

CHEMISTRY
FORM ONE

SECTION A (44 Marks)

1. What is a pure substance? {1

mark}

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he

following terms.

a) Radical. {1 m ark}

b) Acid. {1 m ark}

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2. State **TWO** reasons why most of the chemistry apparatus are glassware.

{2 m arks}

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3. Complete the table below.{2 m arks}

Parameter	Apparatus used to measure	Units
Volume		
Temperature		

4. Name any **TWO** industries that have benefited from the knowledge of chemistry. {2 m arks}

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5. What do the following laboratory signs mean?

a)b)



d).

c).

6. State why it is important to adhere to the following laboratory rules.

{2 marks}

a) Label all the chemicals

.....

b) Never

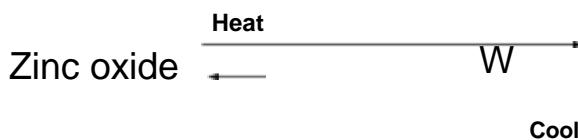
eat anything in the laboratory.

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Write a

chemical equation showing the reaction between sodium metal and excess oxygen. {1 mark}

7. Study the equation below.



a) What is the colour of substance W

..... {½ mark}

b) What type of change is represented in the above equation? {½mark}

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c) Give another example of the change named in b) above.{1mark}

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8. a). Name any **FOUR** apparatus that are necessary to carry out fractional distillation of a mixture containing D distilled water and E thanol

in the laboratory. {2 marks}

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b). Which of the two components of the mixture will form the first fraction.

Explain.{1 mark}

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c). State one industrial application of fractional distillation.{1mark}

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9. Study the equation below.

Hydrated copper (II) sulphate
Colourless



Liquid H

a) Identify the;

i). White solid T

- ii). Colourless liquid H {1 mark}
- b) If Hydrated Copper (II) Sulphate had FIVE water molecules, write the chemical equation for the above reaction. {1 mark}

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10. During the 2012 London Olympic Games, samples from four Decathlon participants (Morgan, Bolton, Jimmy and Jade) were taken and tested for presence of two illegal steroids A and B. Paper chromatography was used for the test.

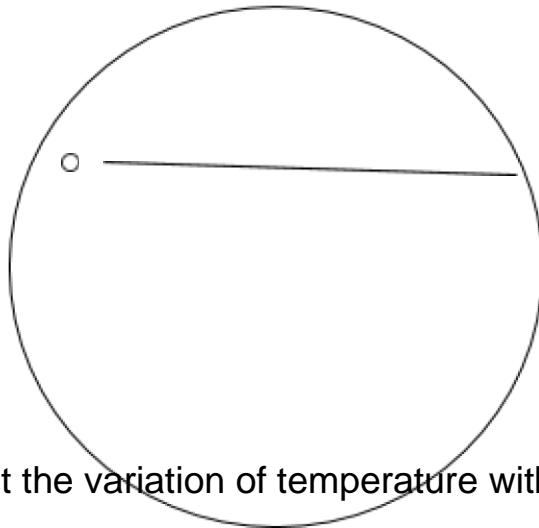


Steroid A Steroid B Morgan Bolton Jimmy Jade

- a) Which athlete(s) tested positive for the illegal steroid? {1 mark}

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- b) On the filter paper representation below, draw the results for the Bolton.{2



11. The curves below represent the variation of temperature with time when marks}

pure and impure samples of a solid were heated.



a) Which of the two curves shows the variation in temperature for pure solid?

Explain.{2 marks}

b) If 300 grams more of the pure substance was added to the sample, show on the graph the time that the sample the pure substance will boil. {½ mark}

c) On the graph above, indicate the boiling point of the pure substance.{½ mark}

12. Arnold, a student from Starehe Boys' Centre Situated 3050m above the sea level boiled 100cm³ of pure water. Another student, Annette, from Mombasa 0 metres above the sea level boiled the same volume of pure water.

i. Which of the two students took the longest time to boil water? {1mark}

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ii. Explain your answer in d, i) above? {1 mark}

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13. Describe the steps you would take to obtain common salts from sand that were mixed accidentally salt.{3 marks}

