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Activity A: Balancing equations Get the Gizmo ready: • Check that the Combination reaction is selected and that all coefficients are set to one. (The coefficients are the numbers in the boxes.) Introduction: The equation  $\text{H}_2 + \text{O}_2 \rightarrow \text{H}_2\text{O}$  is unbalanced because there are two oxygen atoms on the reactants side of the equation, and only one on the products side of the equation.

### **Student Exploration: Balancing Chemical Equations**

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**Chemical Equations Gizmo : ExploreLearning**

This is the fourth lesson in the chemical reaction unit that covers HS-PS1-7, the uses of mathematical representations to support the claim that atoms, and therefore mass, are conserved during a chemical reaction; and DCI-PS1.B, the fact that atoms are conserved, together with knowledge of the chemical properties of the elements involved, can be used to describe and predict chemical reactions.

**Tenth grade Lesson Gizmo: Modeling Chemical Reactions**

Balancing Chemical Equations Balancing chemical equations gizmo answers. Balance and classify five types of chemical reactions: synthesis, decomposition, single replacement, double replacement, and combustion. While balancing the reactions, the number of atoms on each side is presented as visual, histogram, and numerical data Balancing chemical equations gizmo answers.

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Subcategories. A chemical equation represents a reaction--a change of something from one state to another. For example, the burning of hydrogen to form water:  $2\text{H}_2 + \text{O}_2 \Rightarrow 2\text{H}_2\text{O}$  A chemical formula represents a compound, such as  $\text{H}_2\text{O}$  for water. However, muddy water is not a compound. It is water with a...

**5,159 Questions Asked In Chemical Equations - Answers**

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Balancing Chemical Equations Answer Key Vocabulary: coefficient, combination, compound, decomposition, double replacement, element, molecule, product, reactant, single replacement, subscript Prior Knowledge Questions (Do these BEFORE using the Gizmo.) [Note: The purpose of these questions is to activate prior knowledge and get students thinking.

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