

Opto-electronic properties of WS_2

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Declaration

Acknowledgements

Abstract

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1 Introduction

Following the discovery and characterisation of graphene in last decade the focus has been put on other 2D materials. Similar to graphene other bulk layered materials can exist in a monolayer or few layer form. Furthermore these thin layers also exhibit a significant change of properties when number of layers decreases from bulk all the way to monolayer. One of the most popular groups of these materials are transition metal dichalcogenides (TMDC). Their general form is MX_2 where M is a transition metal, and X is a chalcogen atom.

1.1 Properties of TMDCs

TMDCs in their layered form have been known, studied and utilised for a long time. They can be found commonly in use as solid-state lubricants or catalysts. About 60 different TMDCs have been studied and characterised with a general formula of X-M-X where a plane of metal atoms (M) is sandwiched between two chalcogen planes (X). Out of those 40 can be considered layered materials where individual layers are strongly bonded in-plane and weakly bonded out-of-plane in between layers. These weak, interlayer, Van der Waals interactions allow to form a bulk material. These bonds are also what allows for those layers to slide on top of one another similarly to other layered materials like graphite. TMDCs consist of two transition metal and single chalcogen atoms covalently bonded. They can be found in 3 distinct structural polytypes: 1T (tetragonal symmetry, octahedral coordination) with single layer per repeat unit, 2H (hexagonal symmetry,

trigonal prismatic coordination) with 2 layers per repeat unit and 3R (rhombohedral symmetry, trigonal prismatic coordination) with 3 layers per repeat unit [42] as can be seen in Figure 1.

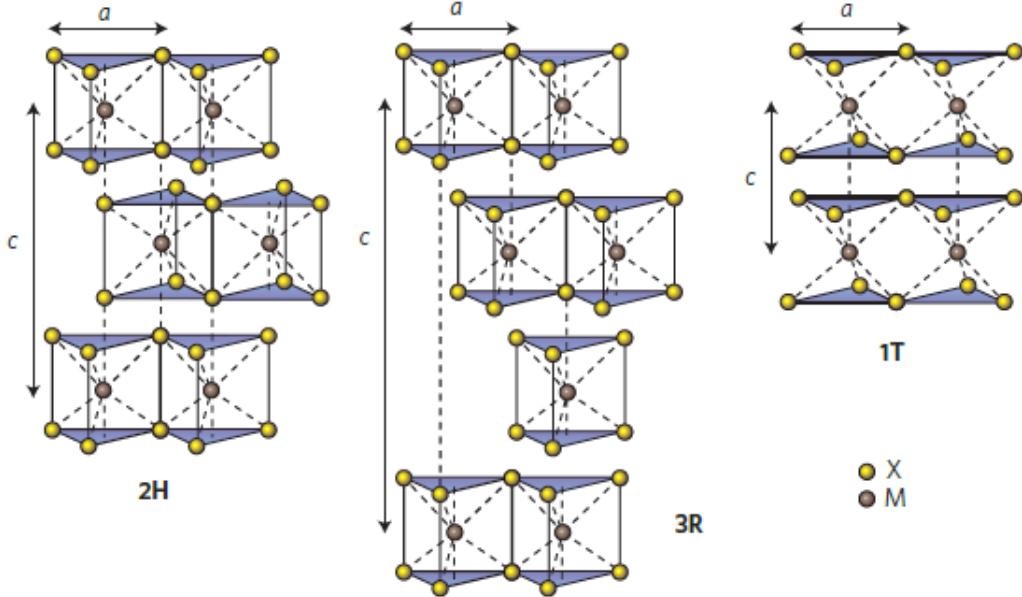


Figure 1: Schematics of the structural polytypes: 2H (hexagonal symmetry, two layers per repeat unit, trigonal prismatic coordination), 3R (rhombohedral symmetry, three layers per repeat unit, trigonal prismatic coordination) and 1T (tetragonal symmetry, one layer per repeat unit, octahedral coordination). The chalcogen atoms (X) are yellow and the metal atoms (M) are grey. The lattice constants a are in the range 3.1 to 3.7 Å for different materials. Adopted from [42]

Since graphene have proven to be difficult to work with in the fields of semiconductors due to its lack of natural finite electronic band gap its role as a successor in electronic and opto-electronic devices remains to be seen. However the techniques learned and effects observed during its characterisation were easily transferred to other layered compounds such as TMDCs. In particular the semiconducting, group VI-based TMDCs, containing sulphur and selenium as chalcogen atoms have proven to be more readily potentially useful as an active material in electronic and opto-electronic devices. This is due to their inherent electronic and optical bandgap in visible-near IR range.

As the number of layers changes from bulk to monolayer the properties of the TMDC undergo a significant change. In most TMDCs the bandgap changes from indirect to a larger direct one.

1.2 Electronic properties

One of the most interesting features that the layered TMDC materials exhibit is the shift in the bandstructure with the changing number of layers. Several studies have shown in simulations and experimentally that TMDCs have very similar electronic band structure as seen in example of WS_2 in Figure 2. In bulk WS_2 the maximum of the valence band (VBM) at Γ point and the minimum of the conduction band (CBM) at Λ form an indirect bandgap. As the number of the layers decreases the CBM at Λ point as well as VBM at K point increases causing the band gap to widen. At 2 layers the K point becomes the actual CBM and a new indirect bandgap forms between Γ point and K point. Finally in a WS_2 monolayer the VBM at K point as well as entire conduction

band increases to form a new greater direct band gap at K point. This means that WS_2 bandgap changes from 1.3 eV indirect bandgap in bulk to 2.1 eV direct bandgap in monolayer.

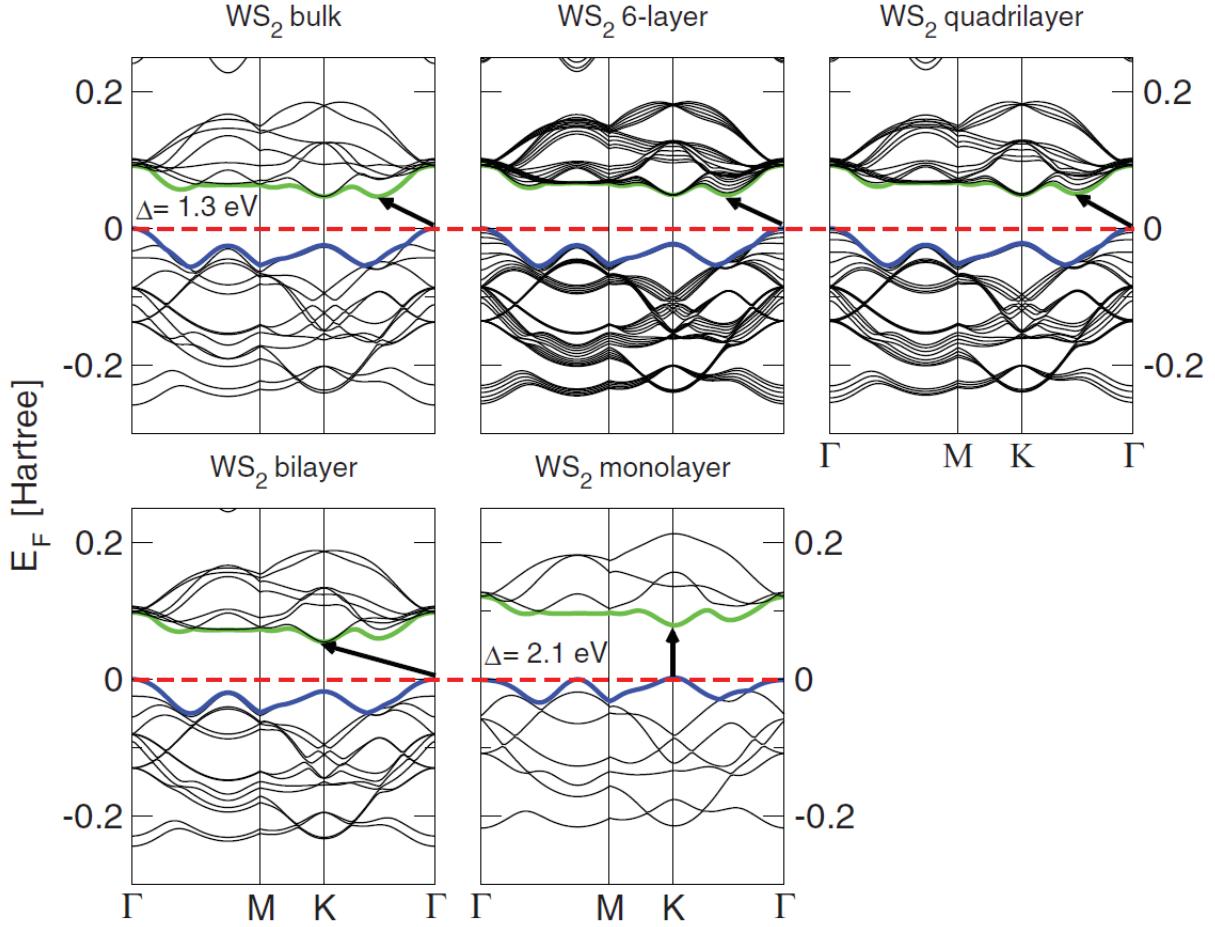


Figure 2: Band structures of bulk WS_2 , its monolayer, as well as, polylayers calculated from the density functional theory (DFT) simulation. The horizontal dashed lines indicate the Fermi level. The arrows indicate the fundamental band gap (direct or indirect) for a given system. The top of valence band (blue) and bottom of conduction band (green) are highlighted. Adopted from Ref. [21]

