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3. Indicate the oxidation number of phosphorus in each of the following compounds: a.  $\text{HPO}_3$  \_\_\_\_ d.  $\text{H}_3\text{PO}_4$  \_\_\_\_ b.  $\text{H}_3\text{PO}_2$  \_\_\_\_ e.  $\text{H}_4\text{P}_2\text{O}_7$  \_\_\_\_ c.  $\text{H}_3\text{PO}_3$  \_\_\_\_ f.  $\text{H}_5\text{P}_3\text{O}_{10}$  \_\_\_\_ 4. Give oxidation numbers for the underlined atoms in these molecules and ions: a.  $\text{Cs}_2\text{O}$  b.  $\text{PtCl}_6^{2-}$  c.  $\text{CaI}_2$  d.  $\text{SnF}_2$  e.  $\text{Al}_2\text{O}_3$  f.  $\text{ClF}_3$  g.  $\text{H}_3\text{AsO}_3$  h.  $\text{SbF}_6^-$  ...

### Worksheet: Oxidation Numbers Name - MARRIC

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produced when \_\_\_\_ grams of hydrogen gas react with the oxygen in the air?

### **8-01,02,03 Note Taking Guide Ep 801 - Welcome to Mole Cafe**

7. In compounds, the elements of groups 1 and 2 as well as aluminum have oxidation numbers of  
+1, +2, and +3 respectively. 8. The sum of the oxidation numbers of all atoms in a neutrals  
compound is 0. 9. The sum of the oxidation numbers of all atoms in a polyatomic ion equals the  
charge of the ion. Answer Key 1. Cl:0 7. Al:3+ 13. N:3- H:1+ 19.

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