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Chemistry Chapter 13- Study Guide *Predict the "like-dissolves-like" rule of solubility in terms of intermolecular forces *Predict relative solubilities of substances from intermolecular forces. *Explain why gases have relatively low solubilities in water. General rule: like-dissolves-like—the simplest way is to determine the intermolecular forces in the molecule.

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Chemistry Chapter 13 study guide questions. what is pressure? a property of all gas. It exerts pressure on (pushes against) the sides. what are the instruments that are used to measure pressure? barometer, what are the units used to describe pressure? atm, mmHg, torr, psi, pascals.

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Diffusion movement of one substance through another elastic ability of a substance or object to readily spring back to its original size or shape after being squeezed, hit, or subjected to similar disturbances kinetic-molecular theory model that explains the physical properties of gases based on the submicroscopic behavior of gas particles; in mathematical form, it [...]

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192 Study Guide for An Introduction to Chemistry Chapter 13 Map Chapter Checklist Read the Review Skills section. If there is any skill mentioned that you have not yet mastered, review the material on that topic before reading this chapter.

Chapter 13 - Gases - An Introduction to Chemistry

AP Chemistry – Chapter 13, Chemical Equilibrium Study Guide & Ch. 16a Ksp. Students should be able to... Describe a . system in equilibrium . Find the reaction . quotient, Q, (K_{eq}) for a reaction equation . Solve for Q given . initial concentrations or pressures. Using . LeChatlier's principle

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about the mole, and in Chapter 13, quantity is added as an additional measurable gas property. The first three big ideas given below are review items from Chapter 4. If you need to review these items, do so before going on. The Chapter 13 big ideas start with Item 4.

Chapter 13

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CHAPTER 12 378 Chapter 12 Study Guide Study Tip Prioritize Schedule your time realistically. Stick to your deadlines. with ChemASAP If your class subscribes to the Inter-active Textbook with ChemASAP, your students can go online to access an interactive version of the Student Edition and

a self-test. Chapter Resources Print

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13. 16. In your textbook, read about using the ideal gas law to solve for molar mass, mass, or density. Use the following terms below to complete the statements. Each term may be used more than once. mass molar mass Volume per unit (20) divided by the and the of the gas must be known Study Guide The number of moles of a gas is equal to the (17 ...

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Chemistry 11 Final Exam Study Guide Chemistry 11 - Final Exam Study Guide Page 15 When Electronegativities of bonding atoms are the same (as they are in diatomic molecules) or close to the same, they SHARE electrons. Bonds formed when atoms share electrons are called Covalent Bonds .

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