Composition Of Potassium Chlorate Lab Answers

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Composition Of Potassium Chlorate Lab

Pre-laboratory Assignment: The Composition of Potassium Chlorate. In Part A of this lab, you will analyze a sample of potassium chlorate to determine the mass percent of oxygen present in it. To perform the analysis, you will decompose the potassium chlorate by heating it.

5: The Composition of Potassium Chlorate (Experiment ...

Page 1 of 4. The Composition of Potassium Chlorate. Objectives The objectives of this laboratory are: a) To experimentally determine the mass percent of oxygen in the compound potassium chlorate (KClO3) via the thermal decomposition of a sample of potassium chlorate.

The Composition of Potassium Chlorate

The purpose of this lab was to calculate the percent composition by mass of oxygen in potassium chlorate. Procedure: 1.The mass of the crucible and the cover was measured. 2.KClO3 was added to the crucible. 3.The mass of the crucible, cover and KClO3 was measured. 4.MnO2 was added in the crucible.

Percent Composition by Mass of Oxygen in Potassium ...

In this lab, students will test the percent composition of KClO 3. "Potassium chlorate is one of several compounds of oxygen that are easily decomposed by heating. Potassium chlorate will decompose into oxygen and potassium chloride.

Lab: Percent Composition of KClO3 | Curriki

The decomposition of potassium chlorate (KClO3) is used as a source of oxygen in the laboratory. How much potassium chlorate is needed to produce 10.2 mol of oxygen? Answer in units of mol asked by DEANA on December 1, 2014; Chemistry. It says 8.

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Mass approximately 2 g of potassium chlorate. Transfer to crucible. Determine the mass of the crucible, contents, and cover. Place the crucible on the tripod, place the cover in place, and heat very gently for 2 minutes. At 2-minute intervals gradually increase the temperature until the crucible is heated strongly.

lab - Angelfire

Page 1 of 4 The Composition of Potassium Chlorate Objectives The objectives of this laboratory are: a) To experimentally determine the mass percent of oxygen in the compound potassium chlorate (KClO 3) via the thermal decomposition of a sample of potassium chlorate. b) To qualitatively demonstrate that the residue resulting from the decomposition of potassium chlorate is potassium chloride.

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