

Chemquest 5 Specific Heat Answer Key

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Chemquest 5 Specific Heat Answer

13. A beaker containing 250.0 g of water is heated with 1500.0 J of heat energy. If the temperature of the water changed from 22.0000°C to 23.4354°C, what is the specific heat of water? $\Delta T = 23.4354^\circ\text{C} - 22.0000^\circ\text{C} = 1.4354^\circ\text{C}$ 14. Heat energy equal to 25,000 J is applied to a 1200 g brick whose specific heat is 2.45 J/g°C.

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CHEMQUEST 5 SPECIFIC HEAT ANSWER KEY PDF

: Specific Heat In questions 1 and 2 above you probably recognized that the temperature change of a substance depends on the mass of the substance. You also have probably experienced the fact that different substances heat at different rates as was discussed in questions 3 and 4 above.

ChemQuest 5 - New Castle Community School Corp.

This is video made for a chem project and it is intended to educate other on specific heat.

Chemquest 5: Specific Heat

16 ChemQuest 5 Name: _____ Date: _____ Hour: _____ Information: Heat and Temperature When a substance absorbs heat energy, the temperature of the substance increases. There are a number of factors (such as mass of the substance) that affect how much the temperature of a substance changes. Critical Thinking Questions 1. Consider two pots of water.

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Chemquest 10 - Specific Heat - Course Hero

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The specific heat of water is 1 cal/g°C. If a 3.1g ring is heated using 10.0 calories, its temperature rises 17.9°C. Calculate the specific heat capacity of the ring. The temperature of a sample of water increases from 20°C to 46.6°C as it absorbs 5650 calories of heat.

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Set the parameters to the following: mass of metal = 120 g temp of metal = 220°C mass of water

= 50 g temp of water = 20°C 3. Record the specific heat of the metal and the water. 4. Click START. 5. Record the final temperature of the water and the metal.

Energy, Heat, Work, Temperature Web Quest - mrmchem.com

ChemQuest #1: Numbers in Chemistry 1 ChemQuest #2: Significant Figures 5 ChemQuest #3: Changes and Matter 10 ChemQuest #4: Phase Diagrams 14 ChemQuest #5: Specific Heat 16 ChemQuest #6: Energy Changes in Matter 19 ChemQuest #7: Density, Mass and Weight 22 ChemQuest #8: Structure of the Atom 24 ChemQuest #9: Average Atomic Mass 28

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Unit 7: Energy and Thermodynamics. STUDY. PLAY. Energy. is the ability to do work. Law of Conservation of Energy. the energy of the universe is constant. ... The temperature of the water rises to 18.3°C. Calculate the specific heat capacity of the metal assuming that all the heat lost by the metal is gained by the water. (p.43) 0.244 J/g°C.

Unit 7: Energy and Thermodynamics Flashcards | Quizlet

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