PART-1 : PHYSICS

SECTION-I (i)

1) A bar of mas $M = 1.00$ kg and length $L = 0.20$ m is lying on a horizontal frictionless surface. A small mass $m = 0.10$ kg is moving on the same horizontal surface with 5.00 m s^{-1} speed on a path perpendicular to the bar. It hits the bar at a distance $L/4$ from the left end and returns back on the same path with speed v. After this elastic collision, the bar rotates with an angular velocity ω. Which of the following statement is correct?

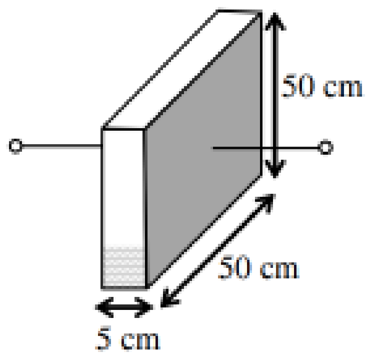
(A) $\omega = \frac{600}{47}$ rad s^{-1} and $v = \frac{165}{47}$ m s^{-1}

(B) $\omega = \frac{300}{47}$ rad s^{-1} and $v = \frac{165}{47}$ m s^{-1}

(C) $\omega = \frac{200}{23}$ rad s^{-1} and $v = \frac{51}{23}$ m s^{-1}

(D) $\omega = \frac{800}{23}$ rad s^{-1} and $v = \frac{51}{23}$ m s^{-1}

2) A container has a base of 50 cm × 5 cm and height 50 cm, as shown in the figure. It has two parallel electrically conducting walls each of area 50 cm × 50 cm. The remaining walls of the container are thin and non-conducting. The container is being filled with a liquid of dielectric constant 3 at a uniform rate of 250 $cm^3\ s^{-1}$. What is the value of the capacitance of the container after 20 seconds ? [Given: Permittivity of free space $\epsilon_0 = 9 \times 10^{-12}\ C^2\ N^{-1}\ m^{-2}$, the effects of the non-conducting walls on the capacitance are negligible]



(A) 27 pF

(B) 63 pF

(C) 81 pF

(D) 135 pF

3) One mole of an ideal gas expands adiabatically from an initial state (P_0, V_0) to final state (P_A, $5V_0$). Another mole of the same gas expands isothermally from a same initial state (P_0, V_0) to the vol. (P_B, $5V_0$). The ratio of the specific heats at constant pressure and constant volume of this ideal gas is γ.

What is the ratio $\frac{P_A}{P_B}$?

(A) $5^{\gamma-1}$

(B) $5^{1-\gamma}$

(C) 5^{γ}

(D) $5^{1+\gamma}$

4) Two satellites P and Q are moving in different circular orbits around the Earth (radius R). The heights of P and Q from the Earth surface are h_P and h_Q, respectively, where h_p = R/3. The accelerations of P and Q due to Earth's gravity are g_P and g_Q, respectively. If $g_P/g_Q = 36/25$, what is the velocity of Q? (M = Mass of earth, G = Universal gravitational constant)

(A) $\sqrt{\frac{5GM}{3R}}$

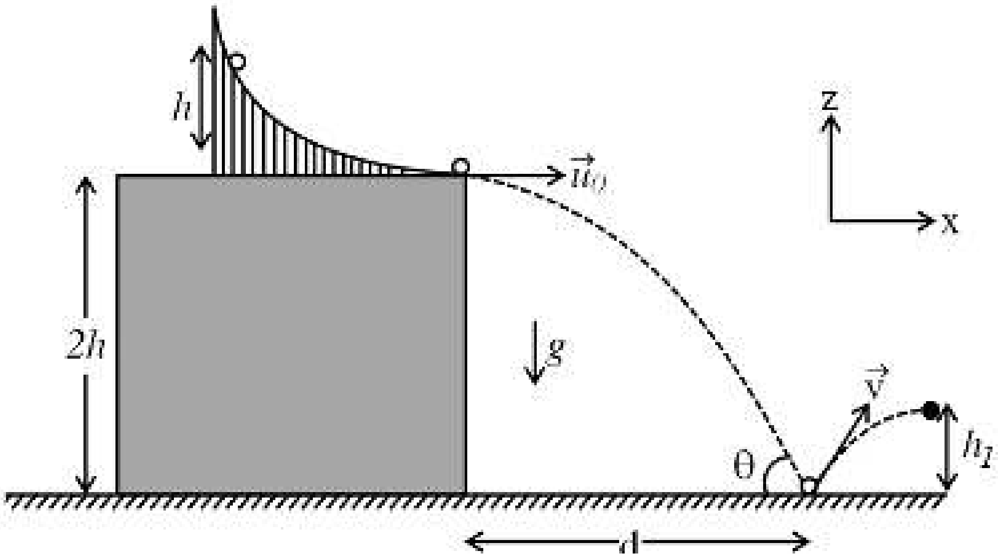
(B) $\sqrt{\frac{5GM}{8R}}$

(C) $\sqrt{\frac{5GM}{6R}}$

(D) $\sqrt{\frac{6GM}{7R}}$

SECTION-I (ii)

1) A slide with a frictionless curved surface, which becomes horizontal at its lower end,, is fixed on the terrace of a building of height 2h from the ground, as shown in the figure. A spherical ball of mass m is released on the slide from rest at a height h from the top of the terrace. The ball leaves the slide with a velocity $\vec{u}_0 = u_0\hat{x}$ and falls on the ground at a distance d from the building making an angle θ with the horizontal. It bounces off with a velocity $\vec{V}$ and reaches a maximum height h_1. The acceleration due to gravity is g and the collision is elastic. Which of the following statement(s) is(are) correct ? [Take h = 20 m, g = 10 m/s^2]. The ground is frictionless.



(A) $\vec{u}_0 = 20\hat{i}$ m/s

(B) $\vec{v} = \left(20\hat{i} + 20\sqrt{2}\hat{k}\right)$ m/s

(C) $\theta = 60°$

(D) $\frac{d}{h_1} = \sqrt{2}$

2) A plane polarized blue light ray is incident on a prism such that there is no reflection from the surface of the prism. The angle of deviation of the emergent ray is $\delta = 60°$ (see Figure-1). The angle of minimum deviation for red light from the same prism is $\delta_{min} = 30°$ (see Figure-2). The refractive index of the prism material for blue light is $\sqrt{3}$

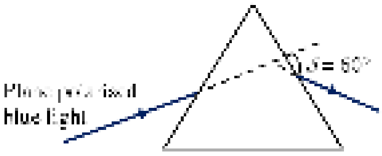


Figure-1

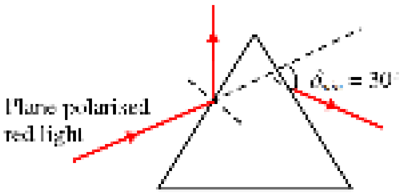


Figure-2

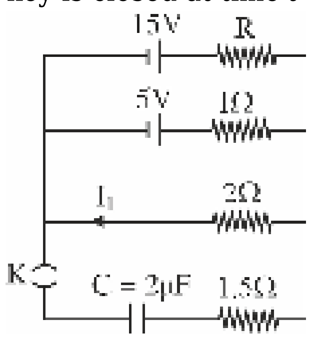
(A) The blue light has its electric field in the plane of incidence

(B) The angle of the prism is 60°.

(C) The refractive index of the material of the prism for red light is $\sqrt{2}$.

(D) The angle of refraction for blue light in air at the exit plane of the prism is 45°.

3) In a circuit shown in the figure, the capacitor C is initially uncharged and the key K is open. In this condition, a current of 3A flows through 1 Ω resistor. The key is closed at time $t = t_0$. Which of the following statement(s) is(are) correct? [Given: $e^{-1} \approx 0.35$]



(A) The value of the resistance R is 1Ω.

(B) Long after the switch is closed, value of current I_1 is 4 A.

(C) At $t = t_0 + 8$ μs, the current in the capacitor branch is approxmately 0.7 A.

(D) For $t \to \infty$, the charge on the capacitor is 16μC.

SECTION-I (iii)

1) List-I shows different radioactive decay processes and List-II provides possible emitted particles. Match each entry in List-I with an appropriate entry from List-II, and choose the correct option.

	List-I		List-II
(P)	${}^{238}_{92}U \rightarrow {}^{234}_{91}Pa$	(1)	three β^- particles and one α particle
(Q)	${}^{214}_{82}Pb \rightarrow {}^{210}_{82}Pb$	(2)	two β^- particles and one α particle
(R)	${}^{210}_{81}Tl \rightarrow {}^{206}_{82}Pb$	(3)	one α particle and one β^+ particle
(S)	${}^{228}_{91}Pa \rightarrow {}^{224}_{88}Ra$	(4)	one α particle and two β^+ particles
		(5)	one α particle and one β^- particle

(A) P → 5;Q → 2;R → 1;S → 3

(B) P → 4;Q → 3;R → 4;S → 2

(C) P → 5;Q → 3;R → 1;S → 4

(D) P → 5;Q → 2;R → 4;S → 1

2) Match the temperature of a black body given in List-I with an appropriate statement in List-II, and choose the correct option. [Given: Wien's constant as 2.9×10^{-3} m-K and $\frac{hc}{e} = 1.24 \times 10^{-6}$ V-m]

	List-I		List-II
(P)	2000 K	(1)	The radiation at peak emission wavelength will result in the widest central maximum of a single slit diffraction.
(Q)	3000 K	(2)	The power emitted per unit area is 1/16 of that emitted by a blackbody at temperature 6000 K.
(R)	5000 K	(3)	The radiation at peak wavelength is visible to human eye.
(S)	10000 K	(4)	The radiation at peak emission wavelength can be used to image human bones.
		(5)	The radiation at peak wavelength can lead to emission of photoelectrons from a metal of work function 4 eV

(A) P → 4;Q → 2;R → 1;S → 5

(B) P → 3;Q → 2;R → 4;S → 1

(C) P → 1;Q → 2;R → 3;S → 5

(D) P → 1;Q → 2;R → 5;S → 3

3) A series LCR circuit is connected to a 45 sin (ωt) Volt source. The resonant angular frequency of the circuit is 10^5 rad s^{-1} and current amplitude at resonance is I_0.. When the angular frequency of the source is $\omega = 8 \times 10^4$ rad s^{-1}, the current amplitude in the circuit is 0.05 I_0. If L = 50 mH, match each entry in List-I with an appropriate value from List-II and choose the correct option.

	List-I		List-II
(P)	I_0 in mA	(1)	2250
(Q)	The quality factor of the circuit	(2)	500

(R)	The bandwidth of the circuit in rad s^{-1}	(3)	18
(S)	The peak power dissipated at resonance in Watt	(4)	400
		(5)	44.4

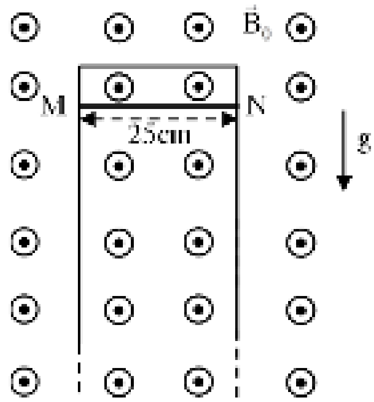
(A) P → 3;Q → 1;R → 2;S → 4

(B) P → 3;Q → 1;R → 4;S → 2

(C) P → 4;Q → 5;R → 3;S → 1

(D) P → 4;Q → 5;R → 1;S → 3

4) A thin conducting rod MN of mass 20 gm, length 25 cm and resistance 10 Ω is held on frictionless, long, perfectly conducting vertical rails as shown in the figure. There is a uniform magnetic field B = 4 T directed perpendicular to the plane of the rod-rail arrangement. The rod is released from rest at time t = 0 and it moves down along the rails. Assume air drag is negligible. Match each quantity in List-I with an appropriate value from List-II, and choose the correct option. [Given : The



acceleration due to gravity $g = 10\ ms^{-2}$ and $e^{-1} = 0.4$]

	List-I		List-II
(P)	At t = 0.2 s, the magnitude of the induced emf in Volt	(1)	1.20
(Q)	At t = 0.2 s, the magnitude of the magnetic force in Newton	(2)	0.07
(R)	At t = 0.2 s, the power dissipated as heat in Watt	(3)	2.00
(S)	The magnitude of terminal velocity of the rod in m s^{-1}	(4)	0.12
		(5)	0.14

