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Chemical Engineering Thermodynamics CHE 3062. M,T,W,R 12:20 to 1:15 Rec Center 3250 (Help Session Wednesdays 3-5 ERC 405) (Alex Mcglasson/Zinhui Sun Help Session Wednesdays 6-9pm ERC 435)

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Introduction to chemical engineering thermodynamics ...

Introduction to Thermodynamics. Conservation of Energy. First Law of Thermodynamics = energy cannot be created or destroyed, only converted between different forms. Example: $\text{CH}_4(\text{g}) + 2\text{O}_2(\text{g}) \rightarrow \text{CO}_2(\text{g}) + 2\text{H}_2\text{O} + \text{energy}$. Reaction gives off energy as heat. Potential energy stored in chemical bonds is lowered. Total energy is unchanged.

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The general study of thermodynamics is similar in that regard. A number of simple relations like this one are coupled together in a way that requires some training to understand. In the absence of chemical reactions, we may also write a mole balance by replacing mass with moles in the balance.

Introductory Chemical Engineering Thermodynamics: Basic ...

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Introduction 1.1 Basic Definitions Thermodynamics is the science that seeks to predict the amount of energy needed to bring about a change of state of a system from one equilibrium state to another. While thermodynamics tells us nothing about the mechanisms of energy transfer, rates of change,

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Chemical Potential, Fugacity, Activity SHR §2.2.1 & Table 2.2 P_i vapor pressure of species i . y_i vapor mole fraction of species i . liquid mole fraction of species i . For phase equilibrium, the chemical potential, fugacity, activity (but not activity coefficient or fugacity coefficient) of any species is equal in all phases.

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