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Chemistry In Focus 5th Edition

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Chemistry - Wikipedia

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Coordination complex - Wikipedia

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Home | W. W. Norton & Company

In order for NaCl to be soluble, the Na⁺ and Cl⁻ ions must break free from the crystal-lattice structure of the solid. When the ions are in solution, they are surrounded by water molecules, and the ions are said to be solvated, or dissolved in an aqueous solution, denoted (aq). Hence, the process of dissolving a NaCl crystal can be described by the following chemical equation (Equation 1):

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News: Assistant Prof. Yoshimitsu Sagara (Research Institute for Electronic Science, Hokkaido Univ.)

who received Ph.D under supervision of Prof. Kato in 2009, received The Young Scientists' Prize from Ministry of Education, Culture, Sports, Science and Technology (MEXT)(April, 2019).

Kato Laboratory

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GHS implementation - Transport - UNECE

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