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### **Chapter Summaries - Chemistry Matter and Change**

Solutions Manual Chemistry: Matter and Change • Chapter 12 239 CHAPTER 12 SOLUTIONS MANUAL Section 12.3 Liquids and Solids pages 415–424 Section 12.3 Assessment page 424 18. Contrast the arrangement of particles in solids and liquids. The particles are closer together in solids than in liquids because of intermolecular attractions.

### **States of Matter - Weebly**

Chemistry Notes & Handouts First semester. TEXTBOOK: Florida Chemistry - Matter and Change. Columbus, OH: McGraw-Hill, 2019. NOTE ON OPENING LINKS: If a link does not open, right-click on the link and save it to a location such as your desktop or documents folder. Then open it from the saved location. "REFRESHING" THE PAGE:

### **Kwanga.net - Chemistry - Notes and Handouts**

Chemistry: Matter and Change Teacher Guide and Answers 8 molar mass  $\text{KCO}_3$  99.11 g/mol  $\text{KCO}_3$  n molar mass of molecular formula/molar mass of empirical formula 198.22 g/mol/99.11 g/mol  $2(\text{KCO}_3)_n$  The molecular formula of the compound is  $\text{K}_2\text{C}_2\text{O}_6$ . Section 10.5 The Formula for a Hydrate 1. hydrate 2. hydration 3. water molecules

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