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14. The mass of is about 14 amu. Does this agree with what you determined in questions 11 and 13? 15. The charge (in the upper right hand corner of the element symbol) is -1 for a fluorine ion. Why isn't it +1 or some other number? 16. What is the charge on every atom? Why is this the charge? 17. How do you determine the charge on an ion? 18.

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The atom is the fundamental unit of matter that makes up all living and non-living things. Explore the history of atomic theory and the experimental data that lead to these incremental changes. Build your chemistry toolkit by learning to read and speak the language of chemistry, interpret the periodic table, and apply key quantitative skills.

AP Chemistry | Practice Questions | Albert

X = chemical symbol of the element. Z = "atomic number" A = "mass number", , and are notations that represent . isotopes. of carbon., and are notations that represent . isotopes. of hydrogen. The part of the atom where the protons and neutrons are is called the . nucleus. Critical Thinking

Questions. 1.

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#20 Introduction to the Mole Golda Steiner, Park Ridge High School, Park Ridge, NJ Introduction ... This unit can be used to introduce the mole concept in any chemistry class, grades 9-12. ... Answer questions in Analysis and Interpretations. 3. Record masses your group measured on a transparency summarizing class data.

#20 Introduction to the Mole - Terrific Science

In this example, propane (C₃H₈) is the “excess reactant” because after the reaction there will be some of it left over. It is important to remember that everything in a chemical equation is related by

mole ratios. If you only know the mass (grams) of the substances, you need to convert to moles.
Critical Thinking Questions

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