Chemical Formulae

 Sb_2O_3

 H_2O

Chemical Equations

 $CO_2 + C \longrightarrow 2CO$ $Hg^{2+} \xrightarrow{I^{-}} HgI_{2} \xrightarrow{I^{-}} [Hg^{II}I_{4}]^{2-}$

Charges

 H^+

 CrO_4^{2-} CrO_4^{2-}

 $[AgCl_2]^ Y^{99+}$ Y^{99+}

Oxidation States

FeFe₂O₄ FeFe₂O₄

Stoichiometric Numbers

 $2H_2O$ $2H_2O$

 $2~\mathrm{H_2O}$

 $\frac{1}{2}\,\mathrm{H}_2\mathrm{O}$

 $1/2 H_2O$

Nuclides, Isotopes $^{227}_{90}\text{Th}^{+}$

$$^{227}_{90}$$
Th $^{+}$

 $_{-1}^{0}n^{-}$

Parenthesis, Braces, Brackets

$${
m (NH_4)}_2{
m S}$$

$$(NH_4)_2S$$

$$\left[\left\{(X_2)_3\right\}_2\right]^{3+}$$

 $[\{(X_2)_3\}_2]^{3+}$

 $\mathrm{CH_4} + 2 \left(\mathrm{O_2} + \frac{7}{2} \mathrm{N_2} \right)$

States of Aggregation

 CO_3^{2-}

 $NaOH(aq, \infty)$

Radical Dots

 $NO^{(2\bullet)-}$

OCO•-

Escaped Modes

 NO_r is the same as NO_r

```
Fe^{n+} is the same as Fe^{n+}
```

Fe(CN)₆

$$CO_{2}^{3}$$

 μ -Cl

Reaction Arrows

 $A \longrightarrow B$

A ← B

11 \ 1

 $A \rightleftharpoons B$

 $A \longleftrightarrow B$

Above/Below Arrow Text

$$A \xrightarrow{\text{Detra}} B$$
 $A \xrightarrow{\text{Above}} B$

 $\begin{array}{c} \text{Below} \\ \text{CH}_3\text{COOH} \xrightarrow[+\text{H}^+]{} \text{CH}_3\text{COO}^- \end{array}$

Below

Precipitation and Gas

 $A \downarrow B \downarrow \longrightarrow B \uparrow B \uparrow$

SO₄²⁻ + Ba²⁺
$$\longrightarrow$$
 BaSO₄ \downarrow

Alignments

- - $A \longrightarrow B$
- $B \longrightarrow C + D$



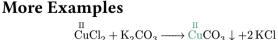
 $+2\,{\rm OH^{-}}$

 $+2 H^{+}$

amphoteres Hydroxid

 $+2\,\mathrm{OH^-}$ $\operatorname{Zn}(\operatorname{OH})_2 \downarrow \xrightarrow{+2\operatorname{OH}} \left[\operatorname{Zn}(\operatorname{OH})_4\right]^{2-}$ $+2H^{+}$

Hydroxozikat



 $Hg^{2+} \xrightarrow{I^{-}} HgI_{0} \xrightarrow{I^{-}} [Hg^{II}I_{4}]^{2-}$