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Balancing Chemical Equations -Answer Key Balance the equations below: 1)  $1 \text{ N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3$  2)  $2 \text{ KClO}_3 \rightarrow 2 \text{ KCl} + 3 \text{ O}_2$  3)  $2 \text{ NaCl} + 1 \text{ F}_2 \rightarrow 2 \text{ NaF} + 1 \text{ Cl}_2$  4)  $2 \text{ H}_2 + 1 \text{ O}_2 \rightarrow 2 \text{ H}_2\text{O}$  5)  $1 \text{ Pb(OH)}_2 + 2 \text{ HCl} \rightarrow 2 \text{ H}_2\text{O} + 1 \text{ PbCl}_2$  6)  $2 \text{ AlBr}_3 + 3 \text{ K}_2\text{SO}_4 \rightarrow 6 \text{ KBr} + 1 \text{ Al}_2(\text{SO}_4)_3$

### **Balancing Chemical Equations Answer Key**

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Stoichiometry: Mixed Problems (KEY) 1)  $N_2 + 3H_2 \rightarrow 2NH_3$  What volume of  $NH_3$  at STP is produced  
if 25.0 of  $N_2$  is reacted with an excess of  $H_2$ ? 3 3 3 2 3 2 2 40.0L  $NH_3$  1mol  $NH_3$  22.4L  $NH_3$  1mol N  
2mol  $NH_3$  28.0g N 25.0g N 1mol N  $\times \times \times = 2$ )  $2KClO_3 \rightarrow 2KCl + 3O_2$  If 5.0g of  $KClO_3$  is decomposed,  
what volume of  $O_2$  is produced at STP? 2

**Stoichiometry: Mixed Problems (KEY)**

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8-14a,b Mixed Problems--Mole-Mole and Mole-Mass wkst- ...

**Answer Key Chemistry If8766 Moles And Mass**

Chemistry IF8766 62 . STOICHIOMETRY: MASS-MASS PROBLEMS 1.  $2KClO_3 \rightarrow 2KCl + 3O_2$  Name How many  
grams of potassium chloride are produced if 25 g of potassium chlorate decompose? / 25 ken How  
many grams of hydrogen are necessary to react completely with 50.0 g of nitrogen in the above  
reaction?

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