14. Complete the table below. {2

marks}

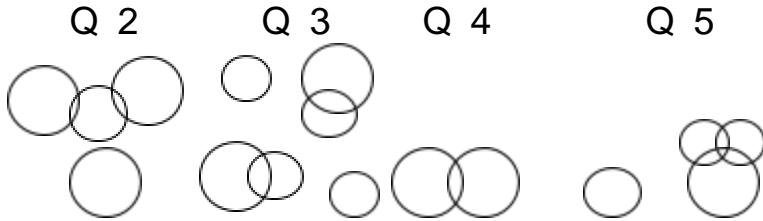
	Solid	Liquid	Gas
Shape			Indefinite
Volume		Fixed	Not fixed
Density	Very high	High	
Packing of molecules	Tight	Apart	Far apart

15. a). Draw a well labelled diagram showing how electrical conductivity of a given solid can be tested in the laboratory.

b). Name one non-metal that conducts electricity {1 mark}

16. Study the diagrams below.

Q1



Which of the following sets of drawing clearly illustrates?

{2 marks}

a) Hydrogen molecule (H_2)

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b) Ammonia molecule (NH_3)

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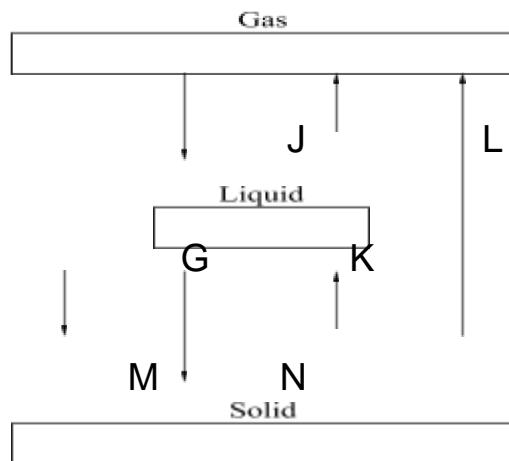
c) Helium molecule (He)

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d) Mixture of Sodium Chloride (NaCl) and Helium (He)

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17. The figure below shows the changes that take place between states of matter.



a) Give the names of the processes J and K.

{1 mark}

J

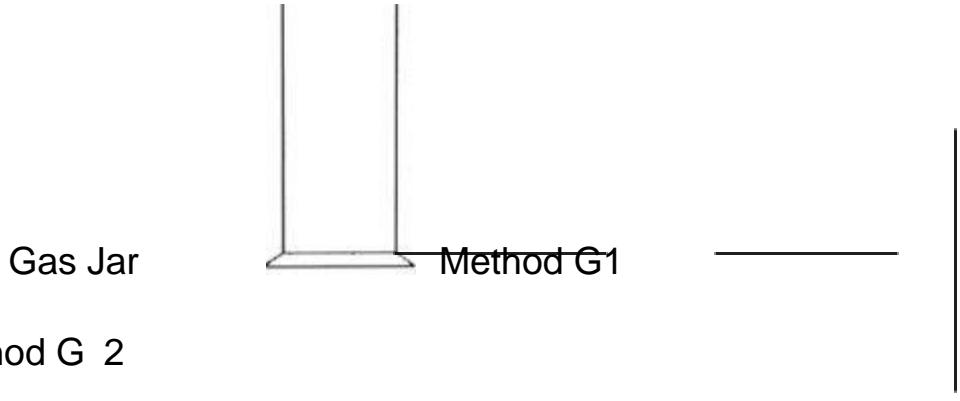
K

b) Name one substance that can undergo process K when left in an open container. {1 mark}

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18. Below are two methods of collecting gases in the laboratory.

Gas in



Method G 2



G as in



- a) Name the methods represented by G 2.

..... {½ mark}

- b) Name an example of gas that can be collected using G 1.

..... {½ mark}

SECTION B : (56 Marks)

19. What is the meaning of the following?

- c) i). Drug. {1 mark}

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ii). Dosage. {1 mark}

- iii). Drug addiction. {1

mark}

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d) Differentiate between Over the Counter (OTC) drugs and Prescription Drugs.

{1 mark}

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e) Mr. Radisha went to a doctor who sent him to a pharmacy to pick some drugs. The pharmacist wrote on the medicine packaging 2X3. Clearly state what 2X3 meant.{1 mark}

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f) State two reasons why it is important to adhere to the doctor's prescription.

