

Chemquest 29 Answer Key

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Chemquest 29 Answer Key

86 ChemQuest 29 Name: _____ Date: _____ Hour: _____ Introduction Questions 1. Write the equation for the reaction of sodium and chlorine to form sodium chloride. $\text{Na} + \text{Cl}_2 \rightarrow 2\text{NaCl}$ 2. Write the equation for the reaction of calcium nitride and sodium chloride to produce calcium chloride and sodium nitride.

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Answer the questions that follow based on this chemical equation: $\text{Na} + \text{MgCl}_2 \rightarrow 2\text{NaCl} + \text{Mg}$. a) Why can't NaMg be produced Na + cannot bond with Mg 2+ because they are both positive. b) Why can't NaCl 2 be produced Na + only requires one Cl - .

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(Base the answer to this question on the number of moles of propane that actually get combusted—which is your answer to part a.) 12 moles. For every mole of propane that combusts 3 moles of CO₂ are produced, so the number of moles of CO₂ that can be produced when 4 moles of propane combusts = $4(3) = 12$.

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Answer the questions that follow based on this chemical equation: $\text{Na} + \text{MgCl}_2 \rightarrow \text{NaCl} + \text{Mg}$. Why can't NaMg be produced? (Hint: Think about the charges) Why can't NaCl₂ be produced? Are NaCl and Mg the only products that can be produced? ... ChemQuest 28 ...

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