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Chapter 15 Chemical Equilibrium

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water molecules: if the attraction of the water molecules for the positive ion is stronger than the attraction of the negative ions near it in the crystal, the ion leaves the crystal and enters solution. After entering solution, the dissolved ion is surrounded completely by water molecules, which tends to prevent the ion from reentering the ...

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and water molecules are also hydrogen bonds (Figure 15.4). Because the attractions between the particles are so similar, the freedom of movement of the ethanol molecules in the water solution is about the same as their freedom of movement in the pure ethanol. The same can be said for the water. Because of this 576 Chapter 15 Solution Dynamics

Chapter 15 Solution Dynamics - An Introduction to Chemistry

Water and Aqueous Systems Bonding and interactions 15.1 Water and its Properties essential Understanding The bonding between water molecules differs when water is in different states, giving it different properties. ... 0132525887_CHEM_WKBK_CH 15.indd 214 4/1/10 10:17:25 AM.

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AP Chemistry Chapter 15 Review Multiple-choice exercise. Choose the correct answer for each question. ... calculate the pH of a buffer solution made by mixing 0.225 mol of HNO₂ and 0.450 mol of NaNO₂ in enough water to make 0.400 liter of solution. ? 3.51 ? 2.98 ? 3.90 ?

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Chemistry \ chemistry ch 14.15.16.18. chemistry ch 14.15.16.18. The one aldehyde and the one ketone with a molecular formula of C₃H₆O are stereoisomers ... They are highly polar molecules. They boil at higher temperatures. Lower members are soluble in water because they can form H-bond with water. Higher members are insoluble in water due to ...

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