{2 marks}

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g) State **THREE** common effects that tobacco smoking and alcohol c

onsumption have. {3 marks}

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20. The diagram below show the apparatus used to prepare oxygen in the laboratory.

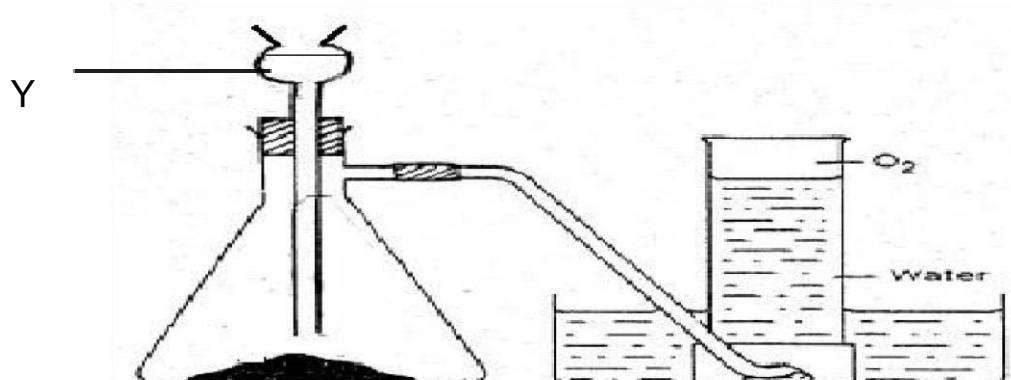
Black Solid X

- a) Name the reagents X and Y {2

marks}

Y X

- b) Why is reagent X used yet reagent Y can decompose to produce O₂ oxygen gas? {1 mark}



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c) Why is it possible to collect oxygen gas using the method above? {1 mark}

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d) State **TWO** other physical properties of oxygen gas {2 marks}

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e) Describe the test you would carry to prove that the gas collected is oxygen. {2 marks}

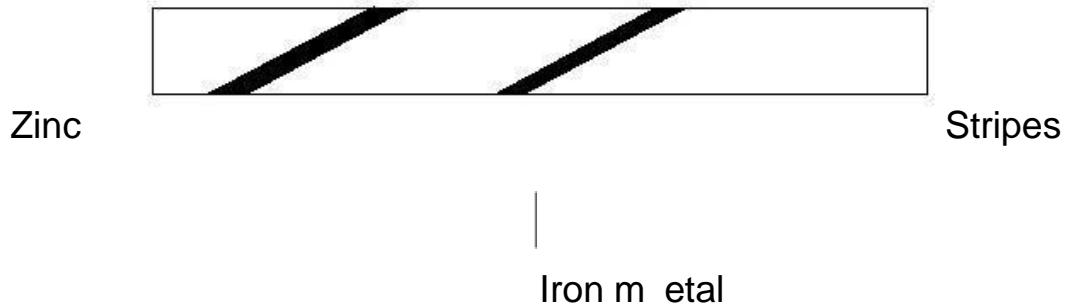
f) Write the chemical name and chemical formulae of rust.
{2 marks}

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Name

Formula.....

- g) The diagram below illustrates one of the methods used to prevent rust, study it carefully.



- h) Which method of rust prevention is shown in the diagram above?

{1 mark}

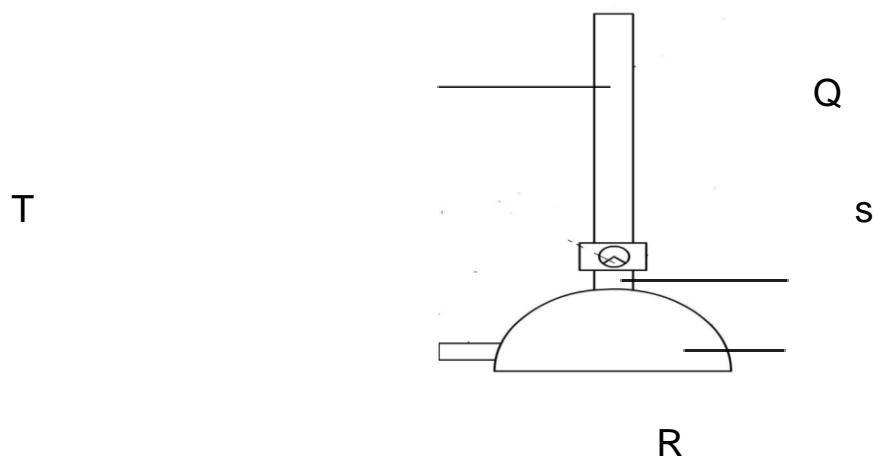
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- i) Why is it possible to prevent rust using the method named in g) above?