(A) P → 5;Q → 2;R → 3;S → 1

(B) P → 1;Q → 4;R → 5;S → 3

(C) P → 4;Q → 3;R → 1;S → 2

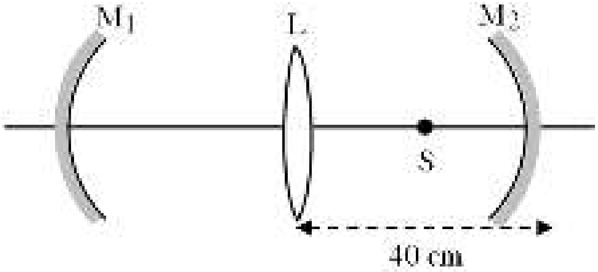
(D) P → 4;Q → 1;R → 5;S → 3

SECTION-II

1) A Hydrogen-like atom has atomic number Z. Photons emitted in the electronic transitions from level n = 5 to level n = 3 in these atoms are used to perform photoelectric effect experiment on a target metal. The maximum kinetic energy of the photoelectrons generated is 1.86 eV. If the photoelectric threshold wavelength for the target metal is 620 nm, the value of Z is _______. [Given: hc = 1240 eV-nm and Rhc = 13.6 eV, where R is the Rydberg constant, h is the Planck's constant

and c is the speed of light in vacuum]

2) An optical arrangement consists of two concave mirrors M_1 and M_2, and a convex lens L with a common principal axis, as shown in the figure. The focal length of L is 10 cm. The radii of curvature of M_1 and M_2 are 20 cm and 24 cm, respectively. The distance between L and M_2 is 40 cm. A point object S is placed at the mid-point between L and M_2 on the axis. When the distance between L and M_1 is x cm, one of the images coincides with S. The value of x is _______. (Consider at least two



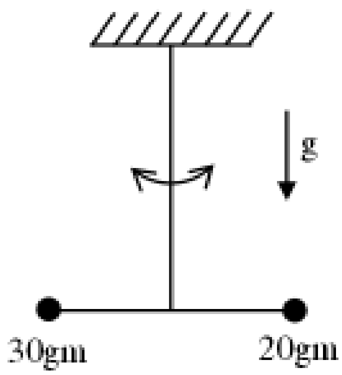
reflections)

3) In an experiment for determination of the focal length of a thin convex lens, the distance of the object from the lens is 10 ± 0.1 cm and the distance of its real image from the lens is 15 ± 0.1 cm. The error in the determination of focal length of the lens is _______ μm.

4) A closed container contains a homogeneous mixture of one mole of an ideal monatomic gas (γ = 5/3) and one mole of an ideal diatomic gas (γ = 7/5). Here, γ is the ratio of the specific heats at constant pressure and constant volume of an ideal gas. The gas mixture does a work of 66 Joule when heated at constant pressure. The change in its internal energy is ________ Joule.

5) A person of height 1.6 m is walking away from a lamp post of height 4 m along a straight path on the flat ground. The lamp post and the person are always perpendicular to the ground. If the speed of the person is 60 cm s^{-1}, the speed of the tip of the person's shadow on the ground with respect to the ground is _______ cm s^{-1}.

6) Two point-like objects of masses 20 gm and 30 gm are fixed at the two ends of a rigid massless rod of length 10 cm. This system is suspended vertically from a rigid ceiling using a thin wire attached to its center of mass, as shown in the figure. The resulting torsional pendulum undergoes small oscillations of amplitude 0.01 rad. The torsional constant of the wire is 1.2×10^{-8} N m rad^{-1}.



The maximum speed of 20 gm mass is _____ μ m/s.

PART-2 : CHEMISTRY

SECTION-I (i)

1) In the scheme given below, **X**, **Y** and **P** respectively, are

$$\text{Metal halide} \xrightarrow{\text{aq. NaOH}} \text{White precipitate } \mathbf{(P)} + \text{Filtrate } \mathbf{(Q)}$$

$$\mathbf{P} \xrightarrow[\text{heat}]{\text{aq. } H_2SO_4,\ PbO_2\text{(excess)}} \mathbf{X} \text{ (a coloured species in solution)}$$

$$\mathbf{Q} \xrightarrow[\text{warm}]{MnO(OH)_2,\ \text{Conc. } H_2SO_4} \mathbf{Y} \text{ (gives blue-coloration with KI-starch paper)}$$

(A) CrO_4^{2-} , Br_2 and MnO

(B) MnO_4^{2-} , Cl_2 and $Mn(OH)_2$

(C) MnO_4^{-} , Cl_2 and $Mn(OH)_2$

(D) $MnSO_4$, HOCl and MnO

2) Plotting $1/\Lambda_m$ (Y-axis) against $c\Lambda_m$ (X-axis) for aqueous solutions of a monobasic weak acid (HX) resulted in a straight line with y-axis intercept of P and slope of S. The K_a is
[Λ_m = molar conductivity
Λ_m^o = limiting molar conductivity
c = molar concentration
K_a = dissociation constant of HX]

(A) $\frac{P}{S}$

(B) $\frac{S}{P}$

(C) $\frac{P^2}{S}$

(D) $\frac{P}{S^2}$

3) On decreasing the pH from 7 to 2, the solubility of a sparingly soluble salt (MX) of a weak acid (HX) increased from 10^{-4} mol L^{-1} to 'x' mol L^{-1}. Then 'x' is ($K_a = 10^{-4}$) :

(A) 10^{-3}

(B) 10^{-4}

(C) 10^{-5}

(D) 10^{-2}

4) In the given reaction scheme, **P** is a phenyl alkyl ether, **Q** is an aromatic compound; **R** and **S** are the major products. $\mathbf{P} \xrightarrow{HI} \mathbf{Q} \xrightarrow[\text{(iii) } H_3O^+]{\text{(i) NaOH, (ii) } CO_2} \mathbf{R} \xrightarrow[\text{(ii) } H_3O^+]{\text{(i) } (CH_3CO)_2O} \mathbf{S}$ The correct statement about **S** is

(A) It primarily inhibits noradrenaline degrading enzymes.

(B) It inhibits the synthesis of prostaglandin.

(C) It is a narcotic drug.

(D) It is *ortho*-acetylbenzoic acid.

SECTION-I (ii)

1) The correct statement(s) related to processes involved in the extraction of metals is(are)

(A) Zone refining of metals is based on the principle of greater mobility of pure metal than that of impurity.

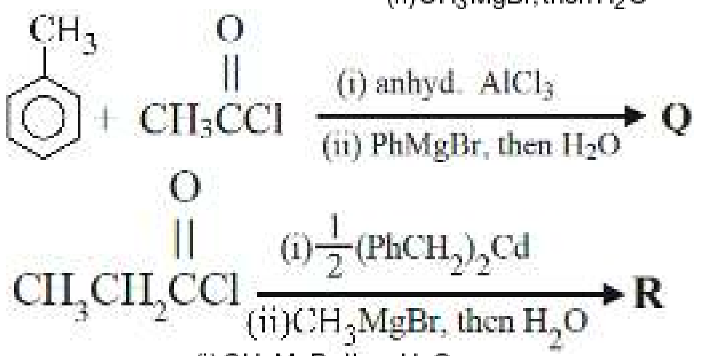
(B) Calcination of calamine produces sphalerite

(C) Copper pyrite is heated with silica in a reverberatory furnace to remove iron

(D) When Malachite is heated in a reverberatory furnace, cupric oxide is formed.

2) In the following reactions, **P**, **Q**, **R**, and **S** are the major products.

$$CH_3CH_2CH(CH_3)CH_2CN \xrightarrow[\text{(ii) } CH_3MgBr\text{, then } H_2O]{\text{(i) } CH_3MgBr\text{, then } H_3O^{\oplus}} \mathbf{P}$$



$$PhCH_2CHO \xrightarrow[\text{(iii) HCN; (iv) } H_2SO_4, \Delta]{\text{(i) } CH_3MgBr \text{ then } H_2O\text{; (ii) } CrO_3\text{, dill.} H_2SO_4} S$$

The correct statement(s) about **P**, **Q**, **R**, and **S** is(are)

(A) Both **P** and **Q** have asymmetric carbon(s).

(B) Both **Q** and **R** have asymmetric carbon(s).

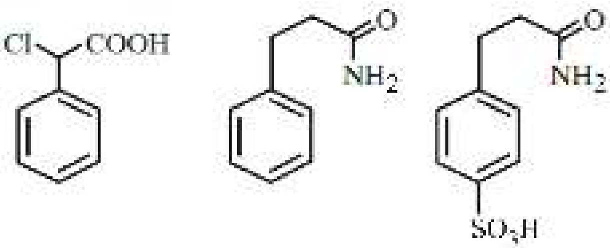
(C) Both **P** and **R** have asymmetric carbon(s).

(D) **P** has asymmetric carbon(s), **S** does **not** have any asymmetric carbon.

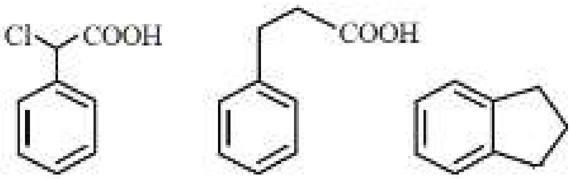
3) Consider the following reaction scheme and choose the correct option(s) for the major products **Q**, **R** and **S**.

$$\text{Styrene} \xrightarrow[\text{(ii) NaOH, } H_2O_2, H_2O]{\text{(i) } B_2H_6} \mathbf{P} \xrightarrow[\text{(ii) } Cl_2\text{, Red Phosphorus; (iii) } H_2O]{\text{(i) } CrO_3, H_2SO_4} \mathbf{Q}$$

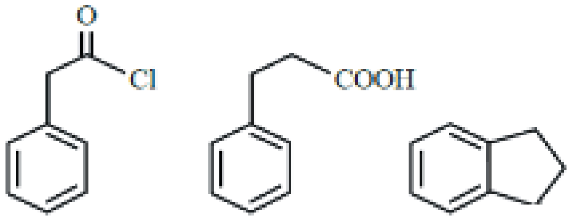
$$\mathbf{P} \xrightarrow[\text{(II) NaCN; (III) } H_3O^+, \Delta]{\text{(I)} SOCl_2} \mathbf{R} \xrightarrow[\text{(II)Zn–Hg/HCl}]{\text{(I)conc.} H_2SO_4} \mathbf{S}$$

(A) 

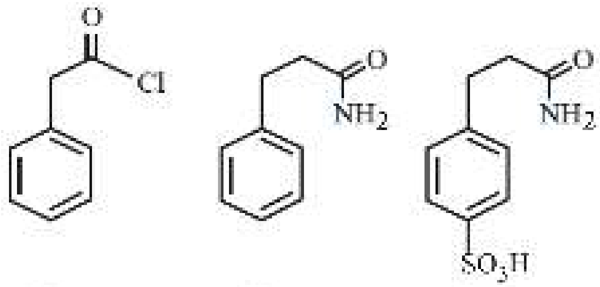
Q R S

(B) 

Q R S

(C) 

Q R S

(D) 

Q R S

SECTION-I (iii)

1) Match the reactions (in the given stoichiometry of the reactants) in List-I with one of their products given in List-II and choose the correct option.

List-I		List-II	
(P)	$PCl_3 + H_3PO_3 \rightarrow$	(1)	H_3PO_2
(Q)	Red P_4 + alkali $\rightarrow$	(2)	$H_4P_2O_5$
(R)	$C_2H_5OH + PCl_3 \rightarrow$	(3)	$H_4P_2O_6$
(S)	$H_3PO_2 + 2H_2O + 4AgNO_3 \rightarrow$	(4)	H_3PO_3
		(5)	H_3PO_4

(A) $P \rightarrow 2; Q \rightarrow 3; R \rightarrow 1; S \rightarrow 5$

(B) $P \rightarrow 3; Q \rightarrow 5; R \rightarrow 4; S \rightarrow 2$

(C) P → 5;Q → 2;R → 1;S → 3

(D) P → 2;Q → 3;R → 4;S → 5

2) Match the electronic configurations in List-I with appropriate metal complex ions in List-II and choose the correct option.
[Atomic Number: Fe = 26, Mn = 25, Co = 27]

List-I		List-II	
(P)	$t_{2g}^6 e_g^0$	(1)	$[NiF_6]^{-2}$
(Q)	$t_{2g}^6 e_g^3$	(2)	$[CoF_6]^{3-}$
(R)	$e^2 t_2^3$	(3)	$[Cu(H_2O)_6]^{2+}$
(S)	$t_{2g}^4 e_g^2$	(4)	$[FeCl_4]^-$
		(5)	$[CoCl_4]^{2-}$

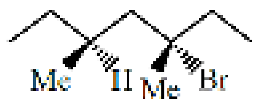
(A) P → 1;Q → 4;R → 2;S → 3

(B) P → 1;Q → 2;R → 4;S → 5

(C) P → 3;Q → 2;R → 5;S → 1

(D) P → 1;Q → 3;R → 4;S → 2

3) Match the reactions in List-I with the features of their products in List-II and choose the correct option.

