FYS2160 - Oblig 2

a. Partition function, $Z$, for the system

$i = 1, 2, 3, 4$

Energies: $\varepsilon_i : \varepsilon_1 = \varepsilon, \ \varepsilon_2 = \varepsilon_3 = 3_4 = 2\varepsilon$

$$Z = \sum_i^4 e^{-\frac{\varepsilon_i}{kT}} = e^{-\frac{\varepsilon_1}{kT}} + e^{-\frac{\varepsilon_2}{kT}} + e^{-\frac{\varepsilon_3}{kT}} + e^{-\frac{\varepsilon_4}{kT}}$$

$$= e^{-\frac{\varepsilon}{kT}} + e^{-\frac{2\varepsilon}{kT}} + e^{-\frac{2\varepsilon}{kT}} + e^{-\frac{2\varepsilon}{kT}}$$

$$= e^{-\frac{\varepsilon}{kT}} + 3e^{-\frac{2\varepsilon}{kT}} \qquad \boxed{-2\frac{\varepsilon}{kT} = -\frac{\varepsilon}{kT} - \frac{\varepsilon}{kT}}$$

$$= e^{-\frac{\varepsilon}{kT}} + 3e^{-\frac{\varepsilon}{kT}} e^{-\frac{\varepsilon}{kT}}$$

$$= e^{-\frac{\varepsilon}{kT}}\left(1 + 3e^{-\frac{\varepsilon}{kT}}\right)$$

b.

$$\bar{E} = \sum_i \varepsilon_i \frac{1}{Z} e^{-\frac{\varepsilon_i}{kT}} = \frac{1}{Z}\left(\varepsilon_1 e^{-\frac{\varepsilon_1}{kT}} + \varepsilon_2 e^{-\frac{\varepsilon_2}{kT}} + \varepsilon_3 e^{-\frac{3_2}{kT}} + \varepsilon_4 e^{-\frac{3_4}{kT}}\right)$$

$$= \frac{1}{Z}\left(\varepsilon e^{-\frac{\varepsilon}{kT}} + 6\varepsilon e^{-\frac{2\varepsilon}{kT}}\right) = \frac{1}{Z}\left(\varepsilon e^{-\frac{\varepsilon}{kT}} + 6\,\varepsilon e^{-\frac{\varepsilon}{kT}}\,\varepsilon e^{-\frac{\varepsilon}{kT}}\right)$$

$$= \frac{\varepsilon \cancel{e^{-\frac{\varepsilon}{kT}}}\left(1 + 6e^{-\frac{\varepsilon}{kT}}\right)}{\cancel{e^{-\frac{\varepsilon}{kT}}}\left(1 + 3e^{-\frac{\varepsilon}{kT}}\right)} = \varepsilon\,\frac{1 + 6e^{-\frac{\varepsilon}{kT}}}{1 + 3e^{-\frac{\varepsilon}{kT}}}$$

c. Heat capacity as a function of temperature.

$$C_V = \left(\frac{\partial E}{\partial T}\right)_{N,V} \rightarrow \frac{\partial}{\partial T}\, \varepsilon\left(\frac{1+6e^{-\frac{\varepsilon}{kT}}}{1+3e^{-\frac{\varepsilon}{kT}}}\right)$$

$$u = 1+6e^{-\frac{\varepsilon}{kT}} \qquad du = \frac{6\varepsilon}{kT^2}\, e^{-\frac{\varepsilon}{kT}}$$

$$v = 1+3e^{-\frac{\varepsilon}{kT}} \qquad dv = \frac{3\varepsilon}{kT^2}\, e^{-\frac{\varepsilon}{kT}}$$

$$\boxed{\frac{du\cdot v - u\cdot dv}{v^2}}$$

$$\rightarrow \varepsilon\left(\frac{\left(\frac{6\varepsilon}{kT^2}e^{-\frac{\varepsilon}{kT}}\right)\left(1+3e^{-\frac{\varepsilon}{kT}}\right) - \left(1+6e^{-\frac{\varepsilon}{kT}}\right)\left(\frac{3\varepsilon}{kT^2}e^{-\frac{\varepsilon}{kT}}\right)}{\left(1+3e^{-\frac{\varepsilon}{kT}}\right)^2}\right)$$

$$\varepsilon\left(\frac{\frac{6\varepsilon}{kT^2}e^{-\frac{\varepsilon}{kT}} + \frac{6\varepsilon}{kT^2}e^{-\frac{\varepsilon}{kT}}\,3e^{-\frac{\varepsilon}{kT}} - \left(\frac{3\varepsilon}{kT^2}e^{-\frac{\varepsilon}{kT}} + \frac{3\varepsilon}{kT^2}e^{-\frac{\varepsilon}{kT}}\,6e^{-\frac{\varepsilon}{kT}}\right)}{\left(1+3e^{-\frac{\varepsilon}{kT}}\right)^2}\right)$$

$$\varepsilon\left(\frac{\frac{6\varepsilon}{kT^2}e^{-\frac{\varepsilon}{kT}} + \cancel{\frac{18\varepsilon}{kT^2}e^{-\frac{2\varepsilon}{kT}}} - \frac{3\varepsilon}{kT^2}e^{-\frac{\varepsilon}{kT}} - \cancel{\frac{18}{kT^2}e^{-\frac{2\varepsilon}{kT}}}}{\left(1+3e^{-\frac{\varepsilon}{kT}}\right)^2}\right)$$

$$= \frac{\frac{3\varepsilon}{kT^2}e^{-\frac{\varepsilon}{kT}}}{\left(1+3e^{-\frac{\varepsilon}{kT}}\right)^2} = \frac{\frac{3\varepsilon e^{-\frac{\varepsilon}{kT}}}{kT^2}}{\left(1+3e^{-\frac{\varepsilon}{kT}}\right)^2} = \frac{3\varepsilon e^{-\frac{\varepsilon}{kT}}}{kT^2\left(1+3e^{-\frac{\varepsilon}{kT}}\right)^2}$$

d. $$Z_R = \sum_{j=0}^{\infty} (2j+1) e^{-\beta E(j)} \quad , \quad \beta = \frac{1}{kT}$$

$$= \sum_{j=0}^{\infty} \underbrace{(2j+1)}_{g(j)} e^{-\beta j(j+1)\varepsilon}$$

e. See script below. Figure 1 shows the plot.

f. When the number of terms that contributes significantly to the partition function, we can write it as an integral

$$Z_R \simeq \int_0^{\infty} (2j+1) e^{-\beta j(j+1)\varepsilon} \, dj$$

In the high-temperature limit, where $Z_R \gg 1$, we can evaluate the integral to show $Z_R(T) = T/\theta_r$

We make the substitution

$$x = j(j+1) = j^2 + j \quad \rightarrow dx = 2j+1$$

$$\rightarrow Z_R = \int_0^{\infty} e^{-\beta x \varepsilon} dx \quad , \quad \beta = \frac{1}{kT} \quad , \quad \varepsilon = \theta_R k$$

$$\rightarrow Z_R = \int_0^{\infty} e^{-\frac{x\theta_R}{T}} dx = -\frac{T}{\theta_R} e^{-\frac{x\theta_R}{T}} \Big|_0^{\infty} = \frac{T}{\theta_R}$$

g. For low temperatures, only a few terms is enough. I'll include three:

$$Z_R = \sum_j (2j+1) e^{-j(j+1)\frac{\theta_r}{T}}$$

$$\rightarrow (2(0)+1)e^{-0} + (2(1)+1)e^{-1(1+1)\frac{\theta_r}{T}} + (2(2)+1)e^{-2(2+1)\frac{\theta_r}{T}}$$

$$= 1 + 3e^{-2\frac{\theta_r}{T}} + 5e^{-6\frac{\theta_r}{T}}$$

h. Energy $E(T)$ of the system for high and low $T$

<u>High T:</u> $(T \gg \theta_r)$

$$U = -\frac{\partial \ln(Z)}{\partial \beta} = kT^2 \frac{\partial \ln Z}{\partial T} = kT$$

Which corresponds to $\frac{1}{2}kT$ per degree of freedom (diatomic molecules has two degrees of freedom)

<u>Low T:</u> $(T \ll \theta_r)$

$$\ln(Z) = \ln\left(1 + 3e^{-2\frac{\theta_r}{T}}\right)$$

$$\overset{\text{Taylor}}{\simeq} \quad 3e^{-2\frac{\theta_r}{T}}$$

$$\frac{\partial \ln(Z)}{\partial T} = \frac{\partial}{\partial T} 3e^{-\frac{2\theta_r}{T}} = 3e^{-2\frac{\theta_r}{T}} \cdot \left(\frac{2\theta_r}{T^2}\right)$$

$$= 3\left(\frac{2\theta_r}{T^2}\right)e^{-\frac{2\theta_r}{T}} = \frac{6\theta_r}{T^2} e^{-\frac{2\theta_r}{T}}$$

$$U = kT^2 \frac{\partial \ln(Z)}{\partial T} = k\cancel{T^2} \frac{6\theta_r}{\cancel{T^2}} e^{-\frac{2\theta_r}{T}}$$

$$= k 6\theta_r e^{-\frac{2\theta_r}{T}}$$

i. Heat capacity $C_v(T)$ for high and low T.

Heat capacity is found by taking the derivative of the energy

High $\left(\frac{T}{\theta_r} >> 1\right)$

$$E(T) = kT \rightarrow \frac{dE}{dT} = k$$

Low $\left(\frac{T}{\theta_r} << 1\right)$

$$E(T) = k6\theta_r e^{-\frac{2\theta_r}{T}}$$

$$\rightarrow \frac{dE}{dT} = k6\theta_r e^{-\frac{2\theta_r}{T}}\left(\frac{2\theta_r}{T^2}\right)$$

$$= 12k\left(\frac{\theta_r}{T}\right)^2 e^{-\frac{2\theta_r}{T}}$$

j. Code is listed below.

k. Figure 2 shows the plot.

l. From the plot, we can see that $Z$ is linear for large values, but lies slightly below the graph where $Z = T/\theta_r$. This happens because the integral produces an error, which underestimates the partition function sum.

m. We can find the energy by

$E = kT^2 \frac{\partial \ln Z}{\partial T}$ If we make the substitution $u = T/\theta_r$

We can write

$$E = k(u\theta_r)^2 \frac{\partial Z}{\partial T} = k u^2 \theta_r^2 \frac{\partial \ln Z}{\partial u} \cdot \frac{\partial u}{\partial T}$$

$$= k u^2 \theta_r \frac{\partial \ln Z}{\partial u}$$

If we know $Z(N, V; u)$ numerically, we can find the energy $E$.

We can find the heat capacity by

$$C_V = \frac{\partial E}{\partial T} = \frac{\partial E}{\partial u} \cdot \frac{\partial u}{\partial T} = \frac{1}{\Theta_r} \frac{\partial E}{\partial u}$$

$$= k \frac{\partial u^2}{\partial u} \cdot \frac{\partial \ln z}{\partial u}$$

h. See plot below.

# Code

## Problem c)

```
N = 100;

T = linspace(.1,20, N);

C_v = zeros(N);

for i = 1:N-1
    C_v(i) = (3*exp(1/T(i))) / ((T(i)^2)*(1+3*exp(1/T(i)))^2);
end

plot(T, C_v)
xlabel('T');
ylabel('C_v');

pause()
```

## Problem e)

```
N = 100;

Ttheta = [0.01 0.025 0.1 0.5 0.75 1.0];
C = ['-r' '-g' '-b' 'c' 'm' 'y'];

F = zeros(N);

for i=1:length(Ttheta)
    j = linspace(1,20,N);
    F = (2*j + 1).*exp(-j.*(j + 1)*Ttheta(i));
    plot(j, F, C(i))
    hold on
end

xlabel('j')
ylabel('T/\theta_r')
legend('T/\theta=0.01','T/\theta0.025', 'T/\=theta0.1','T/\theta=0.5','T/\theta=0.75','T/\theta=1.0')

pause()
```

## Problem j)

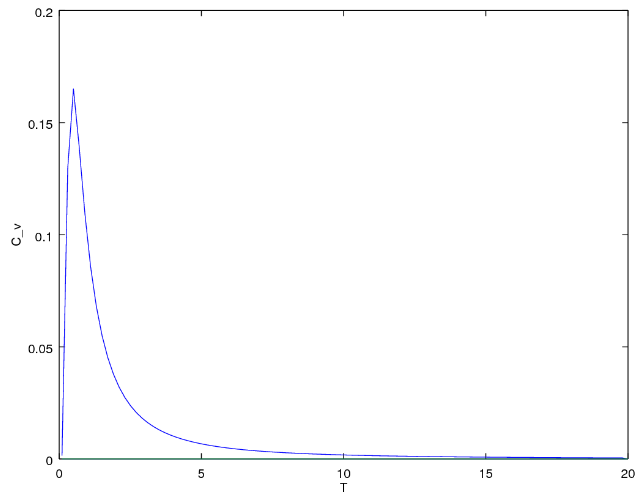
```
u = (0.01:0.01:3.0);

Z = zeros(length(u),1);

j = linspace(0,1000, 1000);

for i = 1:length(u)
    Z(i) = sum( (2*j + 1).*exp(-(j.*(j + 1))/u(i)));
end

plot(u,Z,'-b');
```

```
hold on
plot(u,u,'-r');
xlabel('T/\theta_r');
ylabel('Z');

pause();
```

