 7. Top: Thermal dependence at 10000 G of $\chi_M T$ of **2**. Bottom: Magnetization vs H (2 K) of **2**.

of a $S_T = 3/2$ system with $g = 1.99$. This behavior is in accordance with the susceptibility data.

EPR Spectroscopy. X-band EPR measurements were carried out in powder samples of **2** and are shown in Figure S4 (Supporting Information). In the powder spectrum at 4 K, there are well-resolved peaks (shown with asterisks in Figure 8), which appear to belong to excited states such as the narrow small peak at about $g = 4.5$ as well as the well defined shoulder at $g = 3.14$. Also, for fields greater than 5000 G, it is difficult to observe the contribution of excited states. All these transitions were excluded from the simulation procedure that was carried out with the following Hamiltonian equation for a spin $S = 3/2$:²⁵

$$H = D \left[S_z^2 - \frac{1}{3} S(S+1) \right] + E(S_x^2 - S_y^2) + g_x \mu_B H_x S_z + g_y \mu_B H_y S_y + g_z \mu_B H_z S_z$$

The simulated spectrum is also shown in Figure 8 as a solid line.

An isotropic magnetic-field domain line width was used for the powder spectrum ($1w = 35 \text{ G}$), while the broadness of the spectrum reveals the g -strain effects. Gaussian distributions of the g -principal values were used, and the results are ($\sigma g_x = \sigma g_y = 0.1$, $\sigma g_z = 0.2$). The principal g -values for the powder spectrum are $g_{xy} = 1.96(1)$, $g_z =$

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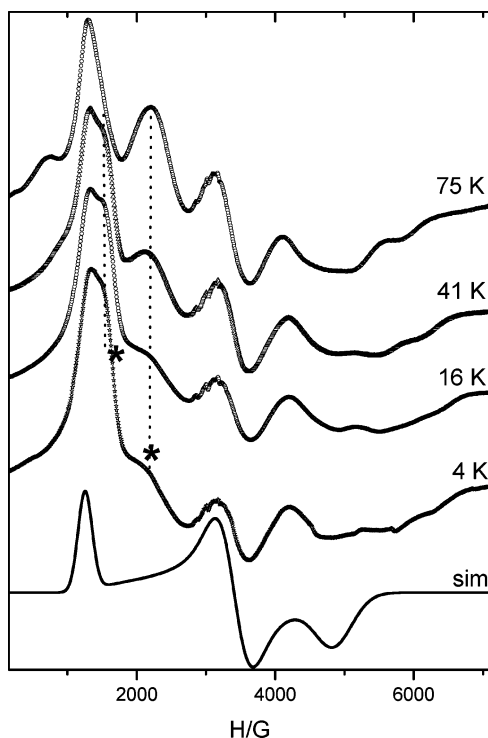


Figure 8. Simulated EPR spectrum (sim line) and temperature dependence of the powder spectra showing with asterisks transitions belonging to excited states. See text for details.

1.98(1). The values for the crystal field terms are $D = 1.5$ – $(5) \text{ cm}^{-1}$ and $\lambda = E/D = 0.3$, revealing the rhombic character of the system. Gaussian distributions of the D and E parameters were also used ($\sigma(D) = 0.3$, $\sigma(E) = 0.03$). Several complexes of Cr^{3+} have large D -values and were calculated through simulations of EPR X-band spectra.²⁶ The simulation showed that the experimental line shape cannot be reproduced by an $S = 3/2$ system with a well-defined value for E/D , even if one employs angular dependent line shape distributions of the type $W^2 = \sqrt{\sum W_i^2 l_i^2}$, where l_i denotes the direction cosines of the applied field. There are well-resolved peaks which appear to belong to excited states and influence the line shape dramatically.

Conclusions

The kinetically inert nature of the Cr^{3+} ion usually requires reactions to be carried out at high temperature to obtain polynuclear clusters of high nuclearity;²⁷ one of the most successful synthetic routes has involved the direct heating

of basic carboxylates $[\text{Cr}_3\text{O}(\text{O}_2\text{CR})_6(\text{H}_2\text{O})_3]^+$. The hydrothermal reaction of complex **1** in ethanol, SCN^- , and bpy produces complex **2** which involves loss of the three coordinated water molecules and four carboxylic acids and the coordination of three thiocyanate ions, three ethoxo bridging ligands, and two bpy molecules. A cluster of higher nuclearity has not been obtained, but a new skeleton of the $[\text{Cr}_3\text{O}]$ core was found.

In order to justify the low-temperature magnetic behavior of complex **1**, the antisymmetric exchange interaction term (G) was added to the HDvV Hamiltonian which represents the spin–orbit coupling attributed to the orbitally degenerated ground term. This antisymmetric term splits the two $S = 1/2$ Kramers doublets and introduces an axial anisotropy in the system which can be clearly observed by EPR spectroscopy: The g_{\parallel} and g_{\perp} values are 1.98 and 1.54, respectively, at 2 K. The magnitude of the J value is similar to the value of the interaction constant observed in analogous $[\text{Cr}_3\text{O}]$ complexes previously reported.

Complex **2** shows different geometry, and one of the interactions between two chromium(III) ions (J_2) is different from the other two (J_1). The J_1 value includes the coupling through the triple-bridging oxygen atom, the benzoate, and the ethanolate bridging ligands. The magnitude of J_2 is the highest, since the interaction is mediated by an oxo bridged with a CrOCr angle of 164.5° for which a high value of about -60 cm^{-1} is expected for the exchange coupling constant by the literature.²⁸ The magnetization data allows us to deduce the ground term of $S = 1/2$, characteristic of equilateral triangular chromium(III), for complex **1** and $S = 3/2$ for complex **2**, which is confirmed by EPR measurements.

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Supporting Information Available: Additional tables and figures. CIF data. CCDC-634867–634868 (**1** and **2**), which can be obtained free of charge from the Cambridge Crystallographic Data Centre via www.ccdc.cam.ac.uk/data_request/cif. This material is available free of charge via the Internet at <http://pubs.acs.org>.

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