List-I		List-II	
(P)	(-)-1-Bromo-2-ethylpentane (single enantiomer) $\xrightarrow[S_N2 \text{ reaction}]{\text{aq. NaOH}}$	(1)	Inversion of configuration
(Q)	(-)-2-Bromopentane (single enantiomer) $\xrightarrow[S_N2 \text{ reaction}]{\text{aq. NaOH}}$	(2)	Retention of configuration
(R)	(-)-3-Bromo-3-methylhexane (single enantiomer) $\xrightarrow[S_N1 \text{ reaction}]{\text{aq. NaOH}}$	(3)	Mixture of enantiomers
(S)	 (Single enantiomer) $\xrightarrow[S_N1 \text{ reaction}]{\text{aq. NaOH}}$	(4)	Mixture of structural isomers
		(5)	Mixture of diastereomers

(A) P → 1;Q → 2;R → 5;S → 3

(B) P → 2;Q → 1;R → 3;S → 5

(C) P → 1;Q → 2;R → 5;S → 4

(D) P → 2;Q → 4;R → 3;S → 5

4) The major products obtained from the reactions in List-II are the reactants for the named reactions mentioned in List-I. Match List-I with List-II and choose the correct option.

List-I		List-II	
(P)	Etard reaction	(1)	Acetophenone $\xrightarrow{\text{Zn–Hg, HCl}}$
(Q)	Gattermann reaction	(2)	Toluene $\xrightarrow[\text{(ii) SOCl}_2]{\text{(i) KMnO}_4\text{, KOH, }\Delta}$
(R)	Gattermann-Koch reaction	(3)	Benzene $\xrightarrow[\text{anhyd. AlCl}_3]{\text{CH}_3\text{Cl}}$
(S)	Rosenmund reduction	(4)	Aniline $\xrightarrow[\text{273–278 K}]{\text{NaNO}_2\text{/HCl}}$
		(5)	Phenol $\xrightarrow{\text{Zn, }\Delta}$

(A) P → 2;Q → 4;R → 1;S → 3

(B) P → 1;Q → 3;R → 5;S → 2

(C) P → 3;Q → 2;R → 1;S → 4

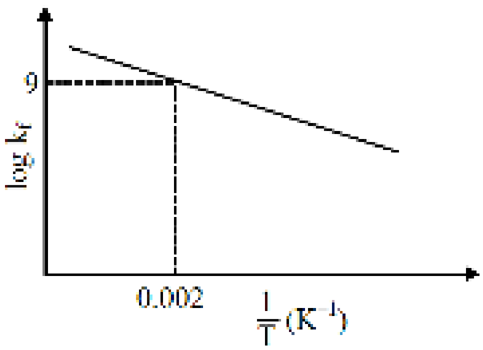
(D) P → 3;Q → 4;R → 5;S → 2

SECTION-II

1) The stoichiometric reaction of 516 g of dimethyldichlorosilane with water results in a tetrameric cyclic product **X** in 50% yield. The weight (in g) of **X** obtained is___.
[Use, molar mass (g mol^{-1}): H = 1, C = 12, O = 16, Si = 28, Cl = 35.5]

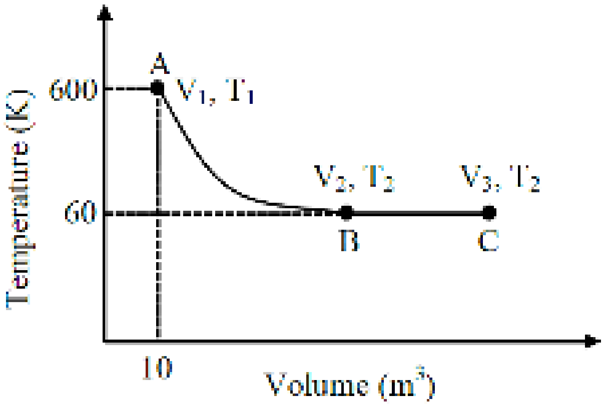
2) A gas has molar volume of 0.4 dm^3 mol^{-1} at a temperature of 800 K and pressure **x** atm. If it shows ideal gas behaviour at the same temperature and pressure, the molar volume is 0.8 dm^3 mol^{-1}. The value of **x** is ___.
[Use: Gas constant, $R = 8 \times 10^{-2}$ L atm K^{-1} mol^{-1}]

3) The plot of log k_f versus $^1/_T$ for a reversible reaction A (g) $\rightleftharpoons$ P (g) is shown.



Pre-exponential factors for the forward and backward reactions are 10^{15} s^{-1} and 10^{11} s^{-1}, respectively. If the value of log K_b for the reaction at 250 K is –5, the value of log K at 500 K is ______.
[K = equilibrium constant of the reaction
k_f = rate constant of forward reaction
k_b = rate constant of backward reaction]

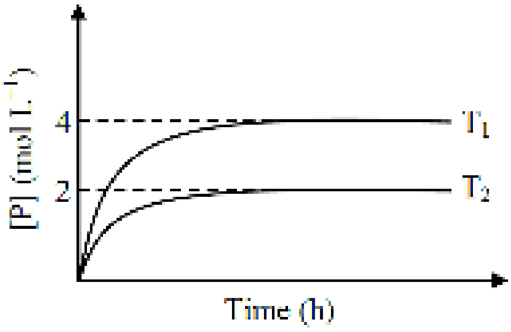
4) One mole of an ideal monoatomic gas undergoes two reversible processes (A → B and B → C) as



shown in the given figure : $A \rightarrow B$ is an adiabatic process. If the total entropy change in the entire process ($A \rightarrow B$ and $B \rightarrow C$) is R ln $\sqrt{10}$, the value of log V_3 is ________.

[Use, molar heat capacity of the gas at constant pressure, $C_{p,m} = \frac{5}{2}R$]

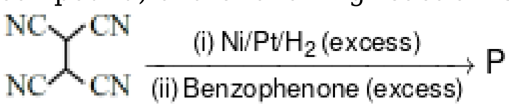
5) In a one-litre flask, 6 moles of A undergoes the reaction $A(g) \rightleftharpoons P(g)$. The progress of product formation at two temperatures (in Kelvin), T_1 and T_2, is shown in the figure:



0

If $T_1 = 2T_2$ and $\Delta H^0 = RT_2 \ln x$, then the value of x is ______.

6) The total number of sp^2 hybridised carbon atoms in the major product **P** (a non-heterocyclic compound) of the following reaction is _______.



PART-3 : MATHEMATICS

SECTION-I (i)

1) Let f : (0, 1) → ℝ be the function defined as f(x) = $\sqrt{n}$ if $x \in \left[\frac{1}{n+1}, \frac{1}{n}\right)$ where $n \in \mathbb{N}$.

Let g : (0, 1) → ℝ be a function such that $\int_{x^2}^{x} \sqrt{\frac{1-t}{t}}\,dt < g(x) < 2\sqrt{x}$ for all x∈(0, 1). Then $\lim_{x \to 0} f(x)g(x)$

(A) is equal to 2

(B) is equal to 1

(C) does **NOT** exist

(D) is equal to 3

2) Let Q be the cube with the set of vertices $\left\{(x_1, x_2, x_3) \in \mathbb{R}^3 : x_1, x_2, x_3 \in \{-1, 1\}\right\}$. Let F be the set of all twelve lines containing the diagonals of the six faces of the cube Q. Let S be the set of all four lines containing the main diagonals of the cube Q; for instance, the line passing through the vertices (–1, –1, –1) and (1, 1, 1) is in S. For lines ℓ_1 and ℓ_2, let $d(\ell_1, \ell_2)$ denote the shortest distance between them. Then the maximum value of $d(\ell_1, \ell_2)$, as ℓ_1 varies over F and ℓ_2 varies over S, is

(A) $\frac{1}{\sqrt{6}}$

(B) $\frac{\sqrt{2}}{\sqrt{3}}$

(C) $\frac{1}{\sqrt{3}}$

(D) $\frac{1}{\sqrt{12}}$

3) Let X = $\left\{(x, y) \in \mathbb{Z} \times \mathbb{Z} : \frac{(x+1)^2}{8} + \frac{(y+3)^2}{20} < 1 \text{ and } (y+3)^2 < 5(x+1)\right\}$. Three distinct points P, Q and R are randomly chosen from X. Then the probability that P, Q and R form a triangle whose area is a positive integer, is

(A) $\frac{83}{220}$

(B) $\frac{71}{220}$

(C) $\frac{73}{220}$

(D) $\frac{79}{220}$

4) Let P be a point on the parabola $y^2 = 4ax$, where a > 0. The normal to the parabola at P meets the x-axis at a point Q. The area of the triangle PFQ, where F is the focus of the parabola, is 90. If the slope m of the normal and a are both positive integers, then the pair (a, m) is

(A) (2, 3)

(B) (3, 2)

(C) (2, 4)

(D) (4, 2)

SECTION-I (ii)

1) Let S = (-2, -1) ∪ (0, 1) ∪ (2, 3) ∪ (4, 5) and T = {0, 1, 2, 3, 4}. Then which of the following statements is(are) true ?

(A) There are infinitely many functions from S to T

(B) There are infinitely many strictly increasing functions from S to T

(C) The number of continuous functions from S to T is at most 786

(D) Every continuous function from S to T is differentiable

2) Let T_1 and T_2 be two distinct common tangents to the ellipse $E : x^2 + 2y^2 - 2x - 8y + 3 = 0$ and the parabola $P : 12x = y^2 - 4y + 16$. Suppose that the tangent T_1 touches P and E at the points A_1 and A_2, respectively and the tangent T_2 touches P and E at the points A_4 and A_3, respectively. Then which of the following statements is(are) true ?

(A) The area of the quadrilateral $A_1A_2A_3A_4$ is 35 square units

(B) The area of the quadrilateral $A_1A_2A_3A_4$ is 36 square units

(C) The tangents T_1 and T_2 meet the x-axis at the point (–2, 2)

(D) The tangents T_1 and T_2 meet the x-axis at the point (–5, 2)

3) Let $f : [0, 1] \to [0, 1]$ be the function defined by $36f(x) = 12x^3 - 36x^2 + 20x + 17$. Consider the square region $S = [0, 1]\times[0,1]$. Let $G = \{(x, y)\in S : y > f(x)\}$ be called the green region and $R = \{(x, y)\in S : y < f(x)\}$ be called the red region. Let $L_h = \{(x, h)\in S : x\in[0, 1]\}$ be the horizontal line drawn at a height $h \in[0, 1]$. Then which of the following statements is(are) true ?

