# Colligative Properties Study Guide Answers

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Study Guide Chapter 12 and 13 Holt, Rinehart and Winston (modified by Ema Gluckmann – 4/11) Page 4 3. An important use of colligative properties is to determine the molar mass of unknown substances. The following situation is an example: 12.0 g of unknown compound X, a nonpolar nonelectrolyte, is dissolved in 100.0 g of melted camphor.

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colligative properties and solutions study guide may 2007. a b; ... colligative property: a property that depends on the number of molecules present, but not on their chemical nature. 3 types of colligative properties: vapor pressure reduction, boiling point elevation, freezing point depression ...

#### Quia - CHEMISTRY COLLIGATIVE PROPERTIES AND SOLUTIONS ...

At the completion of this episode's lesson(s), you should be able to: • Calculate molarity and prepare specific solution concentrations. • Define colligative properties and explain the significance of colligative properties in everyday life.

# Chemistry 1003: Molarity and Colligative Properties ...

CHAPTER 13 REVIEW Ions in Aqueous Solutions and Colligative Properties MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Match the four compounds on the right to their descriptions on the left.

#### 13 Ions in Aqueous Solutions and Colligative Properties

Colligative Properties Problems - Answers . Remember that Colligative properties depend only on the number or concentration of particles in a solution. The properties are, for ideal solutions, independent of the kind or size of the particles, whether ionic, large or small etc. As a consequence:

#### **Colligative Properties Problems - Answers**

Answer Key:  $55.4 \text{ g NaCl x (1mol NaCl }/58.44 \text{ g)} = 0.948 \text{ mol NaCl }42.3 \text{ g KBr x }... \text{ Welcome to "Colligative Properties" a section refreshing your memory on Chapter 10, here you will find all you need to help you with boiling point, and freezing point along with using molarity and molality.$ 

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#### 15 4 Review And Reinforcement Colligative ...

Determine how the physical properties of a solvent are dependent on the number of solute particles present. Measure the vapor pressure, boiling point, freezing point, and osmotic pressure of pure water and a variety of solutions. Compare the effects of four solutes (sucrose, sodium chloride, calcium chloride, and potassium chloride) on these physical properties.

## **Colligative Properties Gizmo: ExploreLearning**

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In this colligative properties worksheet, learners read about freezing point depression and boiling point elevation. They answer three questions about these colligative properties and the effects of salt on icy roads and sea water.

## Colligative Properties Lesson Plans & Worksheets | Lesson ...

172 Guided Reading and Study Workbook SECTION 16.3 COLLIGATIVE PROPERTIES OF SOLUTIONS (pages 487–490) This section explains why a solution has a lower vapor pressure, an elevated boiling point, and a depressed freezing point compared with the pure solvent of that solution. Vapor Pressure Lowering (pages 487–488) 1.

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Section 14.4 Colligative Properties of Solutions. TEACHER GUIDE AND ANSWERS Study Guide – Chapter 14 – Mixtures and 14.3 Solvation and Solubility. 1. solution. 2. solute. 3. solvent. 4. soluble. Solubility of Gases & Colligative Properties Study – Worksheet Hour: \_\_\_\_\_ Date: The solute is dissolved into the solvent. b.

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