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a) If this reaction took place in 2.0 L of solution, what is the concentration of leftover hydrochloric acid after the reaction $\text{mol leftover} = 2.5 - 1.9 = 0.6 \text{ mol} \div 2.0\text{L} = 0.3 \text{ M}$ b) From your answer to part a, verify that the pH of the solution after the reaction is approximately 0.52. $\text{pH} = -\log(0.3) = 0.52$

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a) Given your answers to questions 4-6, what happens to the rate of the reaction if you increase the temperature of the reactants? b) Try to think of two reasons WHY and write them down.

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