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Post lab analysis. 4. For each trial, divide the number of grams of hydrogen peroxide by the total mass of the hydrogen peroxide solution (see step 7 in the Procedure), and multiply the answer by 100. The result is the percent hydrogen peroxide in the commercial antiseptic. Note: Assume the density of the commercial antiseptic solution is 1.00g/mL.

Analysis of Hydrogen Peroxide - A. Sedano - AP Chemistry ...

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2. Multiply the number of moles of permanganate ion by the mole ratio for hydrogen peroxide (see the Pre-Lab Questions) to determine the number of moles of hydrogen peroxide for each trial. 3. Multiply the number of moles of hydrogen peroxide by the molar mass of hydrogen peroxide to determine the number of grams of hydrogen peroxide for each trial. 4.

91253 Hydrogen Peroxide Analysis

Transcript of Analysis of Hydrogen Peroxide Inquiry Lab. Rinse the buret with approximately 10 mL of distilled water and then with 10 mL of the potassium permanganate solution (MnO_4^-). Close the stopcock and fill the buret just about the zero mark with the MnO_4^- solution. Open the stopcock to let out any air bubbles.

Analysis of Hydrogen Peroxide Inquiry Lab by Kerri Little ...

In the Analysis of Hydrogen Peroxide Inquiry Lab Solution for AP ® Chemistry, students use an oxidation-reduction titration to determine the percent composition of hydrogen peroxide, while learning about concepts relating to quantitative analysis and more.

Analysis of Hydrogen Peroxide—Blended Inquiry Lab for AP ...

Lab: Analysis of Hydrogen Peroxide Solutions Fill a burette with 50 mL of 0.2M potassium permanganate solution. Add 10 mL of NEW hydrogen peroxide and 2 mL of 6M sulfuric acid to a flask. Titrate the hydrogen peroxide with the KMnO_4 solution to the purple endpoint of excess MnO_4^- ion.

Lab: Analysis of Hydrogen Peroxide Solutions - BrainMass

The Hydrogen Peroxide Analysis lab would be best implemented after the students have had

experience with the use of a buret (as a volume measuring device), as well as general statistics. This is a good lab to introduce the skill of titrating.

HYDROGEN PEROXIDE ANALYSIS INTRODUCTION

In the Analysis of Hydrogen Peroxide Oxidation and Reduction Laboratory Kit, students test the percent hydrogen peroxide in a common “drugstore” solution. The lab ties together oxidation-reduction concepts and knowledge of stoichiometry.

Analysis of Hydrogen Peroxide—Student Laboratory Kit

If the oxidation-reduction reaction of the permanganate ion is performed in a basic or neutral solution, the product will be an insoluble brown solid. The brown solid will be manganese dioxide (MnO_2). 5. (a) The volume of hydrogen peroxide is necessary to know because it will later affect the results when plugged into the equation.

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