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Pideal = 1.0004 P1 = n\*R\*T/(V - nb) P1 = 1.0029 P2 =  $(a*n^2)/V^2$  P2 = 0.0129 Pwaals = P1 + P2 Pwaals = 1.0158 The ideal gas law predicts a pressure of 1.0004 atmospheres, while the van der Waals model predicts 1.0158 atmospheres.

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