(A) There exists an $h \in \left[\frac{1}{4}, \frac{2}{3}\right]$ such that the area of the red region above the line L_h equals the area of the green region below the line L_h

(B) There exists an $h \in \left[\frac{1}{4}, \frac{2}{3}\right]$ such that the area of the red region above the line L_h equals the area of the red region below the line L_h

(C) There exists an $h \in \left[\frac{1}{4}, \frac{2}{3}\right]$ such that the area of the green region above the line L_h equals the area of the red region below the line L_h

(D) There exists an $h \in \left[\frac{1}{4}, \frac{2}{3}\right]$ such that the area of the green region above the line L_h equals the area of the green region below the line L_h

SECTION-I (iii)

1) Let α, β and γ be real numbers. Consider the following system of linear equations

$x + 2y + z = 4$

$x + \alpha z = 10$

$2x - 3y + \beta z = \gamma + 1$

Match each entry in **List - I** to the correct entries in **List-II**

List-I		**List-II**	
(P)	If $\beta = \frac{1}{2}(7\alpha - 3)$ and $\gamma = 28$, then the system has	(1)	a unique solution

(Q)	If $\beta = \frac{1}{2}(7\alpha - 3)$ and $\gamma \neq 28$, then the system has	(2)	no solution
(R)	If $\beta \neq \frac{1}{2}(7\alpha - 3)$ where $\alpha = 1$ and $\gamma \neq 28$, then the system has	(3)	infinitely many solutions
(S)	If $\beta \neq \frac{1}{2}(7\alpha - 3)$ where $\alpha = 1$ and $\gamma = 28$, then the system has	(4)	$x = 10$, $y = -3$ and $z = 0$ as a solution
		(5)	$x = -16$, $y = 3$ and $z = 0$ as a solution

The correct option is :

(A) P → 3;Q → 2;R → 5;S → 4

(B) P → 3;Q → 2;R → 1;S → 4

(C) P → 2;Q → 1;R → 4;S → 5

(D) P → 2;Q → 1;R → 1;S → 3

2)

Consider the given data with frequency distribution

x_i	3	8	11	10	5	4
f_i	5	2	3	2	4	4

Match each entry in **List-I** to the correct entries in **List-II.**

List-I		**List-II**	
(P)	The mean of the above data is	(1)	2.5
(Q)	The median of the above data is	(2)	5
(R)	The mean deviation about the mean of the above data is	(3)	6
(S)	The mean deviation about the median of the above data is	(4)	2.7
		(5)	2.4

The correct option is :

(A) P → 3;Q → 3;R → 5;S → 5

(B) P → 3;Q → 2;R → 1;S → 5

(C) P → 2;Q → 3;R → 4;S → 1

(D) P → 3;Q → 2;R → 4;S → 5

3) Let ℓ_1 and ℓ_2 be the lines $\vec{r}_1 = \lambda\left(\hat{i}+\hat{j}+\hat{k}\right)$ and $\vec{r}_2 = \left(\hat{j}-\hat{k}\right) + \mu\left(\hat{i}+\hat{k}\right)$, respectively. Let X be the set of all the planes H that contain the line ℓ_1. For a plane H, let d(H) denote the smallest possible distance between the points of ℓ_2 and H. Let H_0 be plane in X for which $d(H_0)$ is the maximum value of d(H) as H varies over all planes in X.
Match each entry in **List-I** to the correct entries in **List-II.**

List-I		List-II	
(P)	The value of $d(H_0)$ is	(1)	$\sqrt{3}$
(Q)	The distance of the point (0,1,2) from H_0 is	(2)	$\frac{1}{\sqrt{3}}$
(R)	The distance of origin from H_0 is	(3)	0
(S)	The distance of origin from the point of intersection of planes $y = z$, $x = 1$ and H_0 is	(4)	$\sqrt{2}$
		(5)	$\frac{1}{\sqrt{2}}$

The correct option is :

(A) $P \to 5; Q \to 4; R \to 3; S \to 1$

(B) $P \to 2; Q \to 4; R \to 5; S \to 1$

(C) $P \to 2; Q \to 1; R \to 3; S \to 2$

(D) $P \to 5; Q \to 1; R \to 4; S \to 2$

4) Let z be complex number satisfying $27|z|^3 + 18z^2 + 12\bar{z} - 8 = 0$, where $\bar{z}$ denotes the complex conjugate of z. Let the imaginary part of z be nonzero.
Match each entry in **List-I** to the correct entries in **List-II.**

List-I		List-II	
(P)	$9\,|z|^2$ is equal to	(1)	12
(Q)	$9\,|z - \bar{z}|^2$ is equal to	(2)	4
(R)	$9\left(|z|^2 + |z + \bar{z}|^2\right)$ is equal to	(3)	8
(S)	$|3z + 1|^2$ is equal to	(4)	10
		(5)	7

The correct option is :

(A) $P \to 1; Q \to 3; R \to 5; S \to 4$

(B) $P \to 2; Q \to 4; R \to 5; S \to 1$

(C) $P \to 2; Q \to 1; R \to 3; S \to 5$

(D) $P \to 2; Q \to 3; R \to 5; S \to 4$

SECTION-II

1) Let $\tan^{-1}(x) \in \left(-\frac{\pi}{2}, \frac{\pi}{2}\right)$, for $x \in \mathbb{R}$. Then the number of real solutions of the equation $\sqrt{1 + \cos(2x)} = \sqrt{2}\tan^{-1}\left(\tan\frac{x}{6}\right)$ in the set $(-9\pi, -3\pi) \cup (-3\pi, 3\pi) \cup (3\pi, 9\pi)$ is equal to

2) Let n ≥ 2 be a natural number and $f : [0, 1] \to \mathbb{R}$ be the function defined by

$$f(x) = \begin{cases} n(1-2nx) & \text{if } 0 \leqslant x \leqslant \frac{1}{2n} \\ 2n(2nx-1) & \text{if } \frac{1}{2n} \leqslant x \leqslant \frac{3}{4n} \\ 4n(1-nx) & \text{if } \frac{3}{4n} \leqslant x \leqslant \frac{1}{n} \\ \frac{n}{n-1}(nx-1) & \text{if } \frac{1}{n} \leqslant x \leqslant 1 \end{cases}$$

If n is such that the area of the region bounded by the curves x = 0 , x = 1 , y = 0 and y = f (x) is 6, then the maximum value of the function f is

3) Let $7\overbrace{5\ldots5}^{r}7$ denote the (r + 2) digit number where the first and the last digits are 7 and the remaining r digits are 5. Consider the sum S = 77 + 757 + 7557 + … + $7\overbrace{5\ldots5}^{100}7$. If $S = \frac{7\overbrace{5\ldots5}^{101}7 + m}{n}$, where m and n are natural numbers less than 3000, then the value of m + n is

4) Let $A = \left\{\frac{133 - i(900\cos\theta + 1686\sin\theta)}{7 - 3i\cos\theta} : \theta \in \mathbb{R}\right\}$. If A contains exactly one positive integer n, then the value of n is

5) Let P be the plane $\sqrt{3}x + 2y + 3z = 16$ and let $S = \left\{\alpha\hat{i} + \beta\hat{j} + \gamma\hat{k} : \alpha^2 + \beta^2 + \gamma^2 = 1 \text{ and the distance of } (\alpha, \beta, \gamma) \text{ from the plane P is } \frac{7}{2}\right\}$. Let $\vec{u}, \vec{v}$ and $\vec{w}$ be three distinct vectors in S such that $|\vec{u} - \vec{v}| = |\vec{v} - \vec{w}| = |\vec{w} - \vec{u}|$. Let V be the volume of the parallelepiped determined by vectors $\vec{u}, \vec{v}$ and $\vec{w}$. Then the value of $\frac{16}{\sqrt{3}}V$ is

6) Let a and b be two nonzero real numbers. If the coefficient of x^7 in the expansion of $\left(ax^3 + \frac{70}{27bx^2}\right)^4$ is equal to the coefficient of x^{-6} in the expansion of $\left(ax^2 - \frac{1}{bx^3}\right)^7$, then the value of 2b is

ANSWER KEYS

PART-1 : PHYSICS

SECTION-I (i)

Q.	1	2	3	4
A.	A	C	B	B

SECTION-I (ii)

Q.	5	6	7
A.	A,B,D	A,B,C	A,B,D

SECTION-I (iii)

Q.	8	9	10	11
A.	A	C	D	B

SECTION-II

Q.	12	13	14	15	16	17
A.	2	30	520	132	100	6

PART-2 : CHEMISTRY

SECTION-I (i)

Q.	18	19	20	21
A.	C	C	A	B

SECTION-I (ii)

Q.	22	23	24
A.	C,D	A,B,C,D	B

SECTION-I (iii)

Q.	25	26	27	28
A.	D	D	B	D

SECTION-II

Q.	29	30	31	32	33	34
A.	148	80	6	3	16	52

PART-3 : MATHEMATICS

SECTION-I (i)

Q.	35	36	37	38
A.	**A**	**B**	**C**	**B**

SECTION-I (ii)

Q.	39	40	41
A.	**A,C,D**	**A,C**	**A,B,C**

SECTION-I (iii)

Q.	42	43	44	45
A.	**B**	**D**	**A**	**C**

SECTION-II

Q.	46	47	48	49	50	51
A.	**9**	**12**	**1245**	**19**	**9**	**3**

PART-1 : PHYSICS

1) $0 = Mv_1 \times \frac{L}{4} - \frac{ML^2}{12} \times \omega$

$Mv_1 - mv = mv_0$

$e = 1 = \frac{v_1 + \frac{\omega L}{4} + v}{v_0}$

$v_1 + \frac{\omega L}{4} + v = 5$

$\frac{7\omega L}{12} + v = 5$

$10\frac{\omega L}{3} = v + 5$

$10\frac{\omega L}{3} - v = 5$

$\frac{7\omega L}{12} + v = 5$

$\frac{47\omega L}{12} = 10$

$\omega = \frac{120}{47 \times 0.2} = \frac{600}{47}$

$v = 5 - 7 \times \frac{0.2}{12} \times \frac{600}{47} = \frac{165}{47} m/s$

2)

In t = 10 sec volume of liquid is

V = 5000 cc

$h = \frac{5000}{50 \times 5} = 20cm$

$C_d = \frac{A_d \varepsilon_0 k}{d}$

$= \frac{50 \times 10^{-2} \times 20 \times 10^{-2} \varepsilon_0 \times 3}{5 \times 10^{-2}} = 6\varepsilon_0$

$C_a = \frac{A_a \varepsilon_0}{d} = \frac{50 \times 10^{-2} \times 30 \times 10^{-2} \varepsilon_0}{5 \times 10^{-2}} = 3\varepsilon_0$

$C = C_a + C_d = 9\varepsilon_0$

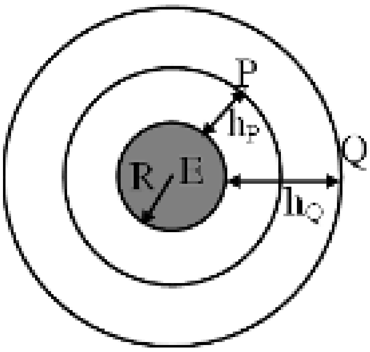
$= 9 \times 9 \times 10^{-12} = 81$ Pf

3) $P_0 V_0^{\gamma} = P_A (5V_0)^{\gamma}$

$\frac{P_A}{P_B} = \frac{P_0}{5^{\gamma} \times P_0}$

$P_0 V_0 = 5 P_B V_0$

$P_B = \frac{P_0}{5}$



4)

$$\frac{g_P}{g_Q}=\frac{\frac{GM}{r_P^2}}{\frac{GM}{r_Q^2}}=\left(\frac{r_Q}{r_P}\right)^2$$

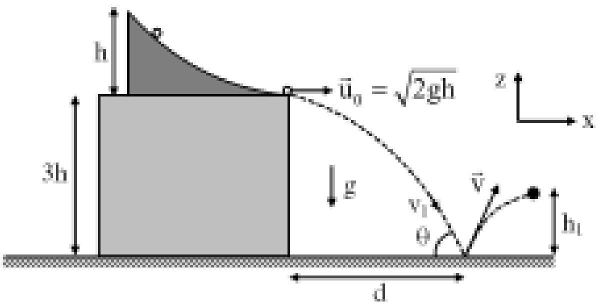
$$\frac{36}{25}=\left(\frac{r_Q}{r_P}\right)^2$$

$$\frac{r_Q}{r_P}=\frac{6}{5}$$

$$r_Q=\frac{6}{5}r_P$$

$$R+h_Q=\frac{6}{5}\left(R+\frac{R}{3}\right)$$

$$v_Q=\sqrt{\frac{5GM}{8R}}$$



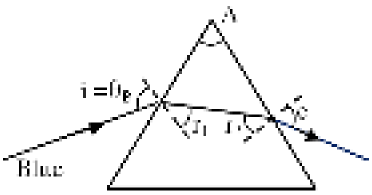
5)

$d = 20\ t$

$40 = 2h = \frac{1}{2}\times 10\times t^2$

$t=2\sqrt{2}$

$d=40\sqrt{2}$



6)

$\tan\theta_B=\mu_B=\sqrt{3}$

$i = \theta_B = 60°$

$1\sin 60^\circ=\sqrt{3}\sin r_1$

$r_1 = 30°$

$r_1 + r_2 = A$

$\delta = (i + e) - A$

$60° = 60° + e - A$

$e = A$

$\sqrt{3}\sin r_2 = 1 \sin e$

$\sqrt{3}\sin(A - 30) = \sin A$

Solving

$A = 60°$

 $e = 60°$

For red light

$$\mu = \frac{\sin\left(\frac{A+\delta_{min}}{2}\right)}{\sin\frac{A}{2}} = \sqrt{2}$$

8) $_{Z_1}Z^{A_1} \rightarrow {}_{Z_2}Y^{A_2} + N_1\ {}_2He^4 + N_2\ {}_1e^0 + N_3\ {}_{-1}e^0$

Conservation of charge

$Z_1 = Z_2 + 2N_1 + N_2 - N_3$ ... (i)

Conservation of nucleons.