Like WS_2 other Mo and W based TMDC undergo similar transitions as seen in Table 1. In all cases the smaller indirect bandgap changes to greater direct bandgap with monolayer bandgap ranging from 1.1 eV to about 2.1 eV. Moreover the VBM at K points exhibits the orbit-spin band splitting at the K point of about 400 meV. This direct bandgap leads to presence of A and B excitons generated by transition between CBM and two VBM at the K point. The conduction band as well as the valence band are dominated by the d-electron orbitals of the transition metal atoms and at the VBM and CBM they hybridize with the p-electron orbitals of the chalcogenide atoms. Because the hybridization happens mostly at the Γ point and the chalcogenide atoms are at the surface of the TMDC layer it leads to strong interactions between the layers. This leads to significant change in the band structure at the Γ and rise of the indirect bandgap as a result of increased number of layers. On the other hand at the K point the d-orbitals of the transition metals remain mostly unaffected due to them being positioned in the middle of the layer [21] [39]

Table 1: Mo and W based TMDC bandgaps comparison

M\X	$-S_2$	$-Se_2$	$-Te_2$
Mo	Semiconducting	Semiconducting	Semiconducting
	1L: 1.8 eV	1L: 1.5 eV	1L: 1.1 eV
	Bulk: 1.2 eV	Bulk: 1.1 eV	Bulk: 1.0 eV
W	Semiconducting	Semiconducting	Semiconducting
	1L: 2.1 eV	1L: 1.7 eV	1L: 1.1 eV
	Bulk: 1.4 eV	Bulk: 1.2 eV	

1.3 Optical properties

TMDCs exhibit a wide array of opto-electronic effects due to their strong light-matter effects. These effects are mostly caused by the abundant presence of excitons, bi-excitons, trions or bound excitons. As a result the change in layer thickness from bulk to monolayer alters the photoluminescence, photoconductivity and absorption in the visible to infrared range.

The primary and most common quasi-particle that forms in such system is an exciton, which is a pair of a negatively charged electron and a positively charged hole bound together by Coulomb forces to form a structure similar to that of hydrogen atom. Such pair is electrically neutral and is of size exceeding size of single cell which makes it a Wannier–Mott exciton. The recombination of these excitons results in a photon emission which can be easily observed during photoluminescence characterisation. On top of excitons other quasi-particles such as trions, bi-excitons or bound excitons can be found. A trion is a group of 2 electrons and a hole or 2 holes and an electron, or otherwise described as a charged exciton. The exact nature of the trion depends usually on the type of intrinsic doping of the TMDC. A bi-exciton is a pair of excitons which is usually only observed in quantum dot systems but can be also seen in excitonically dense systems such as TMDCs. A bound exciton is similar to the free exciton but is trapped by a defect. In a typical photoluminescence spectrum several peaks can be observed depending on specific type of TMDC characterised. In WS_2 monolayer for instance as seen in Figure. 3 the strongest peak (often labelled as an A peak) at about 1.97 eV is caused by the direct transition of single-photon generated exciton. Slightly redshifted by about 30 meV from the A peak a generally weaker peak caused by the trion recombination can be found. At higher energies another peak can be observed due to the presence of bi-excitons. At around the 1.3 eV a much weaker peak (I) can be seen caused by the indirect transition. Additionally a B peak can be observed blueshifted from the A peak which is caused by valley splitting as discussed in chapter 1.2. As seen in Figure 3 as the number of layers increases the main A peak becomes dramatically weaker due to lack of direct transition and redshifted following the pattern discussed in chapter 1.2. At the same time the I peak becomes relatively stronger and eventually dominates the bulk material.

Figure 3: Typical PL spectra of WS_2

Similarly the photoconductivity of the TMDCs is strongly reliant on the number of layers and incident photon energy. The MoS_2 for instance shows 3 times stronger photoconductivity in monolayer around 1.8 eV, where the direct transition is located, than in 2L MoS_2 around 1.6 eV.

Additionally the photoconductivity appears to increase in steps with relation to the photon energy following the direct and indirect transitions. [42].

The sunlight absorption in TMDCs has been shown to be significantly more intense than in commonly used solar cell materials, at about 5-10% which is an order of magnitude greater compared to similar thickness of Si or GaAs. It is also stronger compared to 2-3% of sunlight absorption of graphene. As a result a excitonic solar cell based on MoS_2/WS_2 bilayer shows about 1% power efficiency, about 3 times greater than that of typical ultrathin solar cells [2].

During standard single photon excitation photoluminescence studies the excitons generated can be called "bright" since they appear in PL spectrum. The reason we can observe them easily is because the spin between an electron and a hole is conserved, and thus allowing for photon emission. However another combination is possible, called dark exciton, where both electron and the hole have the same spin. Because of that they cannot recombine by emitting a photon and therefore remain absent from the PL spectrum. Even though they exist much longer than their bright counterparts their presence is of course also more difficult to observe. One way to observe them is to use two photon excitation. Due to two photon selection rule the single photon excitation can be excluded and the dark excitonic states can be observed. In WS_2 the dark excitons result in two peaks at 2.28 eV and 2.48 eV. [Probing excitonic dark states in single-layer tungsten disulphide]

Defect engineering allows to tune the number of charge carriers. In MoS_2 or WS_2 the sulfur vacancies lead to increased number of electrons in the material. Because of that by increasing the number of defects in those materials the level of n-doping can be changed. An easy way to observe the presence of those defects and subsequent quenching of them is to expose the material to varying amounts of oxygen, nitrogen or water. Due to greater electronegativity (???) those species attract the electrons and therefore the electron population in the material decreases. This in turn leads to smaller trion population since trions require an extra electron to form. This then can be observed in PL as a more narrow direct peak, with especially smaller redshifted shoulder. The effect can also be of course reversed by decreasing the amount of oxygen, nitrogen or water in the environment since those exist in already in ambient conditions [Optical control of charged exciton states in tungsten disulfide].

In order to introduce and control the amount of vacancies in the TMDC different method have to be explored. One of the ways of achieving that in already grown material is the use of oxygen plasma. It has been shown that the number of defects can be controlled by limiting the plasma exposure. During the process the oxygen also chemically bonds to the MoS_2 at the defect sites and therefore partially negates the effect of defects on the optical properties. The PL can also be seen to increase in intensity with increasing number of defects with oxygen adsorbed due to the increased yield of bound excitons localised at these defects. [Strong Photoluminescence Enhancement of MoS_2 through Defect Engineering and Oxygen Bonding]

Similar effect has been shown using the 2,3,5,6-tetrafluoro-7,7,8,8-tetracyanoquinodimethane (F_4TCNQ), 7,7,8,8-tetracyanoquinodimethane ($TCNQ$) and (nicotinamide adenine dinucleotide) $NADH$ for chemical doping. Both F_4TCNQ and $TCNQ$ are p-type dopants while $NADH$ is a n-type dopant. By exposing the surface of MoS_2 to these compounds the change in PL intensity and FWHM have been observed. Similar to doping with O_2 , N_2 or H_2O , all of which are p-type dopants, the intensity of PL has increased in presence of F_4TCNQ and $TCNQ$. The effect has been similarly ascribed to lowering the number of defects and therefore the lowering the trion population and subsequently increasing the exciton population increasing the yield. The opposite observation has been made with use of $NADH$ with PL intensity decreasing. Similarly the increase in trion population with lower PL yield is ascribed to the lower PL intensity. [Tunable Photoluminescence of Monolayer MoS_2 via Chemical Doping].

It has also been shown that alloying can be used to fine tune the PL by varying the concentration of alloying material. In monolayer $Mo_{1-x}W_xS_2$ the PL peak position initially decreases from 1.575 eV (PL peak position of pure MoS₂) to 1.56 eV at x=0.21 and then increase up to 1.65 eV (PL peak position of WSe₂) at x = 1. This effect could be attributed to the linearity of VB and non-linearity of CB with regards to change in W composition. The PL position can therefore be engineered on a monotonic range from 1.56 eV to 1.65 eV. In bilayer $Mo_{1-x}W_xS_2$ alloy the position of both direct and indirect transition PL peaks increases monotonically from about 1.49 eV and 1.53 eV for pure MoSe₂ to 1.56 eV and 1.62 eV for pure WSe₂ as the W amount is increased. This opens another relatively easy way of engineering PL position [Two-Dimensional Molybdenum Tungsten Diselenide Alloys: Photoluminescence, Raman Scattering, and Electrical Transport].

Another effect that has been demonstrated that allows for certain degree of control of PL in TMDCs is relation between the helicity of incident light and valley population valley population. It has been shown that by exciting the monolayer MoS_2 with right-polarised light the resulting excitons will fill primarily the VB at K point. Similarly by exciting the MoS_2 with left-polarised light the excitons will fill the VB at K' point. After recombination the resulting photons will exhibit the same circular polarity as the photons that excited the electrons in the first place. [Tightly bound trions in monolayer MoS₂] [Control of valley polarization in monolayer MoS₂ by optical helicity]

The temperature effect on TMDCs has also been investigated. In WSe_2 monolayer it has been shown that as the temperature of the sample increases from room temperature to about 400K the position of the direct transition PL peak redshifts from about 1.65 eV to about 1.58 eV. When the temperature is decreased from room temperature to about 5K the same peak blueshifts to about 1.7 eV. Between 100K and 50K as well 20K and 5K the position of the PL peak does not change. Additionally around 120K another peak appears and as the temperature is lowered it also blueshifts although less than the RT peak. The peak only present at RT is attributed to free excitons whereas the peak appearing at 120K is ascribed to bound excitons. As bound exciton peak appears its intensity increases with lower temperature while the intensity of the free exiton peak decreases. This indicates that the population of free excitons decreases while the population of the bound excitons increases with decreasing temperature [40]

There has been many reports on the spatial distribution of PL in the TMDCs. One of the observed patterns in WS₂ and MoS₂ has been that of much stronger PL intensity at the edges of the flakes. That effect has been primarily observed in small flakes of about 5 μm [12].

