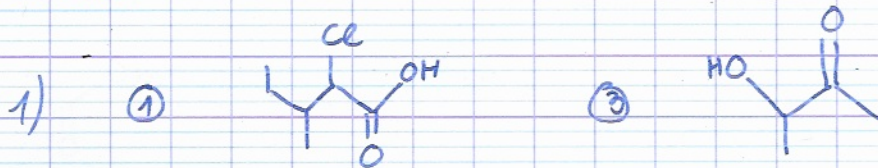


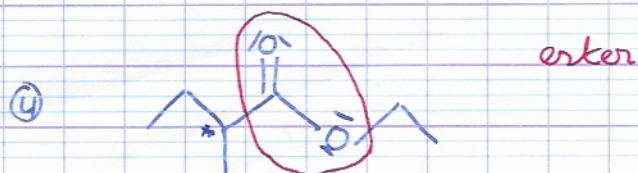
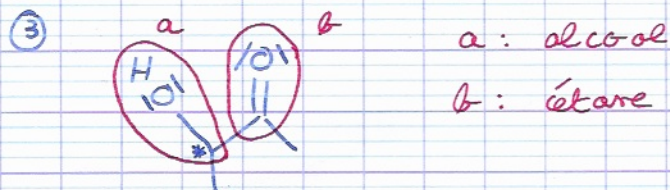
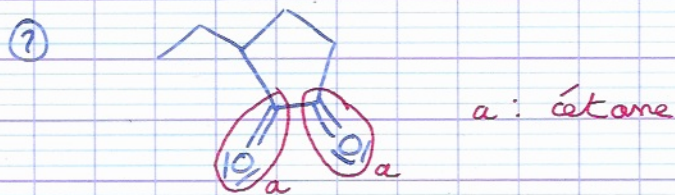
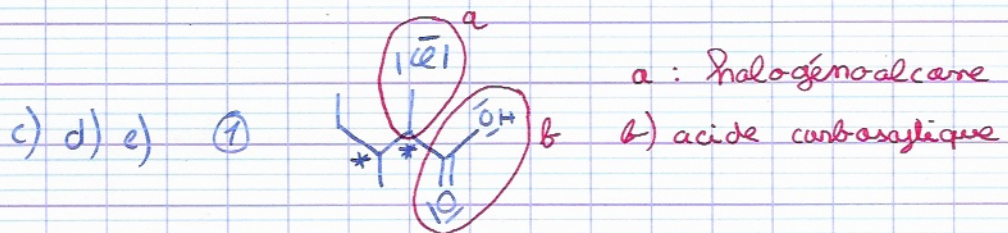
Sauv  
Paul Henry  
TDA2

## Dessin d'Atomistique

### Exercice 1



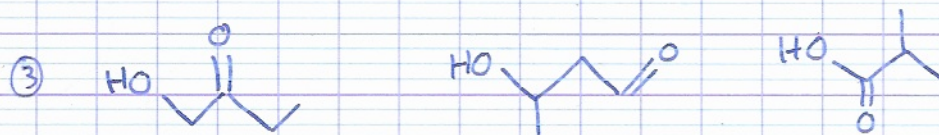
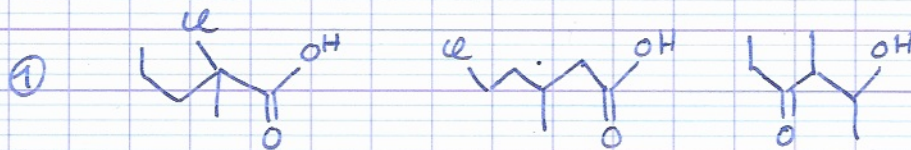
1) a) b)	Molécule	Formule brute	Nombre d'isomérisations
①		$C_6H_{11}ClO_2$	1
②		$C_7H_{10}O_2$	3
③		$C_4H_8O_2$	1
④		$C_7H_{14}O_2$	1



