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Justify your answer. Yes. Addition of an excess of 0.20 M AgNO_3 (aq) will precipitate all of the I^- ion present in the solution because AgI is insoluble, as evidenced by its low value of K_{sp} . 1 point is earned for the correct answer with a valid justification. (ii) The student only has access to one KI

AP Chemistry 2014 Scoring Guidelines - College Board

Questions 4–7 are short free-response questions that require about 7 minutes each to answer and are worth 4 points each. Write your response in the space provided following each question.

A P Chemistry 2014 Free-Response Questions

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1 ISC Specimen Question Paper 2014 CHEMISTRY PAPER – 2 (PRACTICAL) (Three hours) (Candidates are allowed additional 15 minutes for only reading the paper. They must NOT start writing during this time.)----- ALL ANSWERS MUST BE WRITTEN IN THE ANSWER BOOKLET PROVIDED

CHEMISTRY PAPER – 2 - cisce.org

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Physical Setting/Chemistry Regents Examinations

DRAFT RELEASED FRQ ANSWERS AP CHEMISTRY 2014 (numbers in parentheses and red, represent guesstimate points) Question 1 (a) (i) $\text{Pb}^{2+}(\text{aq}) + 2\text{I}^{-}(\text{aq}) \rightarrow \text{PbI}_2(\text{s})$ (1) (ii) Potassium and nitrate ions are in solution at the beginning of the reaction and at the end, do not

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