

Chemquest 33 Limiting Reactants Answers

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(Base the answer to this question on the number of moles of propane that actually get combusted—which is your answer to part a.) ... (Remember—you must base this answer on how many moles of the limiting reactant that reacted.) ... ChemQuest 33 ...

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Chemquest 33 Limiting Reactants Answers Chemquest 33 Limiting Reactants Answers (Base the answer to this question on the number of moles of propane that actually get combusted—which is your answer to part a.) ... (Remember—you must base this answer on how many moles of the limiting reactant that reacted.) ... ChemQuest 33 ... ChemQuest 33 -

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answer on how many moles of the limiting reactant that reacted.) 1.5 mol. 0.75 mol of MgCl_2 is the limiting reactant, so we use the 6:3 ratio of NaCl to MgCl_2 and multiply $(0.75)(6/3) = 1.5$ mol. 10. Consider the double replacement reaction between calcium sulfate (CaSO_4) and sodium iodide (NaI).

Information : Limiting Reactant

100 ChemQuest 33 Name: _____ Date: _____ Hour: _____ Information: Limiting Reactant Again consider the combustion of propane: $\text{C}_3\text{H}_8 + 5\text{O}_2 \rightarrow 3\text{CO}_2 + 4\text{H}_2\text{O}$. If you had 10 moles of propane to burn, you would need 50 moles of oxygen according to the ratio in the balanced equation. If you only had 20 moles of oxygen you could not combust all 10 moles of propane.

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ChemQuest #33: Limiting Reactants 100 ChemQuest #34: Percent Yield 103 ChemQuest #35: Intro. to Gases 105 ChemQuest #36: Gases and Moles 109 ChemQuest #37: Gas Stoichiometry 113 ChemQuest #38: Partial Pressures 115 ChemQuest #39: Concentration 118 ChemQuest #40: Colligative Properties 121

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LIMITING REAGENT Practice Problems 1. At high temperatures, sulfur combines with iron to form the brown-black iron (II) sulfide: $\text{Fe (s)} + \text{S (l)} \rightarrow \text{FeS (s)}$ In one experiment, 7.62 g of Fe are allowed to react with 8.67 g of S. a. What is the limiting reagent, and what is the reactant in excess? b. Calculate the mass of FeS formed. 2. Acrylonitrile ...

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Limiting Reactants. In everyday life, finding the limiting reactant isn't that difficult, as long as you know what to look for. For example, if you have three storage containers but just two lids ...

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