1.4 Phonon dispersion

The vibrational and phononic characteristics of TMDCs have been investigated at length by both theoretical simulations as well as experimental studies. The $2H - MX_2$ crystal structure of the TMDCs belongs to D_{6h}^4 point group and there are 18 lattice dynamical modes at the Γ point. Phonons belonging to these modes can be represented as Eq. 1 [43]:

$$\Gamma = A_{1g} + 2A_{2u} + B_{1u} + 2B_{2g} + E_{1g} + 2E_{1u} + E_{2u} + 2E_{2g} \quad (1)$$

In TMDCs 4 active Raman modes can be observed E_{1g} , E_{2g}^1 , E_{2g}^2 , A_{1g} . These can be seen in Figure 4. The E_{2g}^2 is a shear mode that involves 2 layers vibrating against each other. The E_{1g} is an in-plane vibration of chalcogen atoms but is forbidden in the back-scattering configuration. For monolayers therefore it leaves primarily the E_{2g}^1 which is an in-plane mode involving vibration of both metal and chalcogen atoms as well as A_{1g} which is an out-of-plane mode involving only chalcogen atoms.

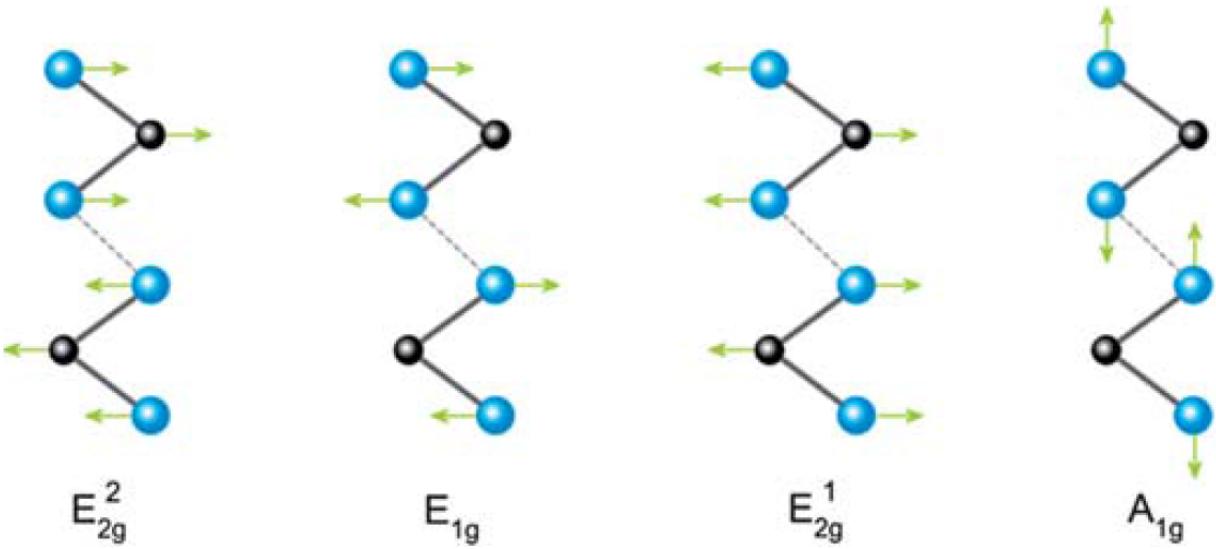


Figure 4: 4 active Raman modes in TMDCs. Metal atoms and chalcogen atoms are black and blue respectively. Adopted from [43]

These two peaks tend to dominate the spectrum of any TMDCs, whether monolayer or few-layer or bulk. The shear mode E_{2g}^2 appears at low Raman shift frequencies and is therefore difficult to observe but can be used to differentiate monolayer from few-layer material. Since E_{2g}^1 is an in-plane mode it tends to be unaffected by the number of layers due to weak van der Waals forces between the layers but can be seen to be slightly redshifted as the number of layers increases. As seen in Figure 5 the E_{2g}^1 peak at about 352 cm^{-1} is overlapping with another stronger peak, a 2LA(M) peak at 350 cm^{-1} which is a longitudinal acoustic mode caused by in-plane collective oscillations of W and S atoms. The second strongest peak at around 416 cm^{-1} is an A_{1g} peak, caused by out of plane vibrations. Because of that it is much more sensitive to the number of layers and is seen to become blueshifted as the number of layers increases. This has been attributed to the restorative forces as well as increase in dielectric screening of the Coulomb forces. Combining both of these shifts in frequency with the changing number of layers the difference between these two peak position can be used to identify the number of layers in TMDCs as seen in Figure 6

Figure 5: Typical Raman spectrum of WS_2

Figure 6: Identification of number of layers by the difference in position of A_{1g} and E_{2g}^1 peaks.

1.5 Heterostructures

1.6 Applications

2 Methods

2.1 Raman spectroscopy theory

2.2 Photoluminescence spectroscopy theory

2.3 XPS theory

3 Role of precursors in growth of monolayer WS_2

In this chapter the CVD growth of WS_2 using different precursors is investigated. The as grown samples were characterised by Raman and PL spectroscopy as well as XPS, XRD, AFM and electrical measurements. As a result it was concluded that using $H_2WO_4 + NaCl$ at 850 °C gives best results. Such grown samples exhibit biggest flakes up to 200 μm in size as well as show the strongest and most narrow PL peak with 36 meV FWHM. The samples grown using $H_2WO_4 + NaCl$ at 950 °C show also best transistor electron mobility in monolayer CVD grown WS_2 while that grown using $WO_3 + NaCl$ at 950 °C shows best transistor electron mobility in bilayer CVD grown WS_2 .

3.1 Introduction

Monolayers of transition metal sulphides and selenides exhibit range of interesting properties such as strong light absorption in the IR and visible range [31][2][53], valley polarisation [30] [15], spin-orbit interactions [47][54], tightly bound excitons [29] or second-harmonic generation [25]. Some of these effects can be attributed to the lack of free dangling bonds and configuration of d-orbitals [44], [24].

Among these materials one of the most promising is the WS_2 . Its visible range bandgap of 2eV as well as an easy and safe manufacturing route via CVD makes it one of the more interesting and studied TMDCs. The typical characterisation by photoluminescence spectroscopy allows to probe the varying synthesis conditions, the grain boundaries or defect population [12] [34] [23] [37]. The PL efficiency in as-grown monolayer WS_2 produced via CVD growth shows ~2-6% efficiency [34][49] [26]. This efficiency is caused mostly by defect-mediated non-radiative recombination centres [1]. LEDs have been successfully produced [18] showing external quantum efficiency up to 10% [51][45]. WS_2 is typically a n-type semiconductor due to the presence of sulphur vacancies [12][38][17]. In order to utilise this material in any potential future applications a reliable and scalable manufacturing method must be developed to ensure a high quality crystal on the wafer scale area. The main method for WS_2 synthesis that satisfies these conditions is Chemical Vapour Deposition (CVD) [13]. The growth of tungsten based TMDCs have been less successful than the equivalent molybdenum based TMDCs and has produced mostly isolated flakes of up to 40 μm [12] [52] [37] [7][48][8][22]. Even larger films of monolayer WS_2 have been shown, however they also exhibit low carrier mobility [19][9]. In the typical CVD synthesis process the sulphur and tungsten oxide are evaporated simultaneously in a tubular furnace with a constant flow of carrier gas like argon at temperatures of at least 900 °C [12][52][37][7][48][8][22]. Such growth is predi-

cated by topotactic transformation leading to low density distribution of domains on an amorphous [12][52][7][8][22] or crystalline substrate [37][48][33] possibly due to low evaporation rates of WO_3 . Since WO_3 requires high temperatures of 950-1000 °C to evaporate while the S becomes volatile at 90 °C the thermodynamics of the process are difficult to control. The low growth dictated by fast evaporation of S leads to limited domain growth and lack of continuous layer. One of the proposed solutions have been to spread the WO_3 on the target substrate [34][23][9][5][50][11][10]. This has however led to low reproducibility, poor control of thickness and stoichiometry and unreacted material left on the substrate. Another approach has been to use more volatile W precursors such as WCl_6 [3] or $W(CO)_6$ [19][6] together with organic compounds as S precursors. Such method while producing a large area domains at lower temperature has led to lower crystal quality and purity.

Here we propose a different method of CVD synthesis that allows for much larger flake growth of up to 800 μm at temperature of 750 degreeC. Such grown material exhibits high electron mobility in one and two layers of WS_2 , higher than other values reported in literature. The photoluminescence peak is also very narrow at 36 meV FWHM at room temperature.

3.2 Results

For the purpose of comparing the CVD synthesis method conditions the several sets of precursors were used: WO_3 , $WO_3 + NaCl$ and $H_2WO_4 + NaCl$. The standard growth procedure involves two separate crucibles placed at distance from each other in a quartz tube. Each of these crucibles is independently heated to ensure that the S and the W precursors evaporation rate is maximised at the same time. The vapours then are deposited on SiO_2/Si (285 nm) substrate which is placed close to the W precursor crucible. The entire process is performed under low vacuum and a supply of Ar gas. By following this method a reproducible deposition of large area flakes can be shown. As seen in Figure 7 the size of the flakes increases from left to right as the precursors used (WO_3 , $WO_3 - NaCl$ and $H_2WO_4 - NaCl$) change as well as demonstrating the lower temperature required to achieve these growths. All of these growths result in formation of triangular flakes with sharp edges and uniformity of colour throughout which suggest a high quality, pristine material across the flake. The growth using only WO_3 results in formation of small flakes of 10 μm at 950 °C while no growth occurs at lower temperatures (Figure 7). This can be explained by the high sublimation temperature of WO_3 .

