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ChemActivity 23, Dipole Moment ChemActivity 23. 1. CO. HI. CIF. 2. See Model 1 of CA 16. The dipole moment is zero. The center of positive charge is at the nucleus of the N atom.

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1. The dipole moment of I =  $2(2d \cdot 1) = 4d \cdot 1$ . The dipole moment of II =  $1(3d \cdot 1) = 3d \cdot 1$ . The dipole moment of I is larger.

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Chemactivity 27 1 no 2 h 2 ol because the molecules. The ketone and the alcohol have dipole-dipole forces. Only the alcohol has hydrogen bonding—which increases the intermolecular forces substantially. ChemActivity 28 1. This is the end of the preview. Sign up to access the rest of the document.

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### **Chemactivity 50 Answers The Electrochemical Cell**

Summing all bond dipole moments for a structure gives the molecule's net dipole moment. If this net dipole moment is zero, the molecule is overall nonpolar. (We will not be doing fancy vector calculations in this course, but be on the lookout for symmetrical molecules whose bond dipole

arrows cancel each other.)

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