

19 K = $1s^2 2s^2 2p^6 3s^2 3p^4 4s^1$ → konfigurasi
↓ 1 → elektron valensi I A

10 Ne = $1s^2 2s^2 2p^6$ → konfigurasi
K L⁸ M¹⁰ N³²
↓ 8 → elektron valensi VIII A

21 Sc = $1s^2 2s^2 2p^6 3s^2 3p^4 4s^2 3d^1$ → konfigurasi
↓ 3 → elektron valensi III

24 Cr = $1s^2 2s^2 2p^6 3s^2 3p^4 4s^1 3d^5$ → konfigurasi
↓ 6 elektron valensi VIB

Tugas.

K L M N
2 8 14

1. Tentukan jumlah proton, elektron dan neutron dari :

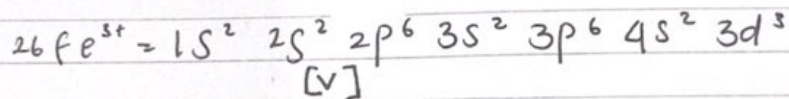
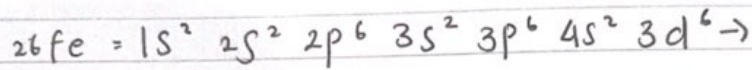
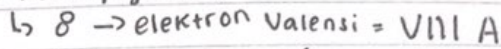
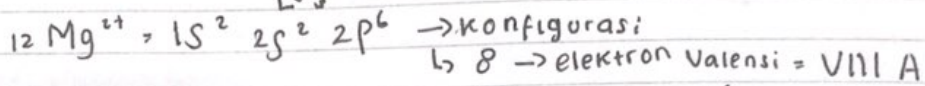
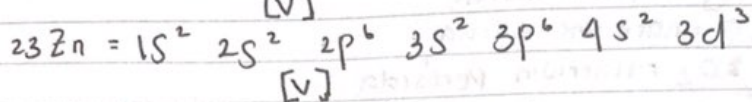
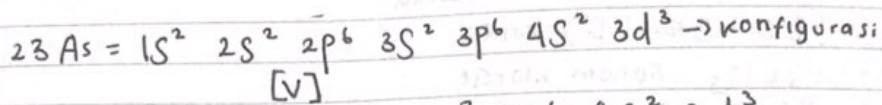
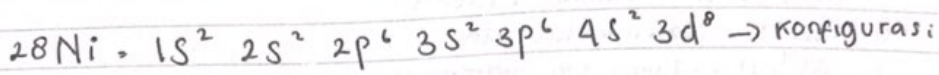
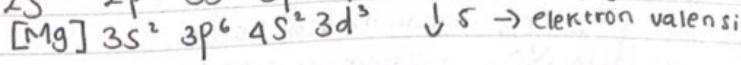
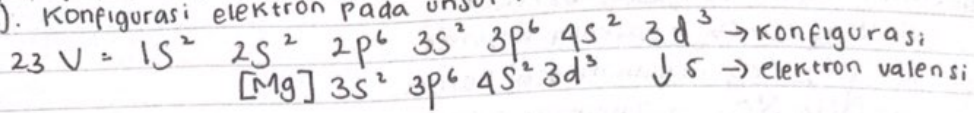
a. $^{84}_{30}\text{Sr}$ Proton : 30
elektron : 30
neutron : $84 - 30 = 46$

b. $^{35}_{17}\text{Cl}^-$ Proton : 17
neutron : 18
elektron : $35 - 17 = 18$

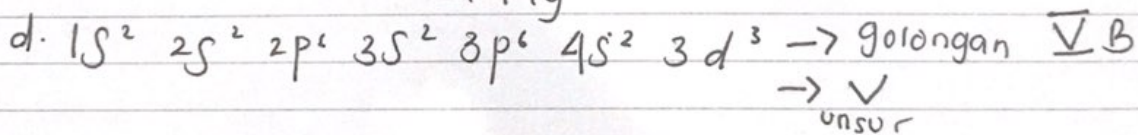
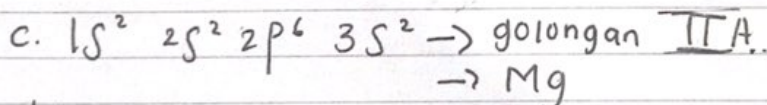
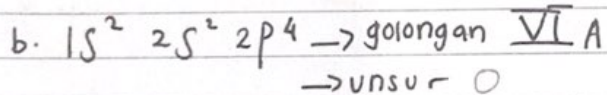
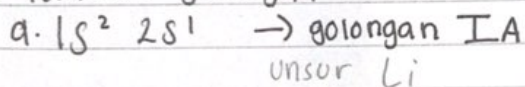
c. $^{56}_{25}\text{Fe}^{2+}$ Proton : 25
elektron : 23
neutron : $56 - 25 = 31$

d. $^{16}_8\text{O}^{2-}$ Proton : 8
elektron : 10
neutron : 16

2). Konfigurasi elektron pada unsur:



3) Tentukan golongan dan periode



4. Berikan nama senyawa:

