

Exploring the Chemistry of Re^I: Physical and Theoretical Investigations

by

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Abstract

The development of Rhenium I photocatalysts has been pursued since Lehn first showed the excellent performance of the Re^I bipyridine tricarbonyl catalyst. Since then, development has modified the organic ligand to demonstrate continued or improved activity with other α -diimine bidentate geometries. Geometry has been limited to κ^2 motifs, with *fac*-(CO)₃ and axial halides. This work will demonstrate the synthesis, characterization, and testing of a new $\kappa^3(L_3)-Re^1(CO)_2X$ (X = Cl, Br, CN, OTf) family of compounds for CO₂ reduction, as well as computational investigations into the mechanism of the reduction of CO₂ to CO and other species.

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Contents

1	Introduction	1
1.1	Photochemistry & Catalysis	1
1.2	Rhenium	3
1.3	CO ₂ Reduction Chemistry	5
2	New Coordination Geometries for Re^I	7
2.1	Introduction	7
2.2	Synthesis of Bidentate and Terdentate Re ^I Complexes	8
2.3	Characterization	11
2.3.1	NMR Analysis	11
2.3.2	Structure Analysis with X-Ray Crystallography and DFT	15
2.3.3	Infrared Spectroscopy	27
2.3.4	Photophysical Properties	30
2.3.5	Fluorescence	34
2.4	Conclusions	35
3	Photocatalysis of CO₂	36
3.1	Introduction	36
3.2	Photocatalytic Reactions with New Compounds	37
3.2.1	Conditions	37
3.2.2	Experimental Results	38
3.2.3	Rationalization of Results	38
3.3	Conclusions	41
4	Mechanism of CO₂ Reduction	42
4.1	Introduction	42
4.2	Mechanism Pathways	43

4.2.1	Eximer Formation and Decomposition of the Sacrificial Amine	45
4.2.2	The ‘Carbonate’ Pathway	49
4.2.3	The ‘Formate’ Pathway	54
4.2.4	The ‘Water-Gas Shift’ Pathway	57
4.2.5	Consequences From κ^2 Terpyridine Complex Inactivity	63
4.3	Comparison Between Mechanistic Pathways	64
4.4	Conclusions	66
5	TurboControl	67
5.1	Development	68
5.2	Usage	69
6	Conclusions	70
A	Experimental Procedures	71
A.1	General Methods	71
A.2	Computational Methods	72
A.3	X-ray Crystallography	73
A.4	(terpy- κ^2 -N,N')Re(CO) ₃ Cl (2.1)	74
A.5	(terpy- κ^3 -N,N',N'')Re(CO) ₂ Cl (2.2)	75
A.6	(terpy- κ^2 -N,N')Re(CO) ₃ Br (2.3)	75
A.7	(terpy- κ^3 -N,N',N'')Re(CO) ₂ Br (2.4)	76
A.8	(terpy- κ^2 -N,N')Re(CO) ₃ CN (3)	76
A.9	(terpy- κ^3 -N,N',N'')Re(CO) ₂ CN (2.6)	77
A.10	(terpy- κ^2 -N,N')Re(CO) ₃ OTf (2.7)	77
A.11	(terpy- κ^3 -N,N',N'')Re(CO) ₂ OTf (2.8)	78
B	Molecular Orbitals Diagrams	79
C	Molecular Orbitals Diagrams	85
D	Reaction Potential Energy Diagrams	94
E	TurboControl and TurboGo Manual	104
E.1	Introduction	104
E.2	System Requirements	104
E.3	TurboGo	105

E.4	TurboControl	106
E.5	Input File Format	107
E.5.1	Keywords	108
E.5.2	Route Card Options	109
E.5.3	Title	110
E.5.4	Charge and Spin	110
E.5.5	Geometry	110
E.5.6	Additional control File Modifications	110
E.5.7	Example Input File	111
E.6	Code Details	111
E.7	Citing TurboControl	113
E.8	License	113
Glossary of Terms		115

List of Tables

2.1	Crystal data and structure refinement for compounds 2.1 , 2.3 , and 2.5	15
2.2	Solvated and Gas Phase Energies of Axial & Trans κ^x -(terpy)-Re(CO) _{5-x} CN (x=2,3)	18
2.3	Selected Distances, Angles, and Torsions for 2.1	20
2.4	Selected Distances, Angles, and Torsions for 2.3	21
2.5	Selected Distances, Angles, and Torsions for 2.5	22
2.6	Selected Distances, Angles, and Torsions for 2.2	25
2.7	Selected Distances, Angles and Torsions for Acetonitrile Adduct of 2.8	26
2.8	Crystal data and structure refinement for compounds 2.2 and 2.8	27
4.1	Gas phase and solvated energies for mechanism reactants and products	48
4.2	Energies for the reaction steps in the excimer formation pathway	48
4.3	Gas phase and solvated energies for the ‘carbonate’ mechanism	51
4.4	Energies for the reaction steps in the ‘carbonate’ pathway	51
4.5	Gas phase and solvated energies for the ‘formate’ mechanism	55
4.6	Energies for the reaction steps in the ‘formate’ pathway	55
4.7	Gas phase and solvated energies for the ‘water-gas shift’ mechanism	59
4.8	Energies for the reaction steps in the ‘water-gas shift’ mechanism	59
4.9	Energies for the reaction steps in the ‘equatorial’ geometry	61

List of Figures

1.1	Two common bidentate complexes using terdentate ligands	4
2.1	Results of TGA analysis on 2.1 and 2.3	9
2.2	The aromatic region of the ^1H NMR spectra of the four bidentate compounds	12
2.3	The aromatic region of the ^1H NMR spectra showing bidentate - terdentate conversion	13
2.4	Proton-explicit skeletal drawing of 2,2':6',2"-terpyridine	13
2.5	The aromatic region of the ^1H NMR spectra of the four terdentate compounds	14
2.6	The ^{13}C NMR spectra of 2.1	14
2.7	X-ray crystal structure representation for 2.1 , 2.3 and 2.5	17
2.8	X-ray crystal structure representation for 2 and 8	23
2.9	FTIR Spectra for complexes 2.1 and 2.2	28
2.10	DFT predicted FTIR spectra for 2.1 and 2.2	29
2.11	FTIR Spectra for complexes 2.7 and 2.8	30
2.12	UV-Vis spectra for all compounds	31
2.13	Plots of the experimental and computed UV-Vis spectra for compound 2.1	32
2.14	Plots of the experimental and computed UV-Vis spectra for compound 2.2	33
2.15	UV-Vis and fluorescence spectra for 2.1 and 2.2	34
3.1	UV-Vis spectra for 2.1 and catalyst-TEOA complex	40
3.2	Structure and absorption spectra of proposed $[\kappa^3\text{-(terpy)}-\text{Re}(\text{CO})_3]^+$. .	41
4.1	DFT calculated structures for the ‘carbonate’ mechanistic pathway . . .	52
4.2	DFT calculated structures for the ‘formate’ mechanistic pathway	56
4.3	DFT calculated structures for the axial ‘water-gas shift’ mechanistic pathway	60
4.4	DFT calculated structures for the equatorial ‘water-gas shift’ mechanistic pathway	62

List of Schemes

2.1	Synthesis of 2.1 and 2.3	8
2.2	Synthesis of 2.2 and 2.4	10
2.3	Anion exchange pathways	10
3.1	Reorganization from catalytic eximer to form 3.1	40
4.1	Overview of mechanistic pathways	44
4.2	Formation of the eximer species via absorption of a photon and oxidation of the sacrificial amine.	46
4.3	Decomposition pathway for the sacrificial amine.	47
4.4	The ‘carbonate’ mechanistic pathway	50
4.5	The ‘formate’ mechanistic pathway	54
4.6	The ‘water-gas shift’ mechanistic pathway	58
4.7	Rearrangement of carbonyl and open site	61

Chapter 1

Introduction

Common distinctions split most chemical compounds into one of two categories: organic and inorganic. Organic molecules contain carbon and hydrogen, with or without additional nitrogen, oxygen, phosphorus, sulfur, and the halides. Inorganic chemistry is, therefore, considered to be the remainder of the molecules possible. While they may include some aspect of organic chemistry (especially in organometallic molecules), the main structural motif or reactive center is a non-organic feature. These inorganic compounds can range from compounds such as lithium or Grignard reagents with significant organic influence, to metallic alloys or mineral compounds. With such a wide range of possibilities, inorganic chemistry has many facets. A widely active research area is the development and testing of transition metal complexes for catalytic, photo-physical, biochemical or manufacturing uses.

1.1 Photochemistry & Catalysis

A report of the first synthesized organometallic complex was published by Zeise in 1831.¹ To form what is now known as Zeise's salt, $\text{K}[\text{PtCl}_3(\text{C}_2\text{H}_4)] \cdot \text{H}_2\text{O}$, he mixed platinum

chloride with ethanol, followed by a reaction with potassium chloride.² After some controversy to the composition of this, it was confirmed by Griess and Martius,³ and later expanded upon by Birnbaum.⁴

The field of organometallics was expanded greatly by Frankland,² and many of his complexes were catalytically active. Further development of this new type of chemistry quickly led to useful catalysts for the conversion of petroleum products and other chemistries, using nearly all of the transition metals. These catalysts take all forms, from simple olefin and halide compounds to multi-metallic complexes with large organic ligands.

Some of the most significant organometallic catalysts since the late 1990s have been the development of earth metal pincer complexes to replace noble metal or early transition metal catalysts, which are often more toxic or expensive to produce. Brookhart and Gibson published a series of papers^{5–8} on the use of iron and cobalt with bis(imino)pyridine ligands to perform ethylene polymerization at rates exceeding those of similar noble metal complexes and metallocenes.⁹ The role of the ligand in the mechanism is still up for debate, but many modified systems have been synthesized and tested since the first work was published.¹⁰

Many of these types of pincer complexes are photochemically active. In transition metal complexes, the interaction between the metal atom(s) and the ligands can cause significant electron mobility upon the absorption of incident photons. The metal atom's *d* orbitals typically lie at or near the Highest Occupied Molecular Orbital (HOMO) energy, while the ligands often have low energy anti-bonding orbitals (π^*) at the Lowest Unoccupied Molecular Orbital (LUMO) levels. When a photon is absorbed and is promoted from the ground state to the excited state, that state is spatially removed from the metal centre, this motion of the electron is labelled a Metal-Ligand Charge Transfer

(MLCT). Formally, the metal atom is oxidized by the photons, this oxidation allows for redox reactivity at the metal centre for as long as the electron remains removed to the ligand. Relaxation (through photon emission via fluorescence or phosphorescence, or via vibrational or other motion processes) can return the electron to the metal to reform the ground electronic state.

1.2 Rhenium

Rhenium compounds in particular are known to have a broad range of applications. Their use in catalysis has been explored, covering olefin hydrogenation,¹¹ epoxidation (with chiral selectivity),¹² and in a host of organic bond formation reactions¹³ (Friedel-Crafts acylation and alkylation,^{14–16} nucleophilic addition,^{17,18} carbon-carbon coupling,^{19,20} as well as other heteroatom-carbon bond formation^{21–25}). Rhenium has also been used in radiopharmaceutical applications, due to the availability of moderately radioactive isotopes.^{26,27} These compounds exhibit other interesting fundamental photophysical properties as well.²⁸

Since the mid-1970's, complexes containing the α -diimine Re^I tricarbonyl core have attracted a great deal of attention due to their attractive photochemical properties. The pseudo-octahedral *fac*-[L₂Re(CO)₃X] and *fac*-[L₂(L')Re(CO)₃]⁺ complexes have been the dominant species.^{29–39} A large family of compounds with these formulations have been accessed by the addition of chelating diimine σ -donor ligands to [Re(CO)₅X] with the quantitative replacement of two *cis* carbonyls in the Re^I starting material.^{29,40–47} Significantly, these reactions form only bidentate coordinated ligands with *facial* tricarbonyl isomers as products even when a potentially tridentate σ -donor, such as bis(imino)pyridine or 2,2':6',2"-terpyridine, are employed in the reaction (see Figure 1.1).^{48–51}

These robust species have been examined for potential applications in organic light-

emitting diodes (OLEDs),^{52,53} chemosensors and biotechnology probes,^{54–58} fluorescence microscopy imaging of cells,^{54,59,60} as chemotherapy agents,⁶¹ and the photochemical reduction of CO₂ to CO.^{62–66} Among the key photophysical features of these α -diimine Re^I compounds is the electron transfer capability of this system and the interplay between the Re center and the well-known non-innocent redox activity of the ligands.^{67,68}

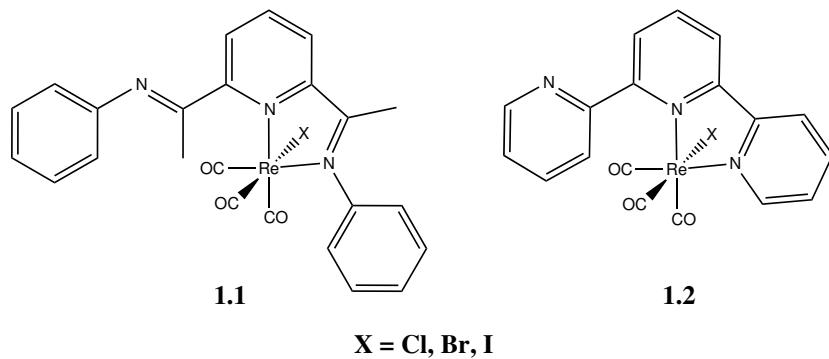


Figure 1.1 Two common *fac*-[L₂Re(CO)₃X] complexes with terdentate σ -donor ligands: L = bis(imino)pyridine (**1.1**) and 2,2':6',2''-terpyridine (**1.2**)

Further development of this chemistry has been restricted by the limited structural and electronic variability of the common pseudo-octahedral *fac*-[L₂ReX(CO)₃] (L₂ = α -diimine) products. While these systems continue to receive considerable attention, studies detailing the coordination chemistry of the meridionally-coordinated tridentate triimine Re^I dicarbonyl core are quite limited.⁶⁹ For example, while κ^3 (terpy)Re(CO)₂Cl was initially reported in 1988,⁷⁰ closer analysis of the reported analytical data (including ¹H NMR) indicate that this compound is more likely κ^2 LRe(CO)₃Cl. A more recent report for this compound provides spectroscopic details of this species as well as the preliminary report for the generation of [κ^3 (terpy)Re(CO)₂L']⁺ cations (L = PPh₃, PEt₃, NC₅H₅, and NCCH₃).⁷¹ Finally, the ¹H NMR data for κ^3 (terpy)Re(CO)₂Br has been reported⁵⁰ but accompanied no other characterization.

In order to fully exploit the potential of this versatile family of compounds, the

limits imposed by the bidentate coordination need to be addressed. Furthermore, it would appear that, on the basis of the tridentate ligands that have been investigated, the concerted effort to produce the tridentate species has been essentially unsuccessful, or requiring harsh conditions.⁵¹ Attracted by this challenge we sought to synthesize, crystallographically authenticate, and investigate the photophysical properties of low-valent rhenium pincer complexes displaying an N,N',N"-chelated terpyridine array.

Complexes of 2,2':6',2"-terpyridine (terpy) are of interest due to the conceptual relationship to established bis(imino)pyridine compounds.^{72,73} This thesis will be a discussion of the development of chemistry of Re^I complexes, their characterization, and comparison of structural and photo-physical properties to computed values. Further exploration of the CO₂ reduction by photo-catalysis of these new complexes will be analyzed. This thesis will also take a more detailed look at specifics of the mechanisms proposed for current Re^I diimine catalysts, and propose new geometries for prior mechanistic steps based on experimental, computational, and literature review work.

1.3 CO₂ Reduction Chemistry

Recent years have seen an increase in the concentration of CO₂ in the atmosphere as a product of combustion of oil, gas and coal in the industry and transportation sectors.⁷⁴ This is of significant concern due to the greenhouse gas properties of CO₂.^{75,76} Attempts to reduce emissions proceed by various pathways, including the utilization of renewable resources such as wind and solar for energy production;⁷⁷ development of more fuel and energy efficient processes; and attempts to capture CO₂ from industrial exhaust streams.^{78–80} It can then be stored in underground depositories, or used as a feedstock in the production of simple molecules or fuels.^{81–83} Due to it being the final product in combustion and its high thermodynamic stability, the reduction of CO₂ is

an energy-intensive task.^{84,85} While plant life naturally performs CO₂ transformation via photosynthesis, no artificial means have proven to be robust and scalable enough for the task in large scale.⁸⁶

The development of catalysts for ‘artificial photosynthesis’ has explored various means, including utilizing metal electrodes,⁸⁷ electrocatalytic semiconductors (such as TiO₂, ZnS, CdS, or As or S doped Ag)^{88,89}, and organometallic species of Co, Ni, and Fe (particularly porphyrin complexes).^{90–98} These typically require the addition of electrons, and current trends look to solar powered electricity production for a ‘green’ CO₂ reduction platform⁹⁹.

Alternate solutions involves the use of photoredox catalysts, including α -imino complexes of Re^I, Ru^{II}, Os^{II}, and Ir^{III}.^{62,100–110} In recent years, the popularity of these α -diimino compounds has increased, owing to their unique MLCT excited states and easily tunable photophysical properties. Of these complexes, Re^I based catalysts are of particular interest. Whether acting as a mono-metallic photocatalyst, or with Ru in multi-metallic complexes, the high quantum yield (up to $\Phi=0.59$)¹¹¹ and exclusive selectivity to production of CO is unparalleled in these catalysts.¹¹² Research focus has turned to tuning photophysical properties for increased use of photons in the visible spectrum (particularly lower energy) and modification of the ligand framework for increased turnover numbers and frequencies.^{85,112,113} This thesis will focus on the synthesis and testing of a novel Re^I catalyst attempting to improve on those points.

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Chapter 2

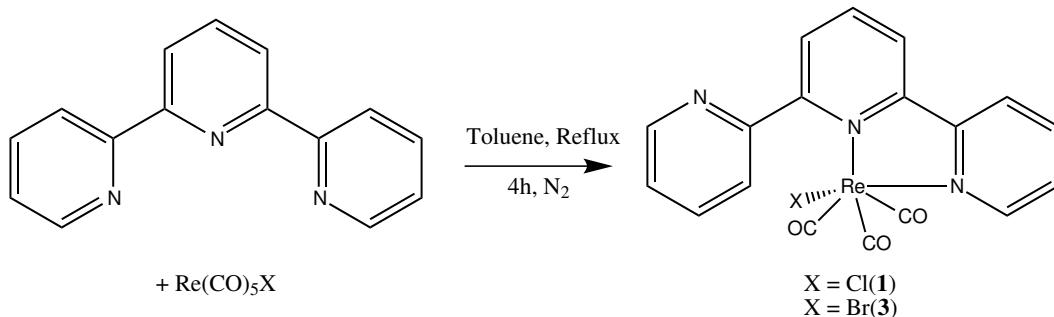
New Coordination Geometries for Re^I

2.1 Introduction

As mentioned in the introduction, Re^I compounds have been typically bidentate (κ^2) compounds, even when using a potentially terdentate (κ^3) ligand such as bis(imino)pyridine or terpyridine (refer to Figure 1.1). The chemistry of this rhenium α -imino complex has been extensively investigated, with over 1700 references appearing in a structure search for that metal-ligand motif. The ejection of an additional carbonyl and the chelation of the pendant arm of the ligand was attempted to extend the conjugated π system of the ligand and its interaction with the metal centre. This was first demonstrated by prior work in our group for the bis(imino)pyridine ligand.⁶⁹

2.2 Synthesis of Bidentate and Terdentate Re^I Complexes

Similar to the prior work, synthesis began with the production of the bidentate complex $\kappa^2(\text{terpy})\text{Re}(\text{CO})_3\text{X}$ ($\text{X} = \text{Cl}, \text{Br}$) by coordination of 2,2':6',2"-terpyridine with a $\text{Re}(\text{CO})_5\text{X}$ starting material in dry toluene at reflux for 4 hours, as shown in Scheme 2.1. A bright yellow powder precipitated from solution and was collected by filtration, washed with cold hexanes, and dried *in vacuo* to a good yield of **2.1** and **2.3** respectively.ⁱ These bidentate compounds were characterized fully and used without further purification to produce $\kappa^3(\text{terpy})\text{Re}(\text{CO})_2\text{X}$ ($\text{X} = \text{Cl}, \text{Br}$) via thermolysis, as well as for anion exchange reactions.



Scheme 2.1 Synthesis of **2.1** and **2.3** from $\text{Re}(\text{CO})_5\text{X}$ and 2,2':6',2"-terpyridine

Conversion of compounds **2.1** and **2.3** to the κ^3 moiety required the release of CO and the subsequent coordination of the free pendant arm. Prior work had identified the thermal lability of the carbonyl, based on a method first described by Buckingham with a osmium complexes.¹¹⁴ In this method, a ceramic sample boat was placed in a tube furnace at elevated temperature, under a flowing atmosphere of N_2 . After some time, the sample is removed and collected at nearly quantitative yield. Determination of

ⁱExperimental details for all compounds can be seen in Appendix A Experimental Procedures

the appropriate thermolysis temperature was performed by Thermogravimetric Analysis (TGA) of the sample. A loss of 6-8 % of starting mass from the sample is consistent with the loss of one carbonyl group from the complex. Results of TGA on **2.1** and **2.3** is shown in Figure 2.1.

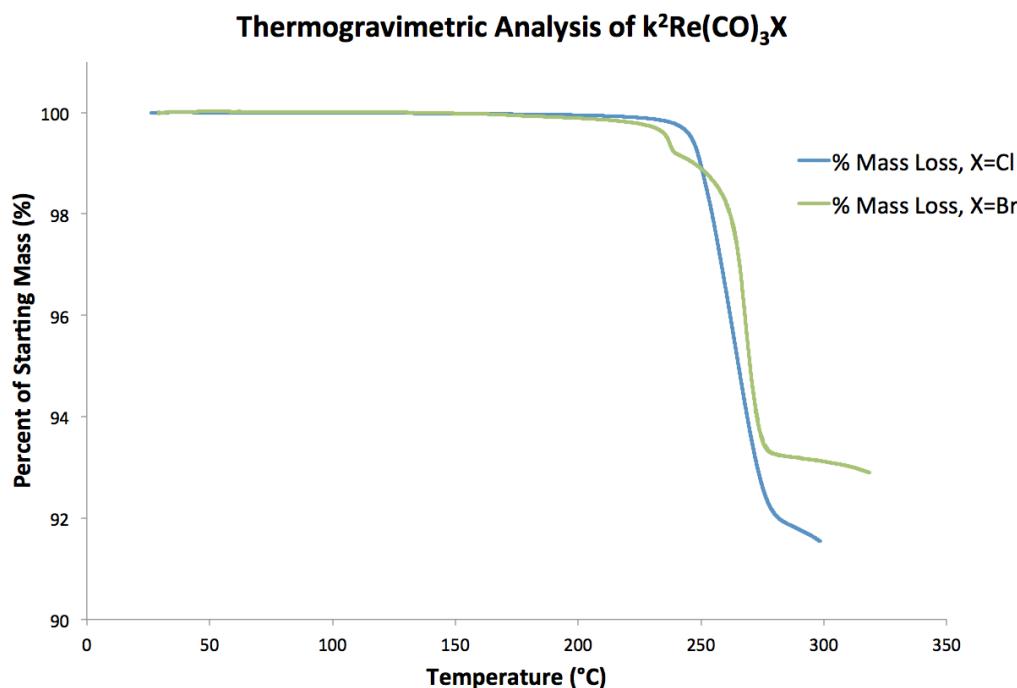
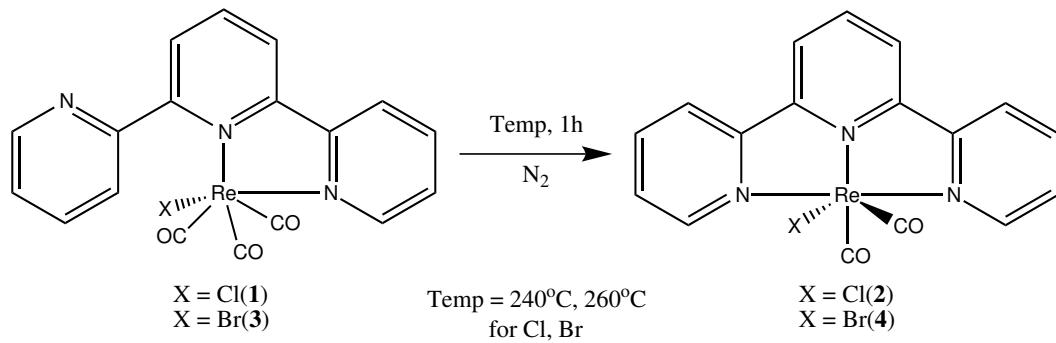


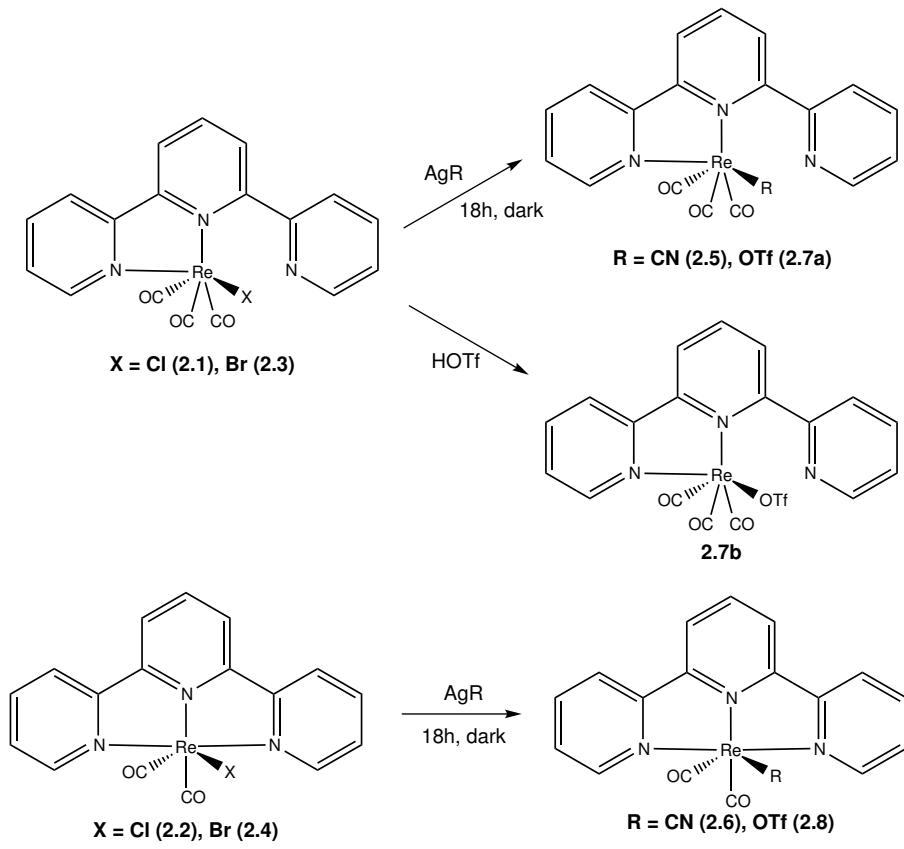
Figure 2.1 Results of TGA analysis on **2.1** and **2.3**

The onset of mass loss in the TGA of **2.1**, and the onset of the main mass loss in the TGA of **2.3** were chosen to identify a thermolysis temperature for each sample. For **2.1**, thermolysis was performed at 240 °C, and for **2.3** thermolysis was performed at 260 °C, yielding **2.2** and **2.4** respectively, at quantitative yields, by the pathway in Scheme 2.2.

Further reactions were carried out on the above products to yield triflate and cyano complexes in bidentate and terdentate geometries. These anion exchange reactions were performed by the addition of the silver salt to the chloride products **2.1** or **2.2**, to precipitate AgCl, leaving $\kappa^2(\text{terpy})\text{Re}(\text{CO})_3\text{CN}$ (**2.5**), $\kappa^3(\text{terpy})\text{Re}(\text{CO})_2\text{CN}$ (**2.6**),



Scheme 2.2 Synthesis of **2.2** and **2.4** by thermolysis of **2.1** or **2.3**, respectively



Scheme 2.3 Anion exchange pathways to synthesize **2.5** - **2.8**

$\kappa^2(\text{terpy})\text{Re}(\text{CO})_3\text{OTf}$ (**2.7a**) and $\kappa^3(\text{terpy})\text{Re}(\text{CO})_2\text{OTf}$ (**2.8a**), as shown in Scheme 2.3.

Reaction with the bromide products proceeds with similar results. As the Ag anion ex-

change reactions result in only moderate yields, **2.7b** was synthesized by the direct addition of neat triflic acid ($\text{CF}_3\text{SO}_3\text{H}$) to **2.1**. HCl was released, the solutions were quenched by addition of dilute aqueous NaCO_3 , and product was collected via separation into chloroform, again at moderate yield.

2.3 Characterization

Full characterization was performed, including Nuclear Magnetic Resonance (NMR) analysis, x-ray crystallography, elemental analysis, as well as UV-Vis and IR spectroscopy. Computational Density Functional afd (DFT) methods were used to solve the geometries, and Time Dependant Density Functional Theorem (TD-DFT) was performed to predict UV-Vis spectra and identify electronic transitions.

2.3.1 NMR Analysis

Proton NMR was performed on each of the samples. Each sample was dissolved completely in deuterated acetonitrile (CD_3CN) and analysis was performed on a Bruker AVANCE 400 MHz spectrometer. Data was processed from the FID signal via the Top-Spin program, and spectra were analyzed using ACD NMR Processor v12.0.

Detailed peak analysis comparing bidentate samples **2.1**, **2.3**, **2.5**, and **2.7** (Figure 2.2) or terdentate **2.2**, **2.4**, **2.6**, and **2.8** (Figure 2.5) show little difference between samples. This is due to the distance between the anion and any protons on the ligand. While anions with different σ -donor strength marginally impact the metal-ligand interactions, these have only small effect on the location of peaks, shifting between samples by typically less than 0.1 ppm. As is shown in Figure 2.2, the characteristic shape of each spectra remains constant, only exact peak locations and some peak order varies

with anion choice.

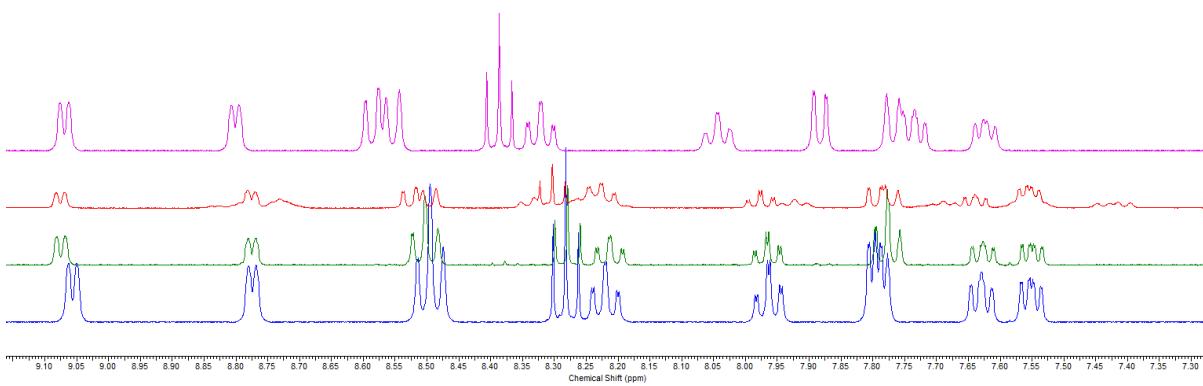


Figure 2.2 The aromatic region of the ¹H NMR spectra for compounds **2.1** (blue), **2.3** (green), **2.5** (red) and **2.7** (purple)

The characteristic feature in the NMR spectra after the transformation from bidentate to terdentate (e.g. sample **2.1** to **2.2**) is the reduction in the total number of the signals in the aromatic region (between 7 and 9 ppm). This simplification is due to the increased symmetrization of the ligand, while the κ^2 -bidentate ligand has a freely rotating pendant group. Prior work in literature⁵⁰ and in our group¹¹⁵ shows the temperature dependence of the rate of rotation of this pendant arm for various ligand species. However, the κ^3 -terdentate species has no free groups, the rigid geometry and higher order symmetry results in the simpler NMR spectrum.

The simplification of peaks due to the symmetrization of the ligand results in the peaks from free pendant arm protons *a*, *b*, *c*, and *d* (see Figure 2.4) with peaks at 9.06, 7.63, 8.22 and 7.79 ppm, respectively, aligning with their mirroring peaks at *a'*, *b'*, *c'* and *d'* (with peaks at 8.77, 7.55, 7.96, and 7.80 ppm). The new symmetrized peaks show integrations of two protons per peak, relative to the single proton (*f*) peak from the central pyridyl ring, *para* to the nitrogen. As well, the presence of metal-ligand interaction in the pendant arms reduces the deshielding effect, shifting the free pyridyl

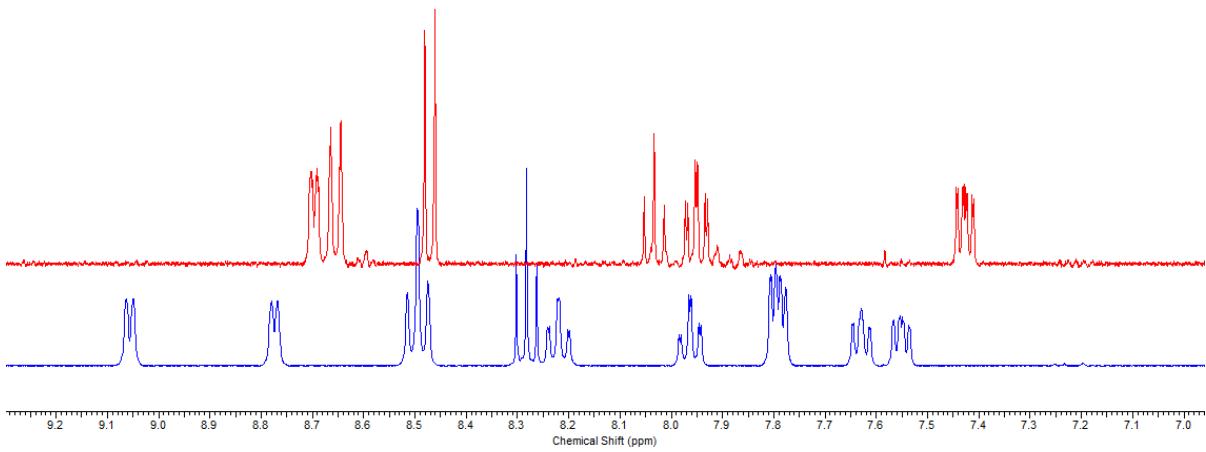


Figure 2.3 The aromatic region of the ${}^1\text{H}$ NMR spectra for compounds **2.1** (blue) and **2.2** (red), showing the simplification of the spectra upon the conversion from bidentate to terdentate

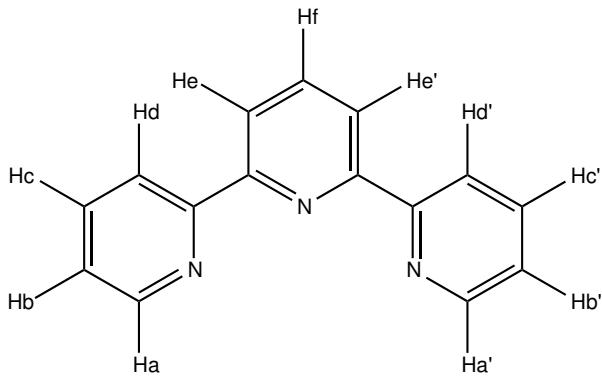


Figure 2.4 Proton-explicit skeletal drawing of 2,2':6',2''-terpyridine

ortho proton from 9.06 ppm to 8.77 in the chelated group, with similar shifts evident for the other pendant protons.

As in the case of bidentate compounds, modification of the anion has only minor effects on the NMR spectra. This is the expected behaviour, the conjugated pi system and the lack of protons near the modified anion combine to be susceptible in only a minor fashion.

Carbon NMR (${}^{13}\text{C}$) was attempted on the complexes as well. Unfortunately, Re^{I}

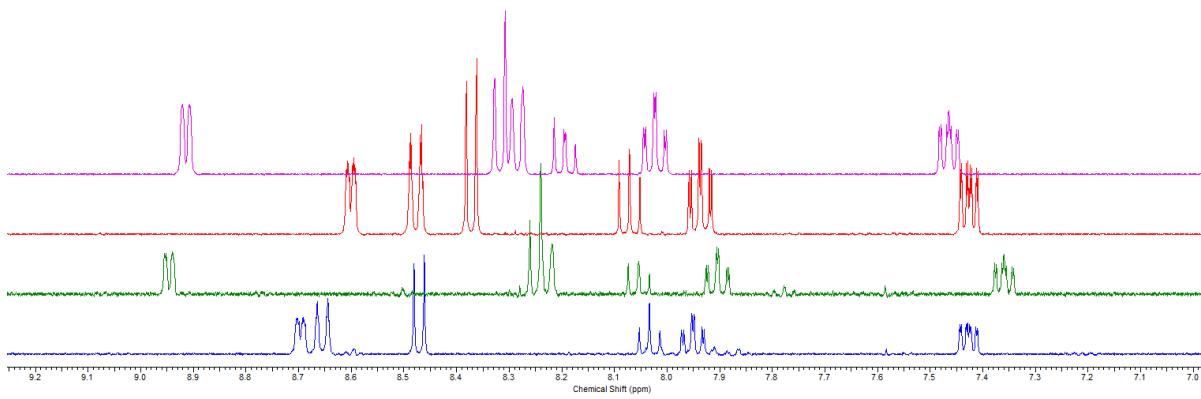


Figure 2.5 The aromatic region of the ¹H NMR spectra for compounds **2.2** (blue), **2.4** (green), **2.6** (red) and **2.8** (purple)

complexes perform poorly in ¹³C NMR experiments, the signal to noise ratio is incredibly poor (if a signal is even visible). The effect of this is a lack of ¹³C NMR analysis of these compounds in literature, with a very few exceptions. Extensive efforts included use of a 500 MHz spectrometer, with 1664 scans to produce the best example, seen in Figure 2.6, with average peak signal to noise ratio (s/n) of only 4 - 5.

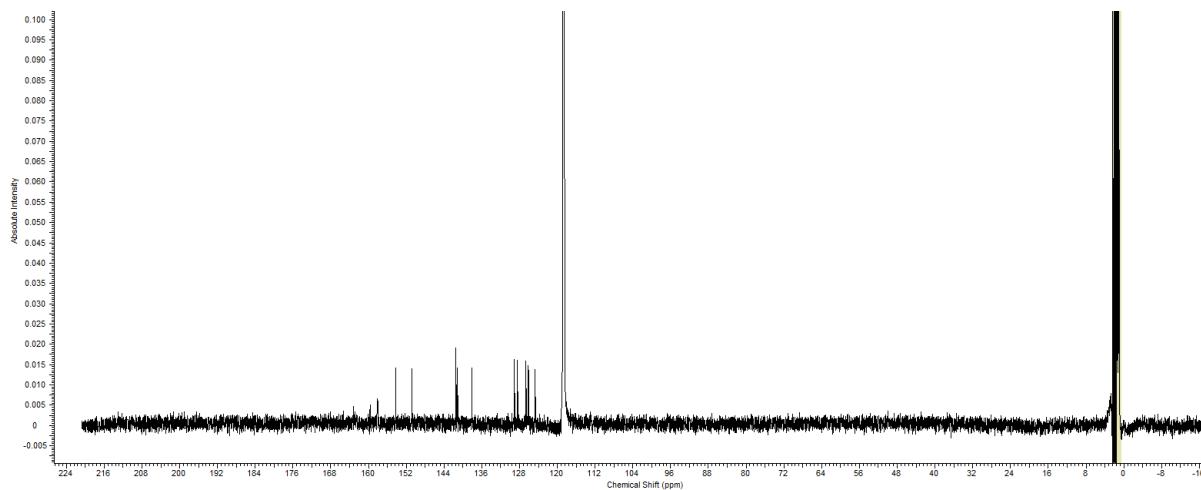


Figure 2.6 The ¹³C NMR spectra of **2.1**. Spectra for other compounds could not be collected

2.3.2 Structure Analysis with X-Ray Crystallography and DFT

Single crystal analysis by x-ray crystallography yielded good structures of compounds **2.1**, **2.2**, **2.3**, **2.5**, and **2.8**. These are the first reported crystal structures of the κ^3 terdentate Re^I compounds. A number of structural characteristics are common between the various bidentate or terdentate complexes. Much analysis has been done on the structures of bidentate complexes in literature^{116–118} the notable characteristic within terpyridine compounds is the rotation of the pendant arm pushing the nitrogen atom away from the plane of the metal-ligand bonds by approximately 100°. Cell parameters and other collection data for compounds **2.1**, **2.3**, **2.5**, and **2.7** are located in Table 2.1.

Table 2.1 Crystal data and structure refinement for compounds **2.1**, **2.3**, and **2.5**

Compound	2.1	2.3	2.5
Empirical formula	C ₁₉ H ₁₁ N ₃ O ₃ ReCl	C ₁₉ H ₁₁ N ₃ O ₃ ReBr	C ₂₀ H ₁₁ N ₄ O ₃ Re
Formula weight (g/mol)	538.96	583.41	530.04
Temperature (K)	200	200	200
Wavelength (Å)	0.71073	0.71073	0.71073
Crystal System	Triclinic	Monoclinic	Triclinic
Space Group	P-1	C2/c	P-1
a (Å)	9.8736(4)	31.1537(7)	9.9196(9)
b (Å)	14.8202(4)	7.1176(2)	14.9902(14)
c (Å)	16.3472(4)	16.8519(4)	16.5187(15)
α (deg)	69.2890(10)	90.000	68.363(2)
β (deg)	80.801(2)	111.0230(10)	80.929(2)
γ (deg)	79.836(2)	90.000	79.975(2)
Volume (Å ³)	2190.00(12)	3488.00	2236.6(4)
Z, r (calc) (Mg/m ³)	2, 1.997	8, 2.222	2, 1.927
Absorption coefficient (mm ⁻¹)	6.063	9.282	5.821
Absorption correction	Semi-empirical from equivalents		
Final R indices [I≥2σ(I)]	R1 = 0.0397, wR2 = 0.0839	R1 = 0.0232, wR2 = 0.0614	R1 = 0.0390, wR2 = 0.0921
R indices (all data)	R1 = 0.0604, wR2 = 0.0951	R1 = 0.0285, wR2 = 0.0642	R1 = 0.0500, wR2 = 0.0961

In addition, the structures of all species were optimized using Gaussian 09¹¹⁹ employing the B3LYP^{120,121} exchange-correlation (XC) functional. The LanL2DZ basis

set/effective core potential¹²² was used on Re, and the all-electron TZVP basis set¹²³ for the remaining lighter atoms. Frequency analysis of all structures was used to confirm the nature of the stationary points. Solvent effects were computed when necessary using the integral equation formalism variant of the Polarizable Continuum Model (PCM) for solvation within Gaussian 09.^{124,125} The results of these calculations are compared to the x-ray crystallography data in the tables below.

The crystal structure of **2.3** had a higher symmetry than the other samples. Details on the exact methods used for structure elucidation are available in section A.3, but all of the structures found are of high quality. The structures of **2.1**, **2.3**, and **2.5** can be seen in Figure 2.7. More views of these structures can be seen in Appendix section A.3. Crystals suitable for x-ray analysis were unable to be collected from compound **2.7**.

Selected bond lengths, bond angles, and torsion angles are listed in Table 2.3, Table 2.4 and Table 2.5 for products **2.1**, **2.3**, and **2.5** respectively. The experimental results agree closely with the computed values for all samples. The structures can be seen in Figure 2.7a, Figure 2.7b and Figure 2.7c, and more views of these structures can be seen in Appendix B, Molecular Orbitals Diagrams.

The Re^I centre in the pseudooctahedral complex is supported by a planar, pincer coordinated ligand defined by the terminal and central pyridyl group of the terpyridine. One of the carbonyl groups lies in this plane trans to the central pyridyl group, while the remaining carbonyl groups and the anionic group or other complexed species lie on an approximately perpendicular plane to the ligand. Bond angles around the Re centre show a significant deviation of up to 15° from the ideal octahedral geometry for all samples analyzed. The typical N-Re-N bond angle of 75° is due to the atomic size of rhenium, comparison to the crystal structure of an analogous compound with a manganese atom¹²⁶ shows an increase in the bonding angle for the Mn complex by approximately 4° due to

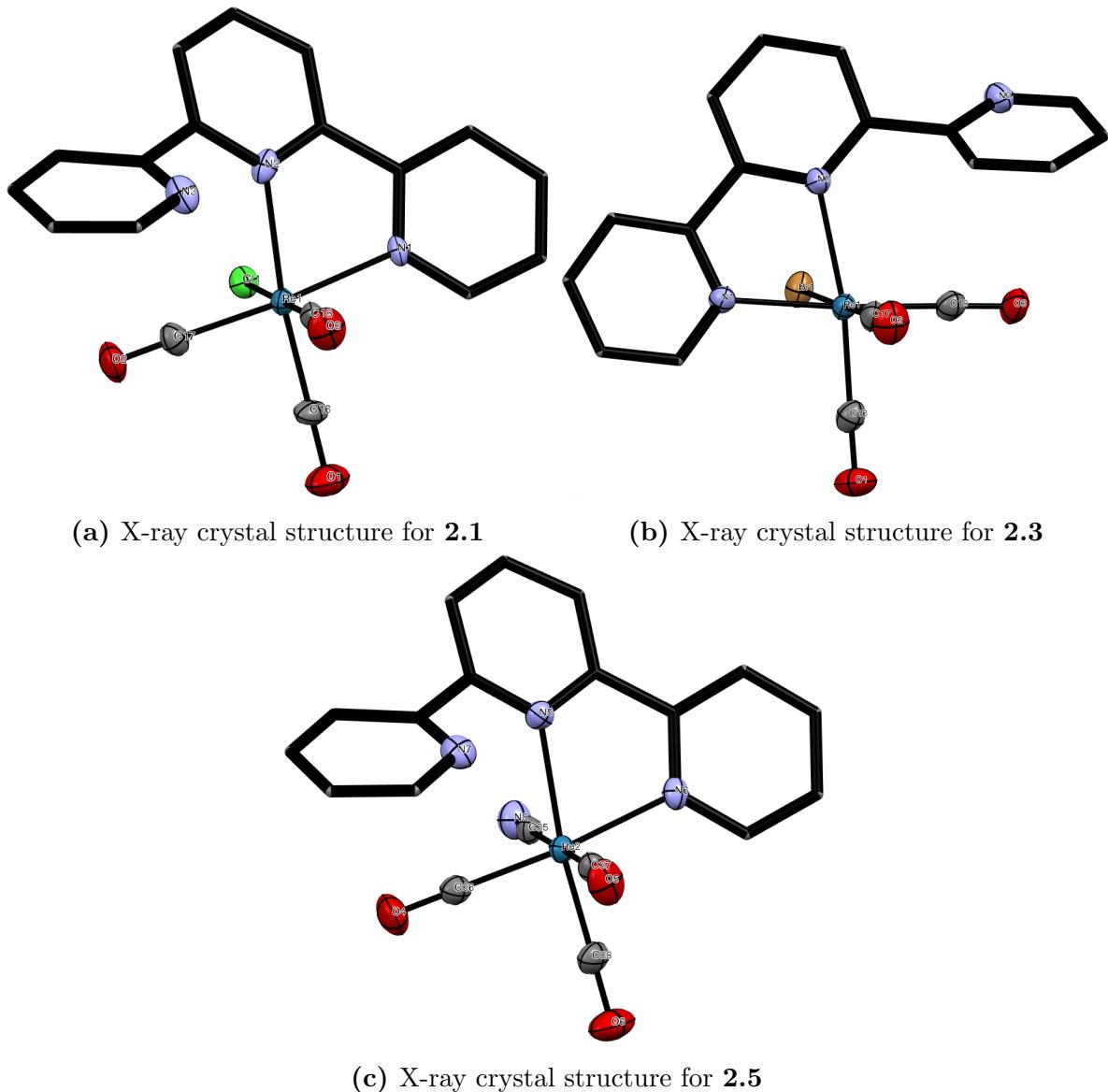


Figure 2.7 X-ray crystal structure representation for **2.1**, **2.3**, and **2.5**. Co-crystallized chloroform, hydrogen atoms, and thermal ellipsoids of ligand carbon atoms are omitted for clarity.

a decrease of bond length from metal to nitrogen of about 0.12 Å for both the central and terminal pyridines.

The deviation from octahedral is further visible in the rotation of the X-Re-CO plane (X=halide, anion, complexed group) by approximately 10 degrees from the right angle

relative to the plane of the ligand. The axial halide, anion, or chelated group typically occupies a position slightly departed from the perpendicular, angled to be over the ligand. This eclipsing is due to some unknown process, no steric interference exists upon that site, analysis of electrostatic or other short-range electronic effects computationally show any interaction between this site and the aromatic rings. In Re^I complexes, the halide or anion is located axial relative to the plane of the ligand. For the acetonitrile complex with triflate counterion, the acetonitrile occupies the axial position. This site occupation holds through the complete set of x-ray crystal structures with a κ^2 -(bipy)Re(CO)₃X core structure motif deposited in the Cambridge Crystallography Data Centre (CCDC) database,¹²⁷ and extends through other α -diimine complexes seen in our lab and in literature.⁶⁹

Table 2.2 Solvated and Gas Phase Energies of Axial & Trans κ^x -(terpy)-Re(CO)_{5-x}CN (x=2,3)

Geometry	Bidentate		Terdentate	
	E(gas) ^a	E(solution) ^b	E(gas) ^a	E(solution) ^b
Axial	-1254.96132	-1254.99059	-1141.57255	-1141.60827
Trans	-1254.93788	-1254.97168	-1141.54660	-1141.58209
Difference	0.02343	0.01891	0.02595	0.02612
Difference (kcal/mol)	14.70	11.87	16.28	16.43

^a B3LYP SCF energy in hartrees.

^b B3LYP SCF energy in hartrees with PCM solvation in acetonitrile.

The crystal structure for compound **2.5** contains two molecules per unit cell, one of which is solved to have the cyano group in the position trans to the ligand. However, careful analysis of the bond lengths, angles, and torsion data in Table 2.5 shows a remarked similarity between all -CO and -CN groups. Additionally, in an x-ray diffraction pattern, -CN and -CO are difficult to differentiate with certainty. Thus, while the structure solved to the two isomers, critical analysis would suggest that this molecule does

not violate the axial position pattern laid out above. The computed structures energies in Table 2.2 show a favouring of the axial position by 12-16 kJ/mol in the gas phase and by PCM in a simulated acetonitrile solvent.

Selected bond lengths, bond angles, and torsions are listed in Table 2.3, Table 2.4 and Table 2.5 for products **2.1**, **2.3**, and **2.5** respectively.

Table 2.3 Selected Distances, Angles, and Torsions for **2.1**

Bond	Distance (Å)	
	Experimental	Calculated
Re(1)-C(16)	1.89(1)	1.916
Re(1)-C(17)	1.934(8)	1.936
Re(1)-C(18)	1.90(1)	1.918
Re(1)-N(1)	2.162(6)	2.197
Re(1)-N(2)	2.236(9)	2.293
Re(1)-Cl(1)	2.496(2)	2.525
C(16)-O(1)	1.16(1)	1.15323
C(17)-O(2)	1.12(1)	1.15039
C(18)-O(3)	1.15(1)	1.15627
Angle	Degrees (°)	
	Experimental	Calculated
C(16)-Re(1)-C(17)	87.6(4)	86.877
C(16)-Re(1)-C(18)	88.3(4)	90.613
C(17)-Re(1)-C(18)	87.3(4)	89.557
C(16)-Re(1)-N(1)	96.4(3)	96.240
C(17)-Re(1)-N(1)	174.9(3)	175.600
C(18)-Re(1)-N(1)	95.9(3)	93.506
C(16)-Re(1)-N(2)	169.3(3)	170.368
C(17)-Re(1)-N(2)	101.1(3)	102.755
C(18)-Re(1)-N(2)	98.3(3)	89.415
N(2)-Re(1)-N(1)	74.5(3)	74.146
C(16)-Re(1)-Cl(1)	91.7(3)	91.453
C(17)-Re(1)-Cl(1)	91.7(3)	94.786
C(18)-Re(1)-Cl(1)	179.9(3)	175.286
N(1)-Re(1)-Cl(1)	84.0(2)	82.058
N(2)-Re(1)-Cl(1)	81.6(2)	87.840
O(1)-C(16)-Re(1)	179.6(9)	178.224
O(2)-C(17)-Re(1)	176.0(8)	176.907
O(3)-C(18)-Re(1)	177.3(9)	179.317
Torsion	Degrees (°)	
	Experimental	Calculated
N(1)-C(5)-C(6)-N(2)	16(1)	15
N(2)-C(10)-C(11)-N(3)	41(1)	139

Table 2.4 Selected Distances, Angles, and Torsions for **2.3**

Bond	Distance (\AA)	
	Experimental	Calculated
Re(1)-C(16)	1.911(3)	1.91740
Re(1)-C(17)	1.890(3)	1.91814
Re(1)-C(18)	1.921(4)	1.93897
Re(1)-N(1)	2.173(3)	2.19687
Re(1)-N(2)	2.232(2)	2.28998
Re(1)-Br(1)	2.6410(4)	2.67953
C(16)-O(1)	1.150(4)	1.15290
C(17)-O(2)	1.157(4)	1.15012
C(18)-O(3)	1.155(5)	1.15591
Angle	Degrees ($^\circ$)	
	Experimental	Calculated
C(16)-Re(1)-C(17)	89.1(1)	90.772
C(16)-Re(1)-C(18)	85.9(1)	86.823
C(16)-Re(1)-N(1)	97.9(1)	96.034
C(17)-Re(1)-N(1)	92.5(1)	93.597
C(18)-Re(1)-N(1)	175.4(1)	175.575
C(16)-Re(1)-N(2)	171.2(1)	170.290
C(17)-Re(1)-N(2)	96.0(1)	89.435
C(18)-Re(1)-N(2)	101.3(1)	102.886
N(1)-Re(1)-N(2)	74.7(1)	74.265
C(16)-Re(1)-Br(1)	92.7(1)	90.399
C(17)-Re(1)-Br(1)	177.6(1)	176.076
C(18)-Re(1)-Br(1)	91.6(1)	94.069
N(1)-Re(1)-Br(1)	85.74(7)	82.555
N(2)-Re(1)-Br(1)	82.07(7)	88.780
O(1)-C(16)-Re(1)	178.6(3)	178.270
O(2)-C(17)-Re(1)	179.5(3)	179.355
O(3)-C(18)-Re(1)	179.9(3)	176.781
Selected Torsions (deg)		
N(1)-C(6)-C(1)-N(2)	-15.4(4)	-14.749
N(2)-C(5)-C(11)-N(3)	141.1(3)	136.119

Table 2.5 Selected Distances, Angles, and Torsions for **2.5**

Axial CN				Planar CN			
Bond	Distance (Å)		Bond	Distance (Å)			
	Exp.	Calc.		Exp.	Calc.		
Re(2)-C(35)	2.148(7)	2.13963	Re(1)-C(19)	2.105(8)	1.98769		
Re(2)-C(36)	1.926(6)	1.94011	Re(1)-C(16)	1.928(5)	2.09197		
Re(2)-C(37)	1.954(7)	1.96758	Re(1)-C(18)	1.96(1)	2.00792		
Re(2)-C(38)	1.902(9)	1.91853	Re(1)-C(17)	1.918(7)	1.90499		
Re(2)-N(5)	2.242(7)	2.28998	Re(1)-N(1)	2.253(5)	2.32197		
Re(2)-N(6)	2.168(5)	2.20279	Re(1)-N(2)	2.176(4)	2.18806		
C(35)-N(8)	1.138(9)	1.16104	C(19)-O(3)	1.17(1)	1.14703		
C(36)-O(4)	1.145(8)	1.15044	C(16)-N(4)	1.149(7)	1.16100		
C(37)-O(5)	1.151(9)	1.15134	C(18)-O(2)	1.14(1)	1.14276		
C(38)-O(6)	1.17(1)	1.15368	C(17)-O(1)	1.130(8)	1.15781		
Angle	Degrees (°)		Angle	Degrees (°)			
	Exp.	Calc.		Exp.	Calc.		
C(36)-Re(2)-C(38)	87.7(3)	87.273	C(16)-Re(1)-C(17)	87.8(3)	90.158		
C(36)-Re(2)-C(37)	88.0(3)	89.890	C(16)-Re(1)-C(18)	87.0(3)	84.822		
C(36)-Re(2)-C(35)	92.1(3)	93.356	C(16)-Re(1)-C(19)	92.5(3)	88.356		
C(38)-Re(2)-C(37)	88.5(3)	90.973	C(17)-Re(1)-C(18)	88.7(3)	88.453		
C(38)-Re(2)-C(35)	90.8(3)	91.628	C(17)-Re(1)-C(19)	90.5(3)	87.745		
C(37)-Re(2)-C(35)	179.2(3)	175.933	C(18)-Re(1)-C(19)	179.1(3)	172.179		
C(36)-Re(2)-N(5)	100.6(3)	102.576	C(16)-Re(1)-N(1)	102.2(2)	98.105		
C(36)-Re(2)-N(6)	174.2(3)	175.708	C(16)-Re(1)-N(2)	175.9(2)	172.047		
C(38)-Re(2)-N(5)	169.3(3)	170.146	C(17)-Re(1)-N(1)	168.3(3)	170.509		
C(38)-Re(2)-N(6)	96.6(3)	96.171	C(17)-Re(1)-N(2)	95.9(3)	97.544		
C(37)-Re(2)-N(5)	98.4(2)	89.360	C(18)-Re(1)-N(1)	97.7(3)	88.487		
C(37)-Re(2)-N(6)	96.0(2)	92.605	C(18)-Re(1)-N(2)	94.8(3)	93.374		
C(35)-Re(2)-N(5)	82.3(2)	87.543	C(19)-Re(1)-N(1)	83.2(2)	96.317		
C(35)-Re(2)-N(6)	83.9(2)	84.008	C(19)-Re(1)-N(2)	85.7(2)	93.899		
N(5)-Re(2)-N(6)	74.7(2)	73.977	N(1)-Re(1)-N(2)	73.9(2)	73.675		
O(6)-C(38)-Re(2)	179.4(7)	178.027	O(1)-C(17)-Re(1)	178.2(7)	177.623		
O(5)-C(37)-Re(2)	175.5(6)	179.414	O(2)-C(18)-Re(1)	172.0(7)	176.452		
N(8)-C(35)-Re(2)	178.0(6)	176.457	O(3)-C(19)-Re(1)	178.0(6)	176.552		
O(4)-C(36)-Re(2)	179.0(7)	177.313	N(4)-C(16)-Re(1)	178.7(6)	178.113		
Torsion	Degrees (°)		Torsion	Degrees (°)			
	Exp.	Calc.		Exp.	Calc.		
N(5)-C(20)-C(25)-N(6)	14.5(9)	13.735	N(1)-C(1)-C(6)-N(2)	12.5(8)	14.777		
N(5)-C(24)-C(30)-N(7)	41(1)	135.774	N(1)-C(5)-C(11)-N(3)	43.7(9)	137.014		

Structural comparisons between the bidentate samples and the terdentate show many similarities. The loss of one carbonyl always accompanies the chelation of the pendant arm of the ligand. The increased coordination forces the ligand to adopt a more rigidly planar geometry, this is visible in the structure of **2.2** (Figure 2.8a) and **2.8** (Figure 2.8b). Suitable crystals for x-ray analysis were unable to be collected for the remaining complexes.

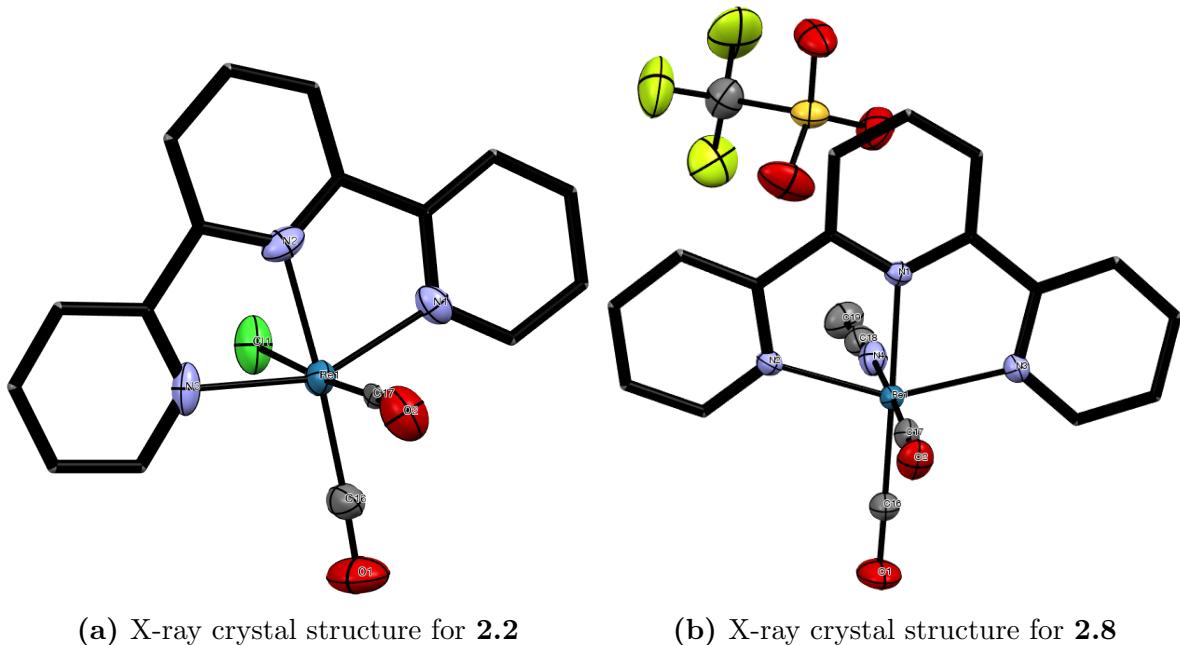


Figure 2.8 X-ray crystal structure representation for **2.2** and **2.8**. Co-crystallized chloroform, hydrogen atoms, and thermal ellipsoids of ligand carbon atoms are omitted for clarity.

Selected bond lengths, bond angles, and torsions are listed in Table 2.6 and Table 2.7 for products **2.2** and **2.8**. Cell parameters and collection data can be found in Table ??.

Both **2.2** and **2.8** structures are pseudooctahedral in geometry, with *mer* coordinated terpyridine ligand. The carbonyl groups and chloride/acetonitrile form a plane approximately perpendicular to the ligand. As in the κ^2 -bidentate samples discussed above, the carbonyls occupy the equatorial and one axial position relative to the ligand, and the

chloride or acetonitrile occupy the remaining axial position. The N-Re-N bond angles remain approximately 75°, and the central pyridyl N to chloride or acetonitrile angle is still approximately 80-85°.

Comparisons between the κ^2 -bidentate and κ^3 -terdentate samples (**2.1** and **2.2**) highlight the geometrical changes experienced in the thermolysis reaction. The distance from Re to the central pyridyl N has shortened from 2.24 to 2.08 Å. This bond shortening of 0.16 Å signifies increased metal-ligand interaction. This comes at the expense of decreased interaction with the carbonyl groups, with the bond to the planar CO increasing from 1.89 to 1.93 Å and the axial CO bond increasing from 1.901 to 1.975 Å. As the carbonyls experience less π backbonding from the metal, the internal C-O bond shortens by as much as 0.1 Å.

The κ^3 -terdentate samples (**2.2** and **2.8**) provide the opportunity to analyze both neutral and cationic species. Due to the weakly coordinated triflate anion in **2.8**, a number of geometric differences arise compared to **2.2**. While the ligand is coordinated by only 0.01 - 0.02 Å closer to the metal atom, the carbonyl groups are 0.04 - 0.1 Å closer, indicating their increased electron donation to the electron poor metal centre. As well, the C-O bonds in the carbonyls are 0.03 - 0.12 Å longer, indicating the increased π -backbonding occurring in the cation.

Table 2.6 Selected Distances, Angles, and Torsions for **2.2**

Bond	Distance (\AA)	
	Experimental	Calculated
Re(1)-C(16)	1.926(9)	1.92438
Re(1)-C(17)	1.975(10)	1.90687
Re(1)-N(1)	2.119(7)	2.13186
Re(1)-N(2)	2.080(7)	2.08705
Re(1)-N(3)	2.126(7)	2.13185
Re(1)-Cl(1)	2.489(3)	2.53337
N(1)-N(3)	4.14(1)	4.14772
C(16)-O(1)	1.14(1)	1.16042
C(17)-O(2)	1.05(1)	1.16341
Angle	Degrees ($^\circ$)	
	Experimental	Calculated
C(16)-Re(1)-C(17)	91.5(4)	89.188
C(16)-Re(1)-N(2)	173.7(4)	172.050
C(17)-Re(1)-N(2)	94.6(3)	98.762
C(16)-Re(1)-N(1)	103.9(3)	102.980
C(17)-Re(1)-N(1)	92.7(3)	93.429
N(2)-Re(1)-N(1)	77.3(3)	76.684
C(16)-Re(1)-N(3)	101.8(3)	102.986
C(17)-Re(1)-N(3)	91.7(3)	93.419
N(2)-Re(1)-N(3)	76.6(3)	76.684
N(1)-Re(1)-N(3)	153.7(3)	153.210
C(16)-Re(1)-Cl(1)	91.8(3)	89.136
C(17)-Re(1)-Cl(1)	176.5(2)	178.324
N(2)-Re(1)-Cl(1)	82.1(2)	82.913
N(1)-Re(1)-Cl(1)	85.4(2)	86.953
N(3)-Re(1)-Cl(1)	88.7(2)	86.953
O(1)-C(16)-Re(1)	177.9(9)	179.079
O(2)-C(17)-Re(1)	173.2(8)	179.182
Selected Torsions (deg)		
N(1)-C(5)-C(6)-N(2)	1(1)	2
N(2)-C(10)-C(11)-N(3)	-4(1)	-2

Table 2.7 Selected Distances, Angles and Torsions for Acetonitrile Adduct of **2.8**

Bond	Distance (Å)	
	Experimental	Calculated
Re(1)-C(16)	1.889(4)	1.93046
Re(1)-C(17)	1.885(3)	1.92844
Re(1)-N(1)	2.091(3)	2.10116
Re(1)-N(2)	2.135(3)	2.15397
Re(1)-N(3)	2.131(3)	2.15392
Re(1)-N(4)	2.160(3)	2.15202
N(2)-N(3)	4.138(4)	4.18483
C(16)-O(1)	1.170(4)	1.15749
C(17)-O(2)	1.171(4)	1.15244
Angle	Degrees (°)	
	Experimental	Calculated
C(16)-Re(1)-C(17)	87.69(16)	88.104
C(16)-Re(1)-N(1)	175.95(12)	176.094
C(17)-Re(1)-N(1)	96.35(12)	95.802
C(16)-Re(1)-N(3)	103.81(13)	103.594
C(17)-Re(1)-N(3)	94.03(12)	92.309
N(1)-Re(1)-N(3)	76.20(10)	76.306
C(16)-Re(1)-N(2)	103.58(13)	103.598
C(17)-Re(1)-N(2)	93.73(12)	92.307
N(1)-Re(1)-N(2)	75.99(10)	76.305
N(3)-Re(1)-N(2)	151.77(11)	152.544
C(16)-Re(1)-N(4)	90.50(14)	88.484
C(17)-Re(1)-N(4)	178.10(12)	176.587
N(1)-Re(1)-N(4)	85.46(10)	87.611
N(3)-Re(1)-N(4)	86.94(10)	88.504
N(2)-Re(1)-N(4)	86.15(10)	88.485
O(1)-C(16)-Re(1)	179.1(3)	178.807
O(2)-C(17)-Re(1)	178.0(3)	178.860
Torsion	Degrees (°)	
	Experimental	Calculated
N(1)-C(1)-C(6)-N(2)	1.7(4)	1.105
N(1)-C(5)-C(11)-N(3)	-1.8(4)	-1.110

Table 2.8 Crystal data and structure refinement for compounds **2.2** and **2.8**

Compound	2.2	2.8
Empirical formula	C ₁₈ H ₁₁ N ₃ O ₂ ReCl	C ₂₁ H ₁₄ N ₄ O ₅ F ₃ SRe
Formula weight (g/mol)	510.95	665.61
Temperature (K)	200	200
Wavelength (Å)	0.71073	0.71073
Crystal System	Triclinic	Triclinic
Space Group	P-1	P-1
a (Å)	8.5275(3)	8.5745(4)
b (Å)	14.2421(5)	11.9805(5)
c (Å)	17.4637(6)	13.0970(5)
α (deg)	77.948(2)	79.748(2)
β (deg)	85.684(2)	81.106(2)
γ (deg)	79.890	88.091(2)
Volume (Å ³)	2041.79(12)	1307.99(10)
Z, r (calc) (Mg/m ³)	4, 2.050	2, 1.993
Absorption coefficient (mm ⁻¹)	6.494	5.094
Absorption correction	Semi-empirical from equivalents	
Final R indices [I≥2σ(I)]	R1 = 0.0636, wR2 = 0.1018	R1 = 0.0294, wR2 = 0.0673
R indices (all data)	R1 = 0.0985, wR2 = 0.1110	R1 = 0.0366, wR2 = 0.0700

2.3.3 Infrared Spectroscopy

Conversion of bidentate to terdentate species was confirmed utilizing Fourier Transform Infrared (FTIR) spectroscopy. A small sample of powder product was placed on the Agilent Cary 630 FTIR spectrometer, with a 2 cm⁻¹ resolution. The instrument is fitted with a diamond ATR for solid sample analysis. Spectra are the Fourier transform of 16 scans.

Analysis of the results in Figure 2.9 shows the significant reduction of one peak in the (ca.) 2100 cm⁻¹ region. This peak is in the CO stretching frequency, the frequency of the peak lost in thermolysis is indicative of a weakly coordinated carbonyl group. A

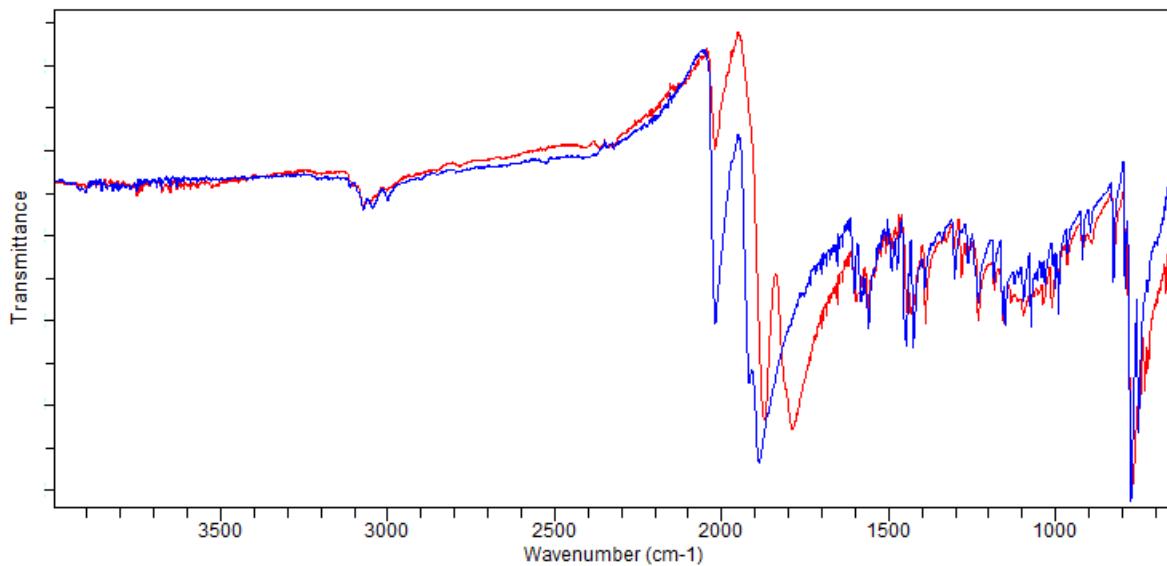


Figure 2.9 FTIR Spectra for complexes **2.1** (blue) and **2.2** (red).

splitting occurs for the other large peak and its shoulder in the conversion from bidentate to terdentate, from 1890 to 1790 cm⁻¹, indicating the further weakening of the metal carbonyl bonds remaining in the complex. This weakened bond is likely the carbonyl co-planar to the ligand, analysis of the x-ray crystal structure shows the CO bond to be 0.1 Å longer than that of the axial carbonyl. This identification is supported by the DFT methods discussed below.

FTIR spectra were predicted using molecular frequency calculations of DFT optimized structures in Gaussian09. Prediction of this spectra was performed as a verification of the optimized structures discussed in subsection 2.3.2 above. The calculation identifies stretching or bending harmonic energies in optimized structures. The computed spectra in ?? shows three peaks for **2.1**, at 2094, 2012, and 1990 cm⁻¹. The relative location of these peaks corresponds to those seen in Figure 2.10 for **2.1**, but are shifted by approximately 100 cm⁻¹ to higher energy relative to the experimental. Similarly, for **2.2**, peaks are seen at 2000 and 1950 cm⁻¹, compared to the 1890 and 1790 peaks seen experimen-

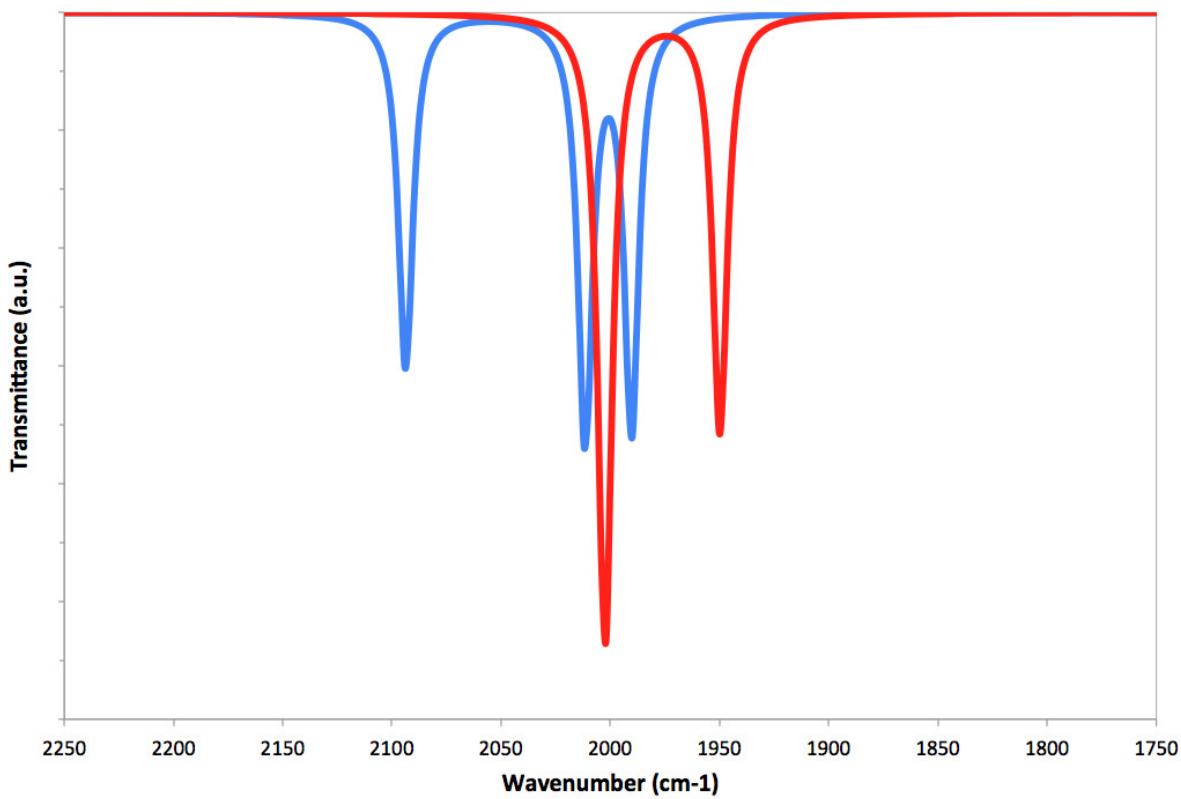


Figure 2.10 DFT predicted FTIR spectra for **2.1** (blue) and **2.2** (red).

tally. This computed value reflects the shift to lower energy carbonyl stretching from **2.1** to **2.2**, and echoes the experimental spectra well.

Further analysis of other spectra were not successful in identification of any additional molecular properties, with the exception of a series of strong peaks appearing in the 1200-1300 cm⁻¹ range, confirming the presence of the triflate anion from the -SO₃ group vibrations for samples **2.7** and **2.8** (Figure 2.11). Additionally, the small peak present at 2270 cm⁻¹ in the terdentate sample corresponds to the weakly coordinated acetonitrile occupying the molecule's axial position.

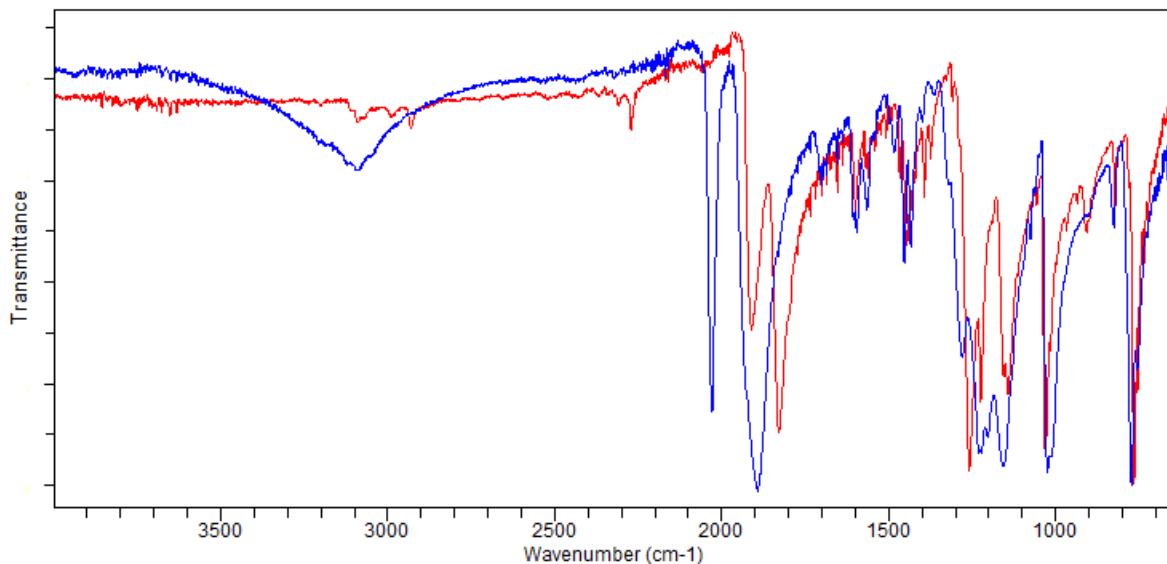
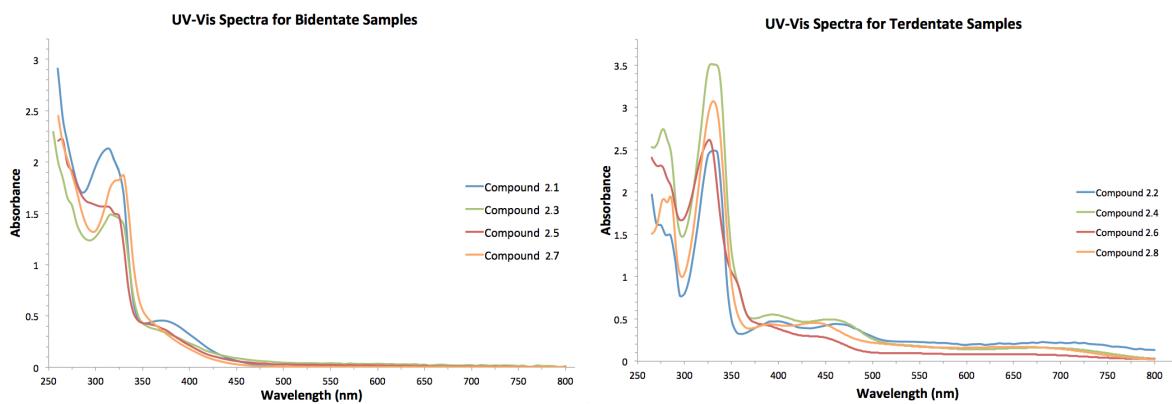


Figure 2.11 FTIR Spectra for complexes **2.7** (blue) and **2.8** (red).

2.3.4 Photophysical Properties

A striking observation upon the conversion of the bidentate species into the terdentate complexes is that these new compounds have a substantially darker colour that reflects a significant change in the photophysical properties. This effect was investigated using a combination of UV-visible spectroscopy and DFT modelling. Spectra were collected on a Agilent Cary 5000 UV-Vis-NIR Spectrophotometer. The stronger absorbance of the terdentate complexes compared to the bidentate precursors is evident in the UV-Vis spectra of these species, and is presented in Figure 2.12. These spectra were obtained in Dimethylsulfoxide (DMSO) with approximate concentrations of 0.01 mM for bidentate, and an order of magnitude lower (0.001 mM) for the terdentate analogues. The terdentate complexes have more intense absorbance for higher energy ligand UV-based $\pi-\pi^*$ transitions (<400 nm) and certainly a richer visible absorption profile that involve the d- π^* transitions. These features are responsible for the colour change observed.

The UV-Vis spectra were modeled using TD-DFT computations within Gaussian 09,



(a) UV-Vis spectra for compounds **2.1**, **2.3**, **2.5**, and **2.7** (b) UV-Vis spectra for compounds **2.2**, **2.4**, **2.6**, and **2.8**

Figure 2.12 UV-Vis spectra for all compounds. Concentrations of bidentate complexes are 0.01 mM and terdentate complexes are 0.001 mM.

using the B3LYP functional with the LanL2DZ basis set and effective core potentials for the rhenium atom, and the TZVP basis set for all lighter atoms. Such functional and basis set choices are common within organometallic literature. Solvent was simulated using the integral equation formalizm variant of the PCM solvation model, with DMSO as the solvent.

The resulting computed spectra were excellent matches to the experimental spectra. The similarity of the spectra between the bidentate species and as well as the terdentate species indicates that parallel electronic transitions appear within these two groups. Specifics will be discussed for the chloro compounds, **2.1** and **2.2**, but strong parallels exist for all species. Plots of experimental and computational data for these two complexes are presented in Figure 2.13 and Figure 2.14, respectively.

Common to all bidentate compounds were peaks at wavelengths of 315-320 nm, and peak shoulders near 350-375 nm (Figure 2.12a). In the case of **2.1**, the computational results suggest that the experimental band centred at $\lambda = 315\text{-}320 \text{ nm}$ ($31750\text{-}31250 \text{ cm}^{-1}$) primarily arises from two electronic transitions. The first is from HOMO-3 to

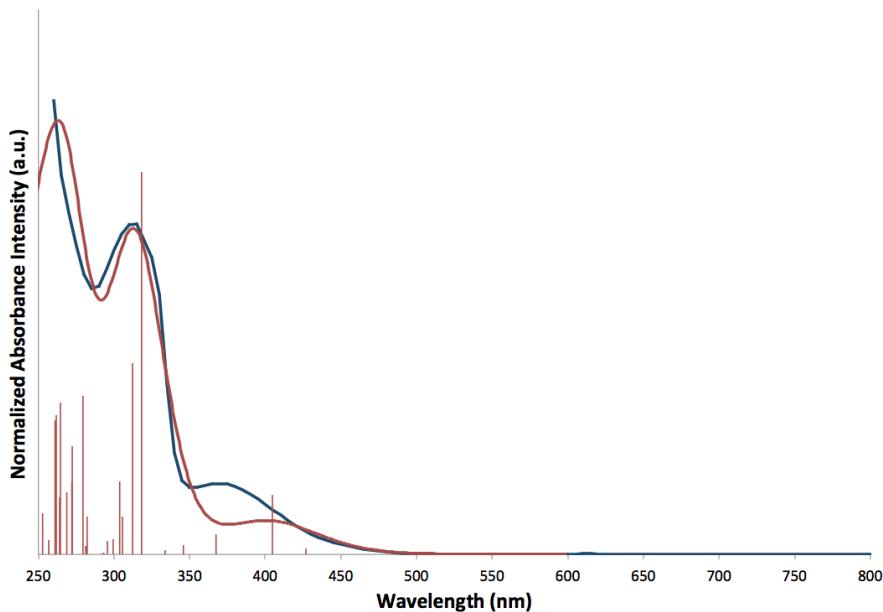


Figure 2.13 Plots of the experimental and computed UV-Vis spectra for compound **2.1**. The blue curve shows experimental result. The red vertical lines show the calculated transitions and relative intensities from the TDDFT calculations, while the red curve shows the Gauss convolution with peak width at half height of 0.250 eV.

LUMO, a $\pi-\pi^*$ transition, while the second major contribution is the excitation from HOMO to LUMO+2, which is a d- π^* transition. The TD-DFT calculations suggest that the experimental band centred at $\lambda = 370 \text{ nm}$ (27030 cm^{-1}) corresponds to a calculated transition at approximately 400 nm, that is a d- π^* (HOMO-1 to LUMO) absorbance. Computed plots of molecular orbitals are included in Appendix C Molecular Orbitals Diagrams, Figures C.1 to C.8.

Like the bidentate samples, the spectra for the terdentate samples are quite similar to one another. In general, all of these species have much higher absorbance coefficients than their bidentate analogues; all contain long trailing absorptions across the wavelengths analysed. For terdentate complex **2.2**, the four experimentally observed UV-Vis absorptions observed are made up of six computed transitions. The computational model provided two equal and strong transitions appearing at 308 and 325 nm to generate the

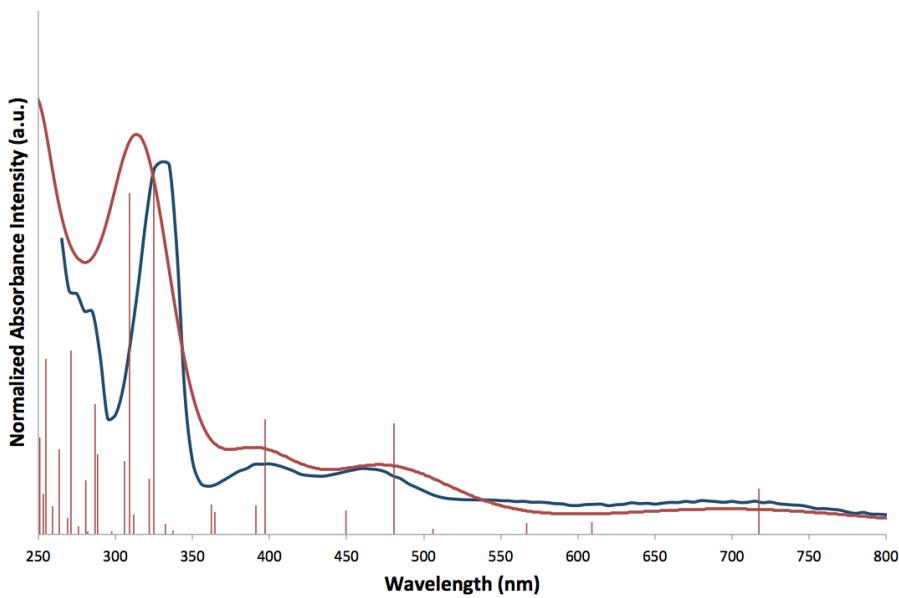


Figure 2.14 Plots of the experimental and computed UV-Vis spectra for compound **2.2**. The blue curve shows experimental result. The red vertical lines show the calculated transitions and relative intensities from the TDDFT calculations, while the red curve shows the Gauss convolution with peak width at half height of 0.250 eV.

experimental band centred at $\lambda = 330 \text{ nm}$ (30300 cm^{-1}). These arise from a transition from HOMO-5, to LUMO, a Cl lone-pair to terpy π^* -orbital absorption, and from HOMO-3 to LUMO, which is a ligand centred $\pi-\pi^*$ transition. All of the lower energy transitions are dominated by MLCT bands. The experimental band centered at $\lambda = 400 \text{ nm}$ (25000 cm^{-1}) arises from an electronic transition from HOMO to LUMO+3 which is d- π^* in nature. The absorbance centred at $\lambda = 460 \text{ nm}$ (21740 cm^{-1}) corresponds to two MLCT d- π^* transitions; HOMO-2 to LUMO and HOMO-1 to LUMO+1. Finally, the broad experimental band 680-715 nm ($14700\text{-}14000 \text{ cm}^{-1}$) arises from the electronic transition appearing at 718 nm, which is the excitation of a d-electron in the Re-centered HOMO to the ligand π^* LUMO.

2.3.5 Fluorescence

Fluorescence data was collected for **2.1** and **2.2** using an Agilent Cary Eclipse Fluorescence Spectrophotometer, using a 5 nm excitation slit and emission slit, using excitation wavelengths selected to correspond to the centre of UV-Vis absorption bands. Data was collected for spectra in N,N-dimethylformamide (DMF) to simulate the photocatalytic reaction environment (see chapter 3). Spectra are shown in Figure 2.15 for **2.1** and **2.2** with their UV-Vis absorption spectra for comparison. Data is normalized to equil concentration for each sample in fluorescence and absorbance.

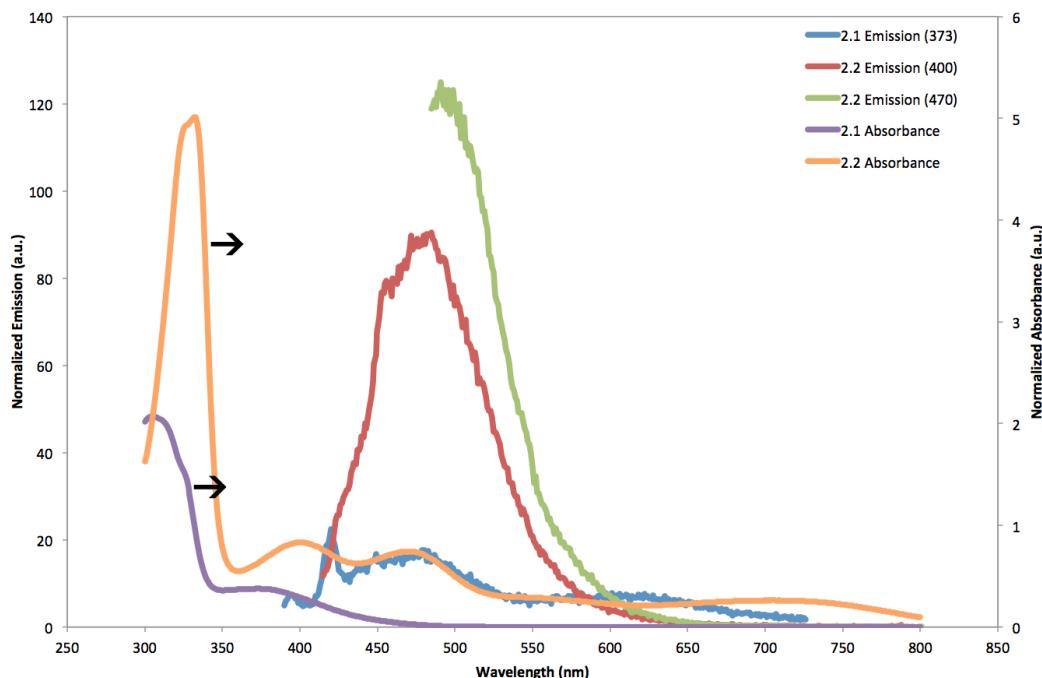


Figure 2.15 UV-Vis and fluorescence spectra for **2.1** and **2.2**. Fluorescence of **2.1** from excitation of 373 nm (blue), excitation of **2.2** by 400 nm (red) and 470 nm (green) are shown, along with the absorption spectra of **2.1** (purple) and **2.2** (orange).

Clearly the transformation from bidentate terpyridine to terdentate ligand has significant effect on the interactions of these Re compounds with visible light. Terdentate samples fluoresce and absorb more intensely at lower energy wavelengths. While the

terdentate sample is excited at various wavelengths (that should correspond to different absorption bands), the emission appears to be from one single band centred at ca. 490 nm.

Interestingly, emission is not seen with the naked eye with 400 nm excitation for the terdentate, while the bidentate emission is very strong. This may be due to self-absorbance in the terdentate samples, emission of 490 nm is easily absorbed by the molecule. In the bidentate, no absorbance bands correspond with emission seen at 470 nm or 600 nm, these appear to the naked eye as a white emission.

2.4 Conclusions

This work reported the first crystallographically authenticated rhenium(I) terpyridine terdentate complexes and thus expanded upon the prior limitations of Re coordination complexes. These terpyridine complexes are accessed via a simple, highly efficient, solid-state thermolysis pathway. They expand upon the known α -diimine photophysical properties, with enhanced metal-to-ligand d- π^* electronic transitions, occurring more frequently and with lower energy than in the associated bidentate compounds. These observations are supported by computational TD-DFT results, affording an expanded understanding of the transition bands. Modification of the bidentate and terdentate species has been shown, the synthetic success of a triflate moiety should facilitate further development of reactivity and provides an opportunity for synthetic and catalytic studies.

Chapter 3

Photocatalysis of CO₂

3.1 Introduction

Only 6 years after Re^I complexes using 2,2'-bipyridine were characterized, Hawecker, Lehn, and Ziessel showed the effectiveness of the compound for the photocatalytic reduction of CO₂.⁶² Since then, many have shown the efficacy of a wide range of α -diimino complexes for the reaction^{63,118,128} and expansion of the systems to bimetallic complexes with ruthenium and osmium as electron transfer agents has produced a wide range of results.^{108,111,129} The mechanism of reduction has been subject of some debate: while mechanisms have been proposed since Lehn et. al. soon after their original publication,⁶³ modifications have been submitted routinely over the past decades.^{85,111,112,130–134} The development of a novel terdentate geometry and the associated increase in photon absorption at lower energies of the catalyst warranted investigation of the CO₂ reduction capabilities, having overcome the criticism of only utilizing high energy photons.¹¹³

3.2 Photocatalytic Reactions with New Compounds

The photocatalytic cycle is, simply, a photon-induced MLCT, followed by the extraction of an electron from a sacrificial reductant. This radical, negatively charged species sheds the anion, opening up a reaction site. Reaction between a CO_2 , a proton (from the decomposition of the reductant or elsewhere), and the catalyst yields any number of CO, H_2O , formate (HCO_2^-), or carbonate (CO_3H^-), depending on the mechanistic pathway. Further discussion and a proposal of a new mechanism geometry based on computational and experimental data can be read in chapter 4.

3.2.1 Conditions

Reaction conditions in use in literature have remained typically unchanged since the original papers. A mixture of DMF with either Triethanolamine (TEOA) or Triethylamine (TEA) at a 5:1 ratio is used to make a 1.0 mM solution of catalyst, with ‘excess’ (depending on reference, a 1.1 to 25 molar ratio) electrolyte salt (typically Et_4NX or $t\text{-Bu}_4\text{NX}$, where X = halide from catalyst) added as a stabilizer. Solutions are degassed by bubbling of CO_2 and a consistent headspace is left to form over the solution. The reaction is monitored via Gas Chromatography (GC) analysis of the headspace, using a HP gas chromatograph with a 15 m CARBONPLOT column with 0.320 mm inner diameter and 1.50 μm film in a 40 °C oven. The instrument is fitted with a Thermal Conductivity Detector (TCD), and, while using He as a carrier gas, is able to resolve CO and CO_2 completely.

3.2.2 Experimental Results

Both bidentate and terdentate κ^n (terpy)Re(CO)_{5-n}X (n=2, 3) **2.1** and **2.2** complexes show no activity for CO₂ reduction. Modification of testing time, light source, product analysis methods, solvent, sacrificial reductant, pH, presence of electrolyte, presence of H₂O, or variation of anion (X=Cl, Br, OTf, CN) shows no change of this inactivity. Testing of κ^2 (bipy)Re(CO)₃Cl under the same reaction and testing conditions shows production of approximately 6 mL CO from CO₂ (20% conversion) in 1 hour of photolysis with visible ($\lambda > 400$ nm) light, verifying the method, isolating the catalyst as the ineffective species.

3.2.3 Rationalization of Results

Darrin, I'm still working on this section (discussed this morning)

Kurz *et al.* demonstrated the requirement for fluorescence for successful catalytic candidates.¹¹⁸ The explanation for this is the requirement for a stable, long-lasting excited state, with lifetime greater than that of the timescale required for electron abstraction from the sacrificial amine. The observed fluorescence demonstrates the lack of non-radiative relaxation pathways, considered to be an analogue for the extended lifetime of the excited state. Sample **2.2** shows only poor fluorescence. While the complex is able to absorb light across the spectrum, and has HOMO to LUMO transitions with high enough energyⁱ for the catalyzed reduction of CO₂ to CO, it appears as if the catalysis is not initiated due to a short excited state lifetime. Fluorescence data presented in chapter 2, at subsection 2.3.5 shows the lack of strong fluorescence in terdentate sample **2.2**.

Explanation for the lack of CO production observed in the attempted photochemical reduction of CO₂ by bidentate sample **2.1** and others must come from another angle.

ⁱElectrochemical reduction of CO₂ in similar environments takes 1.7-2.1 V, equivalent to HOMO-LUMO transitions from 590-750 nm.¹³⁵

These samples are quite fluorescent, emission from the powder sample can be seen with the naked eye with simulation by a 405 nm laser pen under ambient light environments. Importantly, two clues come from the reaction mixture: under intense visible light in the presence of CO₂ the compound bleaches to a very faint yellow-green and colour does not return after storage in the dark, bubbling of new CO₂, or other manipulations. Secondly, a mixture of sacrificial amine, DMF, and catalyst **2.1** in ratios identical to what is required in the reaction mixture changes colour from a yellow to a deep red irreversibly after approximately 5 days at ambient temperature.

This red product was analyzed via UV-Vis spectroscopy in the mother solution.

Blank solvent, solvent with Et₄NCl electrolyte, and a freshly prepared reaction mixture with catalyst were used as blank solutions to attempt to identify spectral characteristics

The most likely cause for the bleaching of the solution is the disassociation of terpyridine ligand from the metal centre. The colours seen in the catalyst are all due to metal-ligand interaction, and the UV-Vis spectrum of free terpyridine in solution is located solely in the UV range ($\lambda^{\max} \approx 285$ nm).¹³⁶ The labilization can be a photoinduced process.¹³⁷

The appearance of the irreversible red product may be the formation of triethanolamine-catalyst adducts.¹³⁸ In the presence of DMF, TEOA has been shown to bind to the open site of the excimer via the amine's oxygen atom. This is susceptible to insertion of CO₂ to form a -OC(O)C-CH₂CH₂N(CH₂CH₂OH)₂ group. DFT studies on these two compounds suggest that they may be a coloured species, with predicted UV-Vis showing lower energy absorption than the catalyst itself (a red shift), demonstrated in Figure 3.1.

The experimental UV-Vis with is overlaid, showing the correlation (or lack thereof) to the Ishitani compounds.

The appearance of this red product may also be due to the formation of a cationic

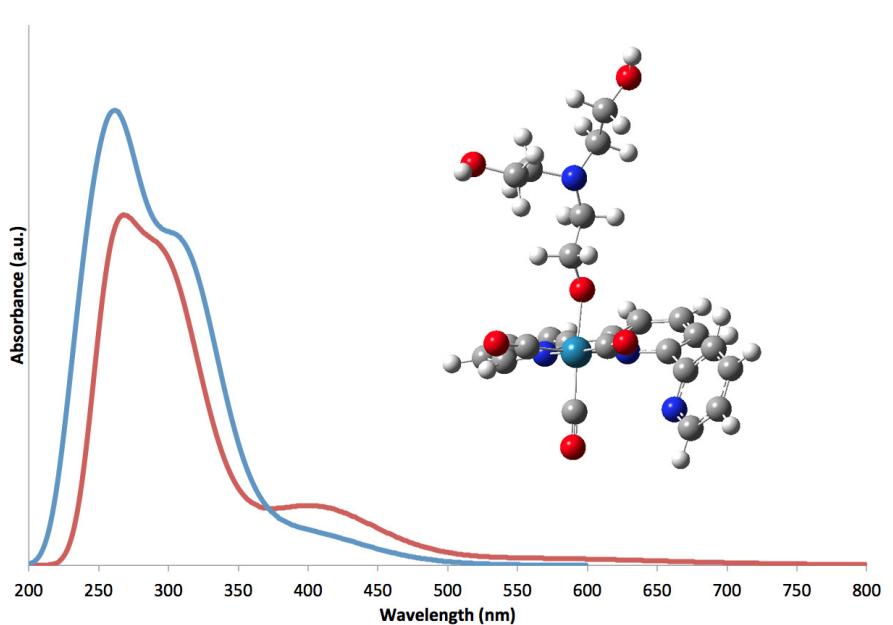
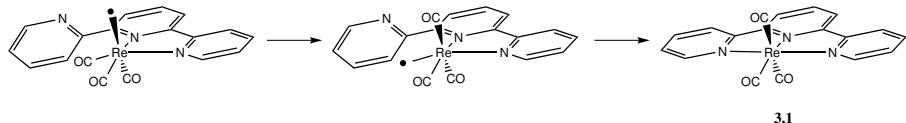


Figure 3.1 Computational structure (inset) and DFT predicted UV-Vis absorption spectra for **2.1** (blue) and the TEOA complex proposed by Ishitani



Scheme 3.1 Formation of **3.1** from catalytic excimer via reorganization of carbonyls and chelation of the pendant arm.

tridentate tricarbonyl complex, seen in Scheme 3.1. This species can be formed by the rearrangement of the carbonyl ligands from the *fac* orientation to a *mer*, migration of one carbonyl to the open site axial to the ligand achieves this reorganization. Once the *mer* species is formed, the open site on the metal centre is oriented towards the pendant arm of the terpyridine ligand, coordination of that group to the metal centre results in compound **3.1**. The UV-Visible spectra of this species is not isolated, however, DFT predicted UV-Vis is shown in Figure 3.2, and compared to the experimental red product.

UV-Vis, and
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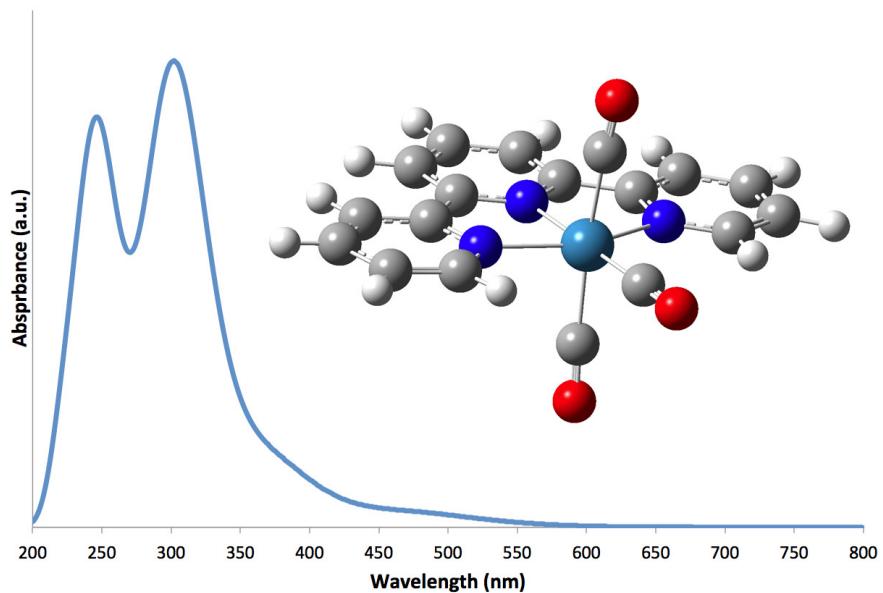


Figure 3.2 Computational structure (inset) and predicted UV-Vis absorption spectra of $[\kappa^3\text{-(terpy)-Re(CO)}_3]^+$

3.3 Conclusions

Experimental data shows the inactivity of these catalysts for photoreduction of CO_2 under a range of experimental conditions. These same conditions show conversion using $\kappa^2(\text{bipy})\text{Re(CO)}_3\text{Cl}$ as the photoreductant, signifying the validity of experimental setup. The inactivity may be due to formation of inactive side-products or ligand labilization (for the case of bidentate catalyst) or due to the short excited state lifetime of the terdentate catalyst.

Chapter 4

Mechanism of CO₂ Reduction

4.1 Introduction

Within two years of the appearance of the originally reported bipyridine rhenium (I) catalyst, experimental studies on the mechanism of the photocatalytic reduction of CO₂ were available in the literature.^{66,113} Studies continue on the mechanism up to the present day,^{hayashi 2003, 64,111,139–144} utilizing new investigative techniques as they become available to elucidate transition states and transient intermediates *in situ*.

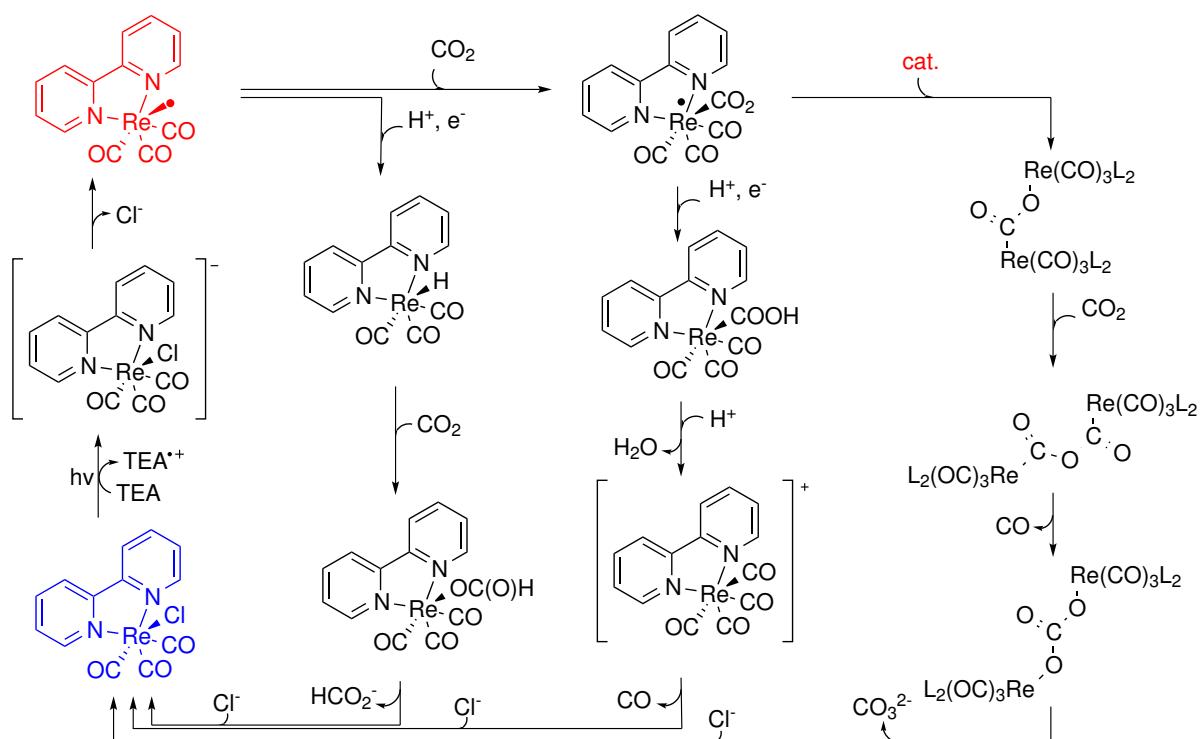
The mechanistic studies performed analyze CO₂ reduction both photocatalytically and electrocatalytically on Re compounds. The electrocatalytic activity was demonstrated by Hawecker *et. al.* only a year after the photochemistry first appeared.¹⁴⁵ Both methods involve similar cycles, and have been treated interchangeably by some authors??, the difference between methods being the production of the excimer species, not the first 2e⁻ reduction at the active site. Electrochemical methods have been employed with the hope of utilizing CO₂ and H₂O together for the direct formation of methane and oxygen, however this problem is highly comple and the target remains elusive.¹⁴⁶

Investigation includes the use of DFT methods to elucidate geometries of intermediates and transition states for of the multi-step cycle. Transition metal catalysis is a non-trivial problem computationally, particularly when considering a metal from the lower period. These elements contain a large amount of electrons, many of which can be involved in non-covalent interactions with the ligands and catalyzed products. Solving for this complex system becomes non-trivial and computationally expensive. For this reason, no overview of the mechanisms, investigated by DFT methods, has ever been made available in the literature.

4.2 Mechanism Pathways

Prior work in the literature has proposed three general mechanistic pathways for the photoreduction of CO₂.^{64,85,130,141} In general, as seen in Scheme 4.1, these pathways result in the formation of CO and H₂O, formate (HCO₂⁻), or carbonate (CO₃H⁻) anions. The formation of carbonate proceeds via the formation of a catalyst dimer over a molecule of CO₂, with the insertion of a second molecule of CO₂ to produce the carbonate and a molecule of CO. Production of formate occurs via insertion of CO₂ to a rhenium hydride bond. The formation of CO without carbonate or formate by-products occurs via the coordination of CO₂ to an open site on the metal, followed by a double proton addition and the release of a molecule of H₂O prior to the loss of one of the four carbonyl groups to open up the axial site for halide re-coordination. This is essentially the Reverse Water-Gas Shift Reaction (RWGSR), wherein protons are made available from decomposition of the sacrificial amine instead of from the decomposition of H₂.¹⁴⁷ These three mechanistic pathways will be referred to as the ‘carbonate’ mechanism, the ‘formate’ mechanism, and the ‘water-gas shift’ mechanism, respectively.

Each of the mechanistic pathways identified in Scheme 4.1 was studied, using DFT



Scheme 4.1 An overview of the mechanistic pathways of photochemical CO_2 reduction. Catalyst is shown in blue, and the excimer species in red

methods. Structures (using 2,2'-bipyridine as the bidentate ligand, and triethylamine as the sacrificial reductant) were optimized to ground or transition states using TurboMole 6.5 software,^{148,149} with the TPSS meta-GGA XC functional.¹⁵⁰ This functional shows good results with organometallic complexes, while maintaining feasible calculation times. The def2-TZVP basis set was used for all atoms.^{123,151} The TurboMole program contains a number of optimizations to the original DFT algorithms,^{152–160} decreasing the calculation time without compromising accuracy. Grimme's dispersion correction (version 3) was included in the calculations.¹⁶¹ Intermediates and transition states were verified by frequency analysis,^{157,162,163} with further verification of transition states by performing dynamic reaction coordinate calculations to determine the Intrinsic Reaction Coordinates (IRCs). The effects of solvation was calculated using the Conductor-like Screening

Model (COSMO) implemented in TurboMole,¹⁶⁴ which is a continuum solvation model implicitly surrounding the solute molecule. Code was developed to assist with managing the computational jobs (see chapter 5).

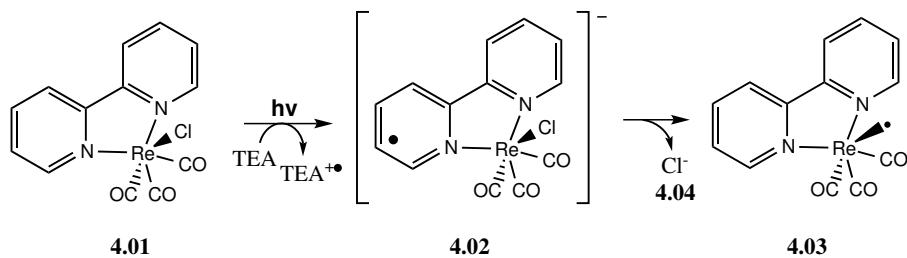
Many of the intermediates have been synthesized in various studies,^{141,165–167} indicating their reasonable stability. While individual portions of the mechanisms have been studied computationally in the past,^{131–133} no overarching study has compared methods relative to each other. Furthermore, while the formation of CO with H₂O is the most anticipated pathway (due to the lack of formation of carbonate or formate in most studies), no literature pathway exists to explain the addition of CO₂ to the open site of the radical catalytic species without a three body reaction step (catalyst, CO₂ and H⁺ together) or without formate reorganization. Furthermore, no mechanism proposed thus far explains the ¹²CO to ¹³CO isotopic exchange demonstrated by Lehn's group in 1986.⁶³

4.2.1 Eximer Formation and Decomposition of the Sacrificial Amine

All of the mechanistic pathways require the formation of a common eximer species, the radical 17e⁻ complex (Scheme 4.2). This occurs through the absorption of an incident photon with enough energy to promote an electron from the metal *d*-orbital to the ligand π^* orbitals of the ground state catalyst, **4.01**, forming the triplet Metal-Ligand Charge Transfer (3-MLCT) complex **4.01**^{3MLCT}. This excitation requires approximately 50 kcal/mol, sourced from absorbed light. The pseudo-oxidized, electron-deficient metal atom extracts an electron from the sacrificial amine present in the reaction solution to return to the Re^I state (**4.02**). However, this complex is formally a radical anion with the lone electron located in the ligand π system, thus halide (**4.04**) is lost to return to the neutral radical eximer species in solution, **4.03**. This extraction to form the radical

anion catalyst and the radical cation amine is thermodynamically expensive in the gas phase, costing over 80 kcal/mol, but in solution phase in DMF it releases 5 kcal/mol. The difference in energies demonstrates the importance of performing calculations in a simulated solution; steps that may have insurmountable energy barriers in the gas phase become possible once solvation is considered.

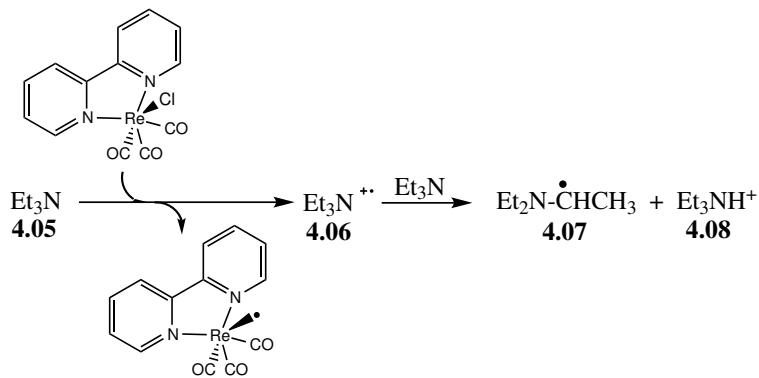
A slight uphill to dissociation of the chloride (15.44 kcal/mol in DMF) allows for the formation of the triplet $17 e^-$ excimer species **4.03**, from which the mechanism pathways discussed in following subsections may diverge. It is important to note that some studies suggest the solvent coordinates with the excimer species.^{85,144} This solvent coordination is expected to stabilize the excimer species in solution prior to reaction with CO_2 .¹⁶⁸ The coordination provides approximately 11 kcal/mol of stabilization (calculated via DFT). However, this event has no bearing on the overall reaction energies, the coordination and subsequent loss of solvent is an energetically neutral occurrence and was not examined in detail as part of this study.



Scheme 4.2 Formation of the eximer species via absorption of a photon and oxidation of the sacrificial amine.

The decomposition route of the sacrificial amine was first identified by Kalyansundaram in 1978,¹⁴⁷ and is summarized in Scheme 4.3. This work showed the decomposition of TEOA but the mechanism for decomposition of TEA is parallel. This decomposition is critical due to the protons it provides to the reaction mixture, and the presence of a

simple second electron abstraction from the decomposition product. Upon absorption of a photon by the catalyst, the amine **4.05** is converted to the radical cationic species ($\text{Et}_3\text{N}^{+\bullet}$, **4.06**). This undergoes a proton transfer to a second molecule of the sacrificial reductant. The transfer removes a proton from the carbon α to the central nitrogen, leaving it a neutral radical species (forming **4.07**). This is then able to react in the catalytic cycle to provide a second electron and form the ethene-diethylamino compound. Triethylammonia is produced as well (**4.08**), this is a proton source for the formate and water-gas shift mechanistic pathways. This step is slightly exothermic, with energy releases of 1 kcal/mol in gas phase, or nearly 3 kcal/mol in DMF.



Scheme 4.3 Decomposition pathway for the sacrificial amine.

Energies of **4.01** - **4.10**, and **4.26** in gas and solution phases (in DMF) is shown in Table 4.1, along with the energy of solvation. Energies of each step of the reaction are listed in Table 4.2. The values show a significant ‘uphill’ series of steps requiring significant energy input. This energy is supplied by the incident photons; this process is a photocatalyzed activation. The reaction typically requires incident light of 400 nm,⁶² this high energy input allows for the excitation pathway to be followed. The energy of each intermediate and the steps are shown in Appendix D, Figure D.9.

Geometries of the catalyst do not change significantly throughout this transforma-

Table 4.1 Gas phase and solvated energies of mechanism reactants and products.

Molecule	Label	E (gas) ^a	E (solution) ^b	E (solvation) ^c
Ground State	4.01	-1374.621419	-1374.651099	18.62
3MLCT Complex	4.01 ^{3MLCT}	-1374.553193	-1374.565998	8.04
Radical Anion	4.02	-1374.684002	-1374.759190	47.18
Open Site Excimer	4.03	-914.3139376	-914.3287245	9.28
Chlorine Anion	4.04	-460.2890817	-460.4058583	73.28
Triethylamine (TEA)	4.05	-292.3051496	-292.3854033	50.36
Radical Cation TEA	4.06	-292.3051496	-292.3854033	50.36
Deprotonated TEA Radical	4.07	-291.9173706	-291.9211226	2.35
Triethylammonia	4.08	-292.9552538	-293.0382729	52.09
Carbon Dioxide	4.09	-188.6945676	-188.6974631	1.82
Carbon Monoxide	4.10	-113.3744946	-113.3754466	0.60
Diethylaminoethene	4.26	-291.3467768	-291.3525868	3.64

^a TPSS SCF energy in hartrees.^b TPSS SCF energy in hartrees with COSMO solvation in DMF.^c TPSS solvation energy in kcal/mol (E(gas) - E(solution)).**Table 4.2** Energies for the reaction steps in the excimer formation pathway

Steps	Energy(gas) ^a	Energy(dmf) ^b
4.01 → 4.01 ^{3MLCT}	42.81	53.40
4.01 ^{3MLCT} + 4.05 → 4.02 + 4.06	81.45	-5.80
4.02 → 4.03 + 4.04	50.82	15.44
4.06 + 4.05 → 4.07 + 4.08	-1.08	-2.92

^a TPSS SCF energy in kcal/mol.^b TPSS SCF energy in kcal/mol with COSMO solvation in DMF.

tion, however, some small changes occur signifying the change in electron localization. One metric analyzed in polyaromatic non-innocent ligand redox reactions is the bonding distance between aromatic rings.¹⁶⁹ From ground state, through triplet MLCT complex to the excited radical, the C-C_(bpy) distance decreases from 1.470 Å to 1.425 Å, ending at 1.416 Å. This 0.06 Å decrease is noted in many previous experiments and calculations for anion radicals.^{169–173} Other key bond lengths and angles include the ligand N-Re bonds and CO-Re bonds, the addition and subtraction of electrons to and from the complex

impact the bonding. As the reaction proceeds, for example, the calculated Re–N distance in solvated structures decreases from 2.18924 Å in the ground state catalyst to 2.14212 Å in the neutral radical excimer. Similarly, the distance from the metal to the axial carbonyl decreases from 1.91950 to 1.88746 Å in the same circumstances. These changes are not significant, the bond order is unchanged. Bond length changes of 0.05 Å are large enough to demonstrate a change has occurred in the electron configuration around the metal centre.

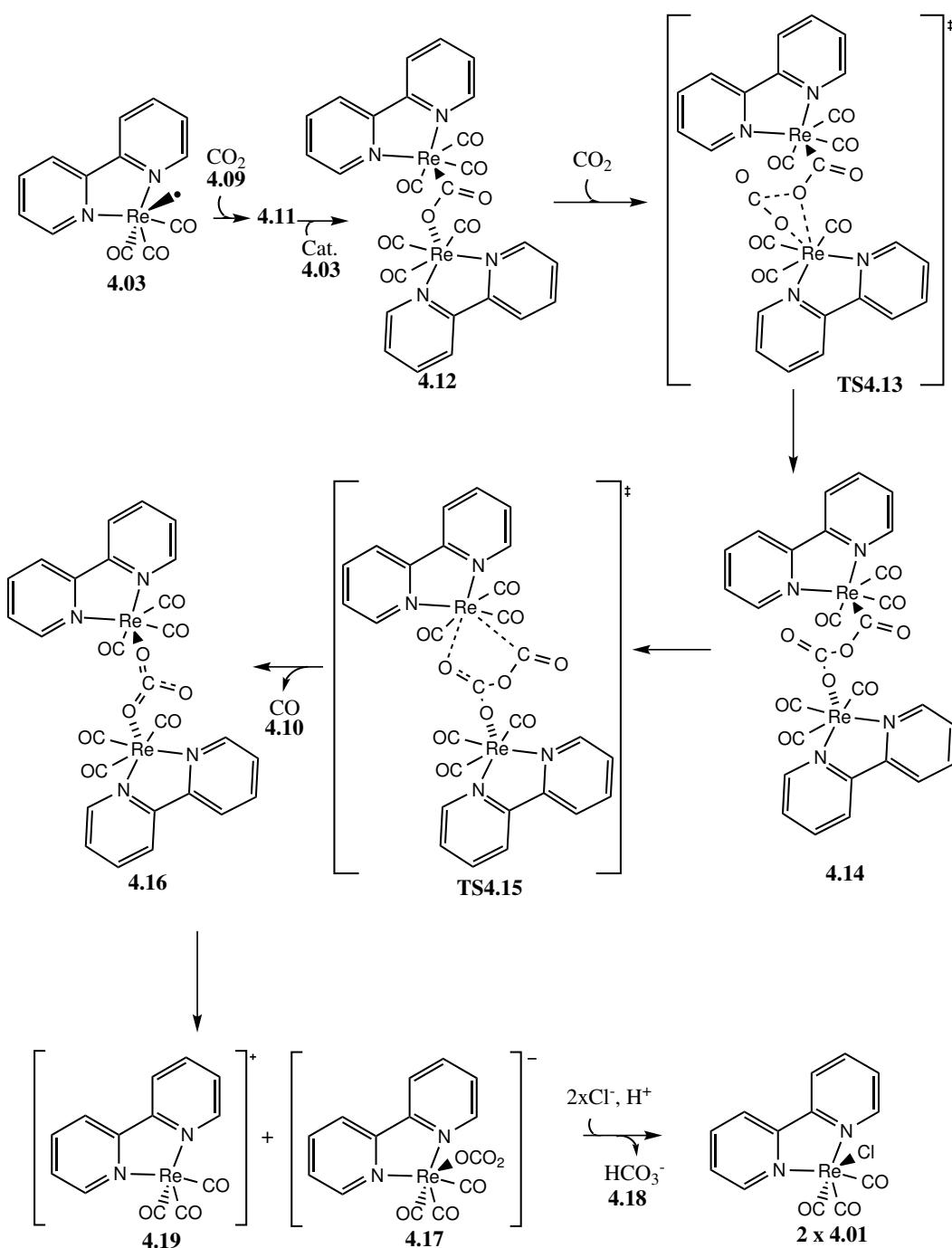
4.2.2 The ‘Carbonate’ Pathway

The carbonate pathway is shown in Scheme 4.4, starting from the excimer species. This pathway has been studied in some detail in the literature, but never as completely as will be shown here. Typical analysis consists of investigation from the formed CO₂ linked dimer, through the release of CO, terminating at the bicarbonate linked dimer. Studies typically build the dimer as a three-body reaction, or start with a Re–Re bound catalyst dimer and the insertion of CO₂. However, the formation of the [L₂Re(CO)₃]₂ is exceptionally slow in the presence of solvent, with a rate constant 8 orders of magnitude below the solvent-stabilized radical •L₂Re(CO)₃(solv) complex,¹⁶⁸ and the [L₂Re(CO)₃]₂ is considered completely unreactive to CO₂.¹³⁰

DFT energies of each of the compounds involved in this mechanism pathway are shown in Table 4.3, along with the energy of solvation. Table 4.4 shows the potential energy change in each step of the reaction. The computed intermediate and transition state structures are shown in Figure 4.1.

The mechanism begins with the addition of a CO₂ molecule to the excimer **4.03**, forming **4.11**. This is a very weakly bound species when solved in a simulated DMF environment; in the gas phase this transition complex will not solve. Energies for the

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Scheme 4.4 The 'carbonate' mechanistic pathway

Table 4.3 Gas phase and solvated energies of compounds, transition states and intermediates in the ‘carbonate’ mechanism

Molecule	Label	E (gas) ^a	E (solution) ^b	E (solvation) ^c
Catalyst - CO ₂	4.11	-1103.009	-1103.016	4.47
CO ₂ Linked Dimer	4.12	-2017.373	-2017.412	24.59
CO ₂ Addition	TS4.13	-2206.121	-2206.165	27.80
C ₂ O ₄ Linked Dimer	4.14	-2206.048	-2206.097	31.23
5 Member Ringed Dimer	TS4.15	-2206.014	-2206.062	29.87
CO ₃ Linked Dimer	4.16	-2092.678	-2092.726	29.75
Bicarbonate Catalyst Cation	4.17	-1178.065	-1178.120	34.24
Bicarbonate Anion	4.18	-264.485	-264.497	7.20
Open Site Cation	4.19	-914.106	-914.187	50.75

^a TPSS SCF energy in hartrees.^b TPSS SCF energy in hartrees with COSMO solvation in DMF.^c TPSS solvation energy in kcal/mol (E(gas) - E(solution)).**Table 4.4** Energies for the reaction steps in the ‘carbonate’ pathway

Steps	Energy(gas) ^a	Energy(dmf) ^b
4.03 + 4.09 → 4.11	-0.25	6.37
4.11 + 4.03 → 4.12	-31.56	-42.39
4.12 + 4.09 → TS4.13	-33.27	-34.66
TS4.13 → 4.14	45.91	42.47
4.14 → TS4.15	21.10	22.46
TS4.15 → 4.16 + 4.10	-24.36	-24.84
4.16 → 4.17 + TS4.19	317.95	262.71
4.17 + TS4.19 + 4.04(×2) → 4.01(×2) + 4.18	-180.55	-239.63

^a TPSS SCF energy in kcal/mol.^b TPSS SCF energy in kcal/mol with COSMO solvation in DMF.

gas phase for this compound are calculated as single point energies from the solvated structure. The DMF solved structure has a Re-C bond length of 2.50654 Å, and O-C-O bonding angle of 142°, when compared to the Re–C distances of rhenium carbonyls of *ca.* 1.9 Å, this is a very weak bond. The formation of this complex requires only 6.37 kcal/mol, the radical species is not satisfied with the addition of CO₂ and requires further electron contribution to become more stable. This unstable complex is able to

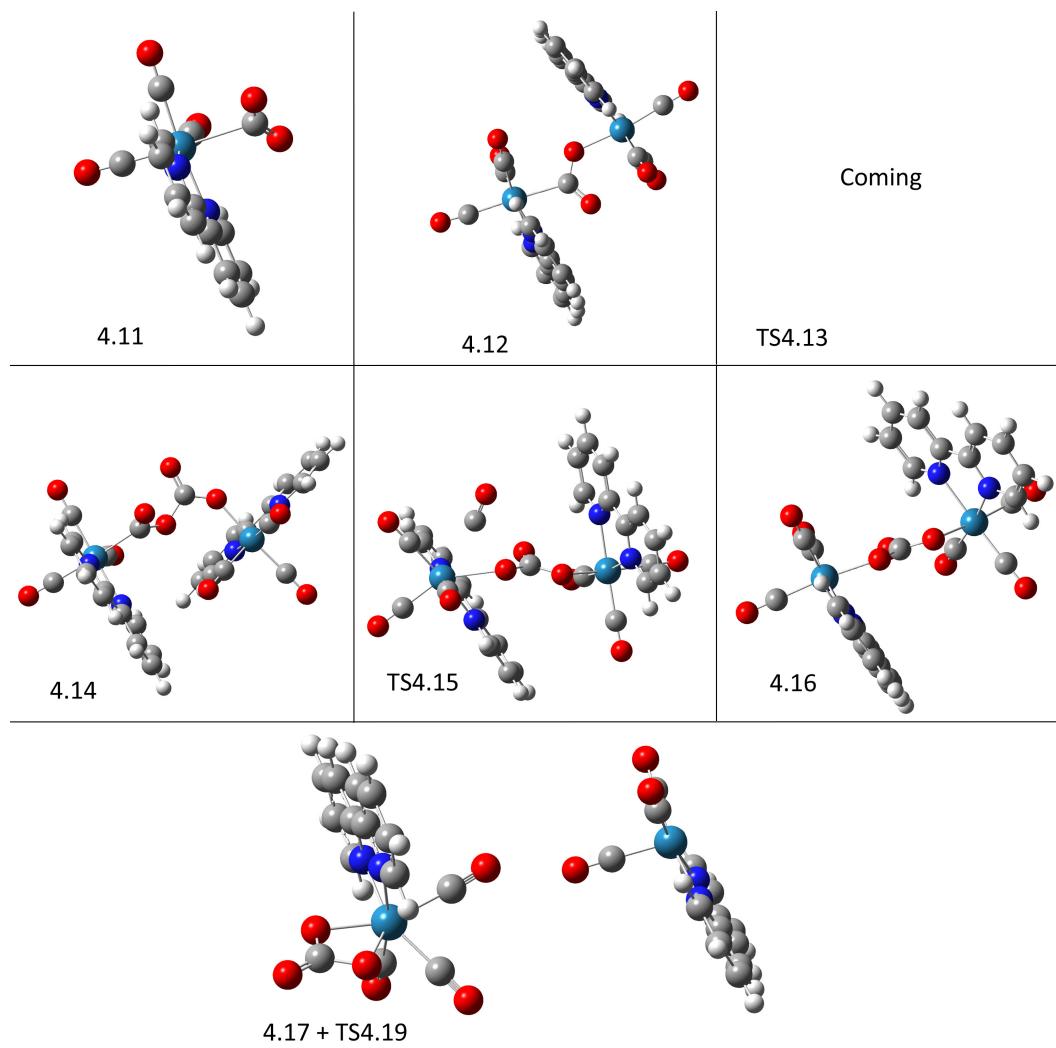


Figure 4.1 DFT calculated structures for the ‘carbonate’ mechanistic pathway

extract a hydrogen from triethylammonium (Et_3NH^+ , **4.08**) to continue with the ‘water-gas’ pathway (see below subsection 4.2.4), or combine with a second molecule of the excimer to form a dimer **4.12**. This dimer formation is favourable; resolution of radical species is provided. The addition of the second catalyst releases 42.5 kcal/mol in DMF. The quenching of the radical forms much stronger bonds between the metal atoms and the linking CO_2 . The Re-C distance has shortened from the 2.50654 Å seen in **4.11** to

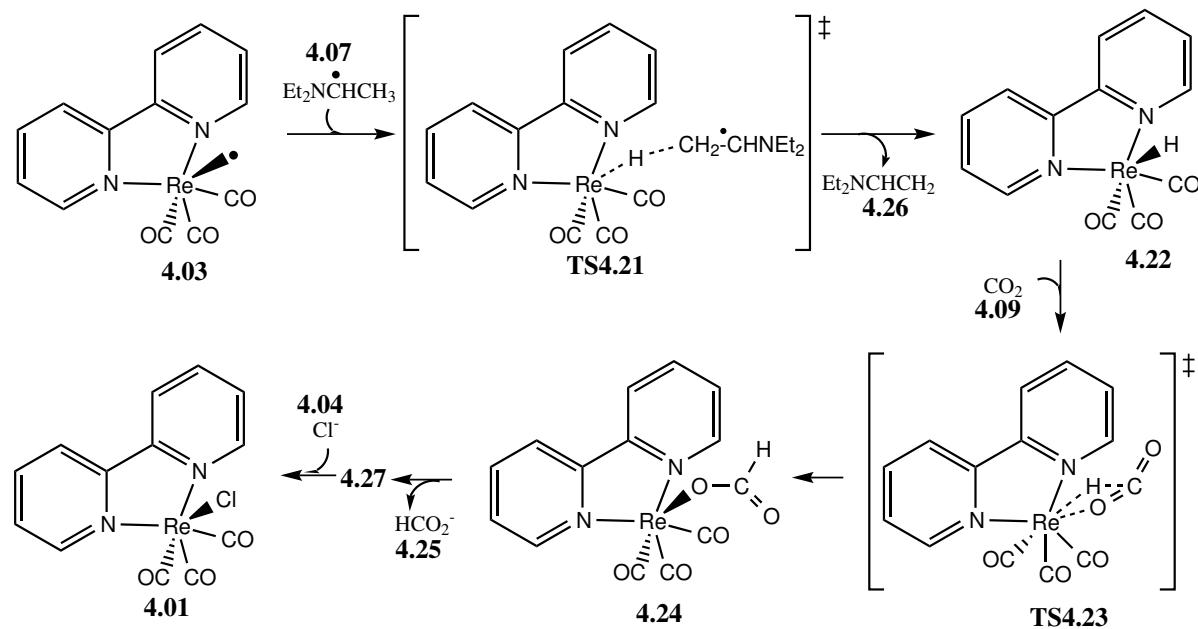
2.25829 Å. This is still longer than the Re-CO bonds, which is expected due to the lack of π back-bonding observed with carbonyl ligands, however, it corresponds with similar published crystal structures of Re-C bond lengths for sp² carbons.¹⁷⁴ The Re-O bond is 2.13 Å, a value that remains constant for Re-O through the intermediates in the reaction pathway.

After the dimer **4.12** has been formed, a second molecule of CO₂ is inserted via transition state **4.13** to form the Re-C(O)-O-C(O)-O-Re complex **4.14**. This linker contains bonds from length 1.28 - 1.44 Å, longer than expected by about 0.1 Å for sp²-carbon oxygen single or double bonds.¹⁷⁵ Bond angles are typically just under the idealized 120° expected as well. Due to the linker, the two catalyst ligand bipyridines have moved from a nearly co-planar geometry to a nearly perpendicular geometry. Re-C and Re-O bonds remain constant in length compared to **4.12**. The formation of a 5 membered ring transition state structure **4.15** costs 22 kcal/mol. This leads to the release of the CO and the formation of a carbonate linked dimer, **4.16**, resulting in a net energy decrease of -2.4 kcal/mol, and returning the catalyst ligands to a more co-planar orientation. The carbonate dimer species is left to decompose to a catalyst cation **4.19** with an open site, and the bicarbonate adduct **4.17**. This carbonate dianion may pick up a proton before or after the disassociation to the catalyst cationic species, resulting in the released of the bicarbonate species to solution (**4.18**) when the catalyst is returned to ground state **4.01** with addition of a chloride. The decomposition of the bicarbonate dimer species has a barrier of 262 kcal/mol in DMF. This is due to the charge separations that occur, while solvation stabilizes these charged species it does not accurately simulate a system with an ionic strength similar to what could be expected in the experimental reaction. The Gibbs energy for this step was calculated to be -11.4 kcal/mol in solution, with an additional -17 kcal/mol for the anion exchange, signifying the thermodynamic favourableness of

this transformation. The presence of excess molar equivalents of an electrolyte such as tetraethylammoniumchloride ensures a surplus of chloride anions are in solution and are an additional force for conversion, according to Le Châtelier's principle.

4.2.3 The ‘Formate’ Pathway

In comparison to the catalytic dimer formed in the carbonate pathway above, the formate formation occurs via a much simpler mechanism. The addition of a hydrogen to the open site axial to the ligand occurs via the simultaneous electron and proton transfer from a by-product of the reduction of the amine. CO_2 inserts into this metal hydride bond, resulting in a metal–oxide bond, with the formate anion.¹⁷⁶ Separation of the weak metal–oxygen bond allows for the reinsertion of the halide to the cationic metal centre.



Scheme 4.5 The ‘formate’ mechanistic pathway

DFT energies of each of the compounds involved in this mechanism pathway are shown in Table 4.5, along with the energy of solvation. Table 4.6 shows the potential

energy change in each step of the reaction. The computed intermediate and transition state structures are shown in Figure 4.2.

Table 4.5 Gas phase and solvated energies of compounds, transition states and intermediates in the ‘formate’ mechanism

Molecule	Label	E (gas) ^a	E (solution) ^b	E (solvation) ^c
Proton Transfer	TS4.21	-1206.302	-1206.327	15.10
Catalyst Hydride	4.22	-914.920	-914.945	15.29
CO ₂ Insertion	TS4.23	-1103.581	-1103.620	24.32
Catalyst Formate	4.24	-1103.635	-1103.665	19.04
Formate Anion	4.25	-189.305	-189.415	69.01
Open Site Cation	TS4.27	-914.106	-914.187	50.75

^a TPSS SCF energy in hartrees.

^b TPSS SCF energy in hartrees with COSMO solvation in DMF.

^c TPSS solvation energy in kcal/mol (E(gas) - E(solution)).

Table 4.6 Energies for the reaction steps in the ‘formate’ pathway

Steps	Energy(gas) ^a	Energy(dmf) ^b
4.03 + 4.07 → TS4.21	-44.98	-48.46
TS4.21 → 4.22 + 4.26	22.43	18.60
4.22 + 4.09 → TS4.23	21.24	14.02
TS4.23 → 4.24	-33.94	-28.66
4.24 → 4.25 + TS4.27	140.39	39.67
TS4.27 + 4.04 → 4.01	-141.77	-36.37

^a TPSS SCF energy in kcal/mol.

^b TPSS SCF energy in kcal/mol with COSMO solvation in DMF.

After formation of the excimer **4.03**, the radical species extracts a hydrogen atom from the oxidized chain of the sacrificial amine **4.07** in transition state **4.21**, a step releasing 45 kcal/mol in the gas phase or 48.5 kcal/mol in DMF. This is due to the satisfaction of the radical states and removal of the charge on the amine. The sacrificial amine involved in this step had previously had one proton extracted by another molecule of the amine in a proton exchange step (see subsection 4.2.1), resulting in a neutral

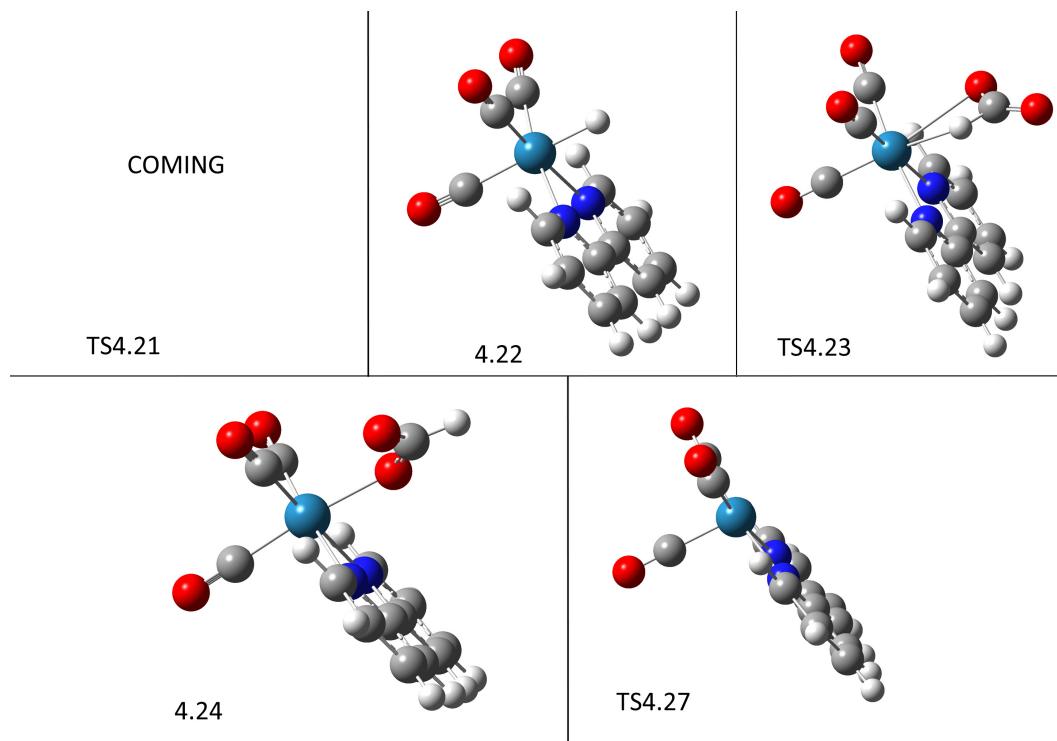


Figure 4.2 DFT calculated structures for the ‘formate’ mechanistic pathway

radical. Extraction of the proton and electron pair allows for the formation of the ethene, completing the decomposition of the amine to the final neutral, singlet molecule **4.26**. Relaxation of this transition state results in the hydrogen extraction from the radical species, yielding the formation of the hydride complex **4.22**.

This hydride complex is able to insert a molecule of CO_2 into the metal-hydrogen bond, in transition step **4.23**, with only a 14.4 kcal/mol barrier to transition state formation.. CO_2 insertion to metal hydrides is commonly observed, most mechanisms in CO_2 reduction in similar ruthenium systems employ this addition.¹⁷⁶ The Re-H bond length is 1.76 Å, compared to the length of the Re-Cl bond from the ground state **4.01** species of 2.51 Å. This bond length difference reflects the observations on anion change from Cl to Br (as discussed in chapter 2, subsection 2.3.2), the anion size is the critical factor in this variation. When a molecule of CO_2 approaches, the transition state of a

pseudo-septacoordinate species **4.23** forms. The Re-H bond increases in length to 2.16 Å, the Re-O bond is 2.88 Å, and the O-Re-H angle is very tight, at only 45.3°. This step completes with 28.66 kcal/mol energy release to form the formato anion complex **4.24**.

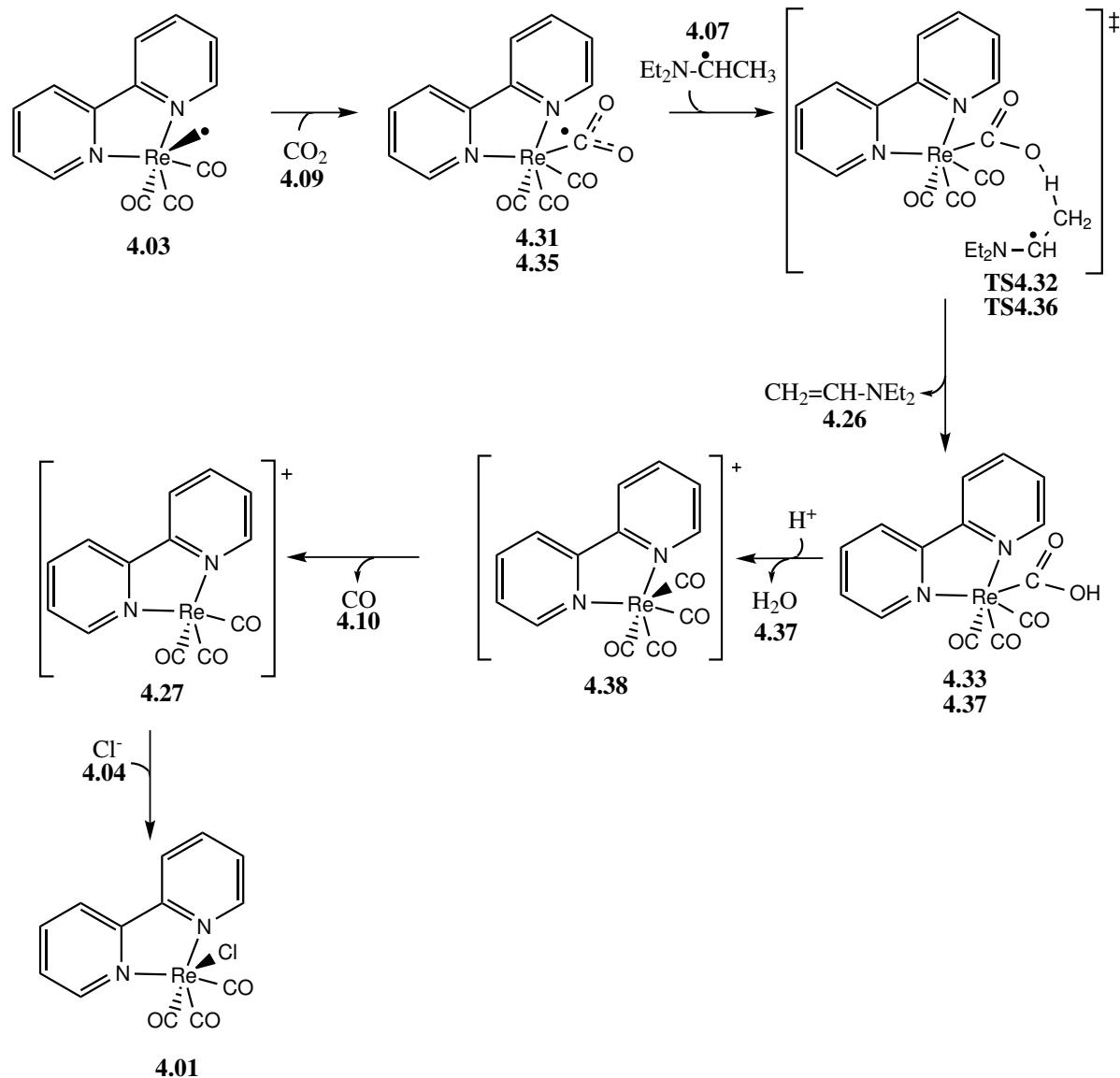
The formato anion **4.24** contains a Re-O bond of 2.15 Å, consistent with previously discussed rhenium - oxygen bonds. This formate dissociates with a chloride addition, the exchange is endothermic by less than 3 kcal/mol overall, but has a barrier of nearly 40 kcal/mol in DMF (or 140 kcal/mol in gas phase) calculations. As discussed with the carbonate mechanism, this large barrier is due to the charge separation that occurs. The calculated Gibbs energy for the formation of the charge separated transition state is -12.7 kcal/mol in solvent, showing the thermodynamic favourableness of this exchange. Further reactions of the formato anion in solution are not investigated, but the anion may remain deprotonated in the slightly basic environment.¹³⁸

Some attempts were made at performing the reaction along alternative pathways. While direct σ bonding from a metal to an oxygen atom in the CO₂ molecule (as η^1 -OCO) has been observed in a few systems (including photoreduction of CO₂ on other metal systems),^{177–179} this geometry is rare.^{180–182} Attempts to coordinate CO₂ in an η^1 -OCO geometry failed to converge both in gas and solution phase, as CO₂ was ejected from the complex. Binding of CO₂ to the metal through π coordination of the C=O bond is more common,^{181,182} but these structures failed to solve in the current DFT system as well, typically via the dissociation of the CO₂.

4.2.4 The ‘Water-Gas Shift’ Pathway

The water-gas shift mechanism involves the addition of two protons from the reductant to a CO₂ molecule bound to the metal centre. The first proton addition yields an acid species, this is dehydrated via the second addition of a proton and the release of one

molecule of H_2O . The resulting tetracarbonyl cationic species is able then to release an axial carbonyl to return to the ground state. While any of the carbonyl groups could be labile, the carbonyl at the axial position is the one actively replaced by the halide to return to the starting catalyst.¹⁶⁵



Scheme 4.6 The ‘water-gas shift’ mechanistic pathway

DFT energies of each of the compounds involved in this mechanism pathway are

shown in Table 4.7, along with the energy of solvation. Table 4.8 shows the potential energy change in each step of the reaction. The computed intermediate and transition state structures are shown in Figure 4.3.

Table 4.7 Gas phase and solvated energies of compounds, transition states and intermediates in the ‘water-gas shift’ mechanism

Molecule	Label	E (gas) ^a	E (solution) ^b	E (solvation) ^c
Catalyst-CO ₂ (Axial)	4.31	-1103.009	-1103.016	4.47
H Transfer to Axial CO ₂	TS4.32	-1394.981	-1395.011	18.86
Catalyst-CO ₂ H (Axial)	4.33	-1103.610	-1103.640	18.90
Catalyst with Migrated Open Site	TS4.34	-914.277	-914.297	12.88
Catalyst-CO ₂ (Equatorial)	4.35	-1102.964	-1102.992	17.92
H Transfer to Equatorial CO ₂	TS4.36	-1394.939	-1394.992	33.05
Catalyst-CO ₂ H (Equatorial)	4.37	-1103.597	-1103.626	17.68
Tetracarbonyl Catalyst Cation	4.38	-1027.546	-1027.619	46.02
Water	4.39	-76.464	-76.476	7.33

^a TPSS SCF energy in hartrees.

^b TPSS SCF energy in hartrees with COSMO solvation in DMF.

^c TPSS solvation energy in kcal/mol (E(gas) - E(solution)).

Table 4.8 Energies for the reaction steps in the ‘water-gas shift’ mechanism

Steps	Energy(gas) ^a	Energy(dmf) ^b
4.03 + 4.09 → 4.31	-0.25	6.37
4.31 + 4.07 → TS4.32	-34.68	-46.72
TS4.32 → 4.33 + 4.26	15.33	11.64
4.33 → 4.38 + 4.39	-6.51	8.88
4.38 → 4.27 + 4.10	40.89	35.57
4.27 + 4.04 → 4.01	-141.77	-36.37

^a TPSS SCF energy in kcal/mol.

^b TPSS SCF energy in kcal/mol with COSMO solvation in DMF.

This mechanistic pathway it thought to start by the same addition of CO₂ that is seen in the carbonate mechanism (see subsection 4.2.2), forming **4.11**. As before, the complex is only weakly coordinated, and requires solvation effects to solve computationally. The

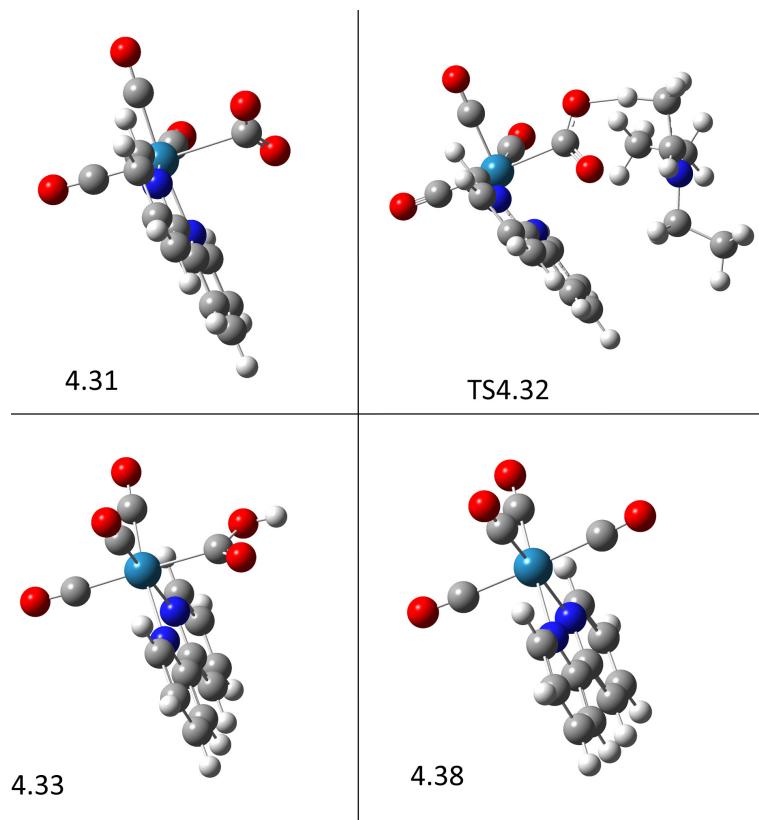
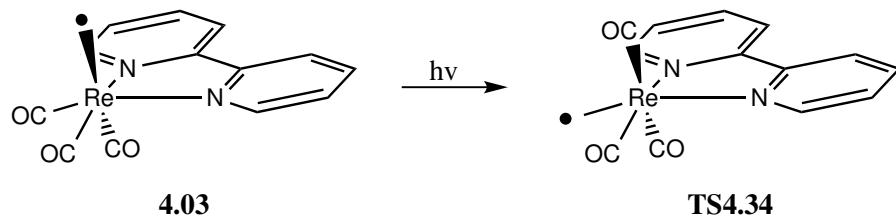


Figure 4.3 DFT calculated structures for the axial ‘water-gas shift’ mechanistic pathway

added CO_2 is able to extract a hydrogen from the previously-reduced sacrificial amine **4.07** with a net energy change of -35.1 kcal/mol in DMF, again allowing the formation of the ethene amine **4.26**. The newly formed acid species **4.32** dehydrates in the presence of a second proton (via **4.33**) to form water **4.37** and the tetracarbonyl cationic species **4.38**. This has an energy cost of 9 kcal/mol in solution, but the ΔG is -15 kcal/mol. This tetracarbonyl cation exchanges a CO molecule for a Cl^- to return to **4.01**, with a 0.8 kcal/mol energy release.

Typically, this reaction had been thought to proceed on the axial site of the catalyst, mirroring the pathways discussed above. However, due to the ease of migration of carbonyl groups in organometallic complexes, it is proposed that the ‘water-gas shift’

mechanism does not occur axial to the ligand, but begins with relocation of a CO₂ ligand to the axial position (see Scheme 4.7), followed by the coordination of CO₂ into the now vacant planar open site, forming **4.35** (energies for this reaction are shown in Table 4.9, and DFT computed structures are shown in Figure 4.4). This CO₂ bound in the plane of the ligand then undergoes hydrogen addition and dehydration to produce a molecule of H₂O, continuing as before. Once reduced to the tetracarbonyl cluster, the catalyst could shed any CO ligand and pick up a chloride anion to return to the neutral ground state. However, labilization of the axial carbonyl is the most favoured, formation of the planar coordinated chloride complex is not expected¹⁶⁵



Scheme 4.7 Rearrangement of carbonyl and open site

Table 4.9 Energies for the reaction steps in the ‘equatorial’ geometry

Steps	Energy(gas) ^a	Energy(dmf) ^b
4.03 → TS4.34	23.37	19.77
TS4.34 + 4.09 → 4.35	4.79	1.56
4.35 + 4.07 → TS4.36	-36.36	-49.13
TS4.36 → 4.37 + 4.26	-3.39	8.34
4.37 → 4.38 + 4.39	-14.52	-0.35
4.38 → 4.27 + 4.10	40.89	35.57
4.27 + 4.04 → 4.01	-141.77	-36.37

^a TPSS SCF energy in kcal/mol.

^b TPSS SCF energy in kcal/mol with COSMO solvation in DMF.

This newly proposed mechanism provides a solution to a previously unexplained

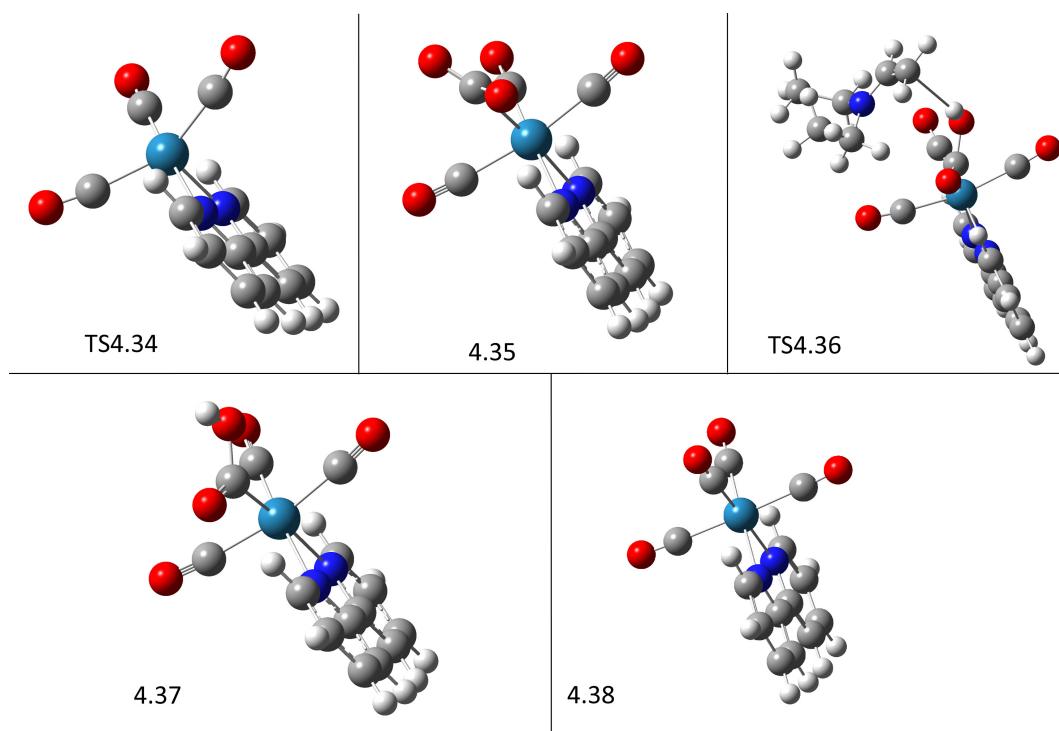


Figure 4.4 DFT calculated structures for the equatorial ‘water-gas shift’ mechanistic pathway

phenomenon. The exchange of carbonyl groups on the catalyst for ^{13}CO when using $^{13}\text{CO}_2$ in the photoreduction is documented as early as Hawecker *et al.* in 1986.⁶³ It was shown that complete exchange occurs with very few catalytic turnovers. Furthermore, Koike *et al.* demonstrated that photochemical ligand substitution occurs at only axial sites relative to the α -imino ligand,¹⁴⁰ no exchange occurs at the equatorial site, nor can the *fac*-(^{12}CO)₂ ^{13}CO be expected to shift the ^{13}CO to the equatorial position in the time-frame of the reaction. Thus the isotopic exchange does not proceed by independent uptake of produced ^{13}CO , but in fact the conversion to the ^{13}CO complex must be inherent to the reduction mechanism.

This modified geometry does not violate any previously published experimental mechanism work. These studies typically focus on photophysical or spectroscopic analysis

to determine intermediates, the analysis methods are not rigid geometry descriptors. Reevaluation of the spectral data within these newly defined parameters does not invalidate the work, mechanistic descriptors published (such as kinetics) are still valid within the new geometry.

Additionally, this provides an explanation for the lack of CO₂ reduction by (*terpy*)- κ^2 -Re(CO)₃Cl discussed in chapter 3. While the carbonyl group is transitioning to the axial site, before a molecule of CO₂ has been coordinated to the complex, the catalyst is primed for chelation of the pendant arm (see Scheme 3.1. It is known that the tridentate catalyst is inactive for CO₂ photoreduction, the formation of a κ^3 complex modifies the photophysics, deactivating the catalyst's reactivity.

4.2.5 Consequences From κ^2 Terpyridine Complex Inactivity

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The lack of reactivity of the κ^2 (*terpy*)Re(CO)₃X motif of complexes contrasting to the activity of the originally published κ^2 (bipy)Re(CO)₃X indicates significant influence of the ligand on the mechanism. While the terdentate complex can be rationalized to be inactive due to its short-lived excited state (as seen in chapter 2),¹⁸³ this explanation does not suffice for the fluorescing bidentate complex. Other substituted bipyridine ligands are known to be active for photocatalytic reduction,^{63,118} identifying the most likely conflicting feature of the terpyridine ligand to be the pendant arm, and its availability for chelation to the metal centre. While in the radical eximer form, the chelation site is sterically blocked by one of the three carbonyl groups. However, reorganization of the substituent carbonyls from a *facial* orientation to a *meridional* could allow for the free pyridine to form the metal-ligand bond, resulting in compound **3.1** (see chapter 3, subsection 3.2.3).

4.3 Comparison Between Mechanistic Pathways

Darrin, I'm still working on this section (discussed this morning)

Previous studies in literature had only analyzed one of the mechanistic pathways (or a subsection thereof), without a fuller analysis of the competitiveness of each pathway relative to the others. Discussion on the tenability of each potential pathway relied on the *in situ* observation of intermediates or transition states, the success (or lack thereof) of synthesis of the intermediates, and the relative production of by-products in the mechanistic trials.

The overall potential energy diagrams for each of the mechanistic pathways in Scheme 4.1 are shown in Figure 4.5. Individual mechanistic pathways, and separate potential energy diagrams for gas phase and DMF solvated structures are listed in Appendix D. This figure shows the difficulty in determining the most likely reaction pathway based on computational analysis alone. The difference in energy consumption between mechanisms is quite small, with the largest barriers belonging to the final disassociation of the product and coordination of chloride.

Experimental data shows that each of the mechanisms is feasible and could occur in certain conditions. Both the formate and water-gas routes have the maximum transition energy barrier for the final disassociation of the product and coordination of chloride. This barrier is artificially high, the charge separation that occurs is impacted significantly by the basic pH and the presence of electrolyte in solution experimentally. These factors are not taken into consideration in the calculations. Gibbs energies were calculated for these steps, the average ΔG values are around -12 kcal/mol for the production of the transition state, signifying a thermodynamically favoured reaction.

Due to the similarity in potential energy profiles for these reactions, analysis of experimental data must be undertaken to accurately predict which pathway dominates

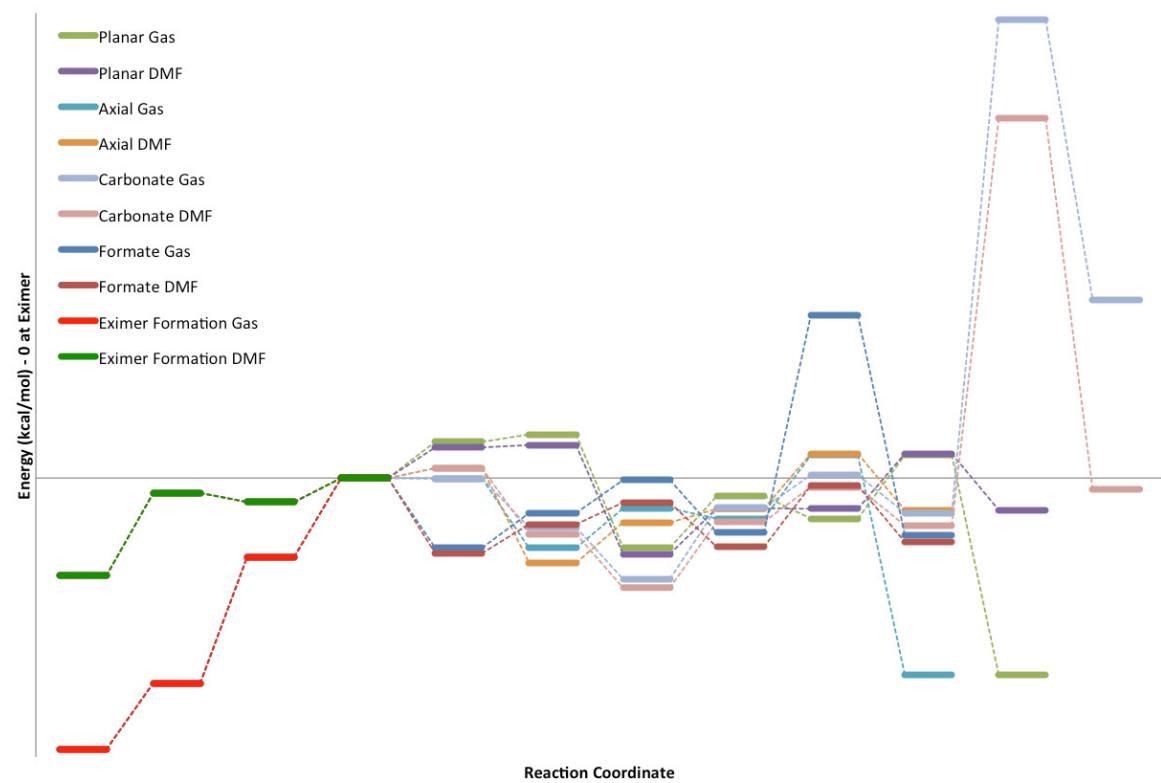


Figure 4.5 An overview of the energies of the three mechanistic pathways of photochemical CO_2 reduction

the photoreduction of CO₂. The energy requirements of each pathway is not the only indicator of mechanistic preference. Studies have shown the rate differs between pathways, with the formation of certain intermediates being so slow as to nearly block the progression of the redox reaction. Utilization of modified reaction conditions (TEA vs. TEOA, changing solvent mixture, presence or absence of H₂O, presence or absence or modification of electrolyte) can provide a range of reduction products, by-products, and intermediates.^{62,63,66,85,130,138}

4.4 Conclusions

Three likely mechanistic pathways for the photocatalytic reduction of CO₂ were investigated in detail using consistent DFT methods, allowing for their comparison. Mechanistic steps not previously elucidated were demonstrated, showing the full mechanistic cycle from ground state excitation through the reaction of the eximer with CO₂ to the final release of product and reforming of catalyst.

A modified geometry for the production of CO and H₂O by the ‘water-gas’ mechanism was proposed. This new geometry does not conflict with previously published experimental data, and provides explanation for previously reported phenomenon. It aligns with the previously demonstrated inactivity of terpyridine catalysts, in contrast to the highly active bipyridine analogue.

Chapter 5

TurboControl

The analysis of the mechanism using computational methods required a significant amount of manual set-up and analysis work of computational input and output files. The calculations required 38 different molecules, intermediates, or transition states to be calculated in multiple different environments. These calculations typically require intervention or set-up of intermediate steps to be able to fully elucidate all of the required information, for example, the set-up of frequency calculations after a geometry optimization to ensure ground state or transition state geometries.

Additionally, while TurboMole contains much faster optimization code for DFT calculations, the user interface for this software requires a significant learning curve and is based in a user-abrasive interactive text prompt environment. This contrasts to Gaussian 09, which is significantly more popular, partially due to the user-friendliness of the GaussView Graphical User Interface (GUI), despite concerns about the speed of optimization of large or complex molecules.

These two factors prompted the development of a new program, TurboControl, written in Python, with the goal of combining the user friendliness of GaussView on top of the optimization efficiency of TurboMole. This program allows a user to prepare the input

files in GaussView, modify them slightly in any text editor available to them, then run a batch calculation on the files. TurboControl monitors the computational jobs, resubmitting for frequency calculations when optimization is complete if requested, and using the TurboMole tools available when required to ensure ground state or first transition state geometries are discovered. TurboControl then outputs statistics files, providing the energy and first frequency vibration of each molecule in a single output file. Additional commands can instruct TurboControl to run in a solution, do single-point energy calculations, or gather more in-depth thermodynamic information from the outputs using TurboMole’s tools.

TurboControl provides this hands-off management of jobs allowing the user to spend more time in data analysis, experimental work, or to be able to produce data at a significantly higher rate compared to the manual setup and monitoring of jobs. TurboControl also contains a single job initiation script entitled TurboGo to allow for the use of the input files without the overhead of the job managing.

5.1 Development

TurboControl began as a script for the simple setup and submission of jobs in TurboMole from Gaussian input files. The jobs were submitted to the ‘wooki’ computing cluster available at the University of Ottawa. This computing cluster consists of approximately 1000 CPU cores, running the CentOS distribution of Linux available in the Rocks software distribution. The GridEngine queuing system ensures fair usage of computing cores for each user, and maximizes throughput by managing computational jobs.

TurboControl initially submitted jobs and monitored their successful completion or failure. Development quickly expanded to include automatic resubmission for frequency analysis of geometry optimization jobs, required to determine the success of finding an

energy minima (stationary point) instead of a transition state. As the complexity of the mechanistic study (see chapter 4) increased to include analysis of thermodynamic data, calculation of transition states, calculation in solution phases, and more, the capabilities of TurboControl increased to meet these demands.

Turbocontrol is now able to handle large batches of input, with jobs of varying complexity and monitor them through successful completion or failure of TurboMole. Analysis of the computational jobs is performed to simply highlight DFT energies, first computed normal mode frequencies, and thermodynamic properties (if requested).

5.2 Usage

TurboControl is available freely on the internet, and is a fully open-source project.¹⁸⁴ The code may be downloaded by anyone and used without prior permission for personal, academic, or commercial purposes, provided that the original software license remains with the code. TurboControl works only with the stated versions of software in the ‘readme’ file available with the code on the internet (and included in Appendix E). The software requires no installation prior to use, and has very few dependencies beyond the typically installed packages on a unix or linux system.

TurboControl and TurboGo are not quantum mechanical packages themselves, they require a properly licensed installation of TurboMole to perform calculations. Additionally, while GaussView may be used to prepare the input files, modification of the files by hand in a text editor is required to access most features, GaussView is not required for the use of TurboControl.

Chapter 6

Conclusions

The target Re^I terdentate terpyridine compounds were successfully synthesized, characterized, and tested for photocatalytic CO₂ reduction. The catalysts show no activity for the reduction, in contrast to the known excellent bipyridine compounds.

The reaction mechanisms were studied successfully with DFT methods, resulting in the proposed new geometry for the production of CO with no carbonate or formate anions. This new geometry does not conflict with known experimental studies, yet avoids three-body mechanistic steps.

expand

Appendix A

Experimental Procedures

Experimental synthesis and characterization data for the compounds discussed in this thesis are shown below by compound number:

A.1 General Methods

Synthesis reactions were set-up in a glovebox under a nitrogen atmosphere and performed under an inert atmosphere. Solvents were sparged with nitrogen and then dried by passage through a column of activated alumina using an apparatus purchased from Anhydrous Engineering. Deuterated chloroform and deuterated acetonitrile was dried using activated molecular sieves. Rhenium starting materials were purchased from Strem Chemicals and used as received. All other chemicals were purchased from Aldrich and used without further purification. NMR spectra were run on Bruker Avance 400MHz spectrometers with CD_3CN or CDCl_3 as solvent and internal standard. Elemental analyses were performed by Midwest Microlab LLC, Indianapolis IN. Solid state reactions were carried out in a Lindberg Blue M Mini-Mite Tube Furnace (model TF55035A-1). Infrared spectra were collected using an Agilent Technologies Cary FT-IR spectrome-

ter using a diamond ATR attachment. UV-Vis spectra were collected using a Agilent Technologies Cary 5000 UV-Vis spectrometer. TGA was performed on a TA Q5000 IR instrument: approximately 10-15 mg of each sample was placed in a ceramic sample pan which was heated at a rate of 5 °C/min up to 150 °C, followed by a rate of 2 °C/min to 300 °C while being purged with N₂ at a flow rate of 25 mL/min. GC was performed using a HP gas chromatograph with a 15 m CARBONPLOT column with 0.320 mm inner diameter and 1.50 μ film in a 40 °C oven. The instrument is fitted with a TCD at 220 °C.

A.2 Computational Methods

For the UV-Vis and experimental correlation study, the structures of all species were optimized using Gaussian 09¹¹⁹ employing the B3LYP^{120,121} exchange-correlation (XC) functional. The LanL2DZ basis set/effective core potential¹²² was used on Re and, the all-electron TZVP basis set¹²³ for the remaining lighter atoms. Frequency analysis of all structures was used to confirm the nature of the stationary points. Solvent effects were computed using the integral equation formalism variant of the PCM solvation model within Gaussian 09 for both the ground state and excited state TD-DFT calculations with DMSO as the solvent.^{124,125} The UV-Vis absorption spectra were extracted using the Chemissian software.¹⁸⁵ In these calculations, a pseudo-Voigt band shape was employed with a default average band width at half-height of 2000cm⁻¹.

For the mechanism study, the ground and transition state structures and energies of all species were obtained by using TurboMole 6.5 software^{148,149} with the TPSS meta-GGA XC functional.¹⁵⁰ The def2-TZVP basis set was used for all atoms.^{123,151} The TurboMole program contains a number of optimizations to the original DFT algorithms,¹⁵²⁻¹⁶⁰ decreasing the calculation time without compromising accuracy. Grimme's dispersion cor-

rection (version 3) was included in the calculations.¹⁶¹ Intermediates and transition states were verified by frequency analysis.^{157,162,163} The effects of solvation was calculated using the Conductor-like Screening Model (COSMO) implemented in TurboMole,¹⁶⁴ which is a continuum solvation model implicitly surrounding the solute molecule.

A.3 X-ray Crystallography

Crystals were mounted on thin glass fibers using paraffin oil. Prior to data collection crystals were cooled to 200.15K. Data were collected on a Bruker AXS SMART single crystal diffractometer equipped with a sealed Mo tube source (wavelength 0.71073 Å) APEX II CCD detector. Raw data collection and processing were performed with APEX II software package from BRUKER AXS53. Diffraction data for sample **3** was collected with a sequence of 0.5° ω scans at 0, 120, and 240° in ϕ . Due to lower unit cell symmetry in order to ensure adequate data redundancy, diffraction data for **1**, **2** and **8** were collected with a sequence of 0.5° ω scans at 0, 90, 180 and 270° in ϕ . Initial unit cell parameters were determined from 60 data frames with 0.3° ω scan each collected at the different sections of the Ewald sphere. Semi-empirical absorption corrections based on equivalent reflections were applied.¹⁸⁶ Systematic absences in the diffraction data-set and unit-cell parameters were consistent with triclinic P $\bar{1}$ ($\mathcal{N}^{\mathcal{O}}2$) for compounds **1**, **2** and **8**, monoclinic C2/c ($\mathcal{N}^{\mathcal{O}}15$) for compound **3**. Solutions in the centrosymmetric space groups for all compounds yielded chemically reasonable and computationally stable results of refinement. The structures were solved by direct methods, completed with difference Fourier synthesis, and refined with full-matrix least-squares procedures based on F^2 .

Solutions for **1** and **2** revealed that both these structures contain two compound molecules per asymmetric unit.

Initial refinement results for the compound **1** suggested presence of two non-merohedrally twinned domains. Two independent orientation matrices were found using CELL_NOW software.¹⁸⁷ Data set was re-integrated with two independent orientation matrices and consecutive model refinement was performed using HKLF5 format reflection data file. Twinning domain ratio coefficient (BASF) was successfully refined to 0.3794.

On the final model refinement stage for compound **2** thermal motion parameters for coordinated CO (-C(33)=O(3)) and Cl (Cl(2)) moieties as well as presence of unusually strong residual electron density peaks in one of the compound molecules suggested a positional CO / Cl disorder not related by symmetry. Disorder was successfully modeled with refined occupation ratio at one position CO / Cl = 70%:30%. Disorder of the second position was inversed in such way that overall occupancy summed up to one full CO and one full Cl ligands in the first coordination sphere of Re metal center. Set of geometrical (SADI) and thermal motion (SIMU, DELU) restraints were applied to achieve acceptable molecular fragment geometries and thermal motion parameter values.

For all the compounds hydrogen atoms positions were initially assigned from the residual electron density peaks coordinates. However, after initial placement all hydrogen atoms were treated as idealized contributions during the refinement. All scattering factors are contained in several versions of the SHELXTL program library, with the latest version used being v.6.12.¹⁸⁸

X-Ray crystal structures are shown in Appendix B

A.4 (**terpy- $\kappa^2\text{-N,N'}$**)Re(CO)₃Cl (**2.1**)

Re(CO)₅Cl (201 mg, 0.556 mmol) and 2,2:6,2 terpyridine (129 mg, 0.553 mmol) were mixed in 60 mL of toluene. The reaction mixture was heated to 100C for 1 hour under N₂. During this time the solution turned a bright red. Upon cooling a yellow precipitate

was formed. The solution was filtered, and the solid was washed with diethyl ether, and dried under vacuum. Compound **1** is a bright yellow powder that was isolated in 70 % yield (208 mg). Crystals were obtained from chloroform with addition of a small amount of hexanes as counter solvent. TGA: 8 % mass loss from 240-280 °C. FTIR: 2019, 1981, 1889 cm⁻¹ (*v* C=O). ¹H NMR (CD₃CN, 400 MHz): δ 9.06 (ddd, J=5.6, 1.7, 0.8 Hz, 1H), 8.77 (ddd, J=4.9 Hz, 1H), 8.49 (td, J=8.2, 1.5 Hz, 2H), 8.28 (t, J=7.9 Hz, 1H), 8.22 (td, J=8.1, 1.6 Hz, 1H), 7.96 (td, J=7.8, 1.8 Hz, 1H), 7.79 (dt, J=7.8, 1.1 Hz, 1H), 7.80 (dd, J=7.7, 1.0 Hz, 1H), 7.63 (ddd, J=7.6, 5.5, 1.2 Hz, 1H), 7.55 (ddd, J=7.6, 5.0, 1.0 Hz, 1H). Elemental analysis calculated (%) for [C₁₈H₁₁ReClN₃O₃]: C 40.11, H 2.06, N 7.80, found C 39.96, H 2.09, N 7.69.

A.5 (terpy- κ^3 -N,N',N'')Re(CO)₂Cl (2.2)

Compound **2.1** (101 mg, 0.187 mmol) was placed in a tube furnace and heated to 240 °C under N₂ flow for 60 minutes. A black solid was collected (90 mg) at 94 % yield based on the formula for **2**. Crystals were obtained from chloroform with addition of a small amount of hexanes as counter solvent. FTIR: 1872, 1788 cm⁻¹ (*v* C=O). ¹H NMR (CD₃CN, 400 MHz): δ 8.70 (ddd, J=4.7, 1.8, 0.9 Hz, 2H), 8.65 (dt, J=8.0, 1.0, 1.0, 2H), 8.47 (d, J=7.8 Hz, 2H), 8.03 (t, J=7.8 Hz, 1H), 7.95 (td, J=7.7, 7.7, 1.9 Hz, 2H), 7.43 (ddd, J=7.5, 4.8, 1.2 Hz, 2H). Elemental analysis calculated (%) for [C₁₇H₁₁ReClN₃O₂]: C 39.96, H 2.17, N 8.22, found C 39.62, H 2.09, N 7.99.

A.6 (terpy- κ^2 -N,N')Re(CO)₃Br (2.3)

Re(CO)₅Br (191 mg, 470 mmol) and 2,2:62 terpyridine (129 mg, 0.553 mmol) were allowed to react under conditions analogous to the preparation of **1**. A bright yellow

powder was obtained, 0.223 g (0.382 mmol, 81 %). FTIR: 2012, 1910, 1886 cm⁻¹ (*v* C=O). ¹H NMR (CD₃CN, 400 MHz): δ 9.07 (ddd, J=5.6, 1.6, 0.9 Hz, 1H), 8.77 (ddd, J=4.6, 1.5, 0.8 Hz, 1H), 8.52 – 8.48 (m, 2H), 8.28 (t, J=7.9 Hz, 1H), 8.21 (td, J=8.0, 1.6 Hz, 1H), 7.97 (td, J=7.8, 1.8 Hz, 1H), 7.80 – 7.75(m, 2H), 7.63 (ddd, J=7.6, 5.5, 1.2 Hz, 1H), 7.55 (ddd, J=7.7, 4.9, 1.1 Hz, 1H). Elemental analysis calculated (%) for [C₁₈H₁₁ReBrN₃O₃]: C 37.06, H 1.90, N 7.20, found C 36.94, H 1.92, N 7.00.

A.7 (terpy- κ^3 -N,N',N'')Re(CO)₂Br (2.4)

Compound **2.3** (182 mg, 0.312 mmol) was placed in an tube furnace and heated to 230 °C under N₂ flow for 60 minutes. A black solid was collected, 0.155 mg (0.279 mmol, 89 % yield). Crystals were obtained from chloroform with addition of a small amount of hexanes as counter solvent. FTIR: 1873, 1794 cm⁻¹ (*v* C=O). ¹H NMR (CD₃CN, 400 MHz): δ 8.95 (d, J=5.4 Hz, 2H), 8.24 (t, J = 8.1Hz, 4H), 8.05 (dd, J=8.2, 7.7 Hz, 1H), 7.90(td, J=7.9, 1.7 Hz, 2H), 7.63 (ddd, J=7.3, 5.5, 1.2 Hz, 2H). Elemental analysis calculated (%) for [C₁₇H₁₁ReBrN₃O₂]: C 36.76, H 2.00, N 7.57, found C 36.66, H 2.00, N 7.50.

A.8 (terpy- κ^2 -N,N')Re(CO)₃CN (3)

To **2.3** (60 mg, 0.103 mmol), AgCF₃SO₃ (40 mg, 0.299 mmol) was added in 10 mL CH₂Cl₂. The reaction was stirred 18 h and kept dark under N₂ atmosphere. Solution was filtered to remove salts, then reduced in volume. Cold diethyl ether was used to precipitate product. Yellow-grey powder (**7a**) was collected by filtration, yielding 48mg (88 %). Crystals were grown from saturated chloroform with hexanes as countersolvent for X-ray crystallography. FTIR: 2013, 1905, 1884 cm⁻¹ (*v* C=O). ¹H NMR (CD₃CN,

400 MHz): δ 9.08 (dd, $J=5.40$, 0.49 Hz, 1 H), 8.78 (dd, $J=5.10$, 0.59 Hz, 1 H), 8.53 (dd, $J=8.18$, 0.93 Hz, 1 H), 8.50 (d, $J=8.23$ Hz, 1 H), 8.30 (t, $J=7.90$ Hz, 1 H), 8.23 (td, $J=7.94$, 1.57 Hz, 2 H), 7.98 (td, $J=7.76$, 1.71 Hz, 1 H), 7.80 (dd, $J=7.79$, 1.03 Hz, 1 H), 7.77 (d, $J=7.74$ Hz, 0 H), 7.64 (ddd, $J=7.50$, 5.50, 1.08 Hz, 1 H), 7.55 (ddd, $J=7.62$, 4.92, 0.98 Hz, 2 H). Elemental analysis calculated (%) for [C₁₉H₁₁N₄O₃Re]: C 43.10, H 2.09, N 10.58, found C 40.06, H 2.06, N 9.85.

A.9 (**terpy- κ^3 -N,N',N''**)Re(CO)₂CN (2.6)

Compound **2.5** (108 mg, 0.204 mmol) was placed in an tube furnace and heated to 220 °C under N₂ flow for 60 minutes. A black solid was collected, 88 mg (0.175 mmol, 86 % yield). ¹H NMR (CD₃CN, 400 MHz): δ 8.60(ddt, $J=4.8$, 1.7, 0.8, 0.8 Hz, 2H), 8.48 (dq, $J=8.0$, 1.0 Hz, 2H), 8.37 (dd, $J=7.9$, 0.8 Hz, 2H), 8.07 (dd, $J=8.0$, 7.6 Hz, 1 H), 7.94 (ddd, $J=7.9$, 7.5, 1.8 Hz, 2H) 7.43 (ddd $J=7.4$, 4.8, 1.2 Hz, 2H). Elemental analysis calculated (%) for [C₁₈H₁₁N₄O₂Re]: C 43.11, H 2.21, N 11.17, found C 40.26, H 2.67, N 9.60.

A.10 (**terpy- κ^2 -N,N'**)Re(CO)₃OTf (2.7)

To **2.1** (80 mg, 0.148 mmol), AgCF₃SO₃ (46 mg, 0.179 mmol) was added in 10 mL CH₃CN. The reaction was stirred 18 h and kept dark. Solution was filtered to remove salts, then reduced in volume. Cold diethyl ether was used to precipitate product. Yellow-grey powder (**7a**) was collected by filtration, yielding 38 mg (40 %). Crystals were grown from saturated chloroform with hexanes as countersolvent for X-ray crystallography. FTIR: 2030, 1895, 1890 cm⁻¹ (ν C=O), 1280, 1228, 1204 cm⁻¹ (ν SO₃). ¹H NMR (CD₃CN, 400 MHz): δ 9.05 (ddd, $J=5.5$, 1.6, 0.8 Hz, 1H), 8.79 (ddd, $J=4.9$, 1.8, 1.1 Hz, 1H), 8.57

(dd, $J=8.1$, 0.9 Hz, 1H), 8.54 (dt, $J=8.2$, 1.1 Hz, 1H), 8.37 (t, $J=7.9$ Hz, 1H), 8.31 (td, $J=8.0$, 1.6 Hz, 1H), 8.03 (td, $J=7.7$, 1.7 Hz, 1H), 7.87 (dd, $J=7.8$, 1.0 Hz, 1H), 7.75 (dt, $J=7.8$, 1.1 Hz, 1H), 7.72 (ddd, $J=7.4$, 5.9, 1.1 Hz, 1H), 7.61 (ddt, $J=7.7$, 4.8, 0.5, 0.5 Hz, 1H). Elemental analysis calculated (%) for $[C_{19}H_{11}ReN_3O_6F_3S]$: C 34.97, H 1.70, N 6.44, found C 31.80, H 1.73, N 5.33.

Alternately, to **1** (72 mg, 0.134 mmol) was added 10 mL CF_3SO_3H (excess) and temperature was increased to 100 °C for 20 minutes. A black solution was neutralized with addition of 5% Na_2CO_3 in H_2O . Product was extracted with $CHCl_3$, then dried under vacuum to yield a brown solid (**7b**) (47 mg, 54 %).

A.11 (*terpy*- κ^3 -N,N',N'')Re(CO)₂OTf (2.8)

To **2.2** (77 mg, 0.143 mmol), $AgSO_3CF_3$ (47 mg, 0.183 mmol) was added in 15 mL CH_3CN . Solution was refluxed for 6 h in the dark under N_2 atmosphere. Solution was filtered, then reduced to minimal volume. Cold diethyl ether was added dropwise to precipitate product. Collected by filtration and washed with additional cold ether, yielding 75 mg (120 mmol, 80 %). Crystals grown from saturated methylene chloride, with hexanes as countersolvent for x-ray crystallography. FTIR: 1910, 1829 cm^{-1} (v C=O), 1259, 1224, 1143 cm^{-1} (v SO₃). ¹H NMR (CD_3CN , 400 MHz): δ 8.91(ddd, $J=5.6$, 1.6, 0.7 Hz, 2H), 8.32 (d, $J=8.0$ Hz, 2H), 8.28 (ddd, $J=8.1$, 1.4, 0.8 Hz, 2H), 8.19 (dd, $J=8.8$, 7.4 Hz, 1H), 8.02 (td, $J=7.9$, 1.5 Hz, 2H) 7.46 (ddd $J=7.6$, 5.6, 1.3 Hz, 2H). Elemental analysis calculated (%) for $[C_{18}H_{11}ReF_3SN_3O_5]$: C 34.62, H 1.78, N 6.37, found C 31.02, H 1.82, N 7.11.

Appendix B

Molecular Orbitals Diagrams

Multiple views of each x-ray crystal structure (including full unit cell) as discussed in chapter 2 are shown in Figures B.1 to B.5.

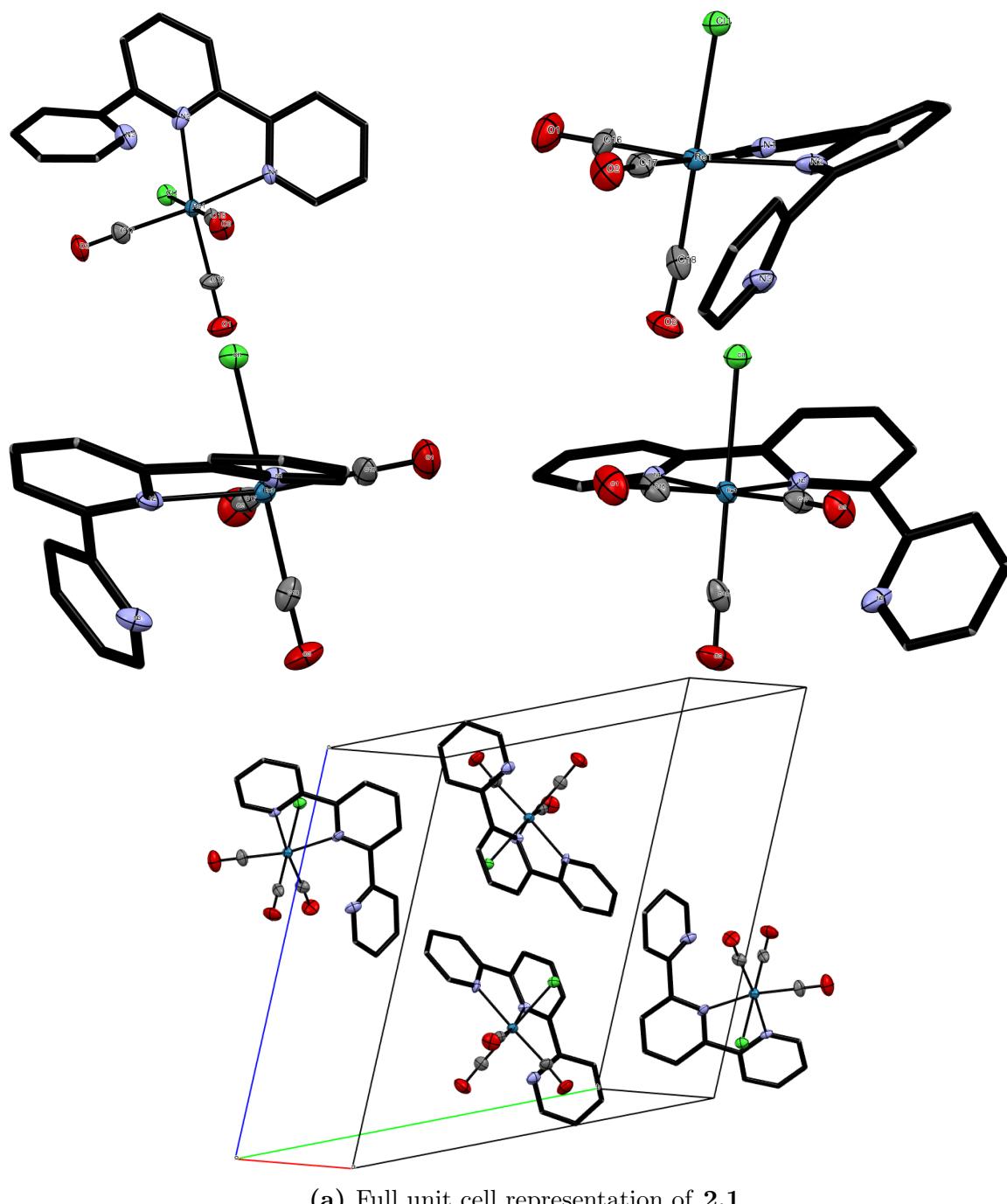


Figure B.1 X-ray crystal structure of **2.1**. Co-crystallized chloroform, hydrogen atoms, and thermal ellipsoids of ligand carbon atoms are omitted for clarity.

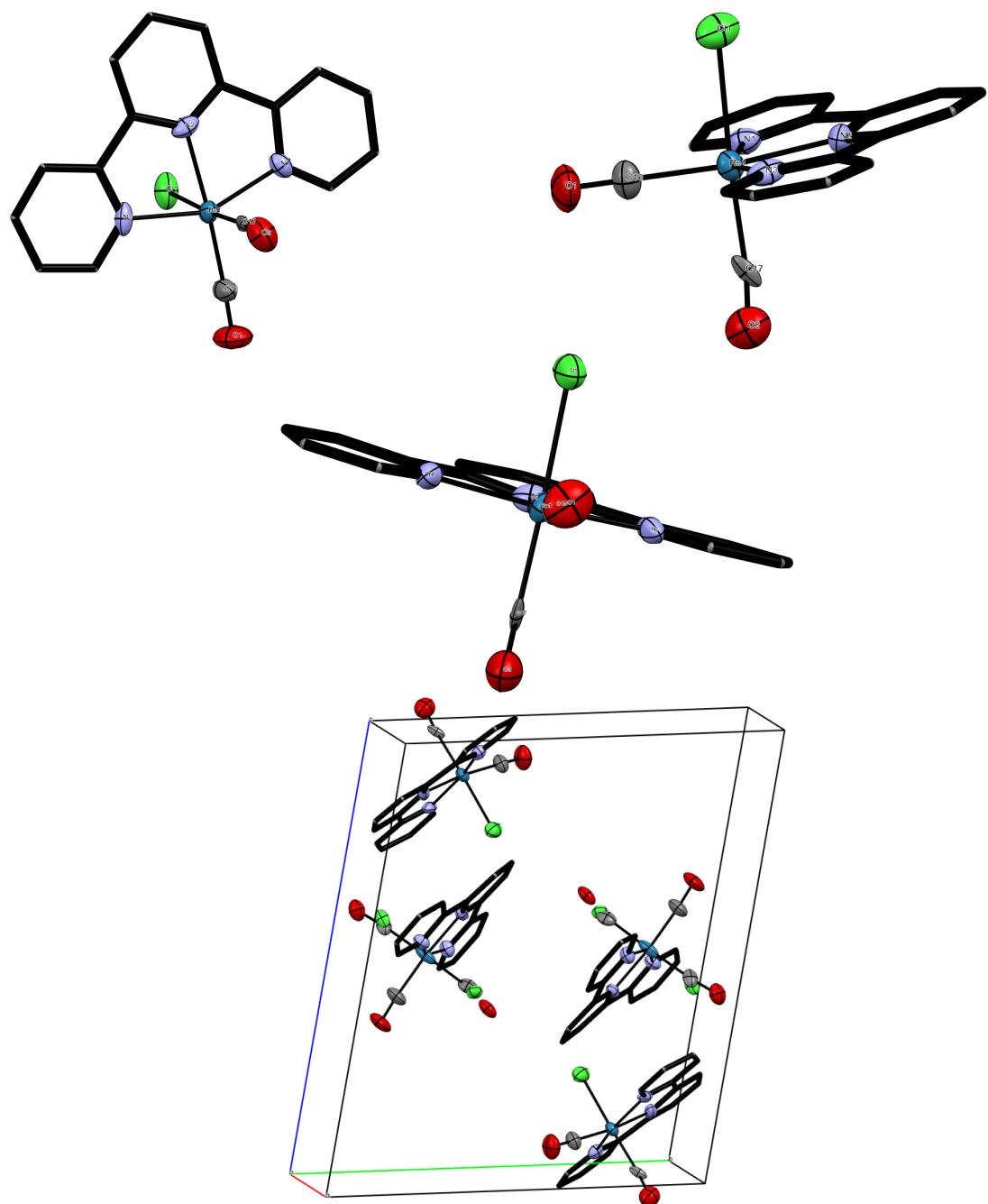
(a) Full unit cell representation of **2.2**

Figure B.2 X-ray crystal structure of **2.2**. Co-crystallized chloroform, hydrogen atoms, and thermal ellipsoids of ligand carbon atoms are omitted for clarity.

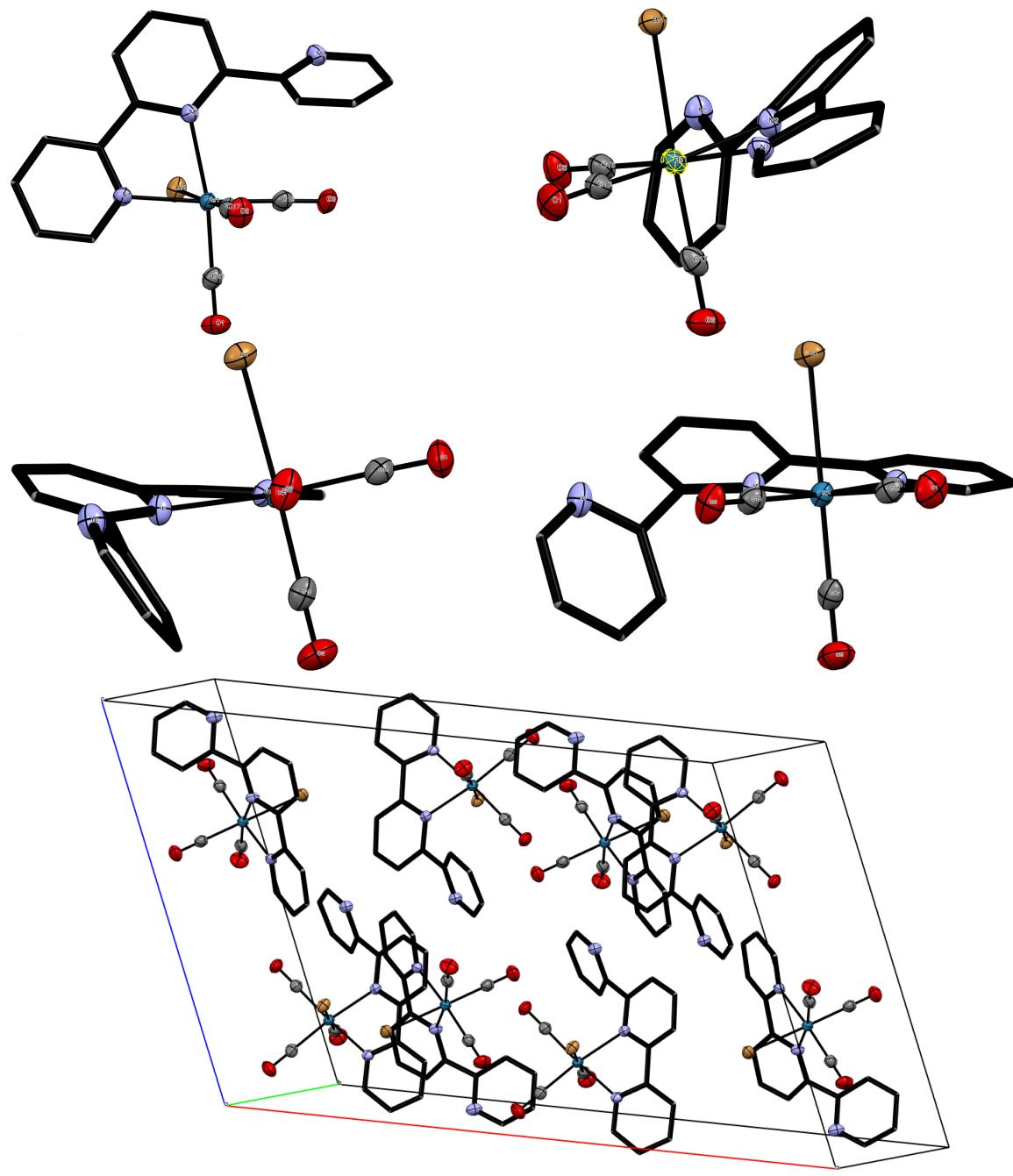
(a) Full unit cell representation of **2.3**

Figure B.3 X-ray crystal structure of **2.3**. Hydrogen atoms, and thermal ellipsoids of ligand carbon atoms are omitted for clarity.

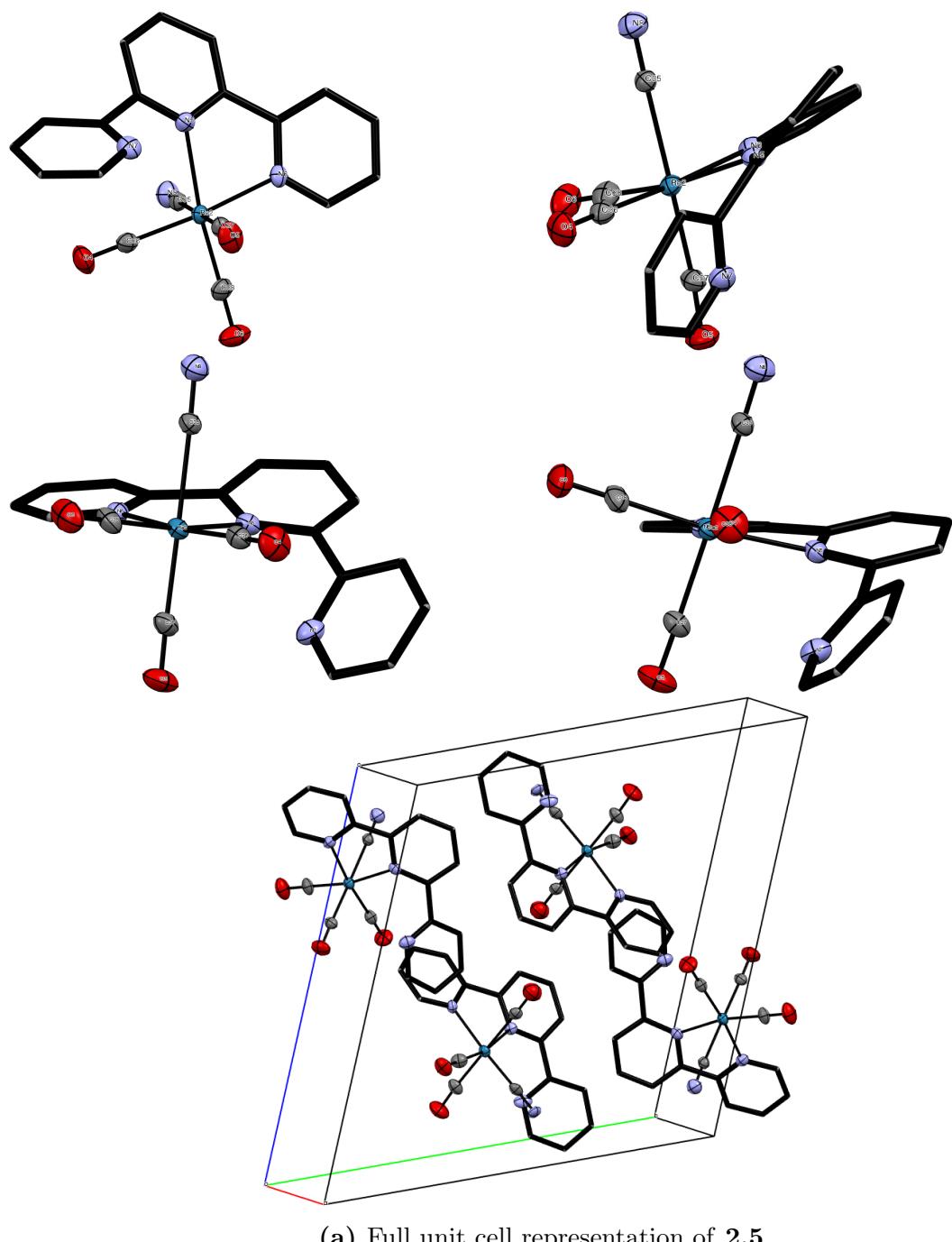


Figure B.4 X-ray crystal structure of **2.5**. Hydrogen atoms, and thermal ellipsoids of ligand carbon atoms are omitted for clarity.

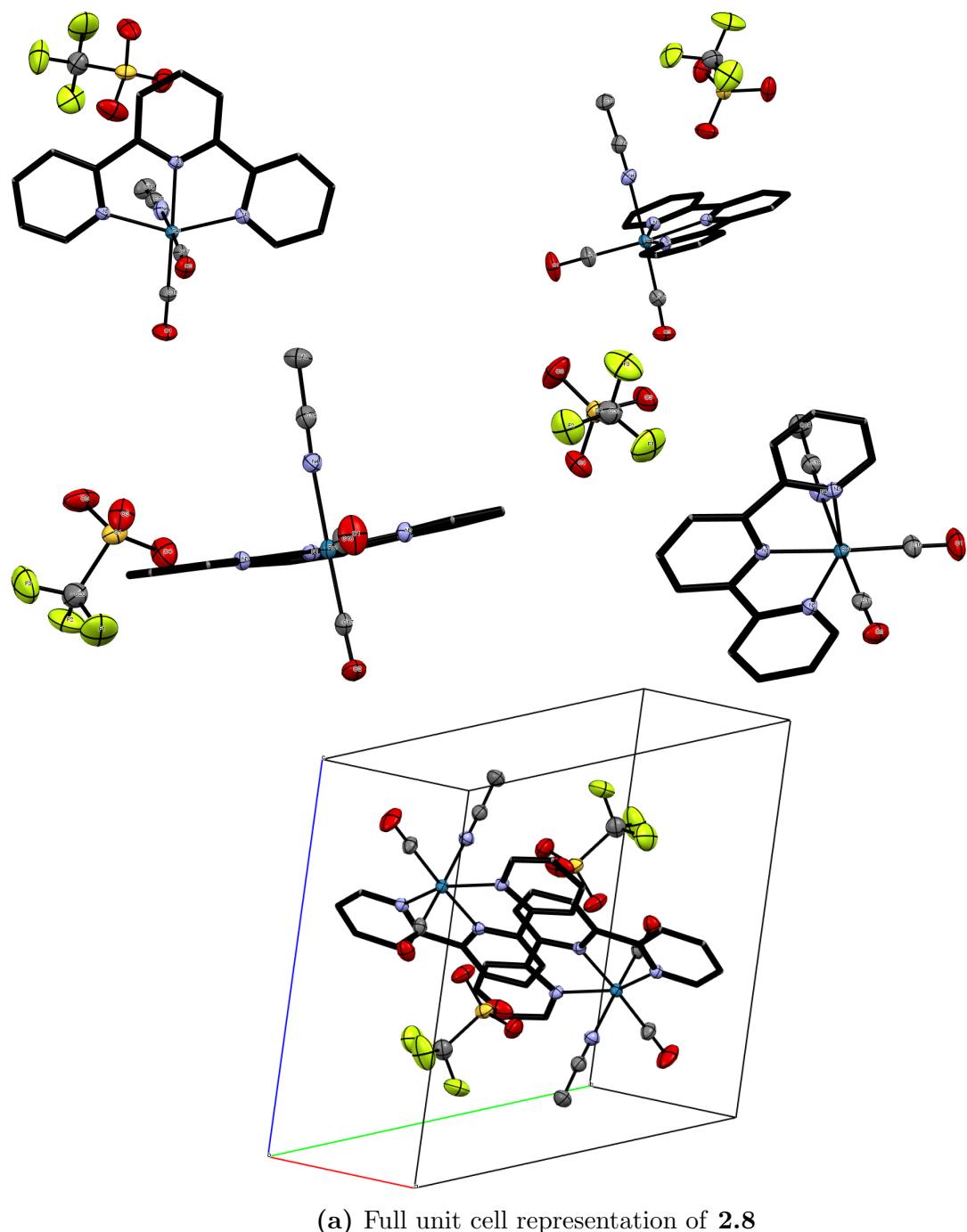


Figure B.5 X-ray crystal structure of **2.8**. Hydrogen atoms, and thermal ellipsoids of ligand carbon atoms are omitted for clarity.

Appendix C

Molecular Orbitals Diagrams

Frontier Molecular Orbitals (MOs) for each compound discussed in this thesis are listed below in Figures C.1 to C.8. These orbitals are generated with the Chemissian program¹⁸⁵ from the output of calculations performed in Gaussian 09.¹¹⁹

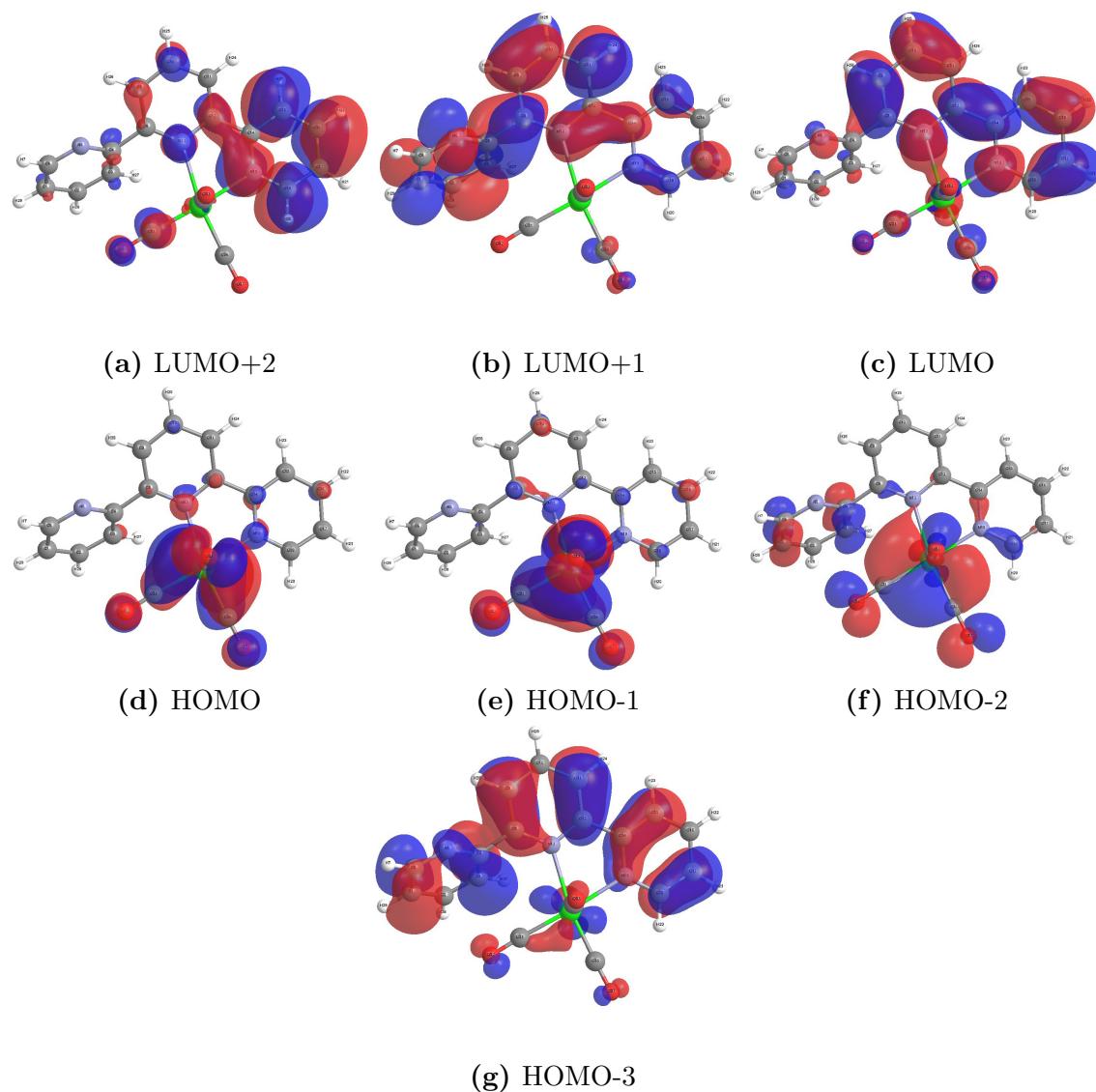


Figure C.1 Isosurface plots of the frontier molecular orbitals HOMO-3 to LUMO+2 of **2.1**

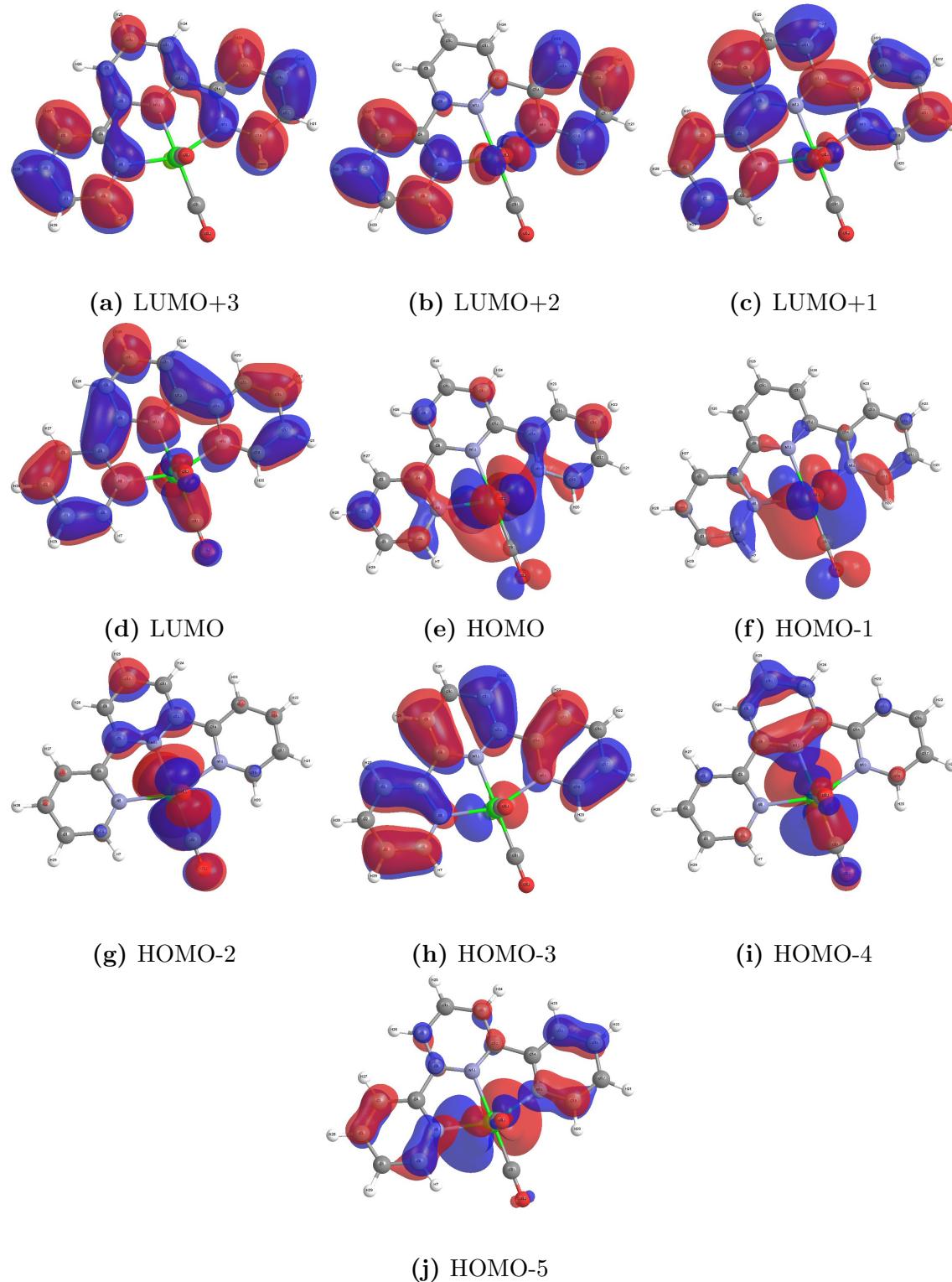


Figure C.2 Isosurface plots of the frontier molecular orbitals HOMO-5 to LUMO+3 of **2.2**

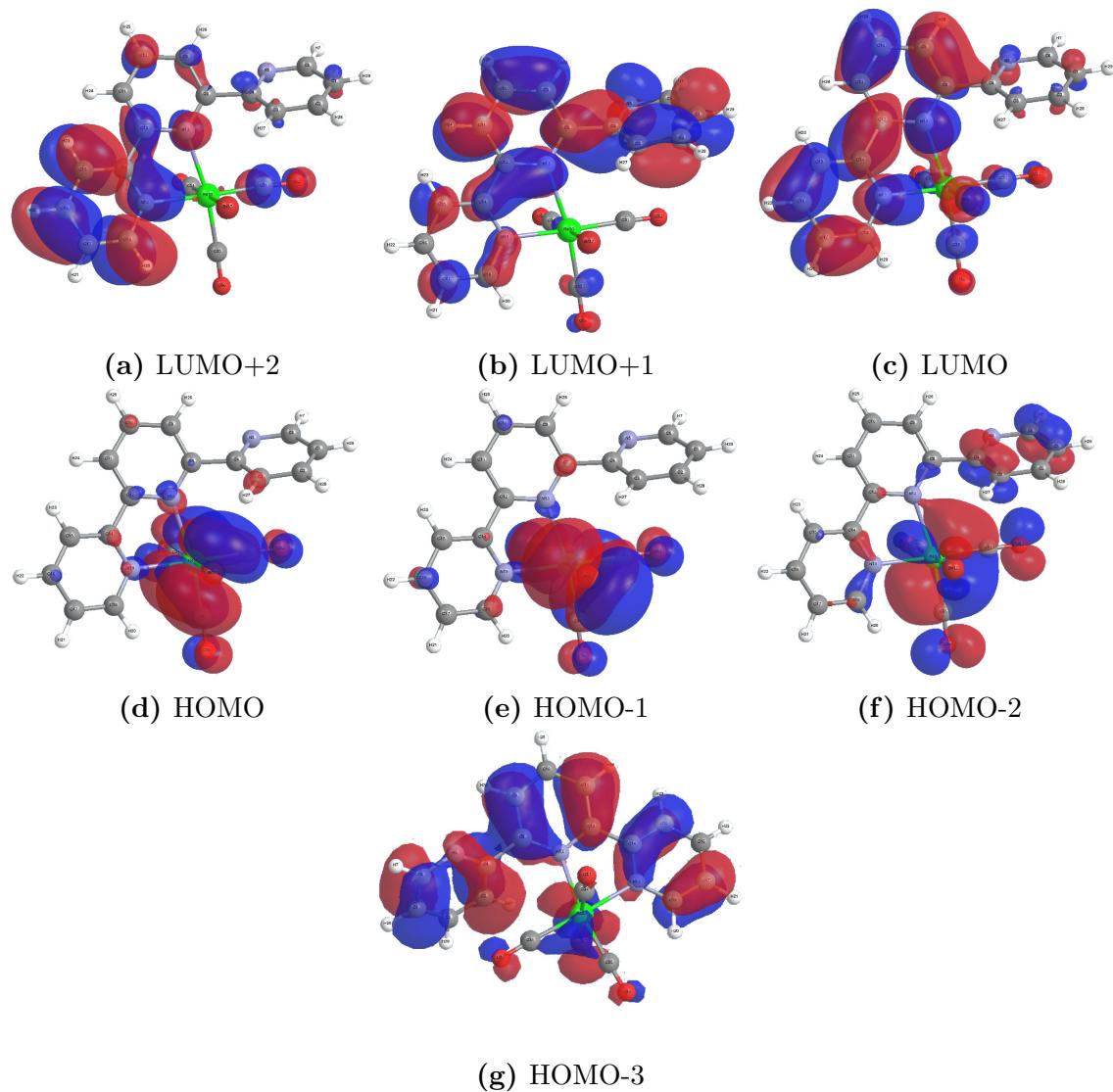


Figure C.3 Isosurface plots of the frontier molecular orbitals HOMO-3 to LUMO+2 of **2.3**

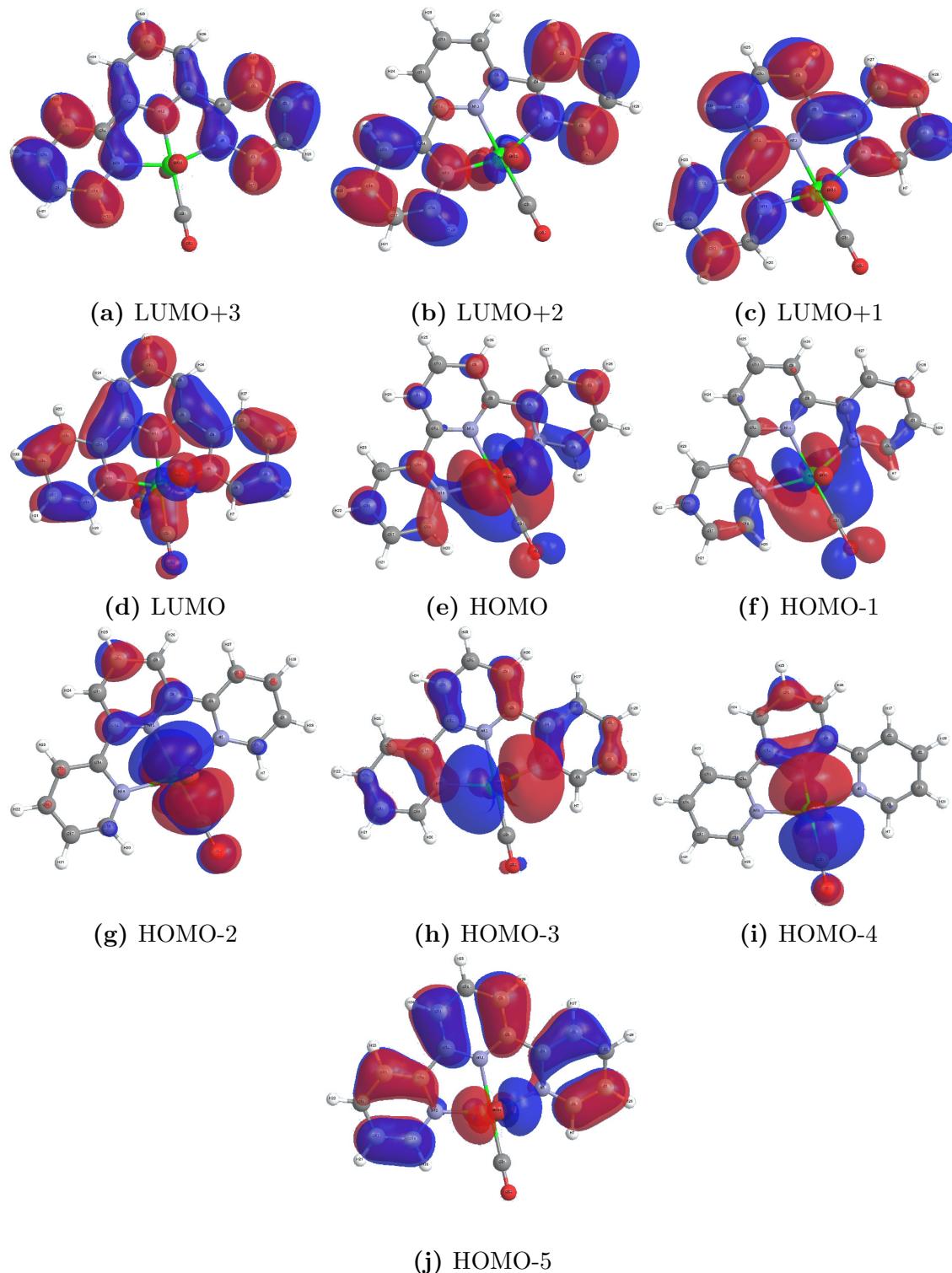


Figure C.4 Isosurface plots of the frontier molecular orbitals HOMO-5 to LUMO+3 of 2.4

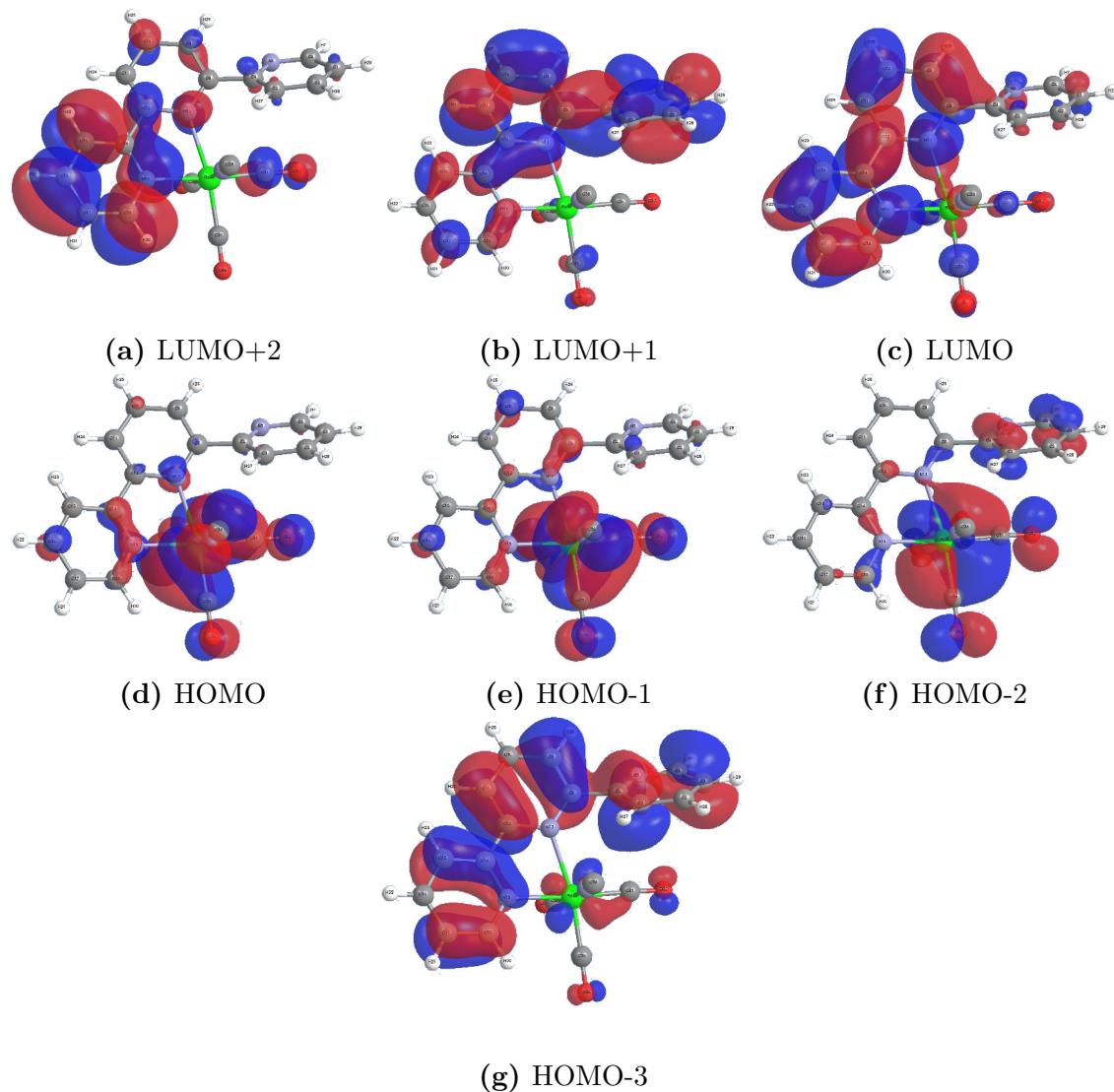


Figure C.5 Isosurface plots of the frontier molecular orbitals HOMO-3 to LUMO+2 of **2.5**

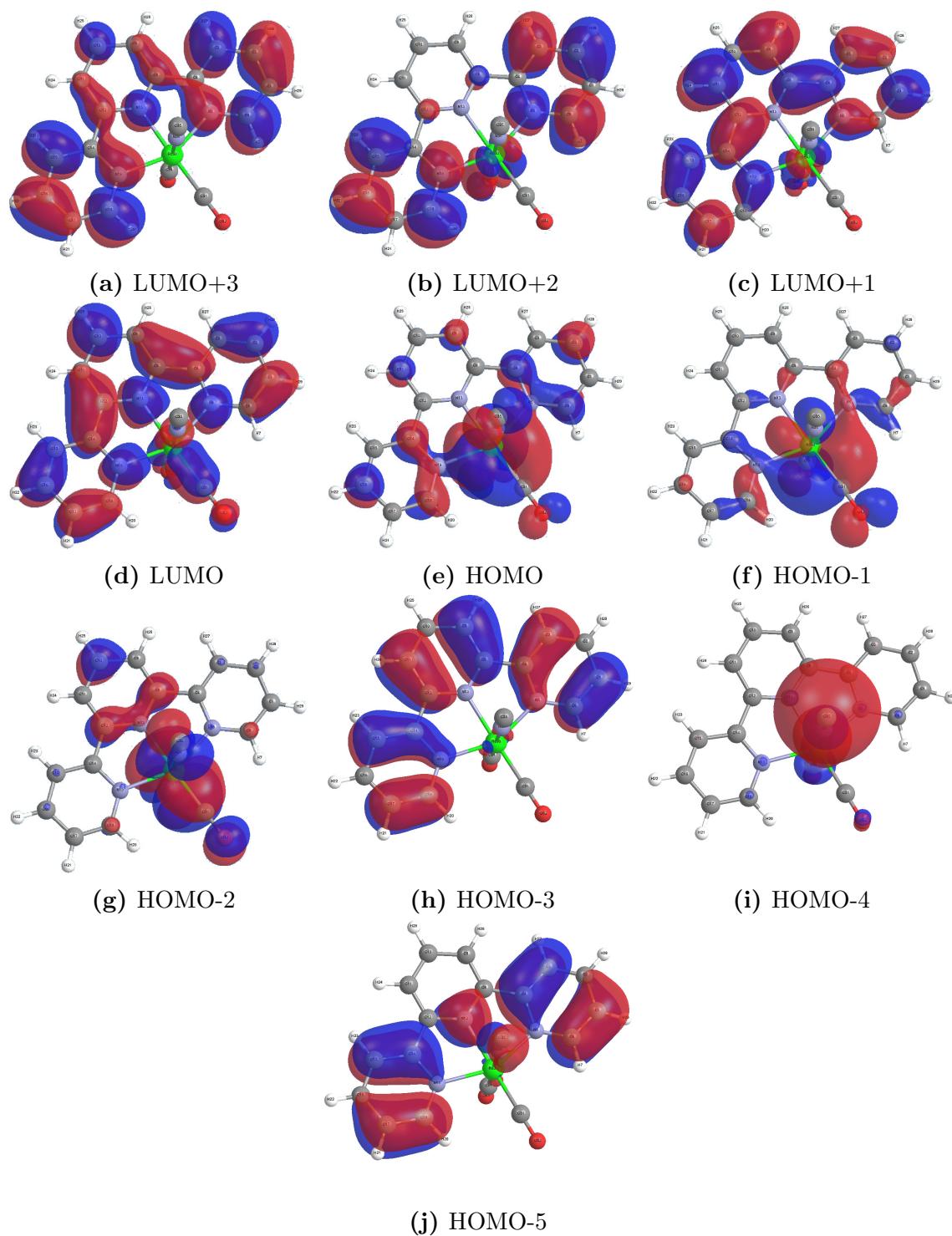


Figure C.6 Isosurface plots of the frontier molecular orbitals HOMO-5 to LUMO+3 of 2.6

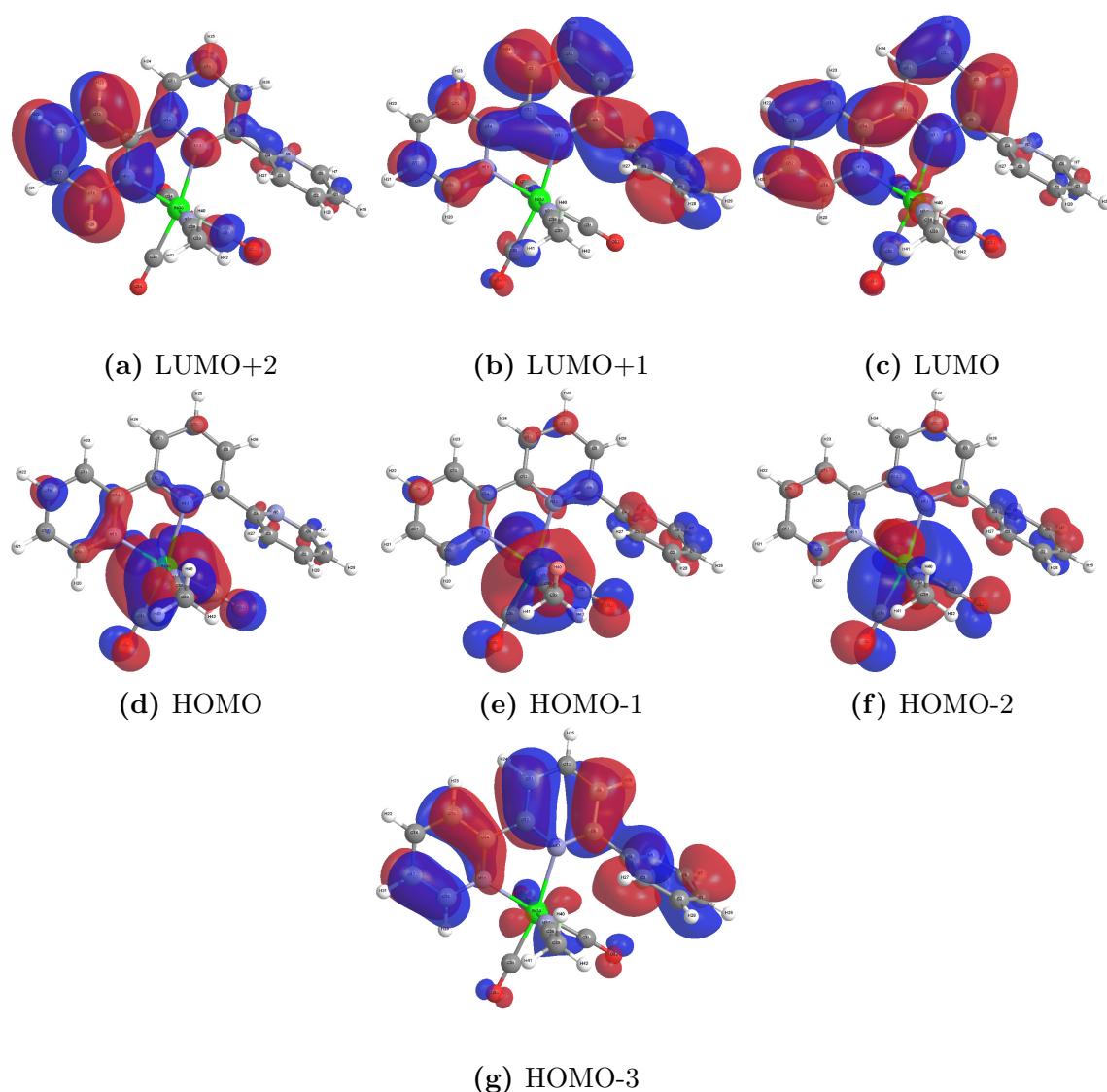


Figure C.7 Isosurface plots of the frontier molecular orbitals HOMO-3 to LUMO+2 of **2.7**

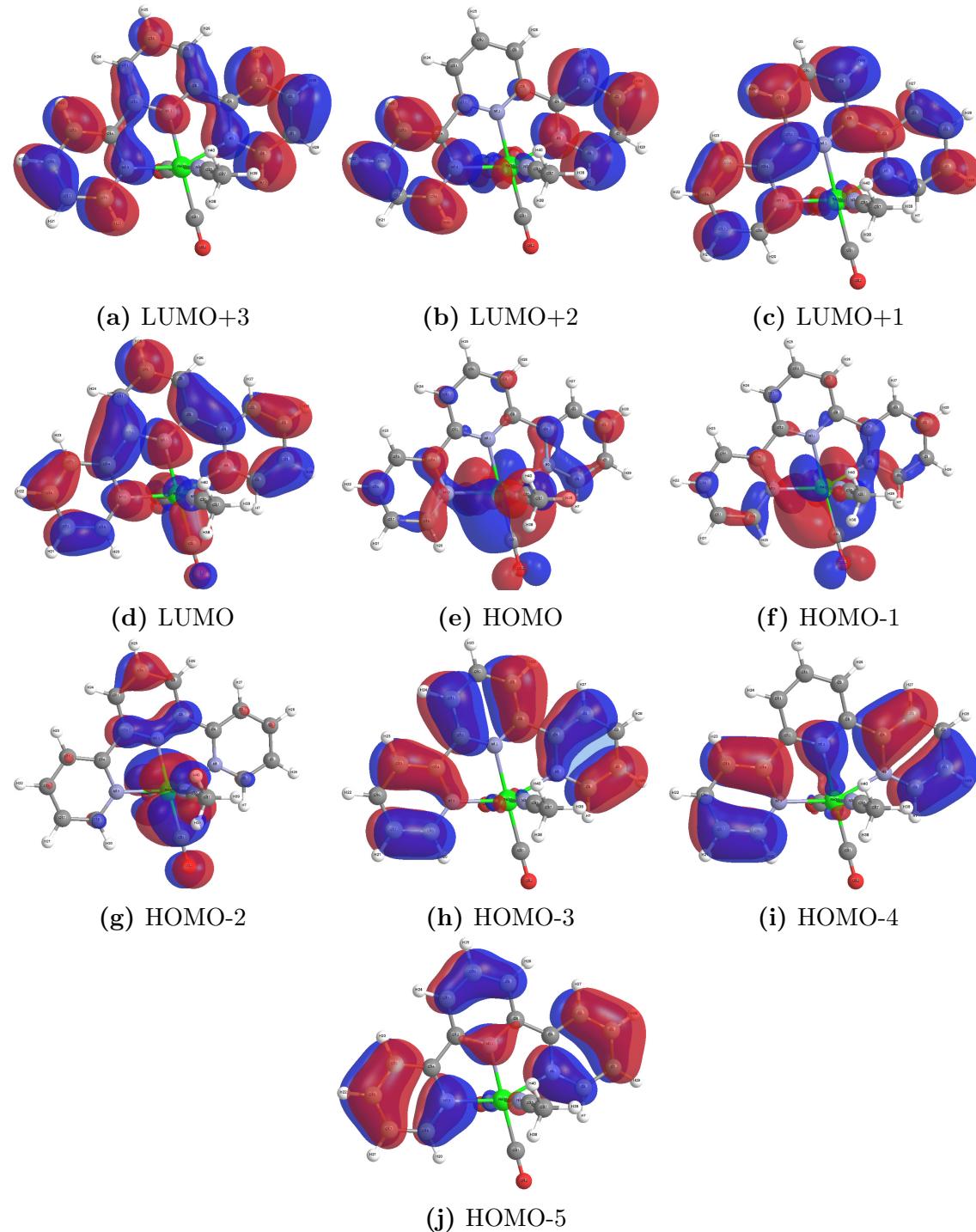


Figure C.8 Isosurface plots of the frontier molecular orbitals HOMO-5 to LUMO+3 of **2.8**

Appendix D

Reaction Potential Energy Diagrams

Potential energy diagrams for the reactions discussed in chapter 4 are listed below in Figures D.1 to D.9.

Reaction Potential Energy Diagrams

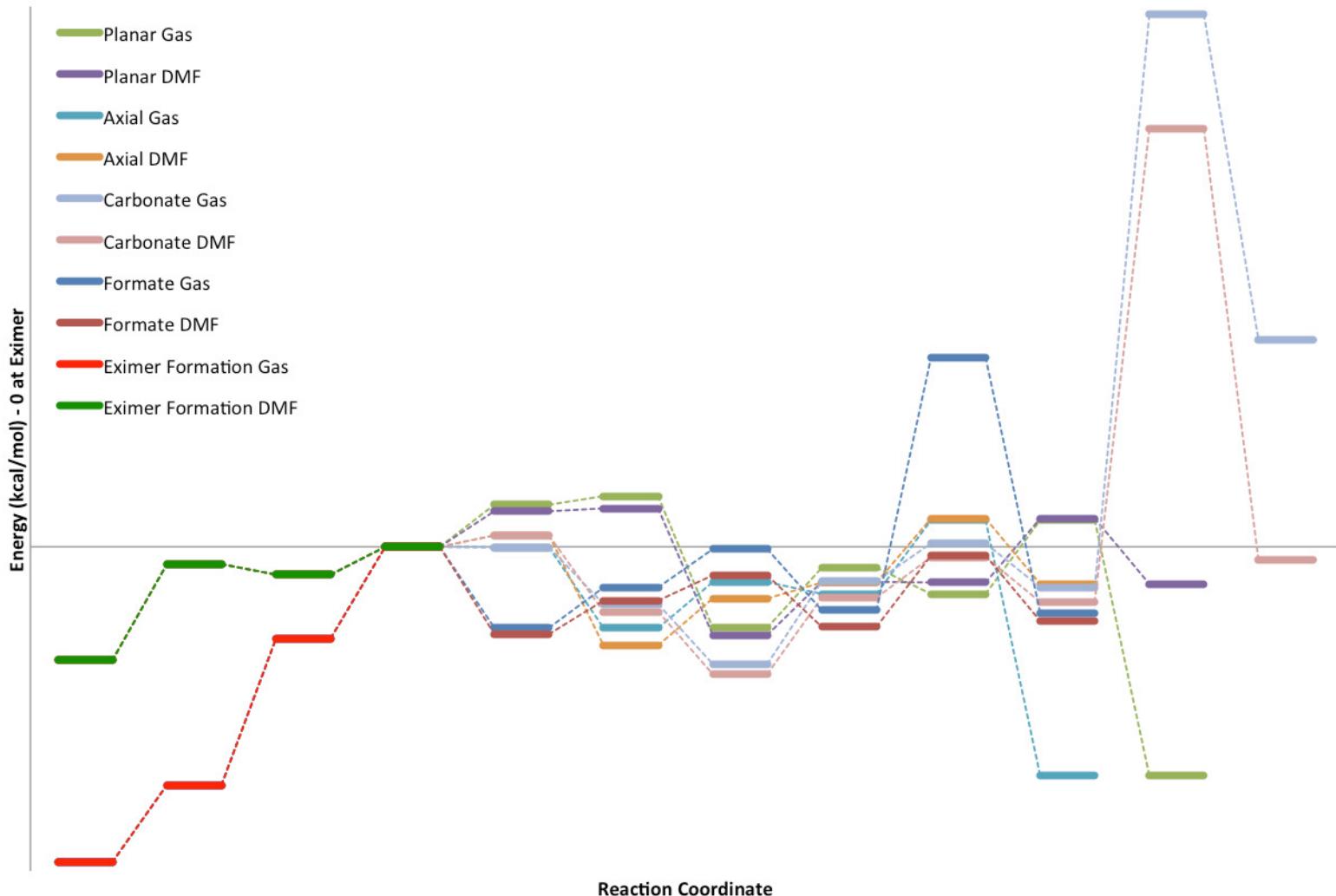


Figure D.1 An overview of the energies of the three mechanistic pathways of photochemical CO_2 reduction

Reaction Potential Energy Diagrams

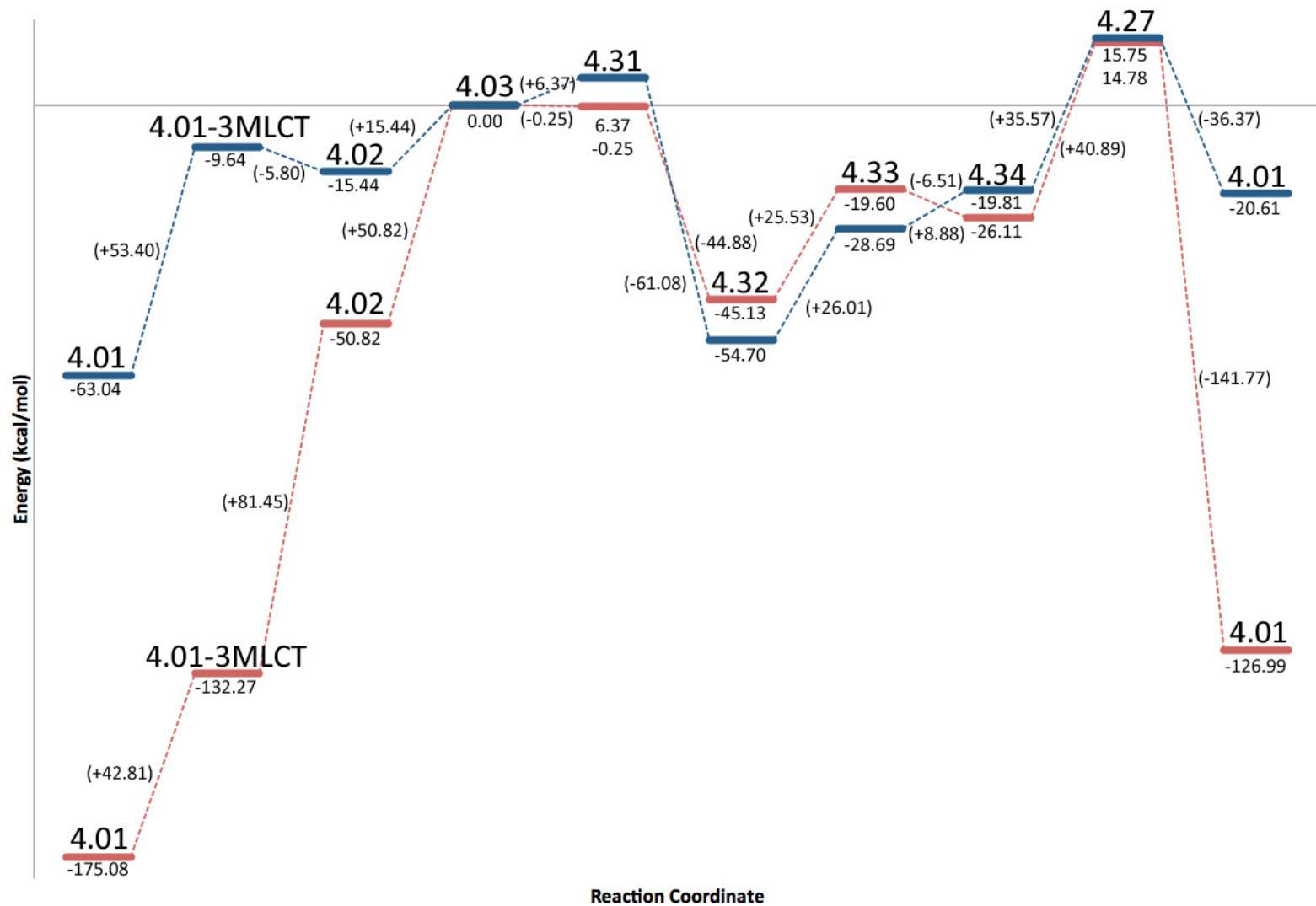


Figure D.2 Potential Energy Surface for the axial geometry of the water-gas shift mechanistic pathway

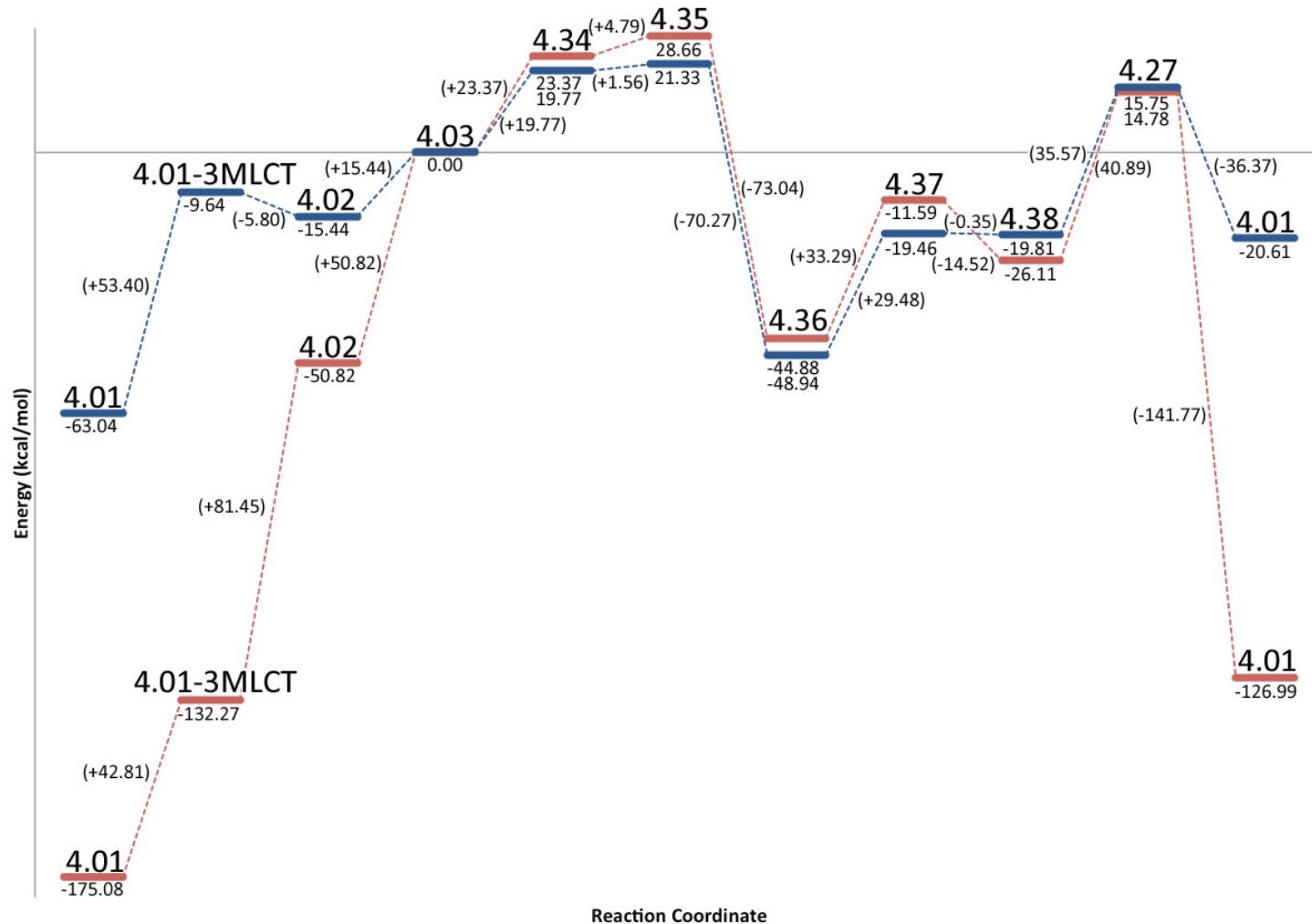


Figure D.3 Potential Energy Surface for the planar geometry of the water-gas shift mechanistic pathway