{1 mark}

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21. Study the diagram below.



- a) The apparatus is used to heat substances in the laboratory.

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- i. Name the parts marked with letters Q, R, S and T (Name on the diagram)

{2 marks}

- ii. Describe **FIVE** Steps followed when lighting the apparatus above.

{2½ marks}

- iii. On what flame should the apparatus be left when not being used in the laboratory? {½ mark}

- iv. State **TWO** reasons for the answer given in iii) above.

{2 marks}

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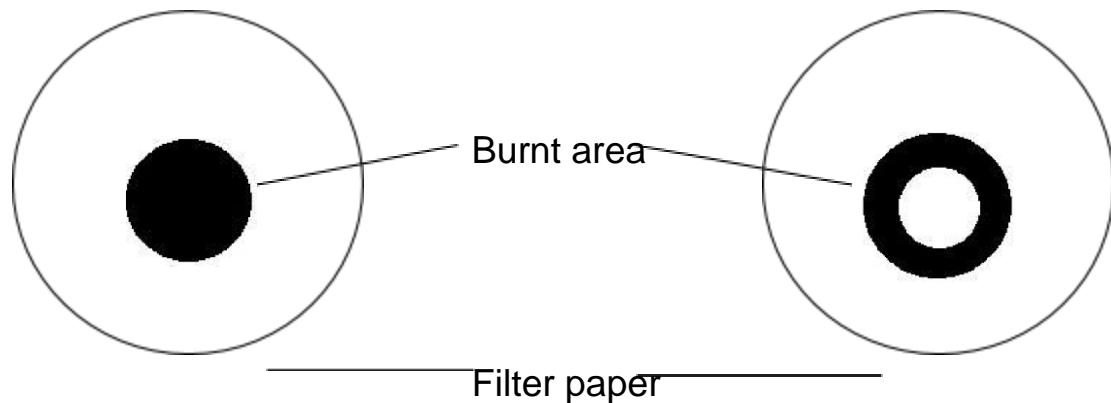
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- b) The diagram below shows the appearances of two pieces of filter papers placed on different parts of a particular flame of a Bunsen burner.



- i. Which flame of the Bunsen burner was used for the experiment?

{1 mark}

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ii. What conclusion can you make from the above experimental results?

{2 marks}

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22. a). Define the following.

i. Element. {1

mark}

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ii. Compound. {1 mark}

b). Write the chemical symbols for the following elements.

{1 mark}

- i. Chlorine ii). Sodium

.....

c). Write the name of the elements represented by the following chemical symbols. {1 mark}

i). K

ii). F

d). How many elements make up the following compounds?

i. Na_2CO_3 , NaHCO_3 , $2\text{H}_2\text{O}$ {1

mark}

ii. NaHCO_3 {1

mark}

e). Write the chemical formulae of the following chemical compounds. (Show your working)

i. Aluminium phosphate. {1

mark}

ii. Copper nitrate. {1

mark}

iii. Magnesium sulphate. {1

mark}

f). Write the number of atoms of each element present in the following compounds.