The growth can also be observed using $WO_3 + NaCl$ at 950 °C and 850 °C with the former showing flakes of size of about 60 μm while the latter showing smaller flakes of about 30 μm in size. To explain this difference in size a Robinson & Robin model can be used which states that at higher temperatures the diffusivity of the adsorbed precursors is favourable to the expansion of the existing domains. On the other hand the desorption of the adsorbed species is high leading which limits the supersaturation and formation of new domains. If the temperature is lowered even further to 750 °C then no growth is observed at all, most likely due to slow evaporation of WO_3 precursor.

With the change of precursors from WO_3 to H_2WO_4 a signifacnt change in size of flakes is observed at 850 °C and higher temperatures with lengths exceeding 200 μm . Additionally the growth have been shown to occur at 750 °C with flakes of the size of 50-200 μm . Moreover continuous monolayer areas of up to 0.8 mm in size has been shown (Figure 16). By increasing the growth pressure (1.6 mbar to 13 mbar) at 950 °C bilayer WS_2 flakes can be preferentially formed (Figure 17).

By replacing WO_3 with H_2WO_4 the lateral size of the WS_2 monolayered domains significantly increases (Figure 1g,h,i). The triangular crystals have edge lengths exceeding 200 μm

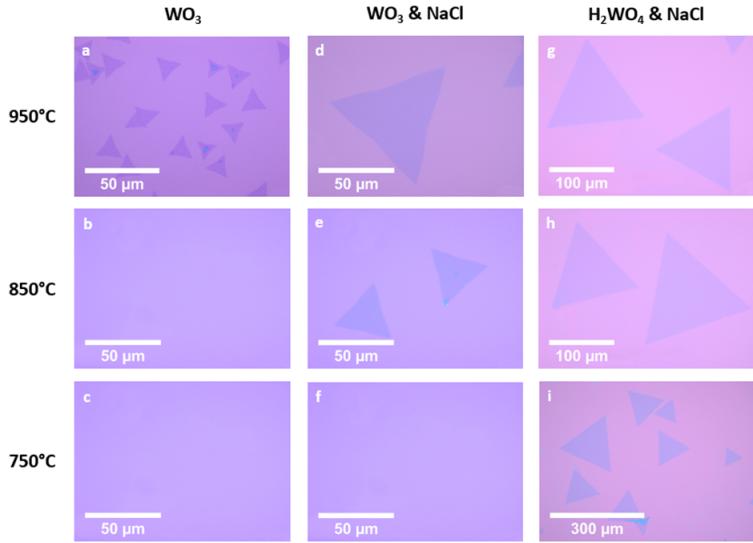


Figure 7: Optical micrographs of WS_2 triangles grown on SiO_2/Si substrates at different temperatures and using different precursors: (a) WO_3 at 950 °C; (b) WO_3 at 850 °C; (c) WO_3 at 750 °C; (d) $WO_3 + NaCl$ at 950 °C; (e) $WO_3 + NaCl$ at 850 °C; (f) $WO_3 + NaCl$ at 750 °C, the OM appears as a bare SiO_2 substrate; (g) $H_2WO_4 + NaCl$ at 950 °C; (h) $H_2WO_4 + NaCl$ at 850 °C; (i) $H_2WO_4 + NaCl$ at 750 °C.

at temperatures higher than 850 °C and between 50-200 μm at 750 °C (Figure 1g,h,i). Continuous polycrystalline monolayer coverage has been obtained over areas of ~ 0.8 mm extension (Figure S2). Increasing the growth pressure (from 1.6 mbar to 13 mbar) at 950 °C bilayered WS_2 flakes are preferentially formed (Figure S3). To understand the facilitated synthesis of WS_2 using H_2WO_4 and $NaCl$, we conducted X-ray diffraction (XRD) analysis of the reaction products between H_2WO_4+NaCl and WO_3+NaCl systems at different temperatures (500 °C, 650 °C and 750 °C) to understand the chemical differences (Figure S4). We found that the main products of the reactions between $NaCl$ and H_2WO_4 are: $NaxWyOz$ and tungsten oxychloride ($WClO_4$ and WO_2Cl_2). The $NaxWyOz$ possesses a high evaporation temperature as it remains in the crucible (Figure S5) after the synthesis of WS_2 is completed. Further, using this compound as precursors for a new growth of WS_2 at 950 °C did not lead to the formation of any WS_2 flakes, confirming the high evaporation temperature. On the bases of previous studies on the synthesis of bulk crystals, the formation of tungsten oxychloride species (WO_2Cl_2 and/or $WOCl_4$) is likely to occur while the formation of metal halides is less favourable (e.g. WCl_6) [] []. Tungsten oxychlorides are volatile already at 200 °C [] and they can be sulfidized in vapour phase and then be deposited onto the target substrate as atomic clusters. $WOCl_4$ has been previously used [43] as precursor for the CVD synthesis of WS_2 bulk films. Despite its strong tungsten oxygen double bonds, $WOCl_4$ proved to be an effective precursor with a clean decomposition pathway in the CVD process without formation of tungsten oxysulfide. We have verified that using this precursor is indeed possible to obtain WS_2 at temperatures as low as 550 °C (Figure S6a). The key role played by the oxyhalide species it becomes apparent if we try to grow WS_2 by using only hydrated tungsten oxide. As this decomposes to form WO_3 , only small WS_2 domains are observed with similar PL characteristics to the WO_3 precursors-growth (Figure S6b). Furthermore, to confirm the key role played by Cl, we replaced $NaCl$ with KCl and we obtained comparable growth results (Figure S6c).

The flakes were investigated using HRTEM to confirm high crystallinity of the material (Figure 8). The lattice constant measured in this way has been found to be 0.3 nm, which is consistent with that of 2H-WS₂ (0.318 nm). The AFM characterisation allowed to confirm the presence of monolayer (Figure 8) and bilayer (Figure 17 flakes with the step of 0.8 nm [?][36].

The Raman spectroscopy characterisation of WS₂ flakes obtained under different growth conditions can be seen in Figure 8. All of these spectra exhibit the main 2 peaks at $\sim(351\pm0.53)$ cm⁻¹ and $\sim(417.6\pm1)$ cm⁻¹. The latter peak corresponds to the A_{1g} vibrational mode while the former peak can be further resolved into two peaks, one related to 2LA vibrational mode and the second to E_{2g}¹. As seen in the Figure 21 and Figure 23 the distribution of intensity of both main peaks is uniform across the flakes. Similarly the difference between peak positions of A_{1g} and E_{2g}¹ is also uniformly distributed as seen in Figure 25 and is equal to $\sim(66.5\pm0.53)$ cm⁻¹ which is indicative of monolayer [46].

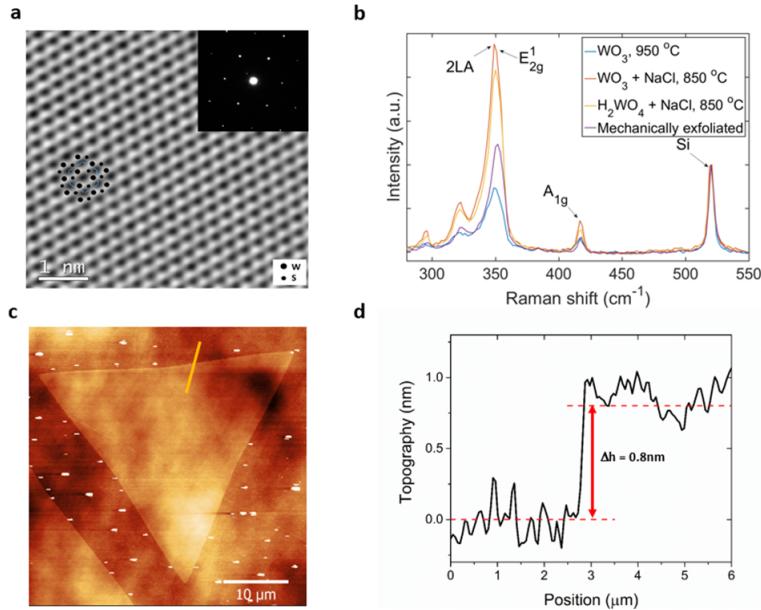


Figure 8: Structural and physical characterization of WS₂ triangles: (a) HRTEM image of the WS₂ lattice grown using $H_2WO_4 + NaCl$, the inset report a selected diffraction area which show a hexagonal pattern. (b) Raman spectra showing the characteristics active modes of WS₂ grown under different conditions and compared with mechanically exfoliated flakes. (c) AFM image and (d) corresponding thickness profile of monolayer WS₂.

As seen in Figure 9 maps of PL intensity has been collected for samples grown using different growth condition. The sample grown using WO_3 at 950 °C appears to be trisected into 3 symmetrical areas. The intensity is lowest along the the trisecting lines and highest in the centre of each of the sub-triangles. The PL peak position is inversely proportional to the intensity of the PL peak as seen in Figure 26. The distribution of the PL peak positions seems to be bimodal with maxima at 1.96 eV and 1.94 eV. The FWHM is distributed mostly uniformly with few small areas of greater and few of smaller FWHM (Figure 27 with the average value of 65 meV. These peak position and FWHM values are comparable with those reported in literature [12][37][14][20]. The PL spectra taken from one of the three smaller parts of the flake are redshifted by about 0.02 eV and are wider than average. This can be explained by the presence of defects, in particular sulfur vacancies, which effectively cause the WS₂ to be n-doped [16][35]. This in turn increases the trion

population which introduces new peak, redshifted by about 30 meV, which results in overall shift and broadening of the PL peak [41][12]. Additionally such localised changes in PL peak intensity and position can be caused by local strain [28][16]. To investigate the potential effect of strain on the observed PL pattern a flake was cut using high power 532nm laser. As seen in Figure 28 the resulting pattern remains the same as before treatment. Therefore the observed pattern is most likely caused by the local variation in defect density [27].