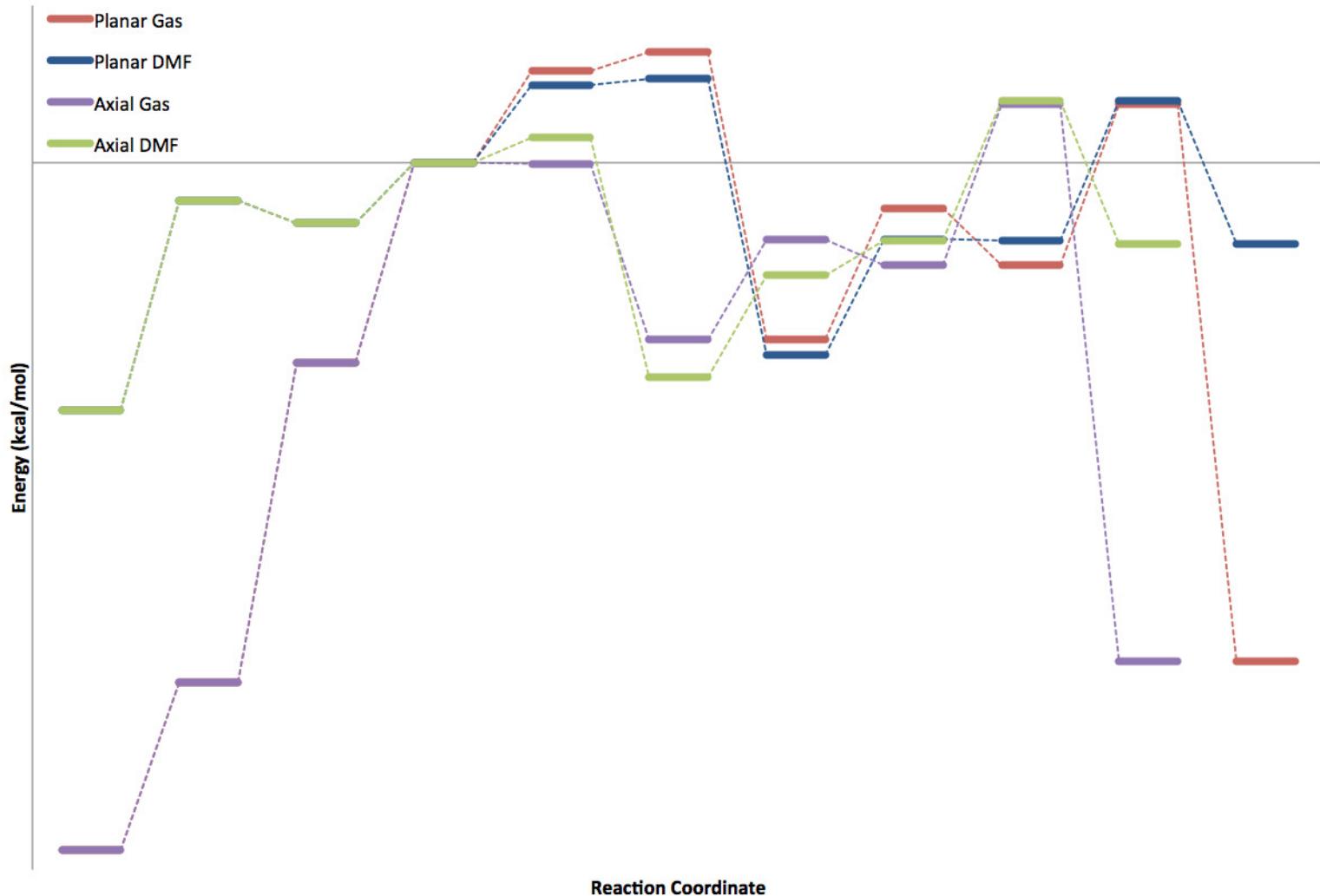


Figure D.4 Potential Energy Surface for the two water-gas shift mechanistic pathway geometries

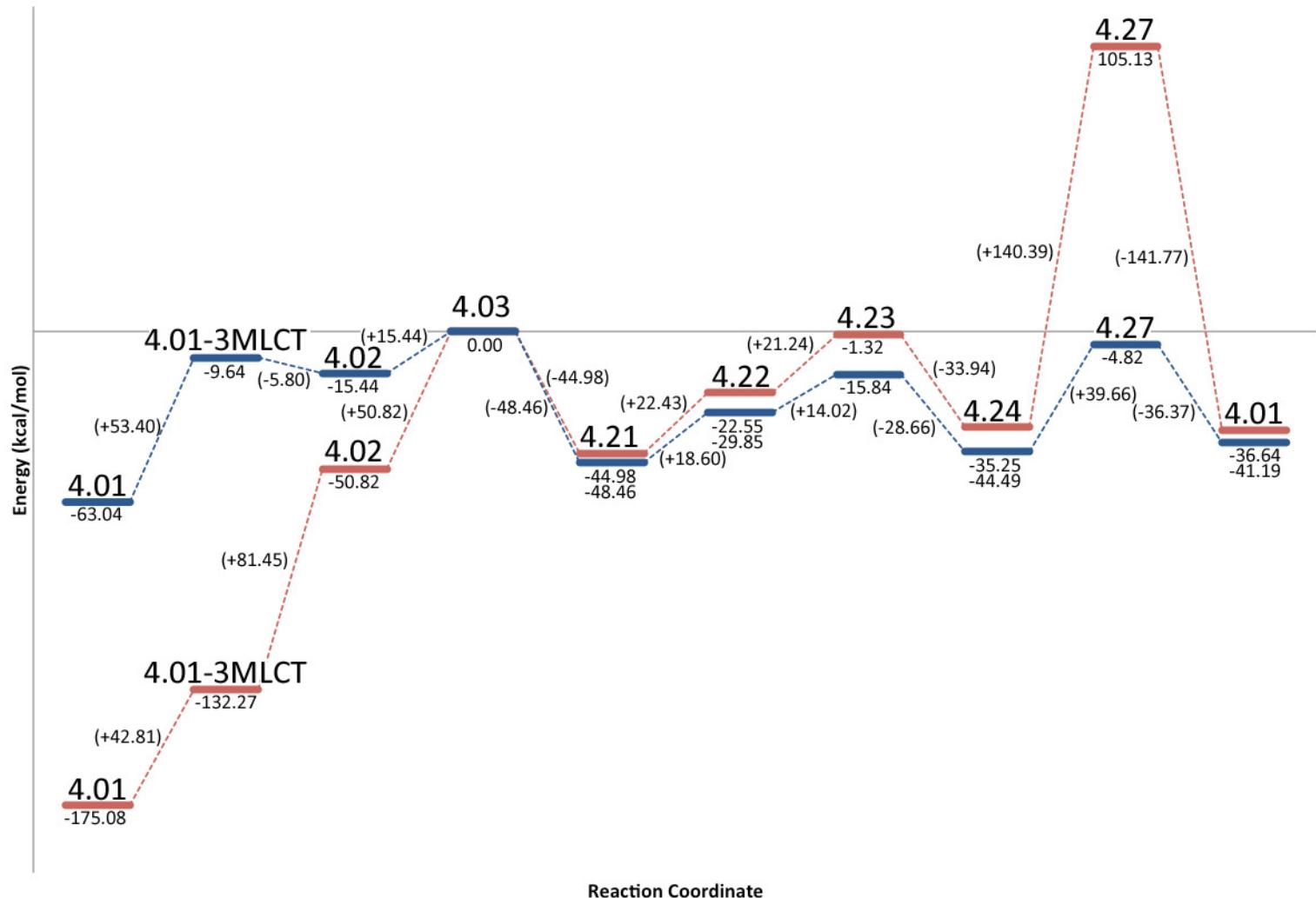


Figure D.5 Potential Energy Surface for the formate mechanistic pathway

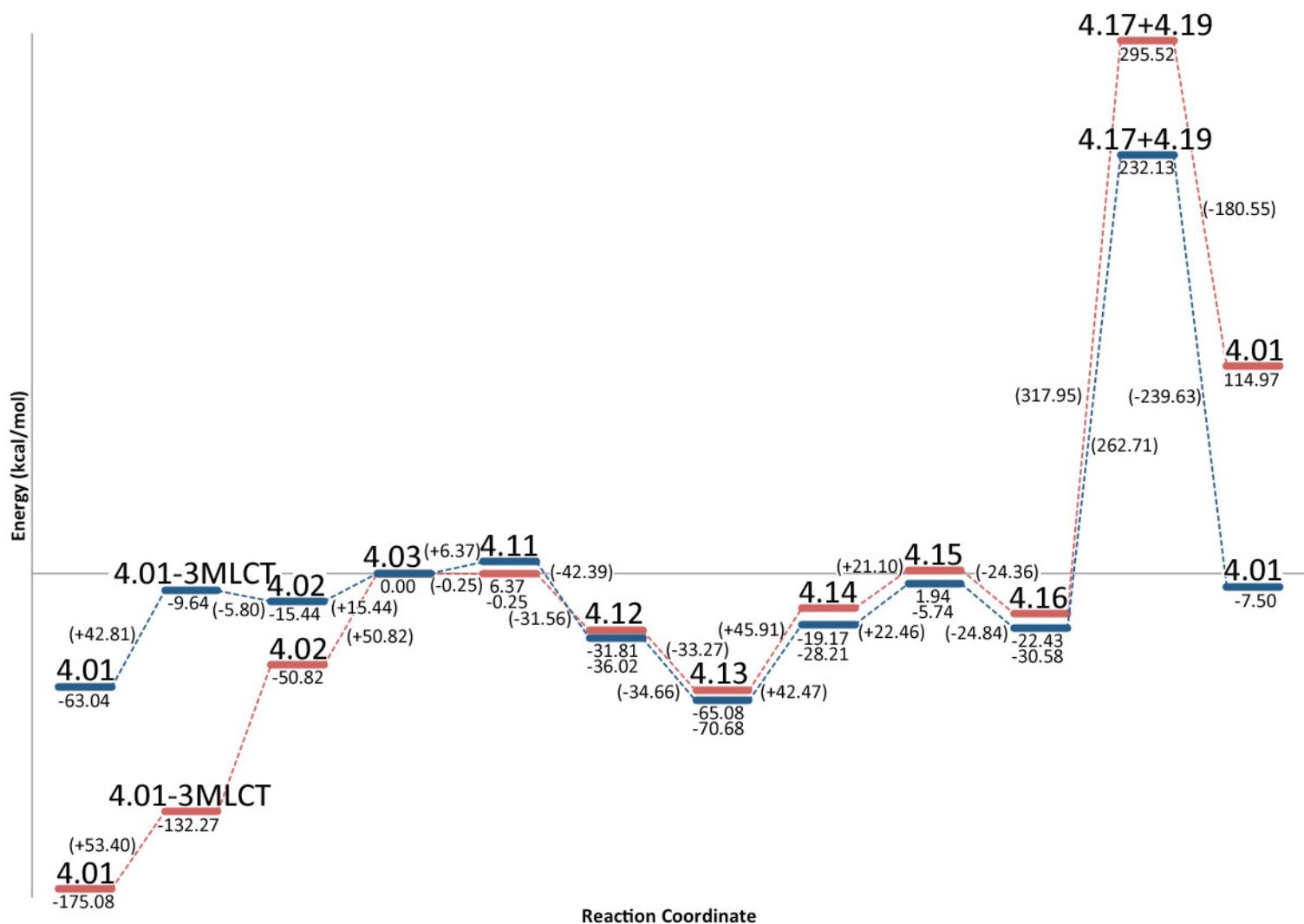


Figure D.6 Potential Energy Surface for the carbonate mechanistic pathway

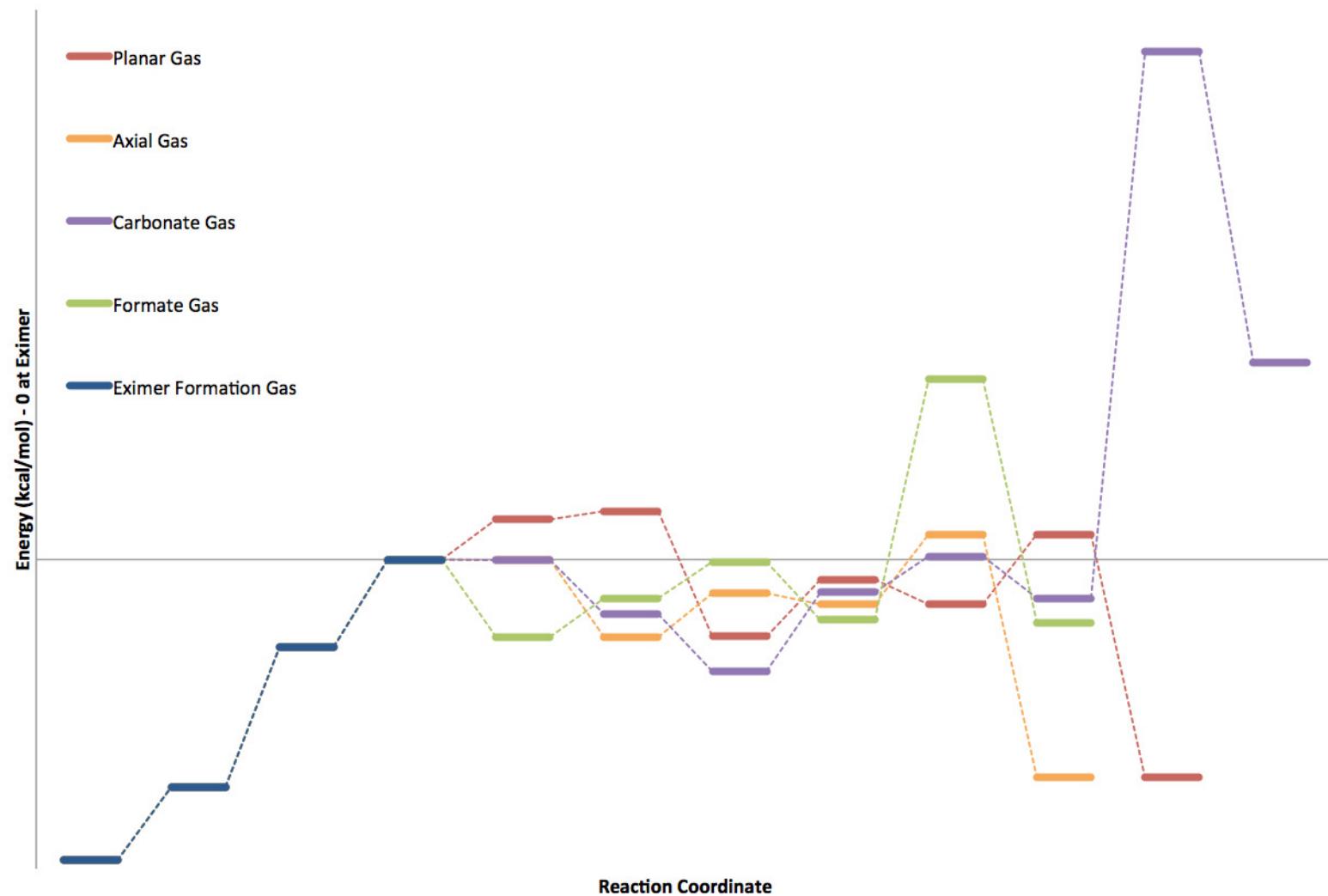


Figure D.7 An overview of the energies of the three mechanistic pathways of photochemical CO_2 reduction in gas phase

Reaction Potential Energy Diagrams

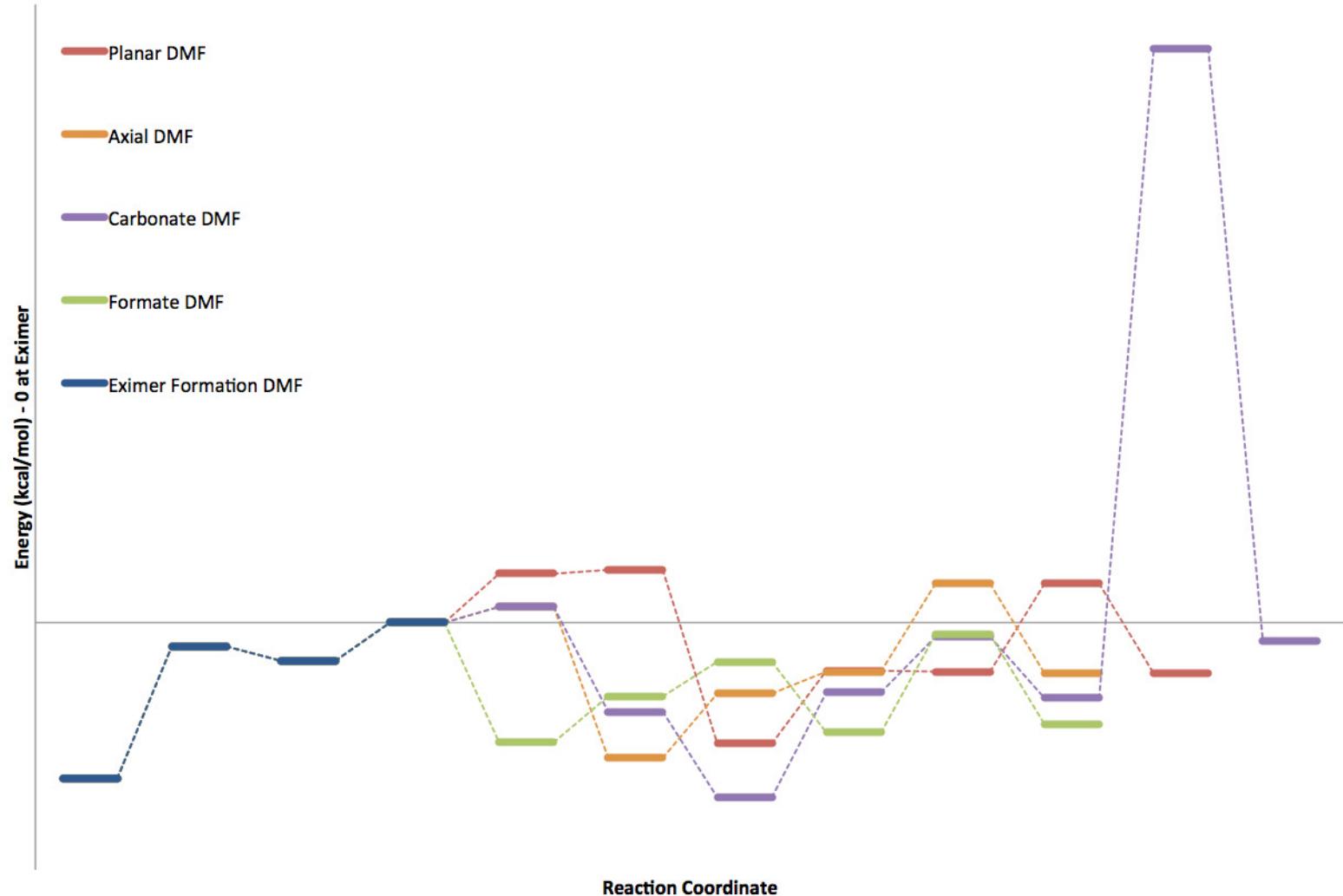


Figure D.8 An overview of the energies of the three mechanistic pathways of photochemical CO₂ reduction in DMF

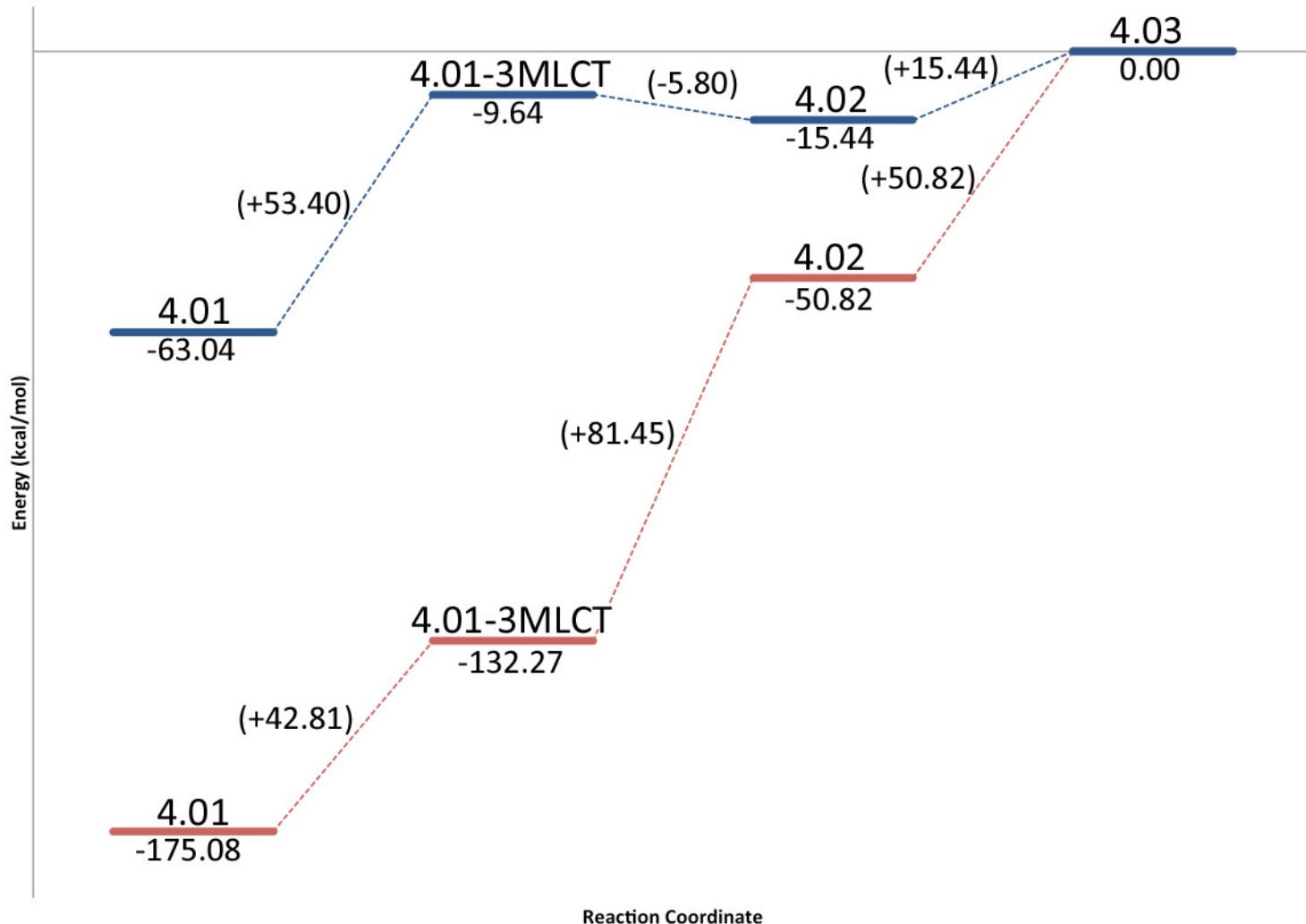


Figure D.9 Potential Energy Surface for the production of the eximer

Appendix E

TurboControl and TurboGo Manual

TurboControl is a series of scripts to run Turbomole jobs from Gaussian style inputs. The following is the user manual included with distributions of TurboControl

E.1 Introduction

Gaussian software is well known for the user friendly GUI it contains (via GaussView). Turbomole, another computational suite, is known for its speed and optimizations, but has a significantly higher learning curve and is less beginner friendly. This software is an attempt to be able to use the user friendly input from Gaussian to smooth over the use of Turbomole.

E.2 System Requirements

There are two user-facing scripts available, both written to work with Turbomole 6.1-6.5 on clusters using Grid Engine queuing software. The only tests of operation are on a system with the following details:

- Rocks 6.1 (Emerald Boa)/CentOS 6.3
- Open Grid Scheduler/Grid Engine 2011.11p1
- Python 2.7.3

Other systems, including different operating systems, different versions of Grid Engine or python, or on other systems, are not supported.

Python dependencies include:

- pexpect 3.2¹⁸⁹
- openbabel^{190,191} (optional)

Prior to running TurboGo or TurboControl, a valid installation of Turbomole must be available. On systems where computational modules must be loaded, Turbomole must have been loaded to the environment. Additionally, running the Turbomole environment configuration is recommended but not required prior to launching TurboGo or TurboControl:

```
$ source $TURBODIR/Config_turbo_env
```

E.3 TurboGo

TurboGo is a script fun on an input file. It generates the inputs required for Turbomole jobs, and submits the job to the GridEngine queue before quitting. TurboGo is run with the following syntax:

```
$ turbogo [-h] [-v] [-q] file
```

positional arguments:

file	Read input from gaussian-type input FILE.
------	---

More info on the input files is available below.

optional arguments:

- **-h, --help** Show this help message and exit
- **-v, --verbose** Run more verbose (show debugging info)
- **-q, --quiet** Run less verbose (show only warnings)

TurboGo saves a log file (`turbogo.log`) in the directory in which it is run. A second logfile (`define.log`) will remain if the setup crashes or is terminated at some points, or if the script is run verbose.

TurboGo writes the final coordinates to `final_geometry.xyz`. If openbabel is installed, it will also write `finalgeom.mol`. The entire optimization is written to `optimization.xyz` for viewing with a molecular viewer, such as vmd.

E.4 TurboControl

TurboControl is a management script called from a parent directory containing sub directories of input files. Each input file must be in its own directory. The input file format must be the same as the input format for TurboGo (listed above), with the extension ‘.in’, ‘.inp’, ‘.input’, ‘.com’, or ‘.gjf’. TurboControl reads the inputs and submits the jobs to the computational cluster queue. It then monitors running jobs to determine when the script has finished. If the job is an Opt-Freq, it prepares the frequency analysis and resubmits to the queue. TurboControl analyzes completed Opt-Freq jobs for true optimization, and attempts to re-run jobs with modified geometries when Transition States are found. TurboControl will not get stuck on the same transition state, but will return a ‘stuck’ job. TurboControl is run with the following syntax:

```
$ turbocontrol [-h] [-v/-q] [-s]
```

Optional arguments:

- **-h, --help** Show this help message and exit
- **-v, --verbose** Run more verbose (show debugging info)
- **-q, --quiet** Run less verbose (show only warnings)
- **-s, --solvent** List available solvents for COSMO and quit

TurboControl outputs information every 3 hours on the status of the jobs. It writes a logfile (`turbocontrol.log`) and may or may not leave other log files in each directory (depending on verbosity level). Ends when the last job finishes or crashes. Requires 1 node or can be run on headnode (minimal resource consumption especially after initial job preparation and submission.)

TurboControl assists with analysis by outputting a `stats.txt` file as jobs complete. This file contains file details, optimization and frequency timing details, energy, and the first frequency. Additional information can be requested by including the `freeh` keyword (see below).

E.5 Input File Format

The input file format is similar to that well known by Gaussian users. A series of keywords, one per line and indicated by a ‘%’, is followed by the ‘route card’ (specific job information). Charge and spin is indicated, then the molecule is shown in Cartesian format. This is followed by optional modifications to the Turbomole Control file. Note the location of blank lines in the example (Section 5.7).

E.5.1 Keywords

Keywords are as follows:

- **%nproc** - number of processors to use for the calculation job.
 - Synonym: **%processors**
- **%arch** - parallelization architecture to use for the job.
 - Synonyms: **%architecture**, **%para_arch**
- **%maxcycles** - number of optimization iterations before failing.
- **%autocontrolmod** - DEFAULT - modify the **control** file to include optimizations to speed up the job.
- **%nocontrolmod** - do not modify **control** file as above.
- **%rt** - specify max expected runtime (for any part of job)in hours. Allows backfilling in gridengine queue to speed up job submission. For example, for a 1 hour opt and 4 hour freq, submit at least a **rt** of 4
- **%cosmo** - use turbomole's COSMO solvation model with the specified solvent or **None** to use the idealized solvent (epsilon = infinity). List of available solvents can be shown by running **turbocontrol -s**

Gaussian args, including **%nosave**, **%rwf=[file]**, **%chk=[file]**, and **%mem=[memory]** are silently ignored.

E.5.2 Route Card Options

Route cards take the form of the following:

```
# [jobtype(s)] [joboption(s)]
```

Job types available:

- **opt** - Perform a geometry optimization
- **freq** - Perform a frequency analysis. Specify method via numforce or aoforce.
default = numforce
- **sp** - Perform a single point energy calculation.
 - Cannot be combined with Opt or Freq
- **ts** - Perform a transition state search to find 1 imaginary vibration.
 - Cannot be combined with Opt or Freq
- **prep** - Prepare the job but do not submit to queue.
 - Cannot be combined with Opt or Freq

Job options available:

- **ri** - Use Turbomole's ri approximation
- **marij** - Use Turbomole's marij approximation
 - Requires **ri**
- **disp** - Use Turbomole's implementation of Grimme's dispersion, version 3
- **aoforce** - Use aoforce for frequency jobs

- **numforce** - Use numforce for frequency jobs
- **freeh** - Use Turbomole's **freeh** thermodynamics data script to extract thermodynamic information after frequency analysis

E.5.3 Title

Following the Route cards, a blank line is added, then a line containing the title of the calculation. This can include any characters, spaces, etc., remaining on only one line. This is followed by a blank line.

E.5.4 Charge and Spin

Charge and spin are listed as two numbers separated by a space: charge spin (eg:0 1)

E.5.5 Geometry

Geometry in xyz coordinate format: Element xcoord ycoord zcoord. Z-matrix geometry is not supported by TurboControl or TurboGo.

E.5.6 Additional control File Modifications

Additional lines to be added or removed from control. Lines automatically added are, as required,:

```
$ricore 0
$paroptions ga_memperproc 9000000000000000 900000000000
$parallel_parameters maxtask=10000
$ricore_slave 1
$maxcor 2048
```

Additional lines may be added, or lines removed, by placing them after the geometry with a \$ (for addition) or -\$ (for removal).

E.5.7 Example Input File

An example input file for benzene in dmf:

```
%nproc=4
%arch=GA
%maxcycles=250
%rt=6
%cosmo=dmf
# opt freq b3-lyp/def2-TZVP ri marij numforce

Benzene Optimization & Frequency

O 1
C 0.000 1.396 0.000
C 1.209 0.698 0.000
C 1.209 -0.698 0.000
C 0.000 -1.396 0.000
C -1.209 -0.698 0.000
C -1.209 0.698 0.000
H 0.000 2.479 0.000
H 2.147 1.240 0.000
H 2.147 -1.240 0.000
H 0.000 -2.479 0.000
H -2.147 -1.240 0.000
H -2.147 1.240 0.000

$disp
-$paraoptions
```

E.6 Code Details

Coverage percentages of code unitests are listed in Table E.1. Results are low for def_op, screwer_op, cosmo_op, freeh_op, turbocontrol, and turbogo because they contain many lines of interacting with GridEngine or TurboMole. Testing is performed via monitoring the status of the scripts as they run in real conditions.

The code style is graded by PyLint and results are shown in Table E.2. PyLint describes coding style, adherence to guidelines, and readability. It does not describe code efficiency or usefulness.

Table E.1 Test Coverage of scripts in Turbocontrol

Name	Statements	Missing	Excluded	Coverage
cosmo_op	106	70	1	34%
cosmo_op_test	17	1	0	94%
def_op	302	226	1	25%
def_op_test	20	1	0	95%
freeh_op	162	55	1	66%
freeh_op_test	27	1	0	96%
screwer_op	71	25	1	65%
screwer_op_test	11	1	0	91%
test_all	18	0	0	100%
turbocontrol	537	319	0	41%
turbocontrol_test	245	24	0	90%
turbogo	343	132	0	62%
turbogo_helpers	383	52	0	86%
turbogo_helpers_test	274	2	0	99%
turbogo_test	98	1	0	99%
TOTAL	2614	910	4	65%

Table E.2 PyLint Scores for Turbocontrol Code

File	Score /10
test_all.py	2.22
turbogo.py	8.80
turbogo_test.py	6.97
turbocontrol.py	8.55
turbocontrol_test.py	7.18
turbogo_helpers.py	8.81
turbogo_helpers_test.py	7.45
def_op.py	8.18
def_op_test.py	5.71
screwer_op.py	7.36
screwer_op_test.py	6.67
freeh_op.py	8.71
freeh_op_test.py	6.79
cosmo_op.py	8.22
cosmo_op_test.py	6.67

E.7 Citing TurboControl

TurboControl, Turbogo, or any other parts of this code may be cited as: Bulsink, Philip. TurboControl, v. 1.1.0. <http://github.org/pbulsink/turbocontrol> (accessed June 2014) Change the version number to match the version that you used, and change the accessed date to when you installed or downloaded TurboControl.

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RIGHT HOLDERS BE LIABLE FOR ANY CLAIM, DAMAGES OR OTHER LIABILITY, WHETHER IN AN ACTION OF CONTRACT, TORT OR OTHERWISE, ARISING FROM, OUT OF OR IN CONNECTION WITH THE SOFTWARE OR THE USE OR OTHER DEALINGS IN THE SOFTWARE.

Glossary of Terms

CCDC Cambridge Crystallography Data Centre

COSMO Conductor-like Screening Model

DFT Density Functional afd

DMF N,N-dimethylformamide

DMSO Dimethylsulfoxide

FTIR Fourier Transform Infrared

GC Gas Chromatography

GUI Graphical User Interface

HOMO Highest Occupied Molecular Orbital

IRC Intrinsic Reaction Coordinate

LUMO Lowest Unoccupied Molecular Orbital

MLCT Metal-Ligand Charge Transfer

MO Molecular Orbital

NMR Nuclear Magnetic Resonance

PCM Polarizable Continuum Model

RWGSR Reverse Water-Gas Shift Reaction

TCD Thermal Conductivity Detector

TD-DFT Time Dependant Density Functional Theorem

TEA Triethylamine

TEOA Triethanolamine

TGA Thermogravimetric Analysis

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Todo list

lim2014 is ASAP	6
zeng2014 is ASAP, get actual once published	6
REMAKE FIG WITH CONCENTRATIONS PUT IN	34
Darrin, I'm still working on this section (discussed this morning)	38
This red product was analyzed via UV-Vis spectroscopy in the mother solution.	
Blank solvent, solvent with Et ₄ NCl electrolyte, and a freshly prepared reaction mixture with catalyst were used as blank solutions to attempt to identify spectral characteristics	39
specific background	39
UV-Vis, and overlay into figures, and discuss	40
add last struct	49
maybe take this section out? covered in rationalization section ch3	63
Darrin, I'm still working on this section (discussed this morning)	64
expand	70