

## *Chemquest 38 Answer Key*

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**Chemquest 38 Answer Key**

Answer the questions that follow based on this chemical equation:  $\text{Na} + \text{MgCl}_2 \rightarrow \text{NaCl} + \text{Mg}$ . a) Why can't NaMg be produced? Na + cannot bond with Mg<sup>2+</sup> because they are both positive. b) Why can't NaCl<sub>2</sub> be produced? Na + only requires one Cl<sup>-</sup>.

**82 ChemQuest 28 Name: - Yumpu**

114 Critical Thinking Questions 4. Verify that for the following reaction to be completed, 4 liters of hydrogen would require at least 2 liters of oxygen gas at the same pressure and temperature.  $2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{g})$  It is a 2:1 ratio of hydrogen to oxygen, so for every two liters of hydrogen one liter of oxygen is required.

**114 Critical Thinking Que - Yumpu**

116 Information: Dry Gas The gas that is collected is not "dry" because there is some water vapor left behind in the beaker. If you attempt to collect pure oxygen, you will actually get mostly pure oxygen with a little water

**Information: Collecting Gas Over Water**

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Answer the questions that follow based on this chemical equation:  $\text{Na} + \text{MgCl}_2 \rightarrow \text{NaCl} + \text{Mg}$ . Why can't NaMg be produced? (Hint: Think about the charges) Why can't NaCl<sub>2</sub> be produced? Are NaCl and Mg the only products that can be produced? ... ChemQuest 28 ...

**ChemQuest 28 - AP Chemistry**

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