$A_1 = A_2 + 4N_1$

$N_1 = \frac{A_1 - A_2}{4}$ ... (ii)

From (i) and (ii)

$$N_2 - N_3 = Z_1 - Z_2 - \left(\frac{A_1 - A_2}{2}\right)$$

(P) $_{92}U^{238} \rightarrow {}_{91}Pa^{234}$

$N_1 = \frac{238 - 234}{4} = 1 \rightarrow 1\alpha$

$N_2 - N_3 = (92 - 91) - \left(\frac{4}{2}\right) = -1 \rightarrow 1\beta^-$

(Q) $_{82}Pb^{214} \rightarrow {}_{82}Pb^{210}$

$N_1 = \frac{214 - 210}{4} = 1 \rightarrow 1\alpha$

$N_2 - N_3 = (82 - 82) - \left(\frac{4}{2}\right) = -2 \rightarrow 2\beta^-$

(R) $_{81}T\ell^{210} \rightarrow {}_{82}Pb^{206}$

$N_1 = \frac{210 - 206}{4} = 1 \rightarrow 1\alpha$

$N_2 - N_3 = (81 - 83) - \frac{4}{2} = -3 \rightarrow 3\beta^-$

(S) $_{91}Pa^{228} \rightarrow {}_{88}Ra^{224}$

$N_1 = \frac{228 - 224}{4} = 1\alpha$

$N_2 - N_3 = (91 - 88) - \frac{4}{2} = 1\beta^+$

9)

⇒ For option (P) temperature is minimum hence λ will be maximum $\beta = \frac{\lambda D}{d}$

⇒ β will also be maximum

⇒ For option (Q) T = 3000

$$\lambda m = \frac{b}{T} = \frac{2.9 \times 10^{-3}}{30000}$$

$$\lambda m = \frac{2.9}{3} \times 10^{-6}$$

$= 0.96 \times 10^{-6}$

$= 966.6$ nm

$P_{3000} = 6A\,(3000)^4$

$P_{6000} = 6A\,(6000)^4$

$$\frac{P_{3000}}{P_{6000}} = \left(\frac{1}{2}\right)^4 = \frac{1}{16}$$

$$P_{3000} = \frac{1}{16}P_{6000}$$

Q – 4

⇒ For (R) T = 5000 K

$$\lambda m = \frac{2.9 \times 10^{-3}}{5 \times 10^{3}} = 0.58 \times 10^{-6}$$

$= 580$ nm

Visible to human eyes

R – 2

⇒ For (S) T = 10,000 → maximum

Hence (3) is wrong as it has minimum (λm)

10)

V = 45 sin wt, L = 50 mH

$$\omega_0 = 10^5\text{rad/s} = \frac{1}{\sqrt{LC}} \Rightarrow C = \frac{1}{L\omega_0^2} = \frac{1}{5 \times 10^{-2} \times 10^{10}}$$

$= 2 \times 10^{-9}$ F

$$I_0 = \frac{45}{R}$$

$\omega = 8 \times 10^4$ rad/s $= 0.8\,\omega_0$

$$I = 0.05I_0 = \frac{I_0}{20} \Rightarrow Z = 20R$$

$$X_L = 8 \times 10^4 \times 5 \times 10^{-2}\Omega = 4k\Omega$$

$$X_C = \frac{1}{8 \times 10^4 \times 2 \times 10^{-9}} = \frac{1}{16} \times 10^5\Omega = \frac{25}{4}k\Omega$$

$$Z^2 = R^2 + (X_C - X_L)^2$$

$$400R^2 = R^2 + \left(\frac{9}{4}k\Omega\right)^2$$

$$R = \frac{\frac{9}{4}k\Omega}{\sqrt{399}} \approx \frac{9}{80}k\Omega = \frac{900}{8}\Omega$$

$$I_0 = \frac{V_0}{R} = \frac{45 \times 8}{900} = \frac{8}{20}A \approx 0.4A = 400mA$$

$$Q = \frac{1}{R}\sqrt{\frac{L}{C}} = \frac{8}{900}\sqrt{\frac{5 \times 10^{-2}}{2 \times 10^{-9}}} = \frac{8}{900}\sqrt{25 \times 10^6}$$

$$Q = \frac{8}{900} \times 5000 = 44.4$$

$$Q = \frac{\omega_0}{\Delta\omega} \Rightarrow \Delta\omega = \frac{\omega_0}{Q} = \frac{10^5}{44.4} = 2250.0$$

$$P_{max} = I_0^2 R = \frac{45^2}{R^2} \times R = \frac{45^2}{R} = \frac{45^2}{900} \times 8 = 18.4W$$

11) From force equation

$$mg - Bi\ell = \frac{mdv}{dt}$$

$$mg - \frac{BBi\ell}{R} \times \ell = \frac{mdv}{dt}$$

$$\frac{mgR}{B^2\ell^2} - v = \frac{mR}{B^2\ell^2}\frac{dv}{dt}$$

$$\frac{B^2\ell^2}{mR}\int_{t=0}^{t} dt = \int_0^v \frac{dv}{\left(\frac{mgR}{B^2\ell^2} - v\right)}$$

Now $\frac{mgR}{B^2\ell^2} = \frac{20 \times 10^{-3} \times 10 \times 10}{16 \times \frac{1}{16}} = 2$

And $\frac{B^2\ell^2}{mR} = \frac{16 \times \frac{1}{16}}{20 \times 10^{-3} \times 10} = \frac{1}{0.2} = 5$

$$\therefore 5t = [-\ell n(2 - v)]_0^v$$

$$-5t = \ell n\left[\frac{2 - v}{v}\right]$$

$\therefore v = 2(1 - e^{-5t})$

At t = 0.2 sec $v = 2(1 - e^{-5 \times 0.2})$

$v = 2(1 - 0.4)$

$v = 1.2$ m/s

(P) Now at t = 0.2 sec

The magnitude of the induced emf = E = $Bv\ell$

$$= 4 \times 1.2 \times \frac{1}{4} = 1.2\text{Volt}$$

Q) At t = 0.2 sec, the magnitude of magnetic force = $BI\ell\sin\theta$

$$= B \times \frac{B\ell v}{R} \times \ell \times \sin 90^\circ$$

$$= \frac{4 \times 4 \times \frac{1}{4} \times 1.3 \times \frac{1}{4}}{10} = 0.12 \text{ Newton}$$

(R) At t = 0.2 sec, the power dissipated as heat

$$P = i^2R = \frac{v^2}{R} = \frac{1.2 \times 1.2}{10}$$

P = 0.144 watt

(S) Magnitude of terminal velocity

At terminal velocity, the net force become zero

$\therefore mg = Bi\ell$

$$mg = B \times \frac{B\ell v_t}{R} \times \ell$$

$$\therefore v_T = \frac{mgR}{B^2\ell^2} = \frac{20 \times 10^{-3} \times 10 \times 10}{16 \times \frac{1}{16}}$$

v_T = 2 m/s

Hence, Answer is (D)

12)

n = 3 n = 5

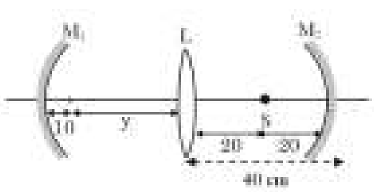
$-1.51Z^2$eV $-0.544\ Z^2$ eV

E = $E_4 - E_3$ = 0.96 Z^2 eV

K_{max} = E – W

$$W = 0.96Z^2 - 1.86 = \frac{hc}{\lambda} = \frac{1240}{620}$$

$\therefore$ Z = 2



13)

Consider reflection 1st at M_2

$$\frac{1}{v} - \frac{1}{20} = \frac{1}{12}$$

$$\frac{1}{v} = \frac{1}{20} - \frac{1}{12} = \frac{3-5}{60}$$

v = -30

$$\frac{1}{20} + \frac{1}{y} = \frac{1}{10}$$

$$\frac{1}{y} = \frac{1}{10} - \frac{1}{20}$$

y = 20

So, the distance is 30 cm.

14) $\delta = \dfrac{10 \times 15}{25} = 6\text{cm}$

df = $36\left(\dfrac{0.1}{10^2} + \dfrac{0.1}{15^2}\right)$

$= 36 \times \dfrac{0.1}{100} + \dfrac{36 \times 0.1}{225}$

$= \dfrac{5.2\text{cm}}{1000}$

15)

At constant pressure

$W = nR\Delta T = 66$

$\Delta U = n(C_V)_{mix}\Delta T$

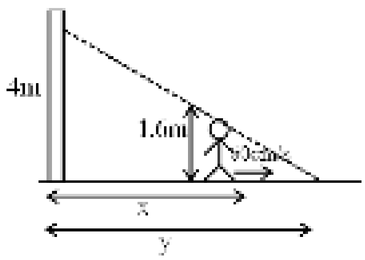
$(C_V)_{mix} = \dfrac{n_1 C_{V_1} + n_2 C_{V_2}}{n_1 + n_2}$

$(C_V)_{mix} = \dfrac{1 \times \frac{3}{2}R + 1 \times \frac{5}{2}R}{2}$

$(C_V)_{mix} = 2R$

$\Delta U = 2\,(nR\Delta T)$

$\Delta U = 2 \times 66 = 132\text{J}$



16)

$\dfrac{4}{y} = \dfrac{1.6}{y - x}$

$4y - 4x = 1.6y$

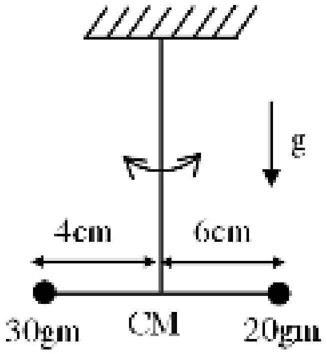
$2.4\,y = 4x$

$X = 0.6y$

$\dfrac{dx}{dt} = 0.6 \times \dfrac{dy}{dt}$

$60 = 0.6 \times \dfrac{dy}{dt}$

$\therefore \dfrac{dy}{dt} = 100$ cm/s

17) 

$$T = \frac{2\pi}{\omega} = 2\pi\sqrt{\frac{I}{C}}$$

$$\Rightarrow \omega = \sqrt{\frac{C}{I}}$$

Where I = moment of inertia

$I = (30)(4)^2 + (20)(6)^2$

$= 1200$ gm-cm^2

$= 1.2 \times 10^{-4}$ kg-m^2

$$\Rightarrow \omega = \sqrt{\frac{1.2 \times 10^{-8}}{1.2 \times 10^{-4}}}$$

$\Rightarrow \omega = (10^{-2})$

$v = A\omega = 6 \text{ cm} \times 0.01 \times 10^{-2} = 6 \times 10^{-6}$ m/s

PART-2 : CHEMISTRY

18)

$$MnCl_2 + NaOH \rightarrow \underset{\substack{(P)\\(\text{white ppt.})}}{Mn(OH)_2} \downarrow + \underset{\substack{(Q)\\(\text{Filterate})}}{NaCl}$$

$$Mn(OH)_2 \xrightarrow[\text{heat}]{PbO_2 + H^+(H_2SO_4)} \underset{\text{Purple}}{MnO_4^-} + Pb^{2+}$$

$$Cl^- \xrightarrow{MnO(OH)_2/\text{conc. } H_2SO_4/\Delta} Cl_2$$

$$Cl_2 \xrightarrow{2I^-} \underbrace{(\text{Starch} + I_2)}_{\text{blue coloration}} + 2Cl^-$$

19) For weak acid, $\alpha = \frac{\Lambda_m}{\Lambda_0}$

$K_a = \frac{C\alpha^2}{1-\alpha} \Rightarrow K_a(1-\alpha) = C\alpha^2$

$\Rightarrow K_a\left(1 - \frac{\Lambda_m}{\Lambda_0}\right) = C\left(\frac{\Lambda_m}{\Lambda_0}\right)^2$

$\Rightarrow K_a - \frac{\Lambda_m K_a}{\Lambda_0} = \frac{C\Lambda_m^2}{(\Lambda_0)^2}$

Divide by 'Λ_m'

$$\Rightarrow \frac{K_a}{\Lambda_m} = \frac{C\Lambda_m}{(\Lambda_0)^2} + \frac{K_a}{\Lambda_0}$$

$$\Rightarrow \frac{1}{\Lambda_m} = \frac{C\Lambda_m}{K_a(\Lambda_0)^2} + \frac{1}{\Lambda_0}$$

Plot $\frac{1}{\Lambda_m}$ vs C Λ_m has

Slope = $\frac{1}{K_a(\Lambda_0)^2}$ = S

y-intercept = $\frac{1}{\Lambda_0}$ = P

20) At pH = 7 ⇒ pure water

solubility = $S_1 = \sqrt{K_{sp}}$

At pH = 2

$$\Rightarrow MX(s) + aq \xrightleftharpoons{K_{SP}} M^+(aq) + X^-(aq)$$

s   s–x

$$X^-(aq) + H^+(aq) \xrightleftharpoons{1/K_a} HX(aq)$$

s–x 10^{-2} $x \simeq s$

Approximation : $s - x \simeq 0$ [X^- is limiting reagent]

