

Chapter 19 Modern Chemistry Answers

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Modern Chemistry 71 Chapter Test Name Class Date Chapter Test B, continued 17. A substance combines with oxygen, releasing a large amount of energy as heat and light, in a(n) . 18. The decomposition of a substance by an electric current is called . 19. A(n) orders the elements by the ease with which they undergo certain chemical reactions. 20.

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Modern Chemistry 4 Oxidation-Reduction Reactions CHAPTER 19 REVIEW Oxidation-Reduction Reactions SECTION 2 SHORT ANSWER Answer the following questions in the space provided. 1. ____ All of the following should be done in the process of balancing redox equations except (a) adjusting coefficients to balance atoms.

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Strategy The key to this assignment is learning how to use Table 19.2. Practice by working the end-of-chapter questions. Answer Questions, Exercises, and Problems 16–20. Include Questions, Exercises, and Problems 21 if you include Section 19.7 in this assignment. Check your answers with those at the end of the chapter.

Chapter 19

CHAPTER 1 REVIEW Matter and Change MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Classify each of the following as a homogeneous or heterogeneous substance. a. sugar d. plastic wrap b. iron filings e. cement sidewalk c. granola bar 2. For each type of investigation, select the most appropriate branch of chemistry from the following

mc06se cFMsr i-vi

Modern Chemistry 35 Chapter Test Name Class Date Chapter Test B, continued Write the orbital notation for the following elements in the space provided. 37. lithium, atomic number 3 38. carbon, atomic number 6 39. neon, atomic number 10 PART VI Write the answers to the questions on the line to the left, and show your work in the space provided.

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Modern Chemistry 3 Oxidation-Reduction Reactions CHAPTER 19 REVIEW Oxidation-Reduction Reactions SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1.

____ All the following equations involve redox reactions except (a) $\text{CaO} + \text{H}_2\text{O} \rightarrow \text{Ca(OH)}_2$. (b) $2\text{SO}_2 + \text{O}_2 \rightarrow 2\text{SO}_3$. (c) $2\text{HgO} \rightarrow 2\text{Hg} + \text{O}_2$. (d) $\text{SnCl}_4 + 2\text{FeCl}_2 \rightarrow 2\text{FeCl}_3 + \text{SnCl}_2$. 2.

CHAPTER 19 REVIEW Oxidation-Reduction Reactions

Modern Chemistry 7 Oxidation-Reduction Reactions CHAPTER 19 REVIEW Oxidation-Reduction Reactions SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. For each of the following, identify the stronger oxidizing or reducing agent. (Refer to Figure 3.2 of the text.) ____ a. Ca or Cu as a reducing agent

CHAPTER 19 REVIEW Oxidation-Reduction Reactions

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Modern Chemistry 1 Solutions CHAPTER 12 REVIEW Solutions Teacher Notes and Answers Chapter 12 SECTION 1 SHORT ANSWER 1. c 2. a 3. b 2. a. alcohol b. water c. the gels 3. The mixture is a colloid. The properties are consistent with those reported in Table 3 on page 404 of the text. The particle size is small, but not too small, and the mixture

CHAPTER 12 REVIEW Solutions - Weebly

CHAPTER 6 REVIEW Chemical Bonding SECTION 4 SHORT ANSWER Answer the following questions in the space provided. 1. b In metals, the valence electrons are considered to be (a) attached to particular positive ions. (c) immobile. (b) shared by all surrounding atoms. (d) involved in covalent bonds.

6 Chemical Bonding - Effingham County Schools / Overview

Holt McDougal Modern Chemistry 1 Chapter Test Assessment Chapter Test A Chapter: The Periodic Law Use the periodic table below to answer the questions in this Chapter Test. In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question. ____ 1.

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