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In a chemical reaction, one or more reactants change into one or more products. Figure 11.1a shows the ingredients for making leavened bread— flour, salt, yeast, and water. Chemical reactions take place when the ingredients are mixed together and heated in the oven. Figure 11.1b shows the product—a loaf of bread.

11.1 Describing Chemical Reactions 11

These are the reactants in a chemical reaction that takes place when the ingredients are mixed together and heated in the oven. Figure 11.1b shows the product—a loaf of bread. The four ingredients (reactants) were changed to bread (the product). Like all chemical reactions, baking bread involves changing substances.

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reactants products yields, gives, or reacts to produce false acetylene oxygen JKL carbon dioxide water mercury(II) oxide JKL mercury oxygen A chemical equation is a representation of a chemical reaction; the formulas of the reactants (on the left) are connected by an arrow with formulas of the products (on the right).

SECTION 11.1 DESCRIBING CHEMICAL REACTIONS (pages 321-329)

11.2 Types of Chemical Reactions> 32 Double-Replacement Reactions This is an example of a double-replacement reaction, which is a chemical change involving an exchange of positive ions between two compounds. •Double-replacement reactions are also referred to as double-displacement reactions.

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Chapter 11

Section 11.1 Assessment. Describe the steps in writing a balanced chemical equation. Write the skeleton equation for the following reactions: Heating solid copper(II)sulfide in the presence of oxygen gas produces pure copper and sulfur dioxide gas. Iron metal and chlorine gas react to form solid iron(III)chloride.

Chapter 11: Chemical Reactions - Henry County Schools ...

•Go Online, Section 11.2 11.2 FOCUS Objectives 11.2.1 Describe the five general types of reactions. 11.2.2 Predict the products of the five general types of reactions. Guide for Reading Build Vocabulary Graphic Organizers Have students draw concept maps entitled "Types of chemical reactions." Have them include all the vocabulary terms ...

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WS 11.1 Describing Chemical Reactions Mrs Rond 2 17. Fill in this table + $\frac{1}{4}$ Used in place of $\frac{1}{4}$ for reversible reactions (s) (l) (g) (aq) Indicates that heat is supplied to the reaction A formula written above or below the yield sign indicates its use as a catalyst (in this example platinum) 18.

College Chemistry Name WS on Section 11.1 Describing ...

Most chemical reactions are actually equilibria that are pushed in one direction or the other depending on the presence of other reactants, catalyzers and external conditions. ... Describing Chemical Reactions (chapter 11)? ... Is it fair for a chemistry professor to forbid calculators for the exams?

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