$\Rightarrow s \simeq x$

$\Rightarrow$ $s(s - x) = K_{sp}$ ...... (1)

$$\frac{s}{(s-x)\,(10^{-2})} = \frac{1}{K_a}$$ ...... (2)

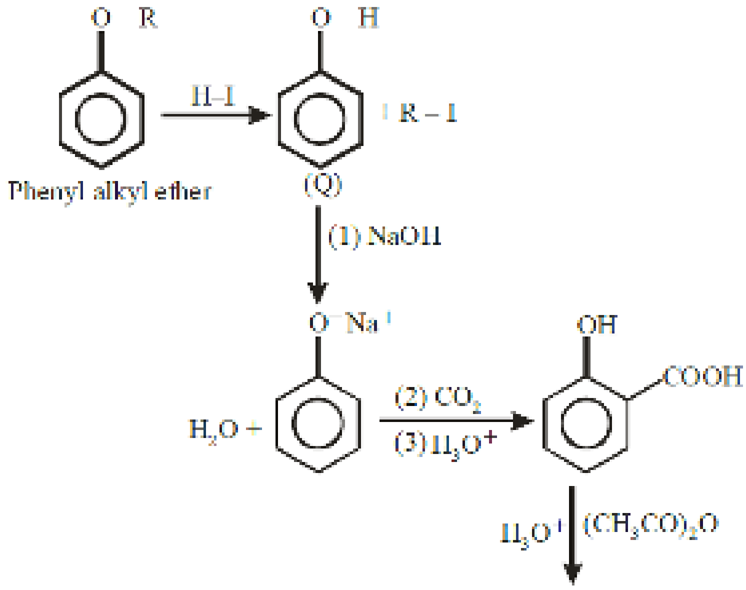
Multiply (1) × (2) $\Rightarrow \frac{s^2}{10^{-2}} = \frac{K_{sp}}{K_a}$

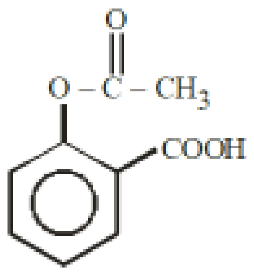
21) P is phenyl alkyl ether

Q is aromatic compound

R and S are the major product

i.e.





(Aspirin)
Pain killer
Acetyl salicylic acid
(Non-narcotic analgesic)

Correct ans is (B)
Aspirin inhibits the synthesis of chemicals known as prostaglandin's.

22) ⇒ Under roasting condition, the malachite will be converted into

$CuCO_3.Cu(OH)_2 \rightarrow 2CuO + CO_2 + H_2O$

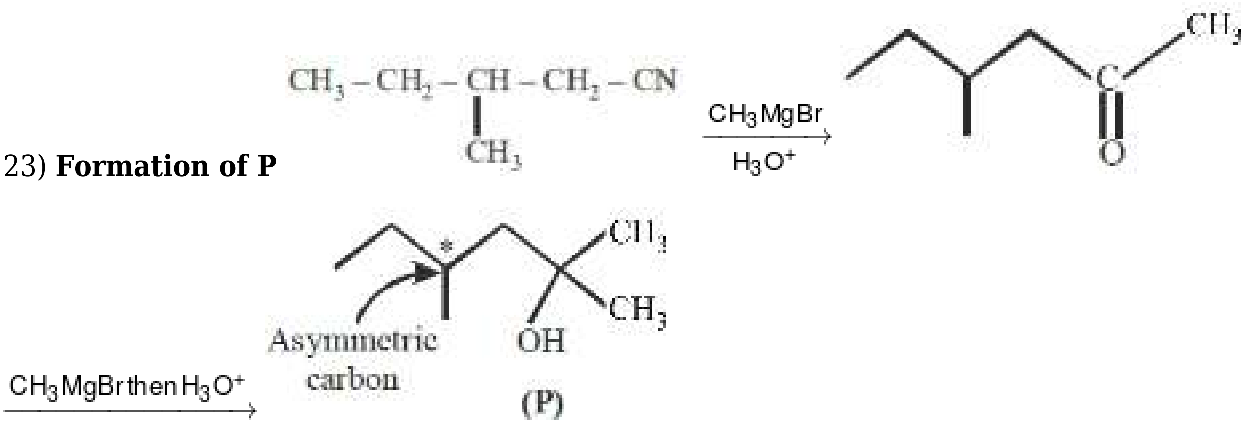
$\underset{\text{(Calamine)}}{ZnCO_3} \rightarrow \underset{\text{(Zincite)}}{ZnO} + CO_2 \uparrow$

⇒ Copper pyrites is heated in a reverberatory furnace after mixing with silica. In the furnace, iron oxide 'slag of' as iron silicate and copper is produced in the form of copper matte.

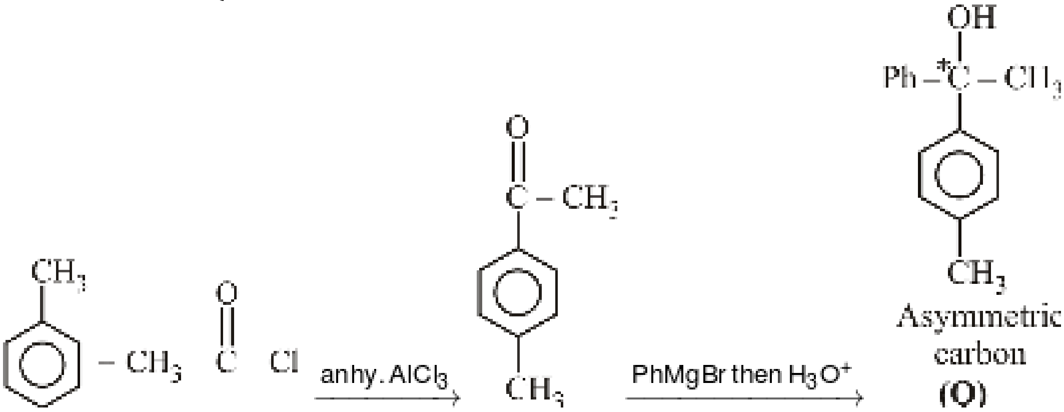
$FeO + SiO_2 \rightarrow \underset{\text{(Slag)}}{FeSiO_3}$

⇒ Malachite $\xrightarrow{\Delta}$ CuO

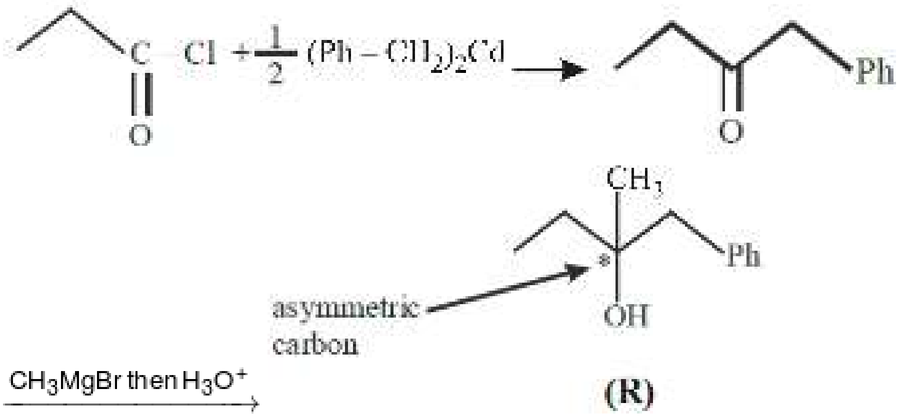
23) **Formation of P**



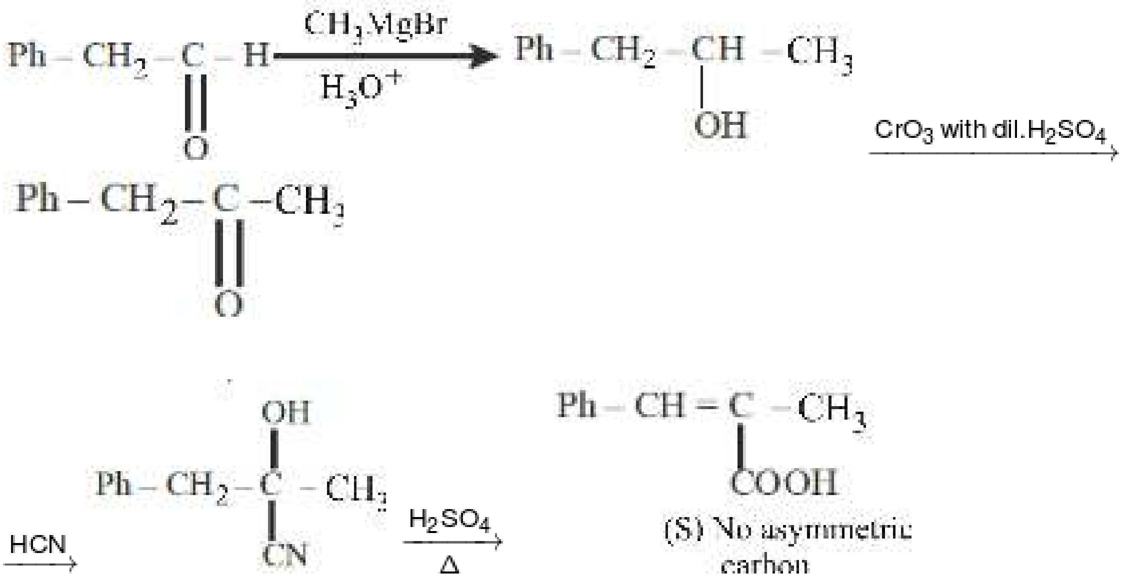
Formation of Q

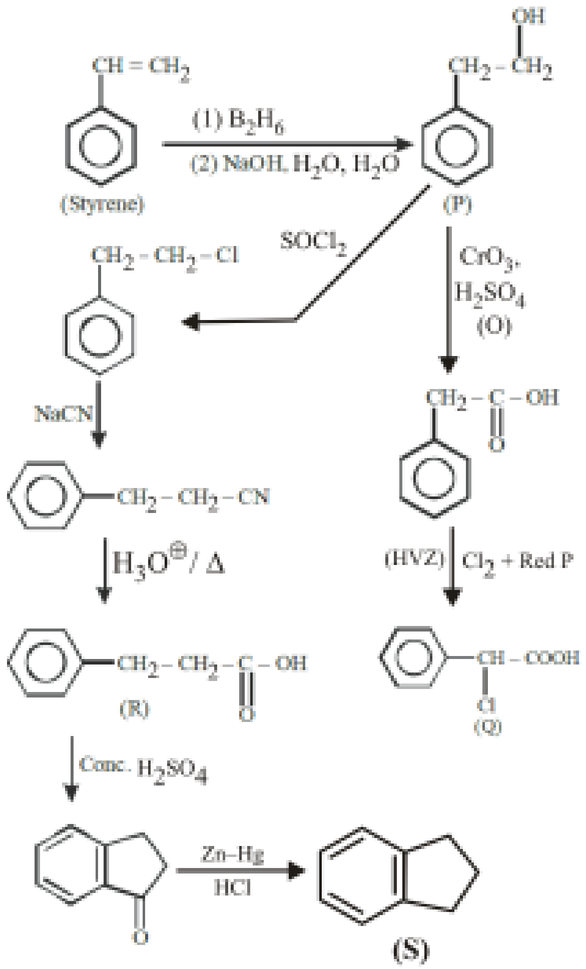


Formation of R



Formation of S

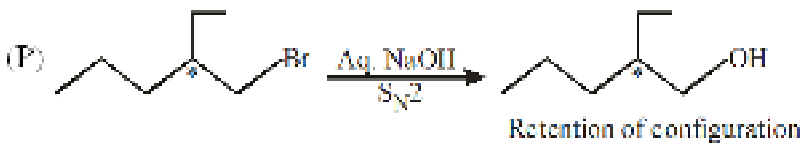


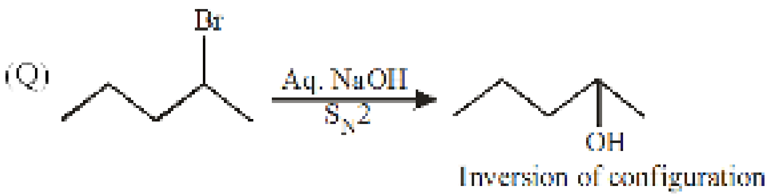


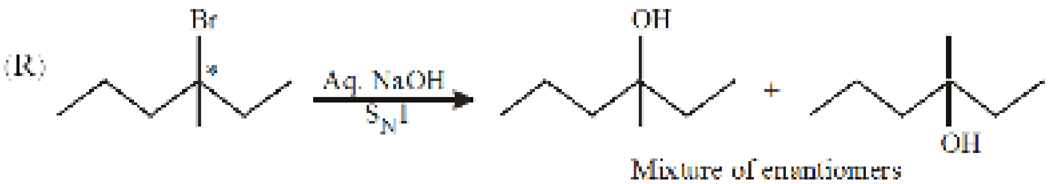
24)