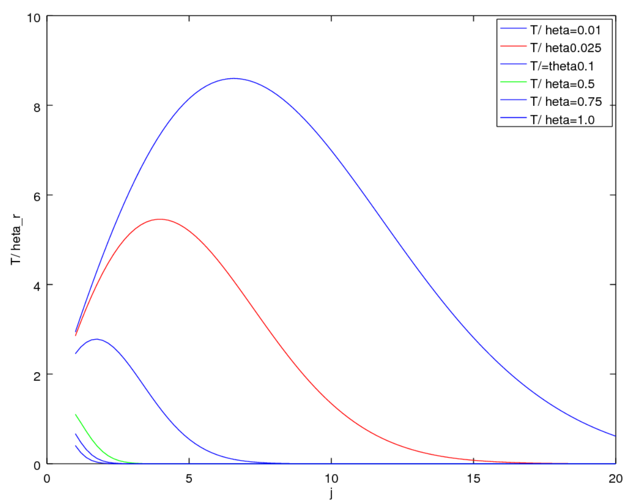
## Problem n)

```
% Calculate partition function Z
u = (0.01:0.01:3.0);

Z = zeros(length(u),1);

j = linspace(0,1000, 1001);

for i = 1:length(u)
    Z(i) = sum( (2*j + 1).*exp(-(j.*(j + 1))/u(i)));
end

% User Z as partition function
du = u(2:end)-u(1:end-1);
u_1 = (u(1:end-1)+u(2:end))*0.5;

logZ = log(Z);

U = u_1.^2.*diff(logZ)'./du;
ddu = 0.5*(du(1:end-1)+du(2:end));
CV = diff(U)./ddu;
u2 = u(2:end-1);

%plot(u_1,U);
%xlabel('u')
%ylabel('U(u)')

plot(u2,CV);
xlabel('u')
ylabel('C_V(u)');

pause();
```

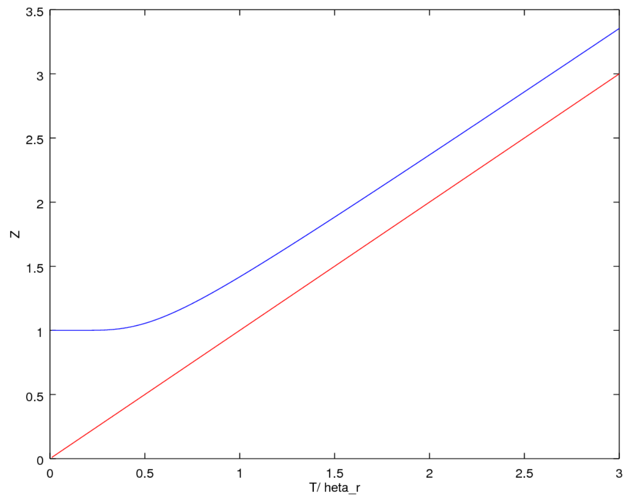
# Plot



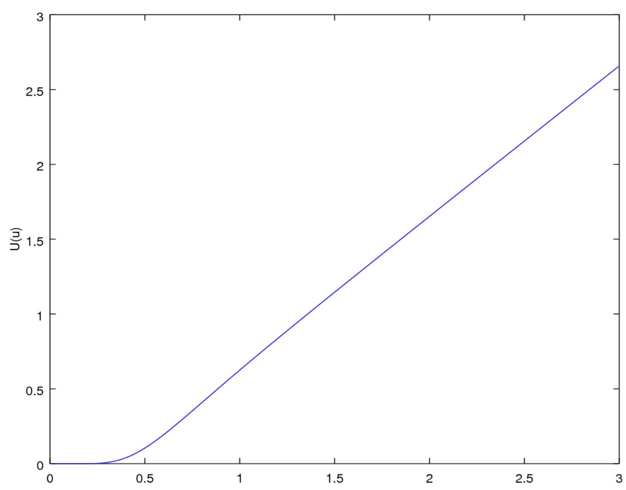

Figur 1: Plot from problem c)



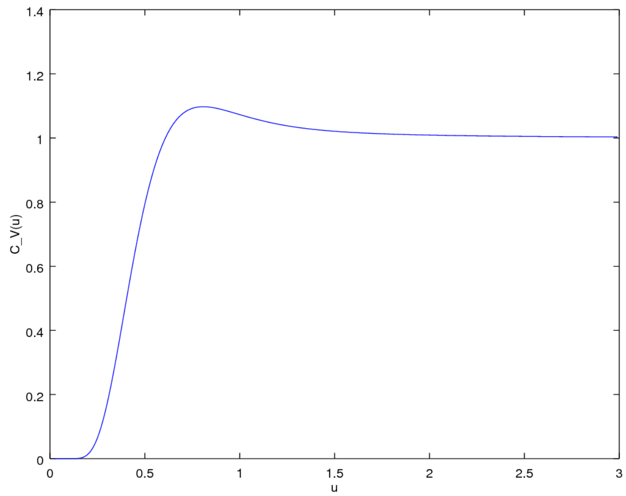

Figur 2: Plot from problem e)




Figur 3: Plot from problem j)




Figur 4: Plot from problem n)




Figur 5: Plot from problem n)