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For calculation  $E^{\circ} = E_{\text{cathode}} - E_{\text{anode}}$ . all metals couple with  $w + 2$  in anode in oxidized. More the reduction... view the full answer

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Explain your answer in terms of your answers for questions 3a and 3b. (hint: the poorer an oxidizing agent a species is, the harder it is to reduce that species. Hint #2: the copper complexes are reduced by gaining an electron.)

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experiment. There will be Chem21labs prelab and postlab assignments due as shown on the schedule and you will need to turn in printed spectra for the room lights, the hydrogen lamp, the six metal salts, and the sparkler at the beginning of your next lab period. There is a prelab video on Blackboard. You experience the excitement of atomic

**5. EMISSION SPECTROSCOPY - chem21labs.com**

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And be careful to use the correct positive or negative enthalpy values depending on the DIRECTION of the reactions that you add together. In your final answer specify whether the enthalpy change occurring in the combined final reaction causes heat to be released or absorbed.

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Hello, my chemistry professor gave my class some questions to answer before our first lab. This is my first Chem course so I have no idea what the answers are, please help. Thank You!! 1: The specific gravity of a urine sample can be measured: a. In a urinometer, while the float is touching the sides b. In a urinometer, while the float is touching neither the bottom nor the sides c.

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