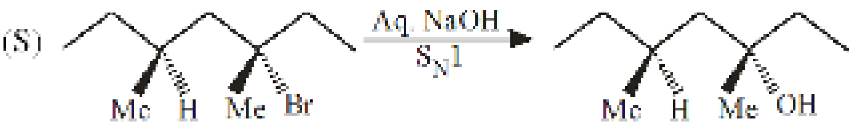
25) (P) $PCl_3 + H_3PO_3 \rightarrow H_4P_2O_5$
(Q) Red P_4 + alkali $\rightarrow H_4P_2O_6$
(R) $C_2H_5OH + PCl_3 \rightarrow [H_3PO_3]$
(S) $H_3PO_2 + 2H_2O + 4AgNO_3 \rightarrow 4Ag + 4HNO_3 + \underline{H_3PO_4}$

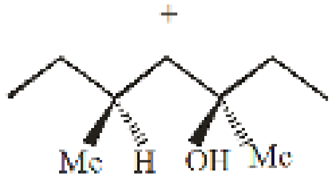
27)









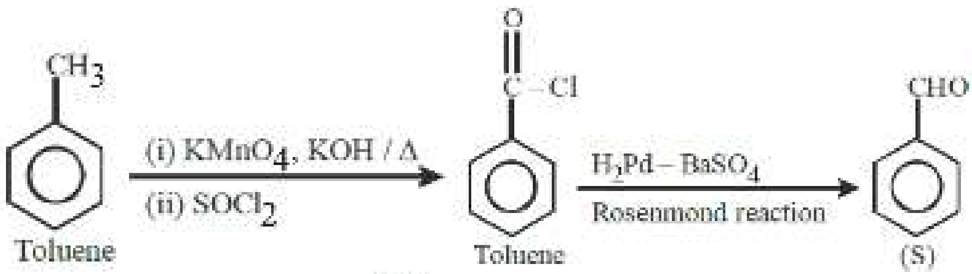


Diastereomeric mixture

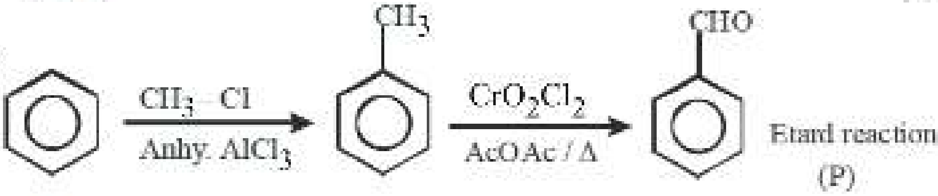
28)

(i)

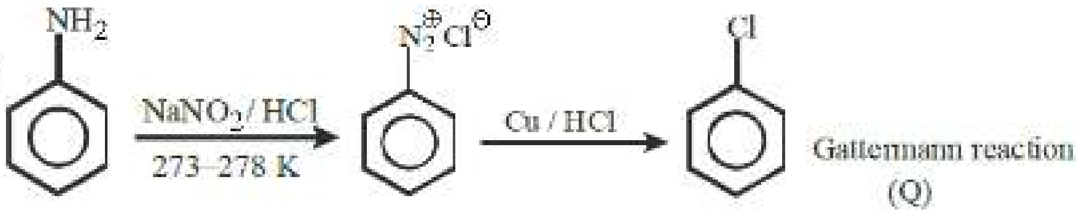
(ii)



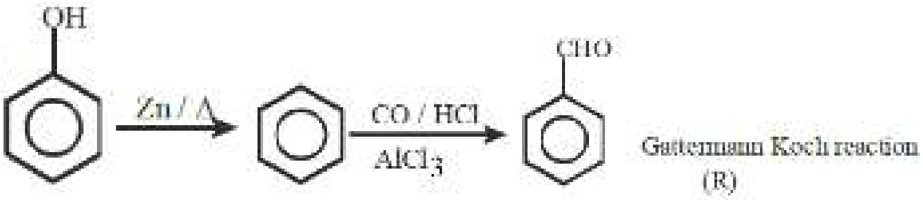
(iii)



(iv)



(v)

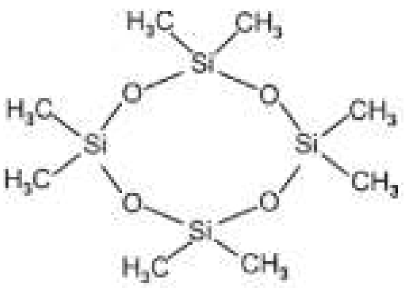


29)

$$4(CH_3)_2SiCl_2 + 4H_2O \xrightarrow{75\%} \underset{(X)}{(CH_3)_8Si_4O_4} + 8HCl$$

w = 516 g

$$\underset{(moles)}{n} = \frac{516}{129} = 4$$



weight = 296 g

% yield = 50

The weight of X(ingram) = $296 \times \frac{50}{100} = 148g$

30) For gas : Z = 0.5, V_m = 0.4 L/mol

T = 800 K, P = X atm.

$\Rightarrow Z = \frac{PV_m}{RT}$

$\Rightarrow \frac{X(0.4)}{0.08 \times 800} = 0.5$

$\Rightarrow$ X = 80

31) For reaction $A(g) \rightleftharpoons P(g)$

$\log k_f = \frac{-E_f}{2.303\,RT} + \log A_f$ [Arrhenius equation for forward reaction]

From plot when, $\frac{1}{T} = 0.002$, $\log k_f = 9$

$\Rightarrow \quad 9 = \frac{-E_f}{2.303\,R}(0.002) + \log(A_f)$

Given : $A_f = 10^{15}\ s^{-1}$

$\Rightarrow \quad 9 = \frac{-E_f}{2.303\,R}(0.002) + 15$

$\Rightarrow \frac{E_f}{2.303R} = \frac{6}{0.002} = 3000$

$\Rightarrow \log k_b = \frac{-E_b}{2.303\,RT} + \log A_b$

At 250 K

$-5 = \frac{-E_{ab}}{2.303R \times 250} + 11$

$\frac{E_{ab}}{2.303R} = 16 \times 250 = 4000$

Now, $K = \frac{k_f}{k_b} = \frac{A_f}{A_b} e^{-(E_f - E_b)/RT}$

$$\log K = -\frac{1}{2.303}\frac{(E_f - E_b)}{RT} + \log\left(\frac{10^{15}}{10^{11}}\right)$$

$\log K = -\frac{(3000-4000)}{500} + 4 = 2 + 4 = 6$

32) For A → B $\quad 600\,V_1^{\gamma-1} = 60V_2^{\gamma-1}$ $(\gamma = 5/3)$

(Reversible adiabatic)

$\Rightarrow$ $600\ (V_1)^{2/3} = 60\ (V_2)^{2/3}$

$\Rightarrow$ $10 = \left(\frac{V_2}{V_1}\right)^{2/3}$

$\Rightarrow$ $10 = \left(\frac{V_2}{10}\right)^{2/3}$

$\Rightarrow$ $V_2 = 10(10)^{3/2} = 10^{5/2}$

$\Delta S = R\ \ln \left(\frac{V_3}{V_2}\right) = R\ \ln \sqrt{10}$

$10^{1/2} = \frac{V_3}{10^{5/2}}$

$V_3 = 10^3$

$\log V_3 = 3$

33) At T_1 K : $A(g) \rightleftharpoons P(g)$

$t = 0$ $\quad$ 6

$t = \infty$ $\quad$ $6 - x$ $\quad$ $x = 4$ (from plot)

$\Rightarrow$ At T_1 K : $K_{P_1} = \frac{4}{2} = 2$

At T_2 K : $A(g) \rightleftharpoons P(g)$

$t = 0$ $\quad$ 6

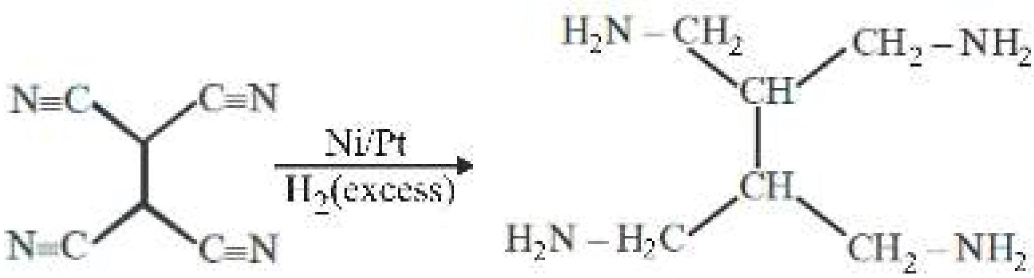
$t = \infty$ $\quad$ $6 - y$ $\quad$ $y = 2$ (from plot)

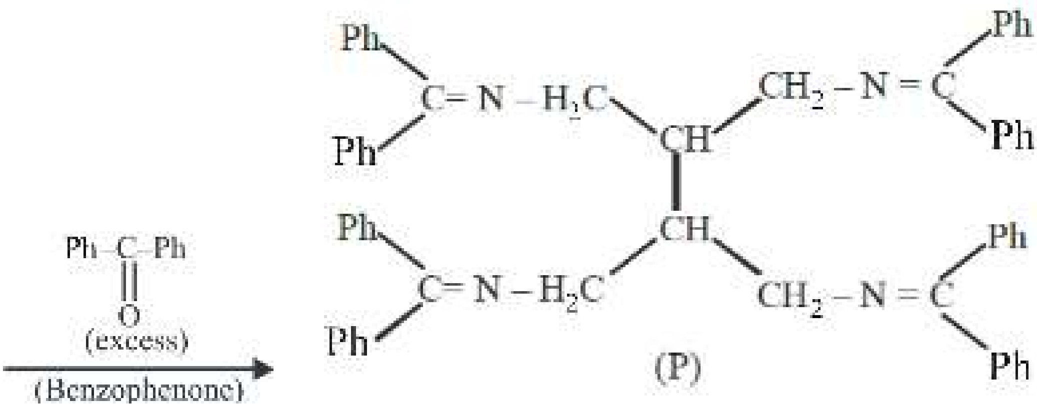
$\Rightarrow$ At T_2 K : $K_{P_2} = \frac{2}{4} = \frac{1}{2}$

$\ell n \frac{K_{P_2}}{K_{P_1}} = \frac{\Delta H}{R}\left(\frac{1}{T_1} - \frac{1}{T_2}\right)$

$-\ell n 4 = \frac{\Delta H}{R}\left[\frac{1}{2T_2} - \frac{1}{T_2}\right]$

$\Delta H = 2T_2\ R \ln 4 = T_2\ R \ln 16$

34) 



Total number of sp^2 hybridised C-atom in
P = 28 + 24 = 52

PART-3 : MATHEMATICS

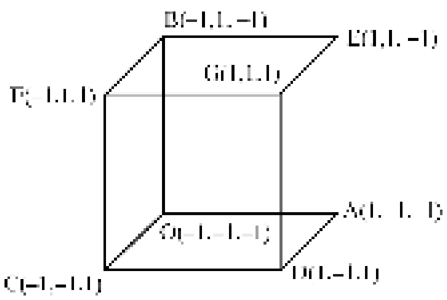
35) $\int_{x^2}^{x}\sqrt{\frac{1-t}{t}}dt.\sqrt{n} \leqslant f(x)g(x) \leqslant 2\sqrt{x}\sqrt{n}$

$\because \int_{x^2}^{x}\sqrt{\frac{1-t}{t}}dt = \sin^{-1}\sqrt{x} + \sqrt{x}\sqrt{1-x} - \sin^{-1}x - x\sqrt{1-x^2}$

$\Rightarrow \lim_{x\to 0}\left(\frac{\sin^{-1}\sqrt{x} + \sqrt{x}\sqrt{1-x} - \sin^{-1}x - x\sqrt{1-x^2}}{\sqrt{x}} \leqslant f(x)g(x) \leqslant \frac{2\sqrt{x}}{\sqrt{x}}\right)$

$\Rightarrow 2 \leqslant \lim_{x\to 0} f(x)g(x) \leqslant 2$

$\Rightarrow \lim_{x\to 0} f(x)g(x) = 2$

36) 

