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For the reaction $a \rightarrow b$, the rate law is: $\text{rate} = k \times [a]$. over the course of 4 minutes, the concentration of a reduces by one-half, from 1.00 m to 0.50 m. how is the rate affected?

For the reaction $a \rightarrow b$, the rate law is: $\text{rate} = k \times [a] \dots$

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