By introducing the NaCl and mixing it with WO_3 precursor WS_2 samples are grown that exhibit higher energy and narrower PL peaks. The spectra are asymmetric and can be therefore deconvoluted to obtain an exciton and trion component (Figure 10). The spatial distribution of the PL position and width is mostly uniform throughout the flake. However the distribution of PL peak positions and FWHM across different flakes grown at the same conditions is bimodal (Figure 10). The PL peak position is found to be ($\sim 1.95 \pm 0.002$ eV and 1.96 ± 0.002 eV) while the FWHM is ($\sim 43 \pm 2.8$ meV and 51 ± 3 meV) which is smaller than most reported works [12][37][14][20]. This can be explained by smaller trion component in the PL spectrum which in turn means that WS_2 samples grown at these conditions contain less structural defects compared to the pure WO_3 growth. A difference in growth mechanism (topotactic or molecular conversion) which results in different defect distribution can be attributed to the differences in samples.

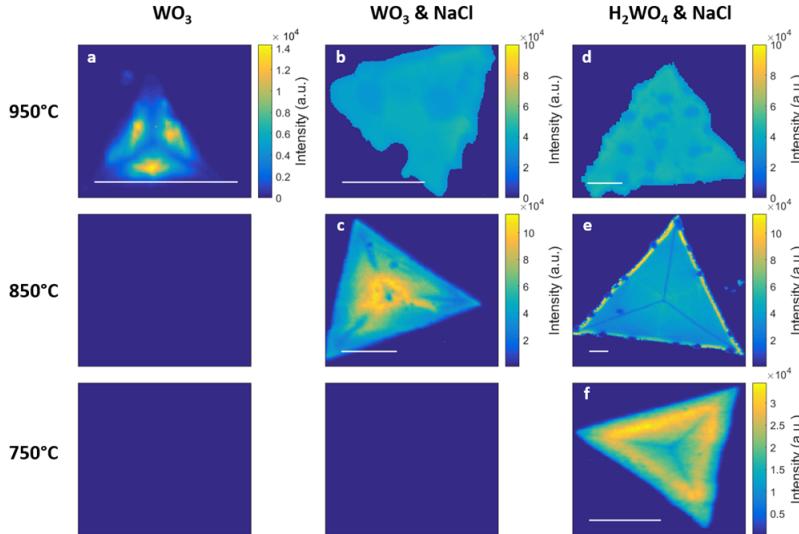


Figure 9: Spatial maps of PL intensity of WS_2 grown in the conditions exemplified in Figure 1. The scale bar length is $10\mu m$.

By changing the WO_3 precursor to the H_2WO_4 precursor a further shift in PL peak position, as well as increase in PL intensity and narrowing of the PL peak is observed as seen in Figure 9, 10. The PL peak position distribution across the flake is narrow and is ($\sim 1.980 \pm 0.005$ eV) which is higher than that of the WO_3 and $WO_3 + NaCl$ systems. The FWHM is found to be ($\sim 36 \pm 3$ meV) which is also smaller than that of the WO_3 and $WO_3 + NaCl$ systems. The PL peak width is also smaller than that reported in literature for CVD grown [12][37][14][20] exfoliated WS_2 [53][38]. At the same time it is comparable to the WS_2 grown on van der Waals substrates [33] as well as mechanically exfoliated WS_2 [18]. Additionally the peak position as well as FWHM variation is also smaller than that of other samples (5meV and 3meV respectively) (Figure 32). The PL peak intensity map (Figure 9) shows that the spatial distribution of intensity is also much more homogeneous with faint weak pattern of three trisecting lines still visible.

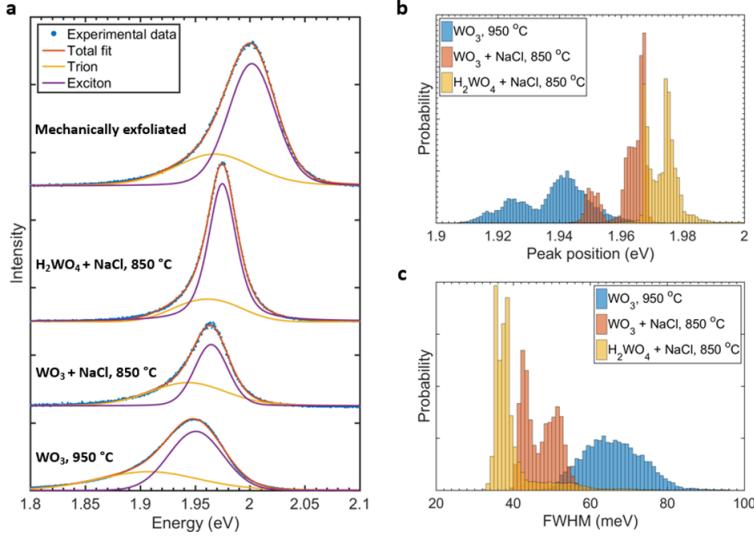


Figure 10: PL spectra characteristics of WS₂ grown using: WO₃ at 950 °C, WO₃+NaCl at 850 °C, H₂WO₄+NaCl at 850 °C: (a) individual spectra (dotted line) and deconvolution in exciton and trion components; (b) distribution of PL peak position and (c) distribution of PL FWHM for several WS₂ grown using the three different precursor systems.

The chemical composition of the flakes grown using $H_2WO_4 + NaCl$ has been investigated using XPS. The W 4f5/2 and W 4f7/2 core levels show peak positions that of the W^{4+} in WS₂ [4][32] (32.7 eV 34.8 eV respectively) with FWHM of 1 eV (Figure 11, the smallest possible using Mg K α x-ray source. It means therefore that the WS₂ is of perfect stoichiometric ratio of W and S. Additionally by integrating the intensity of the W 4f and S 2p core level peaks the same result is achieved. Furthermore the S 2p1/2 and 2p3/2 core levels are also present at the expected position for WS₂ (162.3 eV and 163.4 eV respectively, Figure 11) and small FWHM of 1 eV [32]. Small peaks indicative of W^{6+} (W 4f5/2 and W 4f7/2 at 35.9 eV and 38.1 eV respectively, Figure 11) can be attributed to presence of WO₃ and which partially overlaps with W 5p core level (38.5 eV). These peaks however disappear after the flakes are transferred to another Si/SiO₂ substrate, suggesting that they can be caused by residue WO₃ given that the XPS spot size is relatively big (~ 1 mm). The FWHM of the W 4f core levels was unchanged after the transfer, indicating that the crystallinity and the quality of the flakes was preserved with no extra defects introduced.

The WS₂ samples grown using WO₃ + NaCl have been characterised by XPS as well and the stoichiometric ratio of 2:1 for S:W has been found. However the W^{6+} component, caused most likely by presence of WO₃ (W 4f5/2 and W 4f7/2 at 35.9 eV and 38.1 eV respectively) is more pronounced indicating incomplete sulfurisation. Similarly to the sample grown with H₂WO₄ + NaCl this component disappears entirely upon transferring the sample onto new Si/SiO₂ indicating it is present around the flakes, distributed across the substrate. The W^{4+} 4f core levels peak width is ~ 1.2 eV which indicates higher defect concentration compared to that of the sample grown using H₂WO₄ + NaCl. This FWHM however increases after the transfer, which is most likely caused by the increase in the concentration of the defects caused by the mechanical stress during the transfer. In conclusion the H₂WO₄ is shown to be a better precursor compared to WO₃.

The samples were further characterised for their electrical properties. Bottom-gated field effect transistors were prepared as seen in Figure 12. The FET transfer curve indicates accumulation-type n-channel transistor, where the drain current increases with applied gate bias after a thresh-

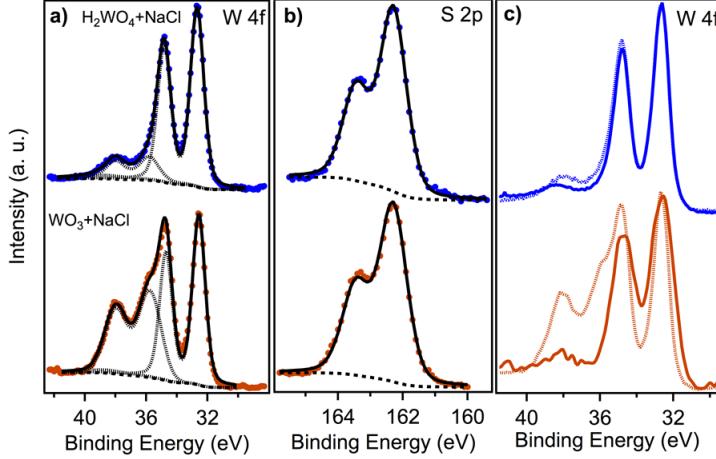


Figure 11: XPS spectra of the W 4f and S 2p core level peak regions. (a) Comparison of W 4f_{5/2} , W 4f_{7/2} and W 5p core levels of WS₂ grown using H₂WO₄+NaCl at 950 °C (blue spectrum) with WS₂ grown using WO₃+NaCl at 950 °C (red spectrum). The deconvolution of W 4f_{5/2} , W 4f_{7/2} and W 5p core levels and overall fit of the spectrum are reported as black dashed and a continuous line respectively. (b) The S 2p_{1/2} and 2p_{3/2} core levels for each of the two growth conditions are reported in the central panel. (c) W 4f_{5/2} , W 4f_{7/2} and W 5p core levels before (dashed line) and after transfer (continuous like) onto a new SiO₂/Si substrate are compared showing the complete disappearance of the residual WO₃ components. The spectra were fit by Doniach-Sunjic function after subtracting a Shirley background (black dashed line).

old is crossed. Once the current-bias curve reaches linear range it can be described by $I_d = \mu_n C_{ox} (W/L)((V_{gs} - V_{th})V_{ds})$, where μ_n is the electron field-effect mobility, C_{ox} is the oxide dielectric and V_{th} is the threshold voltage. As seen in Figure 12 a typical response curve for both the as grown as well as transferred samples of WS₂ grown with H₂WO₄ show an asymmetry around $V_{ds} = 0$ V due to difference in the source and drain potentials. The Schottky barrier at the source is pinned by the gate, while the barrier at the drain diminishes proportionally to the drain bias and vice versa. The contacts however become more "Ohmic", when the bands bend more at the semiconductor/metal interface while the gate bias increase, leading to more significant tunnelling current.

