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Holt ChemFile: Problem-Solving Workbook 122 Percentage Yield ... Is the answer reasonable? Yes; 83% is about 5/6, which appears to be close to the ratio 0.71/0.86. ... Set up a stoichiometry calculation to find the amount of product that can be formed from the given amount of reactant.

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8. actual; theoretical 9. 10. 11. 3.00 g Mg (1 mol Mg/24.30 g Mg)

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Holt ChemFile: Problem-Solving Workbook 57 Mole Concept Name Class Date Problem Solving continued Is the answer reasonable? Yes; 2 g of boron is about 1/5 of the molar mass of boron. Therefore, 2.00 g boron will contain about 1/5 of an Avogadro's constant of atoms. Practice 1. Calculate the number of atoms in each of the following masses:

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Stoichiometry SECTION 2 PROBLEMS Write the answer on the line to the left. Show all your work in the space provided. 1. 4.5 mol The following equation represents a laboratory preparation for oxygen gas: $2\text{KClO}_3(\text{s}) \rightarrow 2\text{KCl}(\text{s}) + 3\text{O}_2(\text{g})$ How many moles of O₂ form if 3.0 mol of KClO₃ are totally consumed?

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Holt ChemFile: Problem-Solving Workbook 1 Conversions ... Being organized is a key to developing good problem solving skills. PLAN What steps are needed to convert grams to milligrams? ... Is the answer reasonable? Yes; the number of milligrams is 1000 times the number of grams.

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