Chemical Formulae

 Sb_2O_3

 H_2O

Chemical Equations

$$CO_2 + C \longrightarrow 2CO$$

 $Hg^{2+} \xrightarrow{I^{-}} HgI_{2} \xrightarrow{I^{-}} [Hg^{II}I_{4}]^{2-}$

Charges

 H^+

 ${\rm CrO_4^{2-}\ CrO_4^{2-}}$

 $[AgCl_2]^ Y^{99+}$ Y^{99+}

Oxidation States

FeFe₂O₄ FeFe₂O₄

Stoichiometric Numbers

 $2H_2O$

 $2 H_2 O$ $2 H_2 O$

 $1/2\,{\rm H}_{2}{\rm O}$

 nH_2O

 $1/2 \, \mathrm{H_2}$ $\frac{1}{2} \mathrm{H_2O}$

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Nuclides, Isotopes
^{227}_{90}Th^{+}
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 $^{227}_{90}Th^{+}$ $_{-1}^{0} n^{-}$ $_{-1}^{0}n^{-}$

Parenthesis, Braces, Brackets

$$(\mathrm{NH_4})_2\mathrm{S}$$

$$(NH_4)_2S$$

 $[\{(X_2)_3\}_2]^{3+}$

 $CH_4 + 2\left(O_2 + \frac{79}{21}N_2\right)$

States of Aggregation

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H_2(aq)
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 ${\rm CO_{3}^{2-}}_{\rm (aq)}$ NaOH(aq, ∞)

Radical Dots

 $NO^{(2\bullet)-}$

OCO•-

Math Modes

 NO_x

 Fe^{n+}

 μ -Cl

 $Fe(CN)_{\frac{6}{2}}$

Escaped Texts

 CO_2^3

Reaction Arrows

 $A \longrightarrow B$

A ← B

 $A \longleftarrow B$ $A \longleftarrow B$

 $A \rightleftharpoons B$

Above/Below Arrow Text

 $\begin{array}{c} {\rm Below} \\ {\rm CH_3COOH} \xrightarrow{+ {\rm OH^-}} {\rm CH_3COO^-} \end{array}$

Relow

Precipitation and Gas

 $SO_4^{2-} + Ba^{2+} \longrightarrow BaSO_4 \downarrow$

 $A \downarrow B \downarrow \longrightarrow B \uparrow B \uparrow$

Alignments

$A \longrightarrow B$

- $B \longrightarrow C + D$

 $+ 2 H^{+}$

+ 2 OH^-

 $+ 2H^{+}$

Hydroxozikat

 $\operatorname{Zn}^{2+} \rightleftharpoons \operatorname{Zn}(\operatorname{OH})_2 \downarrow \rightleftharpoons \operatorname{[Zn(OH)_4]^{2-}}$

amphoteres Hydroxid

More Examples
$$\overset{\text{II}}{\text{CuCl}_2} + \text{K}_2\text{CO}_2 \longrightarrow \overset{\text{II}}{\text{CuCO}_2} \downarrow + 2 \text{KCl}$$

 $Hg^{2+} \xrightarrow{I^{-}} HgI_{2} \xrightarrow{I^{-}} [Hg^{II}I_{4}]^{2-}$