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CHAPTER 9 REVIEW Stoichiometry SECTION 3 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 88% The actual yield of a reaction is 22 g and the theoretical yield is 25 g. Calculate the percentage yield. 2. 6.0 mol of N 2 are mixed with 12.0 mol of H 2 according to the following equation: N 2(g) 3H 2(g ...

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Chapter 3 Mixed Review. 9 Stoichiometry SHORT ANSWER Answer the following questions in the space provided. Much of our knowledge of chemistry is based on the careful quanti- stoichiometry (which you studied in Chapter 3) deals with the mass rela- SECTION REVIEW.

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From above we can see that if we have 12.4 mol H 2 we need 4.13 mol N 2.We don't have that much N 2 so the .892 mol of N 2 must be the limiting reagent. We can now determine how much ammonia will be produced using the mole ratio in the balanced equation :

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CHAPTER 9 REVIEW. Stoichiometry. MIXED REVIEW. SHORT ANSWER Answer the following questions in the space provided. 1. Given the following equation: C3H4(g) + x. O2(g) (3CO2(g) + 2H2O(g) a. What is the value of the coefficient . x. in this equation? b. What is the molar mass of C3H4? c. How many moles are in an 8.0 g sample of C3H4? 2. a. What ...

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