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A conjugate base is the species that results when an acid donates a hydrogen ion. The acid HX donates its hydrogen ion and becomes the conjugate base X⁻. In the reaction shown above, the hydronium ion (H₃O⁺) is the conjugate acid of the base H₂O. The X⁻ ion is the conjugate base of the acid HX.

Chapter 18: Acids and Bases - FCPS

is the base) 9.2 Bronsted-Lowry Acids and Bases. Bronsted-Lowry acid-base reactions involve a transfer of a proton (H⁺) from an acid to a base. o Bronsted - Lowry Acid: proton donor (loses H⁺) o Bronsted - +Lowry Base: proton acceptor (gains H⁺) □ This definition allows for a broader range of bases to be included.

Chapter 9 Acids & Bases - Glendale Community College

Solutions, Acids, and Bases 5 Name Date Class Lab Preview Directions: Answer these questions before you begin the lab. 1. Why should you wear safety glasses while you are doing this lab? 2. Why should a separate teaspoon be used for each substance?

Solutions, Acids, and Bases - Weebly

Study Guide and Reinforcement 3 ANSWER KEY 7. opposes the motion of objects that move through the air, is affected by speed, size, and shape 8. net force 9. microwelds 10. rolling 11. air resistance 12. acceleration 13. sliding 14. parachute 1. Gravity is a force that every object in the

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