3.3 Conclusions

In conclusion a synthesis route has been developed that allows for high quality monolayer and bilayer WS₂ CVD growth. This quality is demonstrated by highest recorded electron FET mobility for both monolayer and bilayer CVD grown WS₂ compared to the literature and comparable to that of mechanically exfoliated ones. Additionally the triangle flakes grown using H₂WO₄ + NaCl are much larger (up to ~200-300 μ m) compared to ones grown using standard WO₃. The CVD growth using H₂WO₄ has also been demonstrated to work at lower temperatures, down to 750 °C compared to more commonly used 950 °C. Finally the PL of the WS₂ grown using H₂WO₄ is uniform throughout the flake and has very FWHM (36 meV) which again points to high crystal quality and lack of defects in form of sulfur vacancies. These findings allow to develop further the TMDCs synthesis method of CVD and can potentially lead to industrially scalable synthesis of

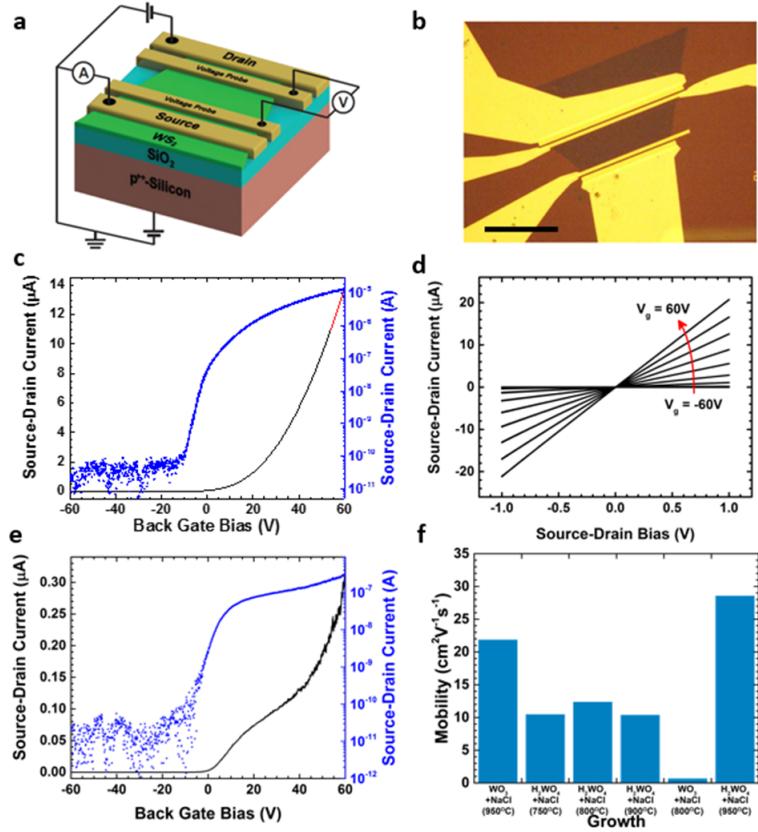


Figure 12: Electrical characteristics of monolayer WS₂: (a) Schematic of the bottom-gated field effect transistors; (b) optical micrograph of the device (scale bar is 20 μm); (c) FET transfer curve for the monolayer WS₂ grown using H₂WO₄+NaCl at 950 °C showing the highest mobility of 28 cm^2/Vs (linear region of the transport graph marked with a red-dashed line); (d) Response curves at different gate biases for a WS₂ triangle grown using H₂WO₄+NaCl; (e) FET transfer curve for the monolayer WS₂ grown using WO₃+NaCl at 800 °C; (f) electron mobilities of monolayer WS₂ grown using different conditions.

monolayer WS₂ or other TMDCs over large area.

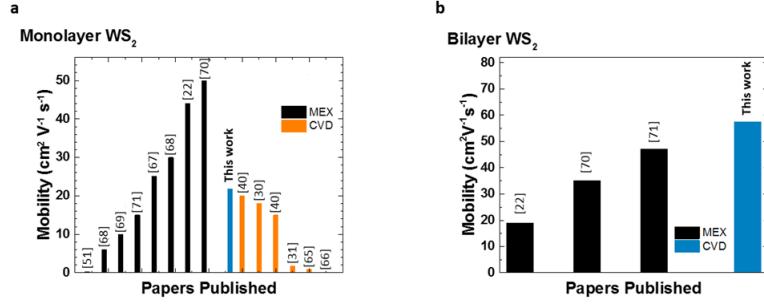


Figure 13: Comparison of our results with the literature of CVD grown material and mechanically exfoliated WS_2 (MEX) : electron mobility for (a) monolayer WS_2 and (b) bilayer WS_2 . The histograms show our record values for both monolayer and bilayer amongst the best values reported for CVD grown WS_2 .

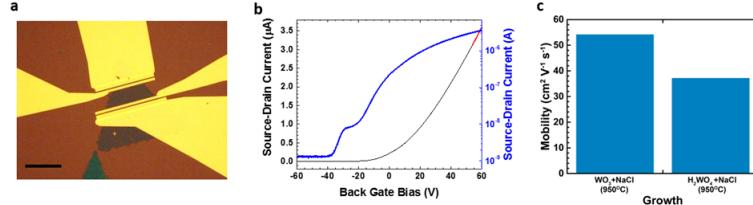


Figure 14: Electrical characteristics of bilayer WS_2 : (a) Optical micrograph of the device (scale bar is $30\mu m$); (b) FET transfer curve for the bilayer WS_2 grown using WO_3+NaCl at $950\text{ }^\circ C$ showing the highest mobility of $52\text{ }cm^2/Vs$ (linear region of the transport graph marked with a red-dashed line); (c) electron mobility of bilayer WS_2 grown by using different precursor systems.

4 Characterisation of WS_2

4.1 Optical microscopy of WS_2 flakes

4.2 Raman spectroscopy of WS_2

4.2.1 Interlayer interactions

4.2.2 Strain

4.2.3 Grain boundaries

4.3 Photoluminescence spectroscopy

4.3.1 PL of WS_2 monolayer

4.3.2 PL variation vs flake size

4.3.3 PL variation vs synthesis conditions

4.3.4 Spatial PL variation

4.3.5 Effects of water and oxygen on PL

5 Transfer of mono- and fewlayer WS_2

5.1 Motivation

5.2 Wet transfer

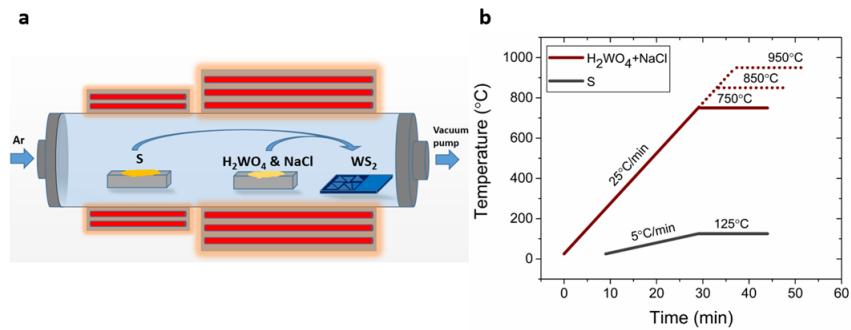


Figure 15: Illustration of: (a) CVD tubular furnace set-up; (b) temperature profile of the sulphur and the W-precursors heaters respectively. The sulphur reaches 125 °C when the metal precursors are at the maximum temperature. The SiO₂/Si wafers used as substrate for WS₂ growth are placed 1-8 cm downstream the W-precursors crucible and they are subjected to the same temperature.

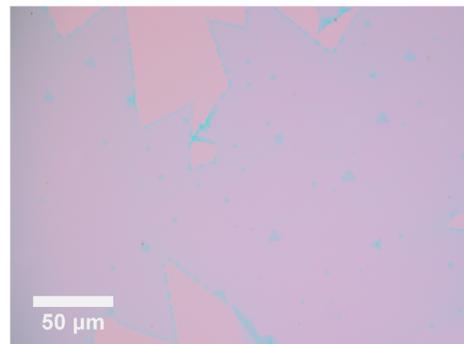


Figure 16: Optical micrograph of a continuous polycrystalline WS₂ monolayer coverage

F. M. Pesci *et al*, "MoS₂/WS₂ heterojunction for photoelectrochemical water oxidation", ACS Catalysis, 2017 - accepted

Conferences

Graphene Week, 13-17 June, 2016. (Best poster)

UK Semiconductors, 14-15 July, 2017.

MRS Boston, 26 November - 1 December, 2017.

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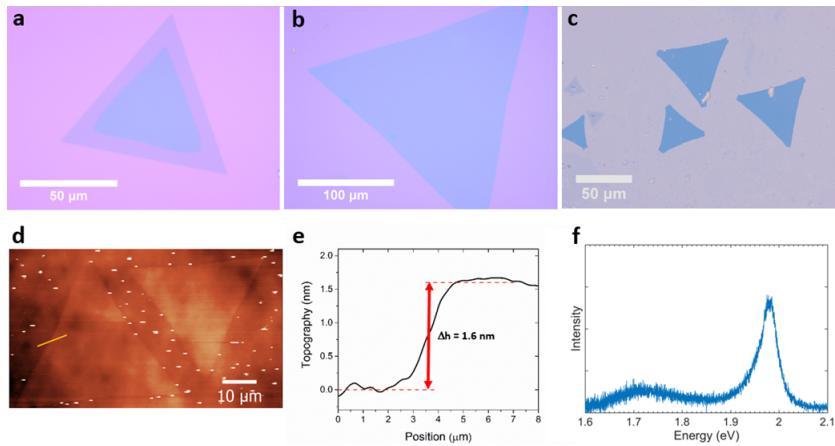


Figure 17: Optical micrographs of (a) second WS₂ layer growing on top of a WS₂ monolayer triangle, (b) uniform bilayer WS₂ triangle, (c) bilayer WS₂ flakes with high nucleation density, AFM (d) image and (e) thickness profile of a bilayer WS₂ triangle, (d) PL spectrum of a bilayer WS₂.

