

Chemistry Nail Lab Answers

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Chemistry Nail Lab Answers

In chemistry we are doing the nail lab experiment. (when you add one nail in a solution of copper (II) chloride. for my moles of copper i got .033 for my moles of iron i got .031 i have to create a BCA table. Is the equation $\text{Fe} + \text{CuCl}_2 = \text{Cu} + \text{FeCl}_2$? how would you do the BCA table for this ? because they both have to be on the products side or the reactants side dont they ? i cannot put the ...

Chemistry Nail Lab question !? | Yahoo Answers

We dropped iron nails in copper chloride with water solution. 1.what are any precautions you need to take? 2.explain what an aqueous solution is and give an example from the lab? 3.when you decanted identify the solution and the precipitate? 4.why did you rinse the precipitate with water then hydrochloric acid and again with water? 5.please write the balanced chemical equation for this ...

Iron Nails Copper Chloride Chemistry Lab ... - Yahoo Answers

Modeling Chemistry 1 U7 nail lab v2.0 Name_____ Chemistry - Nail Lab Purpose The purpose is to determine the ratio of copper produced to iron consumed in a replacement reaction. Procedure Day 1 1. Label, then mass a 250 mL beaker. 2. Put 50.0 mL of copper (II) chloride in the beaker. 3. Mass 2 or 3 nails together to $\pm 0.01\text{g}$. 5.

Name Chemistry - Nail Lab Purpose - Mr. Kleinschrodt

View Essay - Unit 7n Nail Lab Reflection Assignment from SCIENCE Chemistry at Montgomery High, Skillman. Unit 7n: Nail Lab Report Objective: The objective of this lab was to observe the amount

Unit 7n Nail Lab Reflection Assignment - Unit 7n Nail Lab ...

[PDF]Free Chemistry Nail Lab Answers download Book Chemistry Nail Lab Answers.pdf Density - Mass | Volume - PhET Interactive Simulations Mon, 20 May 2019 08:28:00 GMT Why do objects like wood float in water? Does it depend on size? Create a custom object to explore the effects of mass and volume on density. Can you discover the relationship?

Chemistry Nail Lab Answers - foundum.com

Iron Nail And Copper Chloride Lab Answers.pdf Free Download Here Chemistry - Nail Lab - Mr Montero <http://www.mrmmontero.com/U6/NailLab.pdf> Put between 6.0 and 8.5 g ...

Iron Nail And Copper Chloride Lab Answers

Chemistry - Nail Lab Purpose The purpose is to determine the ratio of copper produced to iron consumed in a replacement reaction. Procedure Day 1 1. Label, then mass a 250 mL beaker. 2. Put between 6.0 and 8.5 g of copper(II) chloride in the beaker. To do this, move the rider on the balance beam up by the value you chose, add copper(II) chloride to

Chemistry - Nail Lab - Mr Montero

Mass nails after reaction 3.46 g Mass 250 ml beaker + dry copper 116.25 g 1. Determine the mass of copper produced and the mass of iron used during the reaction. Copper: $120.73\text{ g} - 113.85\text{ g} = 2.4\text{ g Cu}$ Iron: $5.59\text{ g} - 3.46\text{ g} = 2.13\text{ g Fe}$ 2. Calculate the moles of copper and moles of iron involved in this reaction.

Mass nails after reaction 346 g Mass 250 ml beaker dry ...

Honors Chemistry Period 1. Search this site. Navigation. Announcements. Online Quizzes. Class Calendar. Powerpoints. Lab Data/Sample Solutions. Reference Materials. Useful Links. Contact Me. Announcements > Nail Lab Report Format posted Jan 17, 2014, 9:16 AM by Adrian De Alba

Nail Lab Report Format - Honors Chemistry Period 1

I've done the three experiments with copper sulfate, copper (II) chloride and copper nitrate. The effect for each of the three salts is very different. First I did experiments with different types of nails, but these results are very inconsistent. Some nails are copper plated quickly, also in nitrate solution, other nails are not copper-plated.

Nails in Copper (II) Chloride - Chemistry - Science Forums

Iron/Copper Nail Lab . In this lab, an iron nail is placed into a solution of copper (II) chloride on the first day. On the second day, it is removed and the copper that has collected on the nail is scraped off. On the third day, final masses are taken and converted to moles for a comparison. Procedure: Day 1. 1. Label, then mass a 250 mL ...

Modeling Chemistry Unit 7 : simplebooklet.com

Chemical Reactions: Particles and Energy ... Nail Lab Purpose The purpose is to determine the ratio of copper produced to iron consumed in a reaction. Procedure Day 1 1. Label, then mass a 250 mL beaker. 2. Put between 6.0 and 8.5 g of copper(II) chloride in the beaker. 3. Add about 50 mL distilled water to the beaker.

Unit 6 - Chemical Reactions: Particles and Energy

Pre-Lab discussion for the Nail Lab. Project Mc2 DIY LIP BALM Lab! Mix & Make Your Own Lip Balm Sticks! Yummy Scents!

The Nail Lab

A Claim, Evidence, Reasoning (CER) is a short writing piece that allows the students to succinctly discuss their lab activities. The CER must always begin with a Claim (C). The Claim should be a concise statement that summarizes the main point or idea of the laboratory experiment. For instance, "Armada is a great place to live."

Claim, Evidence, Reasoning aka CER - Chemistry

In a zinc-copper voltaic cell, Zinc is oxidized and Copper is reduced, making Zinc the reduction agent and Copper the oxidizing agent. The Zinc loses two electrons becoming Zinc^{+2} as Copper+2 gains two electrons becoming Copper in its elemental form. In this cell, the zinc strip

Electrochemistry Lab Report(s) by Elijah Harris on Prezi

This video shows the setup of the long term reaction of CuCl_2 and an iron nail. www.dlt.ncssm.edu Please attribute this work as being created by the North Carolina School of Science and Mathematics.

Copper Sulfate and Iron Nail Lab: Part 1

So, we did a lab in class called Nail Lab. We had to get three iron nails and put them into copper (II) chloride solution. They had instant rust. We measured and all. But my questions are the conclusion questions: 1) Why did the reaction stop? Which reactant was used up? How do you know? 2) Describe what was happening to the atoms of iron and copper during the reactions.

I have questions about Chemistry? | Yahoo Answers

The mass of two iron nails: At this point, the nails were put into the solution of copper (II) sulfate. Copper started to build up on the surface of the nail, as the iron nail reacted with the copper (II) sulfate. After the reaction was finished, the copper was scraped off the nails, collected, and washed. Mass of iron nails after reaction:

Chemistry Nail Lab Answers

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