i. H_2SO_4 {1 mark}

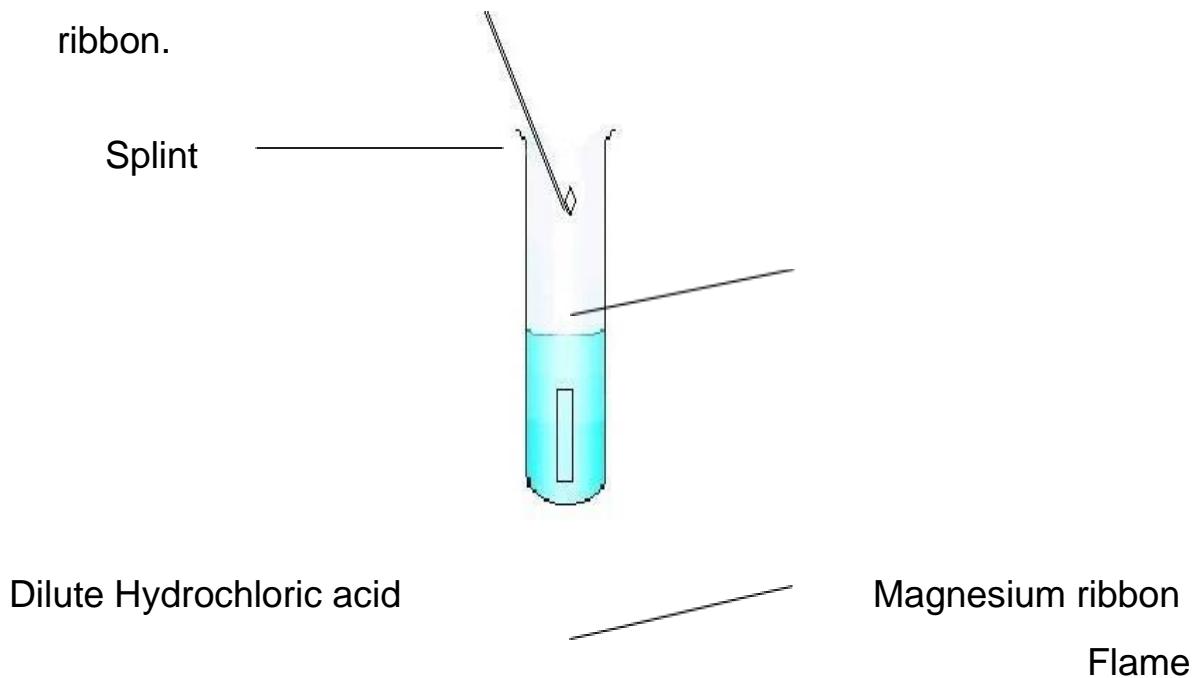
ii. $\text{CuSO}_4 \cdot 10\text{H}_2\text{O}$ {1

mark}

iii. $\text{Na}_2\text{C O}_3$ {1

mark}

23. The diagram below shows the action of dilute acids on a magnesium ribbon.



a) State any **TWO** observations made from the above set up.

{2 marks}

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b) Write a chemical equation for the reaction taking place between the acid and the metal.

{1 mark}

c) The following is a list of some pH values; 2, 4, 5, 7, 9 and 13. Complete the table below indicating the appropriate pH values.

{2 marks}

Substance	pH Value
Dilute hydrochloric acid	
Wood ash Solution	
Orange juice	
Distilled water	

- d) State one advantage of using the universal indicator over flower extract indicators.{2 marks}

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- e) What is a “neutralization reaction”? {1 mark}

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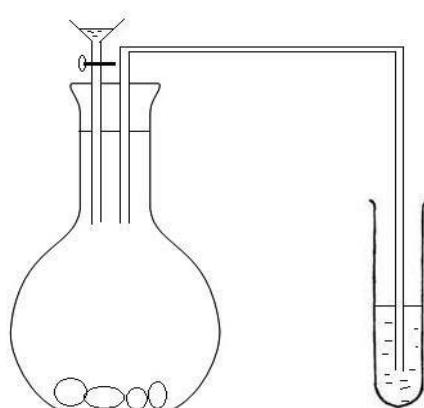
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- f) In another experiment students reacted sulphuric (VI) acid with solid P which is a compound of magnesium. A colourless solution Q was formed with

production of a

H_2SO_4



colourless gas Z.

Solid P ————— Lime water

When the colourless gas Z was bubbled in lime water, it formed a white precipitate.

- i. Identify colourless gas Z.

{1 mark}

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- ii. Identify compound P. {1 mark}

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- iii. Write the chemical formula of compound P. {1 mark}

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- iv. Name colourless solution Q . {1
mark}