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What is the correct acid name for an aqueous solution of HClO_4 ? a. hypochlorous acid b. chlorous acid c. chloric acid d. perchloric acid ____ 9. A substance that increases the concentration of OH^- ions in an aqueous solution is known as a(n) a. Arrhenius acid. b. Arrhenius base. c. Lewis acid. d. Lewis base. ____ 10. An Arrhenius acid in an ...

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a. stronger acid. c. weaker acid. b. stronger base. d. weaker base. ____ 6. Proton-transfer reactions favor production of the a. stronger acid and stronger base. b. weaker acid and weaker base. c. stronger acid and weaker base. d. weaker acid and stronger base. ____ 7. Aqueous solutions of most bases contain a. hydroxide ions and cations. c ...

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4 is considered acid 1, identify the other three terms as acid 2, base 1, and base 2 to indicate the conjugate acid-base pairs. base 2 CO_3^{2-} acid 2 HCO_3^- base 1 NH_3 True b. A proton has been transferred from acid 1 to base 2 in the above reaction. True or False? 3. Consider the neutralization reaction described by the equation: $\text{HCO}_3^-(\text{aq}) + \text{OH}^- \rightarrow \text{CO}_3^{2-}(\text{aq}) + \text{H}_2\text{O}(\text{l})$

14 Acids and Bases - mchsapchemistry.com

Major topics: Arrhenius vs. Bronsted-Lowry definition of acids and bases, conjugate acid/base, acid dissociation constant (K_a), & strong vs weak acids

Chapter 14 (Acids and Bases) - Part 1

Holt Chemistry Chapter 15 Acids and Bases $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl}$. the reaction of the ions that characterize acids (hydronium ions) and the ions that characterize bases (hydroxide ions) to form water molecules and salts. the equivalence point is the point in a

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Acids and Bases Quiz You got: % Correct. You're Associated with Acids and Bases Dave White / Getty Images You are familiar with the basic concepts of acids, bases, and pH, but there's more to learn. Are you a hands-on learner? Try experimenting with homemade pH indicator to test household products. Tired of acids and bases?

Acids and Bases Chemistry Quiz - ThoughtCo

Compare the Brønsted-Lowry definitions of acids and bases with the Arrhenius definitions of acids and bases. Identify conjugate acid-base pairs. Write chemical equations that show how an amphoteric species can behave as either an acid or a base. Acids Vinegar is acidic. So are the juices of the fruits that you see in Figure 1.

CHAPTER

Modern Chemistry 121 Acids and Bases CHAPTER 14 REVIEW Acids and Bases SECTION 3 SHORT ANSWER Answer the following questions in the space provided. 1. Answer the following questions according to the Brønsted-Lowry definitions of acids and bases: ____ a. What is the conjugate base of H_2SO_3 ? ____ b.

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