

## *John Erickson Ionic Compounds Answer*

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Toronto: John Wiley and Sons, 1988. ... and the properties of ionic and molecular compounds; ... answer sheets, ... CHEMISTRY 1-1 - Public Schools of Edison Township / Homepage

### John Erickson Ionic Compounds Answer

An ionic compound is a combination of oppositely charged ions. Ionic compounds generally contain a metal bonded to a non-metal (or non-metals). When naming ionic compounds we simply name the cation (the positive ion) then the anion (the negative ion). The cations generally retain the name of the element. so Na<sup>+</sup> is named sodium.

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© John Erickson, 2005 WS10-4NetIonicEq When two solutions of ionic compounds are mixed, a solid may form. This type of reaction is called a precipitation reaction, and the solid produced in the reaction is known as the precipitate.

### Net Ion Equation - Lauren Sweatt - Net Ionic Equations ...

Ch 7 WS 8 : Empirical Formulas 1. Determine the empirical formula from the compound formula. Circle the ionic compounds. a) C<sub>6</sub>H<sub>6</sub> b) C<sub>2</sub>H<sub>2</sub> c) C<sub>2</sub>H<sub>6</sub> d) Na<sub>2</sub>SO<sub>4</sub> e) C<sub>3</sub>H<sub>8</sub> f) C<sub>6</sub>H<sub>5</sub>N g) Fe<sub>3</sub>(CO)<sub>9</sub> h) P<sub>4</sub>O<sub>10</sub> i) C<sub>2</sub>H<sub>4</sub>O<sub>2</sub> j) Re<sub>2</sub>Cl<sub>6</sub> k) N<sub>2</sub>H<sub>4</sub> l) Se<sub>3</sub>O<sub>9</sub> m) CaBr<sub>2</sub> n) LiCl Determine the empirical formula of the following compounds based on the percent compositions

### CHEMISTRY L E E U N I T S E V E N - CHEMSTEM - Home

Sodium chloride (NaCl), and potassium nitrate (KNO<sub>3</sub>) are two examples of soluble compounds. When these compounds are mixed with water they dissolve, and we describe them as aqueous (aq). There are many ionic compounds that do not dissolve in water, though. These are described as insoluble.

### Solubility Rules Name Chem Worksheet 15-1

Ionic compounds generally contain a metal bonded to a non-metal (or non-metals). When naming ionic compounds we simply name the cation (the positive ion) then the anion (the negative ion). The cations generally retain the name of the element, so Na<sup>+</sup> is named sodium.

### **Naming Ionic Compounds Name Chem Worksheet 8-2**

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Since the compound contains 2 moles of water for every 1 mole of anhydrate the formula is  $\text{CaCl}_2 \cdot 2 \text{H}_2\text{O}$  A hydrate is an ionic compound that contains water molecules in its structure. To determine the formula of a hydrate experimentally, we must calculate the mole: mole ratio of the water portion compared to the anhydrate portion.

### **Determining the Formula of a Hydrate Name Chem Worksheet 11-6**

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A hydrate is an ionic compound that contains water molecules in its structure. To determine the formula of a hydrate experimentally, we must calculate the mole: mole ratio of the water portion compared to the anhydrate portion. An anhydrate is the substance that remains after the water is removed from a hydrate.

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