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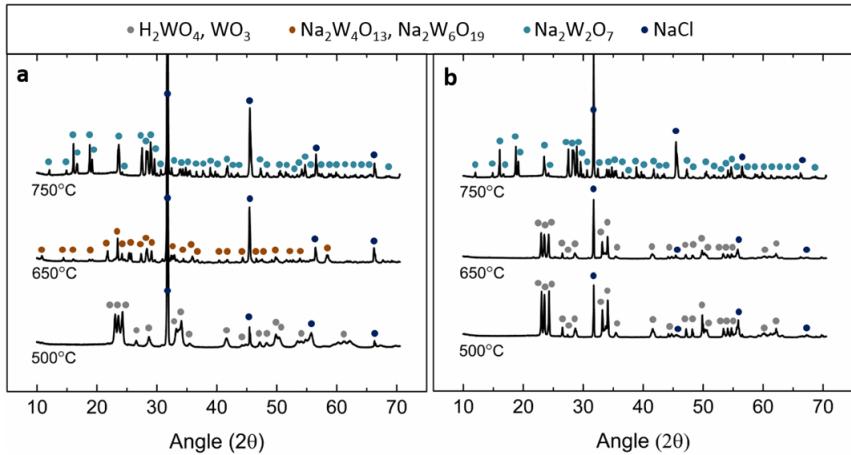


Figure 18: XRD pattern of the residual powder of W-preursors after thermal treatment at 500 °C, 650 °C and 750 °C respectively, using (a) $\text{H}_2\text{WO}_4 + \text{NaCl}$ and (b) $\text{WO}_3 + \text{NaCl}$ as precursor.

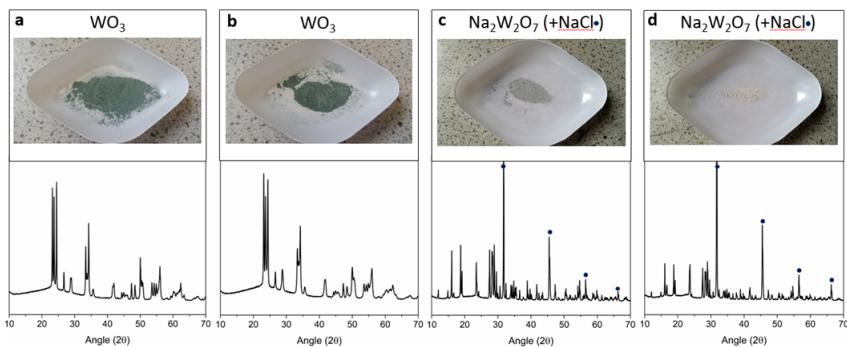


Figure 19: Residual powders of the W-preursors and their XRD patterns after thermal treatment at 750°C using (a) WO_3 , (b) H_2WO_4 , (c) $\text{WO}_3 + \text{NaCl}$, and (d) $\text{H}_2\text{WO}_4 + \text{NaCl}$ as precursor. The blue dots in the XRD patterns indicate the presence of residual NaCl.

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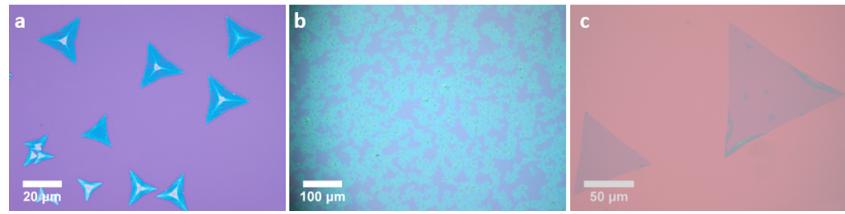


Figure 20: Optical micrographs of WS₂ triangles and amorphous domains grown on SiO₂/Si substrates by using (a) WOCl₄ at 600°C, (b) only H₂WO₄ at 850°C, (c) H₂WO₄+KCl at 850°C.

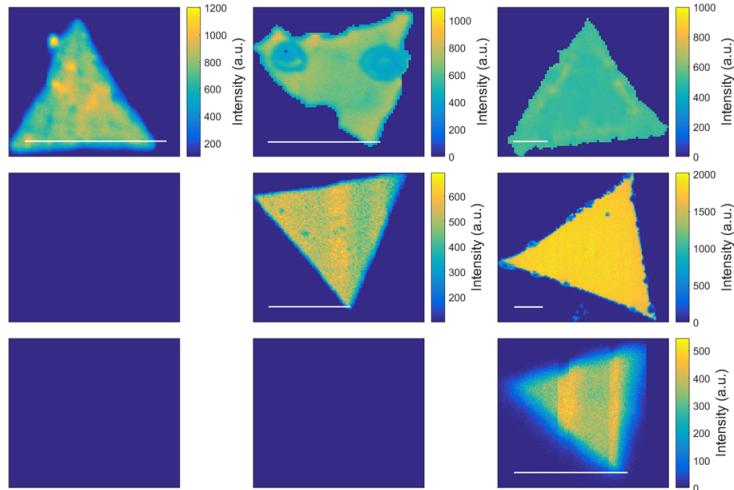


Figure 21: Raman spectroscopy: $2LA + E_{2g}^1$ peak intensity. Scale bar 10 μ m

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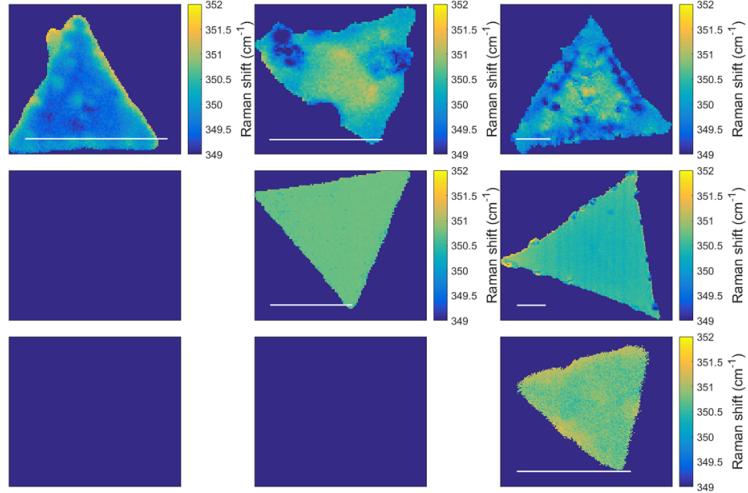


Figure 22: Raman spectroscopy: $2LA + E_{2g}^1$ peak position. Scale bar is $10 \mu\text{m}$

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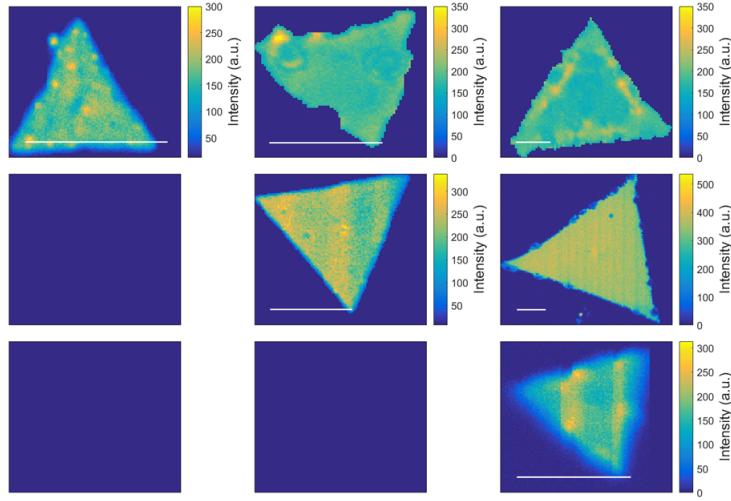


Figure 23: Raman spectroscopy: A_{1g} peak intensity. Scale bar is $10 \mu\text{m}$

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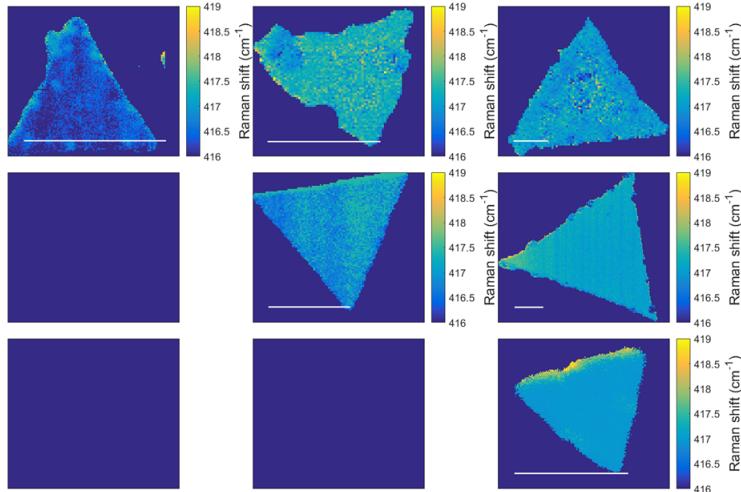


Figure 24: Raman spectroscopy: A_{1g} peak position. Scale bar is $10 \mu\text{m}$

