Santa Barbara City College

721 Cliff Drive Santa Barbara, CA 93109

General Chemistry I

Course Title: Chem 155: **Period:** 09/02/2003 to 12/19/2003

Welcome to Chemistry 155!

Instructor

Name: Orlando Raola
Office: PS 206
Office Phone: 805-965-0581
Email: oraola@sbcconline.net

Course Information

Course: Chemistry
Section #: 3192/ 3193 / 3194/03
Credits: 5.00
Class Room #: PS 128 (Tuesday), PS 130 (Thursday)

Office Hours					
Monday	**Tuesday**	**Wednesday**	**Thursday**	**Friday**	**Sat.**
-	1:00 - 2:00 pm	-	5:00 - 6:00 pm	-	-

Class Meeting Days and Times					
Monday	**Tuesday**	**Wednesday**	**Thursday**	**Friday**	**Sat.**
-	06:00 - 08:05 pm	-	06:00 - 06:50 pm	-	-

Course Description:

This is the first part of a two-semester course and will cover the material in Chapters 1 to 11 in the textbook: Structure of atoms and molecules, redox, stoichiometry, bonding, states of matter, solutions, chemical calculations. This is a CSU/UC transferable course.

Where to leave messages: please use e-mail
Where to leave assignments: use slot Chem 155 in wooden mailbox by the entrance of the stockroom

Pre Requisites/Co-Requisites:

Eligibility for English 103; plus Math 100; plus one year of high school chemistry, or Chemistry 101 or 104.

Course Materials/Supplies:

Chemistry Steven S. Zumdahl, Susan A. Zumdahl 6th edition, Houghton Mifflin, 2003 ISBN: 0-618-22156-5 Chemistry 155 Houseman, B. et al. Thomson Learning Custom Publishing, 2003 ISBN: 0-534-04648-7 Approved safety goggles by September 16

Other Comments:

Discussion session will be held on Tuesdays in PS 128 (Disc A meets at 02:00 pm, Disc B meets at 08:05 pm)

Teacher's Expectations

You are expected to acquire the basic intellectual tools needed for understanding the wide world of chemical interactions from the formation of the simplest molecules in intergalactic space to the inner workings of neurons in the human brain. In order to achieve this, you are expected to come to class every day on time, to read the material suggested in preparation for the lecture, to do all the homework problems assigned, to participate actively in class, to take notes and ask questions, to prepare for the lab discussion and experiment, to complete your lab experiments in a timely fashion and to turn in your completed reports by the beginning of the next lab session, to take all four midterm and the final exam. You are also expected to show respect for your instructor, teaching assistants and departamental staff, to abide by SBCC honor code regarding the individual authorship of all material that you turn in for grade and to express any concern that you might have so that you, your fellow students and your instructor have an easier time working together in order to reach our common goals.

Course Goals and Objectives

Course Objectives (1) Demonstrate mastery of inorganic nomenclature and formula-writing. (2) Calculate stoichiometric relationships in chemical reactions, including mass-mass, thermochemical, gas laws, and solution relationships. (3) Explain the structure of atoms and compounds in terms of modern quantum theory, VSEPR theory, hybridization, the periodic table, and molecular orbital theory. (4) Convert raw laboratory data into laboratory reports that show the mastery of chemical concepts illustrated by experiments. (5) Predict the outcome of proposed chemical reactions from solubility rules and by writing molecular, total ionic and net ionic equations (6) Propose Lewis structures and geometric shapes of molecules/ions. (7) Determine the effect of solutes on the colligative properties of aqueous solutions.

Teacher's Disclaimer Statement

This syllabus is a "living" document and will be adjusted as necessary during the course. Please refer to the web version often (preferably every week) to keep abreast of the changes.

Course Calendar