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WHS AP Chemistry 5 The Gas Laws CHARLES'S LAW Charles' Law states the volume of a gas varies directly with the Kelvin temperature, assuming the pressure is

AP ws Charles Law key - Conejo Valley Unified School District

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WHS AP Chemistry 5 The Gas Laws BOYLE'S LAW Boyle's Law states that the volume of a gas varies inversely with its pressure if temperature is held constant.

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Chemistry IF8766 82 ©Instructional Fair, Inc. LE CHATELIER'S PRINCIPLE Name _ Le Chatelier's Principle states that when a system at equilibrium is subjected to a stress, the system will shift its equilibrium point in order to relieve the stress.

LE CHATELIER'S PRINCIPLE Name

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pH and pOH The pH of a solution indicates how acidic or basic that solution is. pH range of 0-7 acidic 7 neutral 7-14 basic Since $[H^+][OH^-] = 10^{-14}$ at $25^\circ C$, if $[H^+]$ is known, the $[OH^-]$ can be calculated and vice versa. $pH = -\log [H^+]$ So if $[H^+] = 10^{-6} M$, $pH = 6$ $pOH = -\log [OH^-]$ So if $[OH^-] = 10^{-8}$, $pOH = 8$ Together, $pH + pOH = 14$.

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Chemistry IF8766 62 . STOICHIOMETRY: MASS-MASS PROBLEMS 1.2KC103 + _302 Name How many grams of potassium chloride are produced if 25 g of potassium chlorate decompose? / 25 ken How many grams of hydrogen are necessary to react completely with 50.0 g of nitrogen in the above reaction?

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Naming Ionic Compounds - Website

Word Equations Answer Key 1. Zinc and Lead (II) nitrate react to form Zinc Nitrate and Lead. $\text{Zn} + \text{Pb}(\text{NO}_3)_2 \rightarrow \text{Zn}(\text{NO}_3)_2 + \text{Pb}$ Single Replacement

Word Equations Answer Key Zinc and Lead (II) nitrate react ...

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COMPOSITION OF HYDRATES Name A hydrate is an ionic compound with water molecules loosely bonded to its crystal structure. The water is in a specific ratio to each formula unit of the salt.

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Writing Formulas From Names - Website

pH and pOH CONTINUED 1. $0.01 \text{ M HCl } [\text{H}^+] = 0.01 \text{ M } \text{pH} = -\log(0.01) = 2$ 2. $0.0010 \text{ M NaOH } [\text{OH}^-] = 0.0010 \text{ M } \text{pOH} = -\log(0.0010) = 3 \rightarrow \text{pH} = 11$ 3. 0.050 M NaOH

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