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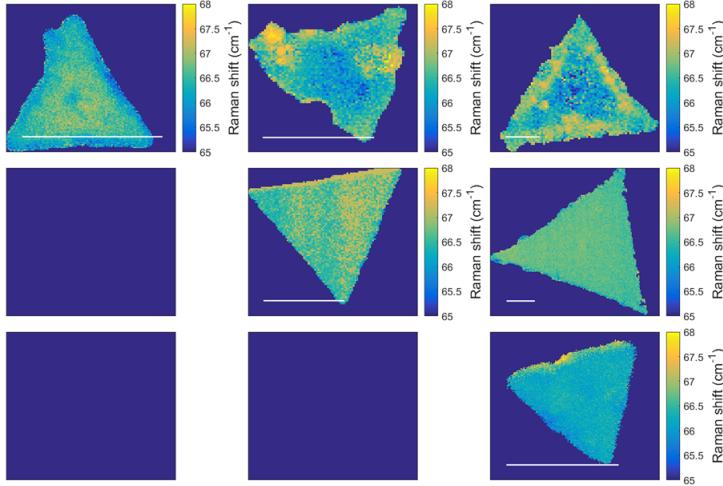


Figure 25: Raman spectroscopy: $2LA - A_{1g}$ energy differences. Scale bar is $10 \mu\text{m}$

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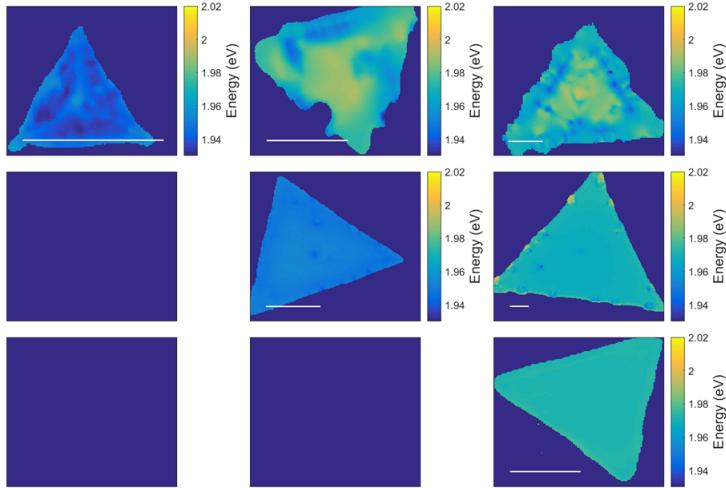


Figure 26: PL spectroscopy: PL peak position. Scale bar is $10 \mu\text{m}$

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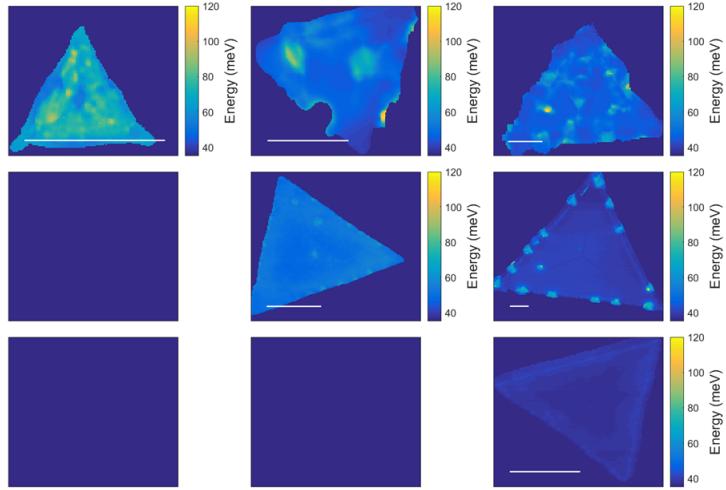


Figure 27: PL spectroscopy: PL FWHM. Scale bar is $10 \mu\text{m}$

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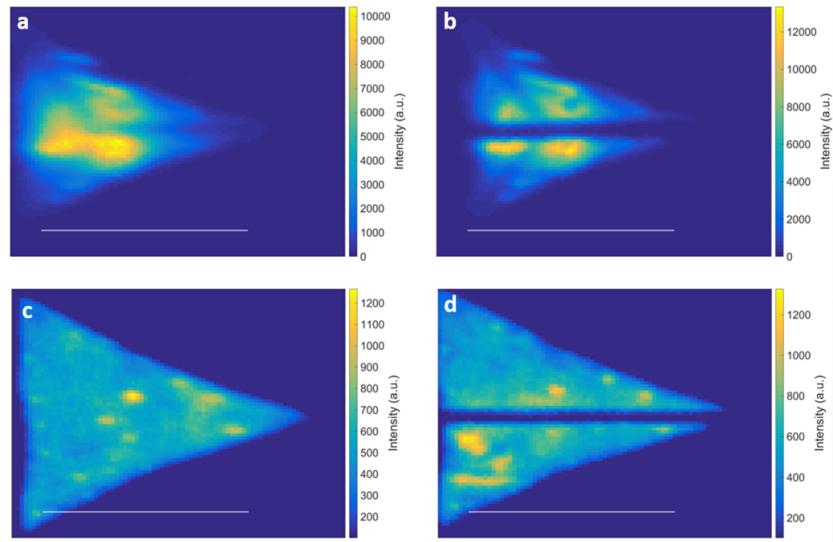


Figure 28: PL intensity (a) before and (b) after the cutting. $2LA + E_{2g}^1$ Raman intensity (c) before and (d) after the cutting. Scale bar is 10 μm

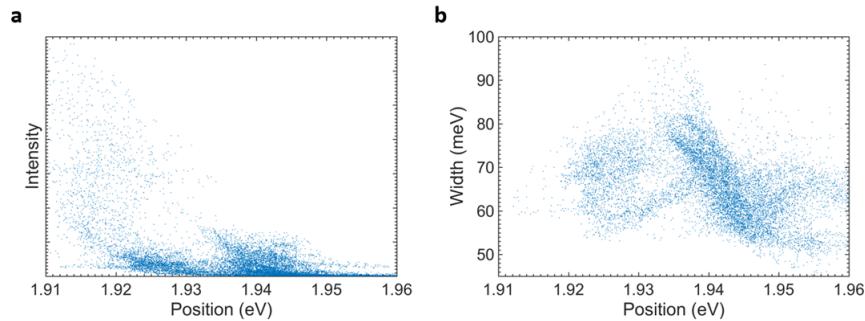


Figure 29: (a) PL Intensity vs position, (b) PL FWHM vs position of WS_2 grown using WO_3 at 950 $^\circ\text{C}$

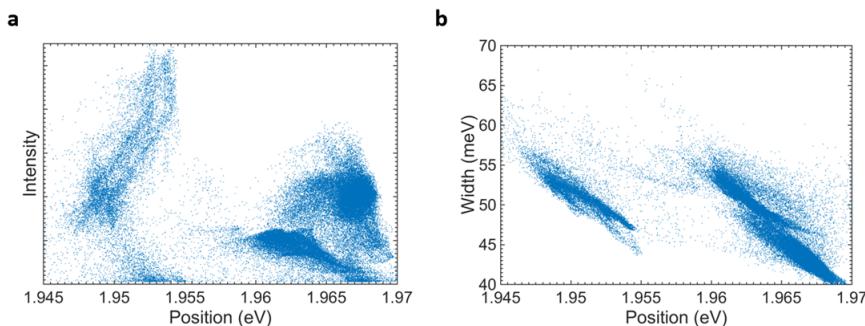


Figure 30: PL Intensity vs position, (b) PL FWHM vs position of WS_2 grown using $WO_3 + NaCl$ at 850 $^\circ\text{C}$

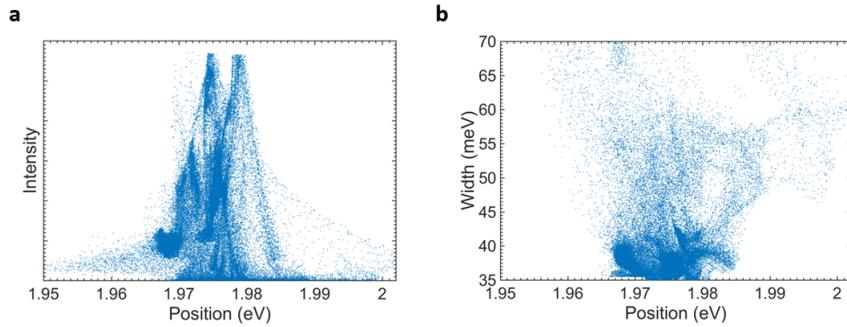


Figure 31: (a) PL Intensity vs position, (b) PL FWHM vs position of WS_2 grown using $H_2WO_4 + NaCl$ at 850 °C

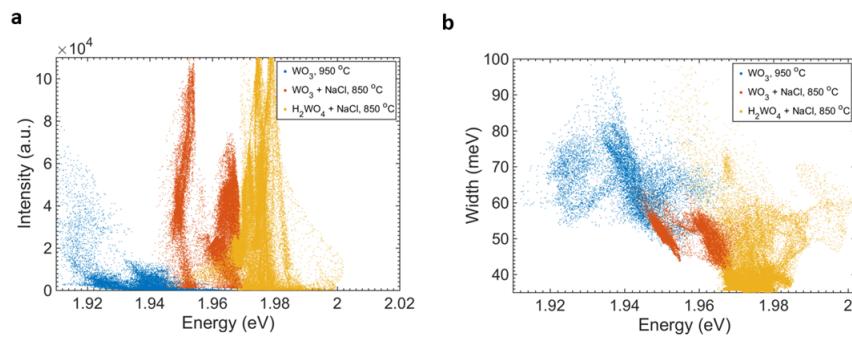


Figure 32: (a) PL Intensity vs position, (b) PL FWHM vs position of WS_2 grown using WO_3 , $WO_3 + NaCl$, $H_2WO_4 + NaCl$