

## *Isotopes And Average Atomic Mass Answer Key*

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### Isotopes And Average Atomic Mass

Atomic mass is really a weighted average of all the element's isotopes based on their natural abundance on Earth. If you're given a list of isotopes to calculate, look for the exact mass of each isotope (that will be in decimal form, but if all you have are whole mass numbers, use those).

### How to Find Average Atomic Mass | Sciencing

Average Atomic Mass. To find the average atomic mass of an element the mass of each individual isotope must be taken into account in addition to how much of that isotope exists. For example, the average atomic mass of Hydrogen is closest to 1, not 2 or 3, because 99.98% of the atoms of hydrogen that you would find in our universe exists as Hydrogen-1.

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The atomic mass of an element is the weighted average of the atomic masses of the naturally occurring isotopes of that element. The average atomic masses are the values we see on the periodic table.  $[0.7577 \left( 34.969 \right) + 0.2423 \left( 36.966 \right) = 35.453]$

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