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pH and pOH CONTINUED 1. 0.01 M HCl $[H^+] = 0.01M$ $pH = -\log(0.01) = 2$ 2. 0.0010 M NaOH $[OH^-] = 0.0010M$ $pOH = -\log(0.0010) = 3 \rightarrow pH = 11$ 3. 0.050 M NaOH

pH and pOH CONTINUED - newburyparkhighschool.net

pH and pOH The pH of a solution indicates how acidic or basic that solution is. pH range of 0-7 acidic 7 neutral 7-14 basic Since $[H^+][OH^-] = 10^{-14}$ at $25^\circ C$, if $[H^+]$ is known, the $[OH^-]$ can be calculated and vice versa. $pH = -\log [H^+]$ So if $[H^+] = 10^{-6} M$, $pH = 6$ $pOH = -\log [OH^-]$ So if $[OH^-] = 10^{-8}$, $pOH = 8$ Together, $pH + pOH = 14$.

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Part 2: For each of the problems below, assume 100% dissociation. 1. A. Write the equation for the dissociation of hydrochloric acid. B. Find the pH of a 0.00476 M hydrochloric acid solution. 2. A.

pH and pOH - Ms. Mogck's Classroom

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Sometimes you are asked to calculate pOH rather than pH. Here's a review of what pOH is, plus an example calculation. Menu. Home. How to Find pOH in Chemistry. Search. Search the site GO. ... How to Find pOH in Chemistry Chemistry Quick Review of How to Find pOH . Share Flipboard Email Print If you know pH, it's easy to calculate pOH. David ...

Chemistry Review of pOH Calculations - ThoughtCo

Calculating pH and pOH worksheet W 335 Everett Community College Tutoring Center Student Support Services Program 1) What is the pH of a 0.0235 M HCl solution? 2) What is the pOH of a 0.0235 M HCl solution? 3) What is the pH of a $6.50 \times 10^{-3} M$ KOH solution? (Hint: this is a basic solution – concentration is of OH^-)

Calculating pH and pOH worksheet

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This lesson introduces the pH scale and discusses the relationship between pH, $[H^+]$, $[OH^-]$ and pOH. ... Chemistry 12.3a pH and pOH IsaacsTEACH. Loading... Unsubscribe from IsaacsTEACH?

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