

## ***Iron Nail In An Aqueous Solution Answer***

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### Iron Nail In An Aqueous

Answers. The chemical formula for copper sulfate is  $\text{CuSO}_4$ , or, 1 copper atom, 1 sulfur atom, and 4 oxygen atoms to each molecule. When the iron is placed in the solution of copper sulfate, it replaces the copper in the solution, turning copper sulfate into iron sulfate ( $\text{FeSO}_4$ ) and pure copper collects on the iron.

### When a clean iron nail is placed in an aqueous solution of ...

Answers. Copper metal. This is a redox reaction. The Copper ions are oxidizing the Iron metal. The chemical equation is:  $\text{CuSO}_4 + \text{Fe} \rightarrow \text{FeSO}_4 + \text{Cu}$  The reason this happens is that Copper accepts electrons better than Iron can. The Copper ions take two electrons from the Iron atoms and this makes them Iron ions.

### Chem Help Please!!When a clean iron nail is placed in an ...

When a clean iron nail is placed in an aqueous solution of copper(II) sulfate, the nail becomes coated with a brownish-black material. (d) Write the balanced equation for the reaction.

### When a clean iron nail is placed in an aqu... | Clutch Prep

When a clean iron nail is placed in an aqueous solution of copper (II) sulfate, the nail becomes coated with a brownish black material. Therefore, the material coating the iron is ferrous sulphate ( $\text{FeSO}_4$ ). Chapter 21, Problem 51P is solved. This is an alternate ISBN.

### Solved: When a clean iron nail is placed in an aqueous ...

If a nail made of elemental iron [ $\text{Fe(s)}$ ] is placed in an aqueous solution of a soluble palladium(II) salt [ $\text{Pd}^{2+}(\text{aq})$ ], the nail will gradually disappear as the iron enters the solution as  $\text{Fe}^{3+}(\text{aq})$  and palladium metal [ $\text{Pd(s)}$ ] forms.

### If a nail made of elemental iron [Fe(s)] i... | Clutch Prep

Iron nails were dipped in an aqueous solution of copper soleplate. After about 30 minutes, it was observed that the color of the solution changed from A. colorless to light green.

### Iron nails were dipped in an aqueous solution of copper ...

Write two observations that you will make when an iron nail is kept in an aqueous solution of copper sulphate. Write the chemical equation for this reaction.

### Write two observations that you will make when an iron ...

Question: If a nail made of elemental iron ( $\text{Fe(s)}$ ) is placed in an aqueous solution of a soluble palladium ... If a nail made of elemental iron ( $\text{Fe(s)}$ ) is placed in an aqueous solution of a soluble palladium (II) salt ( $\text{Pd}^{2+}(\text{aq})$ ), the nail will gradually disappear as the iron enters the solution as  $\text{Fe}^{3+}(\text{aq})$  and palladium metal ( $\text{Pd(s)}$ ) forms. Is this a redox equation? Balance and show the net ionic equation that describes this reaction.

### Solved: If A Nail Made Of Elemental Iron (Fe(s)) Is Placed ...

The iron nails were fully polished with the steel wool. One nail was placed inside each of the small test tubes. All of the small test tubes were placed in the test tube rack. 10mL of 0% sodium chloride solutions were measured in a small measuring cylinder, and then poured into the test tube.

### Effect of Sodium Chloride (NaCl) on Rust: Lab Explained ...

When Iron nail ( $\text{Fe}$ ) kept in Copper sulphate ( $\text{CuSO}_4$ ) there takes reaction between them. As Iron is more reactive than Copper it can replace Copper from  $\text{CuSO}_4$ .  $\text{CuSO}_4$  is in blue Colour and  $\text{FeSO}_4$  is in light green colour. So, there will be change in the composition and colour of the solution (from blue to light green).

### What changes in colour of iron nails and copper sulphate ...

Posted December 31, 2005. When you place a steel (iron) nail in a  $\text{CuCl}_2$  solution, the iron first displaces the copper and forms  $\text{FeCl}_2$ , because iron is more reactive than copper (take a look at an

activity series of metals).

### **Nails in Copper (II) Chloride - Chemistry - Science Forums**

When an iron nail is dipped in copper sulphate solution, a brown coating of copper is formed on the surface of iron and the colour of copper sulphate solution changes from blue to light green.

### **Reaction of Iron Nails with Copper Sulphate Solution in Water - MeitY OLabs**

The Model: Standard Reduction and Oxidation Potentials In general, the reduction of a metal cation in aqueous solution to the elemental state of the metal can be written as eqn 1.  $M^{m+}(aq) + m e^{-} \rightarrow M(s)$  (1) The reduction half-reaction has a standard reduction potential,  $E^{\circ}$  volts (V).

### **ALE 24. Voltaic Cells and Standard Cell Potentials**

Get an answer for 'Write the ionic equation for the following reaction: Iron+copper(2) sulfate solution.' and find homework help for other Science questions at eNotes

### **Write the ionic equation for the following reaction: Iron ...**

An excess of copper (II) sulfate solution (to make sure that all the iron is reacted) will be added to a known amount of iron. The metallic copper produced will be weighed. These weighings will be used to calculate the moles of iron used and the moles of copper formed. If equation (1) is correct, the moles of copper should equal the moles of iron.

### **STOICHIOMETRY: The Reaction of Iron with Copper (II) Sulfate**

Iron Oxides and Aqueous Chemistry. Rusting of iron consists of the formation of hydrated oxide,  $Fe(OH)_2$  or  $FeO(OH)$ , and is an electrochemical process which requires the presence of water, oxygen and an electrolyte - in the absence of any one of these rusting does not occur to any significant extent.

### **Iron chemistry - wwwchem.uwimona.edu.jm**

13 AIM To study the chemical reaction of an iron nail with aqueous copper sulphate solution; and to study the burning of magnesium ribbon in air. (a) Chemical reaction of iron nail with copper sulphate solution in water.

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