

Chemistry Chapter 9 Stoichiometry Answers

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- particles - the relative amounts of atoms, ions, unit formulas or molecules in various reactants or products ... Let's try a real chemistry ...

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e. molar mass of $\text{Fe}(\text{NO}_3)_3 = 241.9 \text{ g}$ 272 g $\text{Fe}(\text{NO}_3)_3$ 1 mol 241.9 g = 1.12 mol $\text{Fe}(\text{NO}_3)_3$ f. molar mass of $\text{FeSO}_4 = 151.9 \text{ g}$ 12.7 mg FeSO_4 1 g 1000 mg 1 mol 151.9 g = 8.36 10^{-5} mol FeSO_4 4 g. molar mass of $\text{LiOH} = 23.95 \text{ g}$ 6.91 103 g LiOH 1 mol 23.95 g = 288.5 = 289 mol LiOH 14. a. molar mass of $\text{CaCO}_3 = 100.1 \text{ g}$ 2.21 10^{-4} mol CaCO_3 100.1 g 1 mol = 0 ...

Chapter 9 Chemical Quantities - Francis Howell High School

Chapter 9: Standard Review Worksheet 1. Answers will vary. An example is included below: $2\text{H}_2\text{O}_2(\text{aq}) \rightarrow 2\text{H}_2\text{O}(\text{l}) + \text{O}_2(\text{g})$ This describes the decomposition reaction of hydrogen peroxide. Microscopic: Two molecules of hydrogen peroxide (in aqueous solution) decompose to produce two molecules of liquid water and one molecule of oxygen gas.

Chapter 9: Standard Review Worksheet

PEP - Chemistry 2 8. Interpret the following equation in terms of (a) relative numbers of representative particles, (b) numbers of moles, and (c) masses of reactants and products.

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Stoichiometry Worksheet #1 Answers 1. Given the following equation: $2 \text{C}_4\text{H}_{10} + 13 \text{O}_2 \rightarrow 8 \text{CO}_2 + 10 \text{H}_2\text{O}$, show what the following molar ratios should be. a. $\text{C}_4\text{H}_{10} / \text{O}_2$ b. O_2 / CO_2 c. $\text{O}_2 / \text{H}_2\text{O}$ d. $\text{C}_4\text{H}_{10} / \text{CO}_2$ e. $\text{C}_4\text{H}_{10} / \text{H}_2\text{O}$ 2. Given the following equation: $2 \text{KClO}_3 \rightarrow 2 \text{KCl} + 3 \text{O}_2$ a. How many moles of O_2 can be produced by ...

Stoichiometry Worksheet #1 Answers

Modern Chemistry 73 Stoichiometry CHAPTER 9 REVIEW Stoichiometry SECTION 1 SHORT ANSWER Answer the following questions in the space provided. 1. ____ The coefficients in a chemical equation represent the (a) masses in grams of all reactants and products. (b) relative number of moles of reactants and products.

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