

## *Chapter 18 Review Chemical Equilibrium Answers Section 3*

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Modern Chemistry 149 Chemical Equilibrium CHAPTER 18 REVIEW Chemical Equilibrium SECTION 4 SHORT ANSWER Answer the following questions in the space provided. 1. Match the solution type on the right to the corresponding relationship between the ion product and the K<sub>sp</sub> for that solution, listed on the left. \_\_\_\_ The ion product exceeds the K<sub>sp</sub>.

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sometimes referred to as the chemical-equilibrium expression. The H<sub>2</sub>, I<sub>2</sub>, HI Equilibrium System Consider the reaction between H<sub>2</sub> and I<sub>2</sub> vapor in a sealed flask at an elevated temperature. The rate of reaction can be followed by observing the rate at which the violet color of the iodine vapor diminishes, as [C]<sup>x</sup>[D]<sup>y</sup> [A]<sup>n</sup>[B]<sup>m</sup> 556 CHAPTER 18

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### 18 Chemical Equilibrium - Clark County School District

1 Chapter 18...Reaction Rates and Equilibrium RATES OF REACTION Speeds of chemical reactions can be extremely fast or extremely slow. The rate is a measure of the speed of any change that occurs within an interval of time (which could range from fractions of a second

### Chapter 18 Reaction Rates and Equilibrium

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CHAPTER 18 REVIEW Chemical Equilibrium MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Consider the following equilibrium equation:  $\text{N}_2(\text{g}) + 2\text{O}_2(\text{g}) \rightleftharpoons 2\text{NO}_2(\text{g})$ ;  $\Delta H = 33 \text{ kJ/mol}$  At equilibrium, which reaction is favored when

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CHAPTER 17 EQUILIBRIUM: THE EXTENT OF CHEMICAL REACTIONS ... The equilibrium constant for the reverse reaction is the inverse of the original constant. When a ... 17.18 The concentration of solids and pure liquids do not change, so their concentration terms are not written in the

### CHAPTER 17 EQUILIBRIUM: THE EXTENT OF CHEMICAL REACTIONS

560 Chapter 18 Chemical Equilibrium H 2 Time Concentration NH 3 N 2 Figure 18-1 The concentrations of the reactants ( $\text{H}_2$  and  $\text{N}_2$ ) decrease at first while the concentration of the product ( $\text{NH}_3$ ) increases. Then, before the reactants are used up, all concentrations become constant.

### Chapter 18: Chemical Equilibrium - Jayne Heier

Chapter 18 "Reaction Rates and Equilibrium" ... chemical), it is part of the equilibrium • so cooling an exothermic reaction will produce more product, and heating it would shift the reaction to the reactant side of the equilibrium: ... Section 18.3 Solubility Equilibrium ...

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