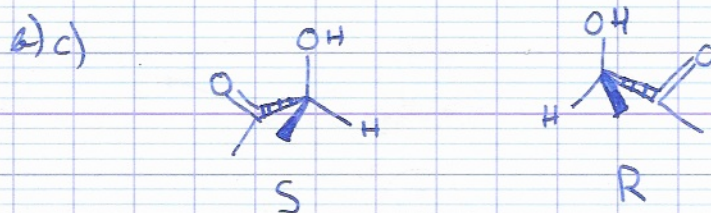


3) C'est : ethyl-2-methylbutanoate

4) 5) Chaîne Position Fonction



6) a) Elle possède un carbone asymétrique, i.e. un centre stéréogène. Elle est donc chirale.



L'ordre de priorité selon les règles CIP est :

H : 4

OH : 1

CH<sub>2</sub> : 3

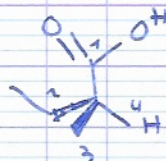
CO : 2

7) a) H : 4

CC : 2

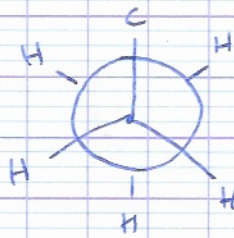
H<sub>3</sub>C : 3

HOOC : 1

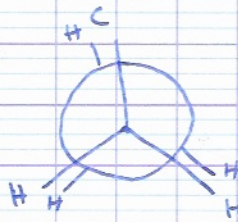




b) c) Conformation décalée (la plus stable)



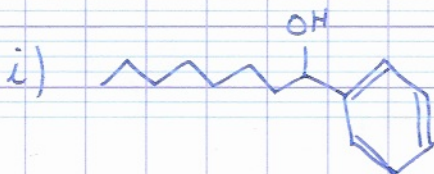
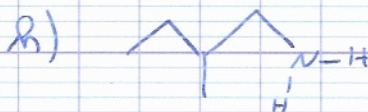
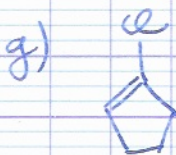
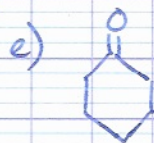
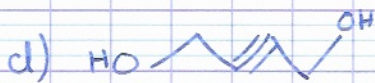
Conformation éclipée (la moins stable)



Bonus 1: Chaque carbone asymétrique crée 2 stéréoisomères de configuration. La molécule en a 2, donc  $2^2 = 4$  stéréoisomères de configuration.

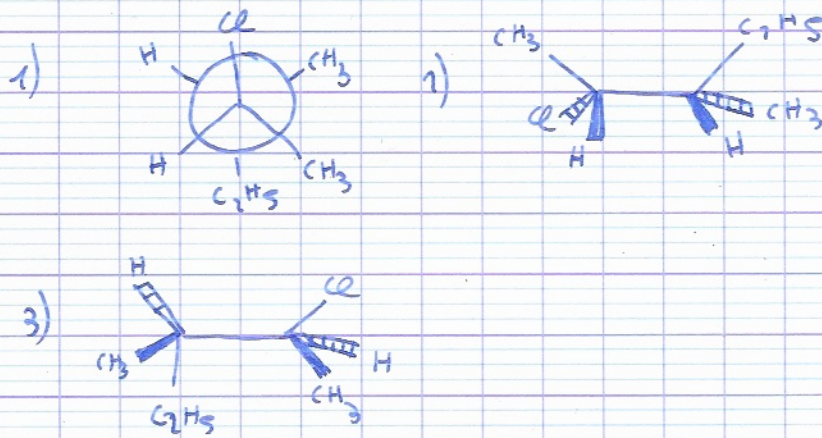
Bonus 2: 3-éthylcyclopentane-1,2-dione.

Exercice 2





### Exercice 3



### Exercice 4

