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Chemistry - Wikipedia

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Acid-base property of the resulting solution from a neutralization reaction depends on the remaining salt products. A salt containing reactive cations undergo hydrolysis by which they react with water molecules, causing deprotonation of the conjugate acids. For example, the acid salt ammonium chloride is the main species formed upon the half neutralization of ammonia in hydrochloric acid solution:

Acid salt - Wikipedia

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Nitrous acid - New World Encyclopedia

Any alkali added to the solution is consumed by the hydronium ions. These ions are mostly regenerated as the equilibrium moves to the right and some of the acid dissociates into hydronium ions and the conjugate base. If a strong acid is added, the conjugate base is protonated, and the pH is almost ...

Buffer solution - New World Encyclopedia

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Atomic Layer Deposition Chemistry: Recent Developments and ...

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Kinetics of Fluorescence Quenching by Electron and H-Atom ...

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