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Balancing Redox Equations Directions: Balance the equations below using the half-reaction method. 1. $\text{Sn}^0 + \text{Ag}^+ \rightarrow \text{Sn}^{2+} + \text{Ag}^0$ 2. $\text{Cr}^0 + \text{Pb}^{2+} \rightarrow \text{Cr}^{3+} + \text{Pb}^0$... Chemistry IF8766 Page 93 Instructional Fair, Inc. Title: Microsoft Word - pg 93 - Balancing Redox Equations.doc

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Sample Study Sheet: Balancing Redox Equations Using the Oxidation Number Technique . Tip-off If you are asked to balance an equation and if you are not told whether the reaction is a redox reaction or not, you can use the following procedure. General Steps Step 1: Try to balance the atoms in the equation by inspection, that is, by the standard technique for balancing non-redox equations.

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Balancing Chemical Equations -Answer Key Balance the equations below: 1) $1 \text{ N}_2 + 3 \text{ H}_2 \rightarrow 2 \text{ NH}_3$ 2) $2 \text{ KClO}_3 \rightarrow 2 \text{ KCl} + 3 \text{ O}_2$ 3) $2 \text{ NaCl} + 1 \text{ F}_2 \rightarrow 2 \text{ NaF} + 1 \text{ Cl}_2$ 4) $2 \text{ H}_2 + 1 \text{ O}_2 \rightarrow 2 \text{ H}_2\text{O}$ 5) $1 \text{ Pb(OH)}_2 + 2 \text{ HCl} \rightarrow 2 \text{ H}_2\text{O} + 1 \text{ PbCl}_2$ 6) $2 \text{ AlBr}_3 + 3 \text{ K}_2\text{SO}_4 \rightarrow 6 \text{ KBr} + 1 \text{ Al}_2(\text{SO}_4)_3$

Balancing Chemical Equations Answer Key

Balancing Chemical Equations Directions: rewrite and balance the equations below. 1. $\text{P} + \text{O}_2 \rightarrow \text{P}_2\text{O}_5$ ____ 2. $\text{KClO}_3 \rightarrow \text{KCl} + \text{O}_2$... Chemistry IF8766 Page 58 Instructional Fair, Inc. Title: Name: Author: ASD Created Date:

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