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Pearson Chemistry Chapter 8. In covalent bonding, electrons sharing usually occurs so that atoms attain the electron configurations of noble gases. Combinations of atoms of the nonmetals and metalloids in groups 4A, 5A, 6A and 7A of the periodic table are likely to form covalent bonds. The combined atoms usually acquire a total of eight electrons or an octet.

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Single Covalent Bonds. When only one pair of electrons is shared, such as in a hydrogen molecule, it is a single covalent bond. The shared electron pair is often referred to as the bonding pair. For a hydrogen molecule, shown in Figure 8.4, each covalently bonded atom equally attracts the pair of shared electrons.

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The Covalent Bonding chapter of this Prentice Hall Chemistry Companion Course helps students learn the essential lessons associated with covalent bonding.

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- covalent bond: a bond formed by the sharing of electrons between atoms
- molecule: a neutral group of atoms joined by covalent bonds
- diatomic molecule: a molecule consisting of two atoms

Glossary Terms

Covalent Bonding - Pittsfield High School

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