

Experiment 2: Kinetics of acid catalysed hydrolysis of ethyl acetate

Reagents and solutions used: ethyl acetate, 0.5 N HCl, 0.2 N NaOH, phenolphthalein indicator, ice cubes

Apparatus: 50 mL burette, 5 mL pipette, 10 mL pipette, conical flasks, wash bottle, 250 mL reaction bottle

Tabulation and calculations:

S. No.	Time min	Volume of 0.2 N NaOH mL	$(V_{\infty} - V_t)$ mL	$\log (V_{\infty} - V_t)$	$K = \frac{2.303}{t} \log \frac{(V_{\infty} - V_0)}{(V_{\infty} - V_t)}$ min ⁻¹
1	0	27.3	24.3	1.386	0
2	10	27.8	23.8	1.377	0.00208
3	20	28.3	23.3	1.367	0.00210
4	30	28.5	23.1	1.364	0.00168
5	40	29.1	22.5	1.352	0.00192
6	50	29.5	22.1	1.344	0.00189
7	∞	51.6	-	-	-

**Plot the graph
between time and
 $\log (V_{\infty} - V_t)$**

Time min	$\log (V_{\infty} - V_t)$
0	1.386
10	1.377
20	1.367
30	1.364
40	1.352
50	1.344

Calculations

$$1. K_{10} = \frac{2.303}{10} \log \left(\frac{24.3}{23.8} \right) = 0.00208$$

$$2. K_{20} = \frac{2.303}{20} \log \left(\frac{24.3}{23.3} \right) = 0.00210$$

$$3. K_{30} = \frac{2.303}{30} \log \left(\frac{24.3}{23.1} \right) = 0.00168$$

$$4. K_{40} = \frac{2.303}{40} \log \left(\frac{24.3}{22.5} \right) = 0.00192$$

$$5. K_{50} = \frac{2.303}{50} \log \left(\frac{24.3}{22.1} \right) = 0.00189$$

Result:

The rate constant for the hydrolysis of an ester from

1. Calculated value = 0.00193
2. Graphical value = 0.00193
3. Molecularity of the reaction = 2 (bimolecular reaction)
4. Order of the reaction = first order (pseudo first order)

$$\begin{aligned}\text{Average } K &= \frac{0.00208 + 0.00210 + 0.00168 + 0.00192 + 0.00189}{5} \\ &= \frac{0.00967}{5} \\ &= 0.00193 \text{ min}^{-1}\end{aligned}$$

Graphical value of $K = 2.303 \times \text{slope}$

$$\begin{aligned}&= 2.303 \times \frac{y_2 - y_1}{x_2 - x_1} \\ &= 2.303 \times \frac{1.386 - 1.344}{50 - 0} \\ &= 2.303 \times 0.00084 \\ &= 0.00193 \text{ min}^{-1}\end{aligned}$$

$\log(v_0 - v_t)$

