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Chapter Test B A. Matching Match each term in Column B with the correct description in Column A. Write the letter of the correct term on the line. B. Multiple Choice Choose the best answer and write its letter on the line. ____ 9. How many oxygen atoms are in 10 formula units of $\text{Al}_2(\text{SO}_4)_3$?
a. 10 atoms O c. 70 atoms O b. 7 atoms O d. 120 atoms O ____ 10.

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CHEMICAL QUANTITIES CHAPTER B TEST ANSWERS

CHAPTER 10: Chemical Quantities BASICS: • The basic unit that is used to determine the amount of a chemical substance is called a mole • A mole(mol) of a substance is equivalent to 6.02×10^{23} particles of that substance • The mole was founded by a scientist named Avagadro, and he decided to use the

CHAPTER 10: Chemical Quantities

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Chapter 10 Chemical Quantities 241 Section Review Objectives • Relate Avogadro's number to a mole of a substance • Calculate the mass of a mole of any substance • Describe methods of measuring the amount of something • Compare and contrast the atomic mass of an element and its molar mass Vocabulary Key Equations • moles representative particles • representative particles moles

Objectives Vocabulary Key Equations

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Percent Composition and Chemical Formulas 11) A compound analyzed in a chemistry laboratory consists of 5.34 g of carbon, 0.42 g of hydrogen, and 47.08 g of chlorine.

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