

## *Intermolecular Forces Lab Answers*

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### **Intermolecular Forces Lab Answers**

Answer 4. Identify the structure as having hydrogen bonding, dipole-dipole moments or London dispersion forces (LDF). 3. State whether the molecule is polar or non polar. H<sub>2</sub>O: Covalent CH<sub>4</sub>: Covalent CO<sub>2</sub>: Covalent NH<sub>3</sub>: Covalent H<sub>2</sub>O: Hydrogen bonding CH<sub>4</sub>: LDF CO<sub>2</sub>: LDF NH<sub>3</sub>: Hydrogen

### **03.07 Intermolecular Forces: Lab by Caymen Conley on Prezi**

Intermolecular Forces: Lab: Dependency of Boiling Points on IMFs. The purpose of this lab was to determine the boiling range of different molecules and observe how the strength of their IMFs affect the temperature change involved with evaporation of the compounds.

### **Lab: Boiling points - Intermolecular Forces**

Laboratory: Intermolecular Forces (IMF) Report Requirement: Answer all of the questions/do all the computations requested in italics. Questions not in italics do NOT need to be answered. You do NOT have to write a formal lab report. You should write your answers into a word processing program and save the file. Go into the Lab-IMF Report and ...

### **Laboratory: Intermolecular Forces (IMF) - Parkland College**

Chaylen Andolino Partner: Bethany Rudolph Make-up Lab Date: 5/25/2012 Evaporation and Intermolecular Attractions Lab Purpose: The purpose of this lab was to observe temperature changes caused by the evaporation of ethanol, 1- propanol, 1- butanol, n- pentane, methanol, and n- hexane, and then relate the temperature changes to the strength of intermolecular forces of attraction.

### **Evaporation and Intermolecular Attractions Lab ...**

Lab #17 - The Effect of Intermolecular Forces on Evaporation Introduction and Background Information In this experiment, we will determine what type of molecule has stronger intermolecular force between molecules by measuring evaporation.

### **Lab #17 - The Effect of Intermolecular Forces on ...**

Exploring Intermolecular Forces Lab Background: Intramolecular forces are forces acting on atoms within ionic crystals or molecules. Intramolecular forces are responsible for many macroscopic properties such as electrical conductivity, hardness, and luster. Other properties of matter such as boiling point,

### **Exploring Intermolecular Forces Lab - Northern Highlands**

Lab 8. Evaporation and Intermolecular Attractions Name Prelab Questions Team Date Section Instructions: Complete the following questions and hand in at the start of your lab period. 1. Draw the structure of each compound below. Use your knowledge of intermolecular forces of attraction to predict

### **Lab 8. Evaporation and Intermolecular Attractions**

Lab - Evaporation and Inter Molecular Attractions - Free download as Word Doc (.doc), PDF File (.pdf), Text File (.txt) or read online for free. ... The more hydrogen present and the higher the molecular weight of a substance, the higher the intermolecular forces should be. 1-butanol Lower than propanol Therefore, ...

### **Lab - Evaporation and Inter Molecular Attractions ...**

Intermolecular Forces Exercises. Answer the following to the best of your ability. Questions left blank are not counted against you. When you have completed every question that you desire, click the "MARK TEST" button after the last exercise. A new page will appear showing your correct and incorrect responses.

### **Intermolecular Forces Exercises**

Intermolecular Forces Worksheet Answers are on page 3 & 4. Do the problems on your own BEFORE looking at the answers. 1. Predict the molecular shape of each of the following: a. H<sub>2</sub>S b. CCl<sub>4</sub> c.

SO 2 d. BrF d. PCI 5 2. List all types of IMFs that would occur in each of the following (you should have a good enough understanding of ...

**Intermolecular Forces Worksheet - Ms. Mogck's Classroom**

Chemistry 20: Intermolecular Forces Worksheet 1. Each of the following statements describes and intermolecular force. For each of the statements, indicate if it describes London Forces (LF) , Dipole Forces (DD) or Hydrogen Bonding (HB) \_\_\_\_ occurs in all molecules \_\_\_\_ is the strongest force \_\_\_\_ occurs in polar molecules

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