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From above we can see that if we have 12.4 mol H₂ we need 4.13 mol N₂. We don't have that much N₂ so the .892 mol of N₂ must be the limiting reagent. We can now determine how much ammonia will be produced using the mole ratio in the balanced equation :

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Modern Chemistry 2 Stoichiometry CHAPTER 9 REVIEW Stoichiometry MIXED REVIEW SHORT ANSWER Answer the following questions in the space provided. 1. Given the following equation: $C_3H_4(g) + xO_2(g) \rightarrow 3CO_2(g) + 2H_2O(g)$ ____ a. What is the value of the coefficient x in this equation? ____ b. What is the molar mass of C₃H₄?

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Chemistry Worksheet on Stoichiometry Mixed Review. Assume all reactions go to completion. Write the formula equation, balance the equations, and solve the problems. Draw a rectangle around the answer and don't forget the units. Methane (CH₄) combines with oxygen to form carbon dioxide and water. Balanced equation:

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