

## *Chemical Quantities Answers Core Teaching Resources*

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### **CORE TEACHING RESOURCES ANSWERS CHEMICAL QUANTITIES**

Chapter 10 Chemical Quantities. mrs. wright. STUDY. PLAY. Terms in this set (...) standard temperature and pressure. 0 C and 101.3 kPa. the volume occupied by a mole of any gas at STP (22.4 L) molar volume. the mass in amu of an atom of an element. atomic mass. the number of representative particles in a mole of a substrate.

### **Chapter 10 Chemical Quantities Flashcards | Quizlet**

Chapter 10 Chemical Quantities Chapter Quiz Answer Key.  
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### **Chapter 10 Chemical Quantities Chapter Test A Answer Key**

Chemical Quantities 287 Print • Guided Reading and Study Workbook, Section 10.1 • Core Teaching Resources, ... Answers might include the following: A mole contains Avogadro's number of particles. The mass of one mole of a substance is its molar mass, which is found

### **10.1 The Mole: A Measurement of Matter 10**

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### **10.3 Percent Composition and Chemical Formulas 10**

Chapter 10 Chemical Quantities 241 Section Review Objectives • Relate Avogadro's number to a mole of a substance • Calculate the mass of a mole of any substance • Describe methods of measuring the amount of something • Compare and contrast the atomic mass of an element and its molar mass Vocabulary Key Equations • moles representative particles • representative particles moles

### **Objectives Vocabulary Key Equations**

CHAPTER 10: Chemical Quantities BASICS: • The basic unit that is used to determine the amount of a chemical substance is called a mole • A mole(mol) of a substance is equivalent to  $6.02 \times 10^{23}$  particles of that substance • The mole was founded by a scientist named Avagadro, and he decided to use the

### **CHAPTER 10: Chemical Quantities**

.ppt of Chapter 10 - Chemical Quantities.ppt of Chapter 10 - Chemical Quantities ... and has  $6.02 \times 10^{23}$  atoms. What is the mass of a mole of a compound? To answer this you must know the formula of the compound. The formula tells you the number of atoms of each element in a

representative particle of that compound. ... Teaching Techniques ...

### Chapter 10 - Chemical Quantities - SlideShare

256 Core Teaching Resources Chapter Test B ... B. Multiple Choice Choose the best answer and write its letter on the line. \_\_\_\_ 9. How many oxygen atoms are in 10 formula units of  $\text{Al}_2(\text{SO}_4)_3$ ?  
a. 10 atoms O c. 70 atoms O b. 7 atoms O d. 120 atoms O ... CHEMICAL QUANTITIES 10  
05\_CTR\_ch10 7/9/04 3:29 PM Page 256.

**05 CTR ch10 7/9/04 3:29 PM Page 256 Chapter 10 Review ...**

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