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Chemistry – Chapter 9 Assessment Answers. 9.1 Assessment. 3. The positive charge of a group A metal is equal to its group number. The charge of a group A nonmetal is the group number minus 8. 4. from the number of electrons lost. 5. -ite or -ate. 6. Group 1A metals, 1+; Group 3A (aluminum), 3+ Group 5A nonmetals, 3-

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e. molar mass of  $\text{Fe}(\text{NO}_3)_3 = 241.9 \text{ g}$  272 g  $\text{Fe}(\text{NO}_3)_3$  1 mol 241.9 g = 1.12 mol  $\text{Fe}(\text{NO}_3)_3$  f. molar mass of  $\text{FeSO}_4 = 151.9 \text{ g}$  12.7 mg  $\text{FeSO}_4$  1 g 1000 mg 1 mol 151.9 g = 8.36  $10^{-5}$  mol  $\text{FeSO}_4$  4 g. molar mass of  $\text{LiOH} = 23.95 \text{ g}$  6.91 103 g  $\text{LiOH}$  1 mol 23.95 g = 288.5 = 289 mol  $\text{LiOH}$  14. a. molar mass of  $\text{CaCO}_3 = 100.1 \text{ g}$  2.21  $10^{-4}$  mol  $\text{CaCO}_3$  100.1 g 1 mol = 0 ...

**Chapter 9 Chemical Quantities - Francis Howell High School**

Solutions Manual Chemistry: Matter and Change • Chapter 9 141 Section 9.1 Reactions and Equations pages 282–288 ... Section 9.1 Assessment page 288 7. Explain why it is important that a chemical ... Explain your answer.  $3\text{Ni} + 2\text{AuBr}_3 \rightarrow 3\text{NiBr}_2 + 2\text{Au}$  The reaction does occur because nickel is more

**Chemical Reactions Chemical Reactions**

Chapter 9 Practice Test - Naming and Writing Chemical Formulas Matching Match each item with the correct statement below. Match each item with the correct statement below. a. monatomic ion f. cation b. acid g. binary compound c. base h. anion d. law of definite proportions i. polyatomic ion e. law of multiple proportions \_\_\_\_ 1.

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