DR'S of OG = 1, 1, 1
DR'S of AF = –1, 1, 1
DR'S of CE = 1, 1, –1
DR'S of BD = 1, –1, 1

Equation of OG $\Rightarrow \frac{x}{1} = \frac{y}{1} = \frac{z}{1}$

Equation of AB $\Rightarrow \frac{x-1}{1} = \frac{y}{-1} = \frac{z}{0}$

Normal to both the line’s

$$= \begin{vmatrix} \hat{i} & \hat{j} & \hat{k} \\ 1 & 1 & 1 \\ 1 & -1 & 0 \end{vmatrix} = \hat{i} + \hat{j} - 2\hat{k}$$

$\overrightarrow{OA} = 2\hat{i}$

S.D. = $\dfrac{\left|2\hat{i}.(\hat{i} + \hat{j} - 2\hat{k})\right|}{\left|\hat{i} + \hat{j} - 2\hat{k}\right|} = \dfrac{2}{\sqrt{6}}$

37) Shift origin to (–1, –3). Thus, we get

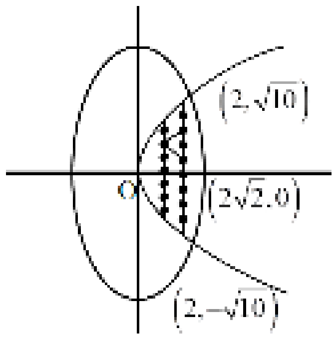
$\frac{x^2}{8} + \frac{y^2}{20} < 1$ & $y^2 < 5x$

Solving corresponding equations

$\frac{x^2}{8} + \frac{y^2}{20} = 1$ & $y^2 = 5x$

$\Rightarrow \begin{Bmatrix} x = 2 \\ y = \pm\sqrt{10} \end{Bmatrix}$

X = {(1,1), (1,0), (1,–1), (1,2), (1,–2), (2, 3), (2,2), (2,1), (2,0), (2,–1), (2,–2), (2,–3)}



Let S be the sample space & E be the event n(S) = ${}^{12}C_3$

<u>For E</u>

Selecting 3 points in which 2 points are either or x = 1 & x = 2 but distance b/w then is even

Triangles with base 2 :

= 3×7 + 5×5 = 46

Triangles with base 4 :

= 1×7 + 3×5 = 22

Triangles with base 6 :

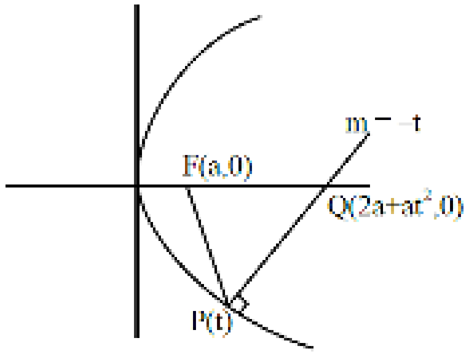
= 1×5 = 5

$P(E) = \dfrac{46 + 22 + 5}{{}^{12}C_3} = \dfrac{73}{220}$

38) Let point P $(at^2, 2at)$ normal at P is $y = -tx + 2at + at^3$

$y = 0, x = 2a + at^2$

$Q(2a + at^2, 0)$



Area of ΔPFQ = $\left|\frac{1}{2}\left(a+at^2\right)(2at)\right| = 120$

 $m = -t$

 $a^2[1 + m^2]\, m = 90$

$(a, m) = (3, 2)$ will satisfy

39) Number of functions :

Each element of S have 5 choice

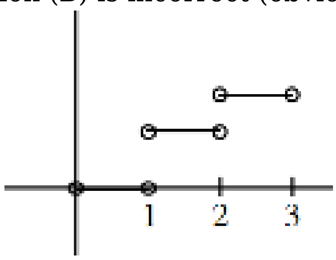
Let n be the number of element in set S.

Number of function = 5^n

Here $n \to \infty$

⇒ Option (A) is correct.

Option (B) is incorrect (obvious)



(C)

For continuous function

Each interval will have 5 choices.

⇒ Number of continuous functions

$= 5^4 = 625$

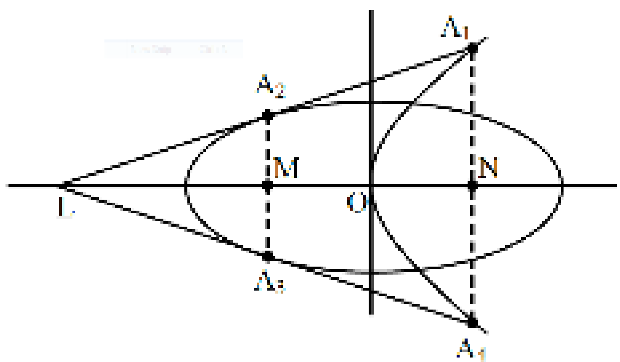
⇒ Option (C) is correct.

(D) Every continuous function is piecewise constant function

⇒ Differentiable.

Option (D) is correct.

40) Shift origin to (1, 2) ⇒ E : $\frac{x^2}{6} + \frac{y^2}{3} = 1$, P : $y^2 = 12x$



$y = mx + \frac{3}{m}$

$C^2 = a^2m^2 + b^2$

$\frac{9}{m^2} = 6m^2 + 3 \quad \Rightarrow m^2 = 1$

T_1 & T_2

$y = x + 3,\ y = -x - 3$

Cuts x-axis at (–3, 0)

$A_1(3, 6) \quad A_4(3, -6)$

$A_2(-2, 1) \quad A_3(-2, -1)$

$A_1A_4 = 12,\ A_2A_3 = 2,\ MN = 5$

Area $= \frac{1}{2}(12+2) \times 5 = 35$ sq.unit

Ans. (A, C)

41) $f(x) = \frac{x^3}{3} - x^2 + \frac{5x}{9} + \frac{17}{36}$

$f'(x) = x^2 - 2x + \frac{5}{9}$

$f'(x) = 0$ at $x = \frac{1}{3}$ in [0, 1]

A_R = Area of Red region

A_G = Area of Green region

$A_R = \int_0^1 f(x)dx = \frac{1}{2}$

Total area = 1

$\Rightarrow A_G = \frac{1}{2}$

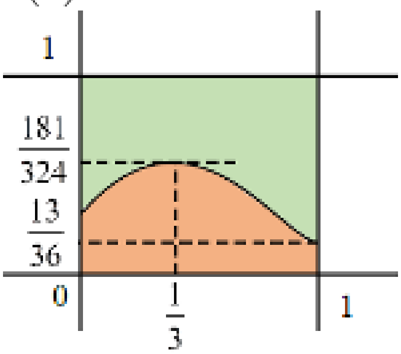
$\int_0^1 f(x)dx = \frac{1}{2}$

$A_G = A_R$

$f(0) = \frac{17}{36}$

$f(1) = \frac{13}{36} \approx 0.36$

$f\left(\frac{1}{3}\right) = \frac{181}{324} \approx 0.558$



(A) Option (D) is remaining coloured part of option (C), hence option (A) is also correct.

(B) Correct when $h = \frac{1}{4}$

$\Rightarrow$ (B) is correct

(C) When $h = \frac{181}{324}, A_R = \frac{1}{2}, A_G < \frac{1}{2}$

$h = \frac{13}{36}, A_R < \frac{1}{2}, A_G = \frac{1}{2}$

$\Rightarrow A_R = A_G$ for some $h \in \left(\frac{13}{36}, \frac{181}{324}\right)$

$\Rightarrow$ (C) is correct

(D) Correct when $h = \frac{3}{4}$ but $h \in \left[\frac{1}{4}, \frac{2}{3}\right]$

$\Rightarrow$ (D) is incorrect

42) Given $x + 2y + z = 4$ .... (1)

$x + \alpha z = 10$ .... (2)

$2x - 3y + \beta z = \gamma + 1$ .... (3)

Now, $\Delta = \begin{vmatrix} 4 & 2 & 1 \\ 10 & 0 & \alpha \\ \gamma+1 & -3 & \beta \end{vmatrix} = \begin{vmatrix} 0 & 2 & 0 \\ 10 & 0 & \alpha \\ \gamma+7 & -3 & \beta+\frac{3}{2} \end{vmatrix}$

$2\gamma\alpha + 14\alpha - 20\beta - 30$

□ if $\beta = \frac{1}{2}(7\alpha - 3)$

$\Rightarrow \boxed{\Delta = 0}$

Now, $\Delta_x = \begin{vmatrix} 7 & 2 & 1 \\ 11 & 0 & \alpha \\ \gamma & -3 & \beta \end{vmatrix}$

$= 21\alpha + 14\alpha - 20\beta - 30$

□ if $\gamma = 28$

$\Rightarrow \Delta_x = 0$

Similarly

□ if $\gamma = 28$

$\Rightarrow \Delta_y = 0$

If $\gamma = 28$

$\Rightarrow \boxed{\Delta_z = 0}$

□ if $\gamma = 28$ and $\beta = \frac{1}{2}(7\alpha - 3)$

$\Rightarrow$ system has infinite solution

and if $\gamma \neq 28$

$\Rightarrow$ system has no solution

$\Rightarrow P \rightarrow (3)$; $Q \rightarrow (2)$

Now if $\beta \neq \frac{1}{2}(7\alpha - 3)$

$\Rightarrow \Delta \neq 0$ $\Rightarrow$ system has unique solution

and for $\alpha = 1$ clearly

$y = -3$ is always be the solution

□ if $\gamma \neq 28$

System has a unique solution

if $\gamma = 28$

$\Rightarrow x = 10, y = -3$ and $z = 0$ will be one of the solution
 R → 1 ; S → 4
 option 'A' is correct

43) x_i 3 4 5 8 10 11
f_i 5 4 4 2 2 3
(P) Mean
(Q) Median
(R) Mean deviation about mean
(S) Mean deviation about median

x_i	f_i	x_if_i	C.F.	$\|x_i - \text{Mean}\|$	$f_i\|x_i - \text{Mean}\|$	$\|x_i - \text{Median}\|$	$f_i\|x_i - \text{Median}\|$
3	5	15	5	3	15	2	10
4	4	16	9	2	8	1	4
5	4	20	13	1	4	0	0
8	2	16	15	2	4	3	6
10	2	20	17	4	8	5	10
11	3	33	20	5	15	6	18
	$\Sigma f_i = 20$	$\Sigma x_if_i = 120$			$\Sigma f_i\|x_i - \text{Mean}\| = 54$		$\Sigma f_i\|x_i - \text{Median}\| = 48$

(P) $\text{Mean} = \frac{\Sigma x_i f_i}{\Sigma f_i} = \frac{120}{20} = 6$

(Q) $\text{Median} = \left(\frac{20}{2}\right)^{th}$ observation = 10^{th} observation = 5

(R) Mean deviation about mean = $\frac{\Sigma f_i |x_i - \text{Mean}|}{\Sigma f_i} = \frac{54}{20} = 2.70$

(S) mean deviation about median = $\frac{\Sigma f_i |x_i - \text{Median}|}{\Sigma f_i} = \frac{48}{20} = 2.40$

44) $L_1 : \vec{r}_1 = \lambda\left(\hat{i} + \hat{j} + \hat{k}\right)$
$L_2 : \vec{r}_2 = \hat{j} - \hat{k} + \mu\left(\hat{i} + \hat{k}\right)$
Let system of planes are
$ax + by + cz = 0$ ..... (1)
 It contain L_1
 $a + b + c = 0$ .... (2)
For largest possible distance between plane (1) and L_2 the line L_2 must be parallel to plane (1)
 $a + c = 0$ .... (3)
$\Rightarrow \boxed{b = 0}$
 Plane H_0 : $\boxed{x - z = 0}$
Now $d(H_0) = \perp$ distance from point (0, 1, –1) on L_2 to plane.
$\Rightarrow d(H_0) = \left|\frac{0+1}{\sqrt{2}}\right| = \frac{1}{\sqrt{2}}$

 $P \to 5$

for 'Q' distance = $\left|\dfrac{2}{\sqrt{2}}\right| = \sqrt{2}$

 $Q \to 4$

 (0, 0, 0) lies on plane

 $R \to 3$

for 'S' x = z ; y = z ; x = 1

 point of intersection p(1, 1, 1).

 $OP = \sqrt{1+1+1} = \sqrt{3}$

 $S \to 2$

 option [B] is correct

45) Replace 3z by z

$\because |z|^3 + 2z^2 + 4\bar{z} - 8 = 0$ .... (1)

Take conjugate both sides

$\Rightarrow |z|^3 + 2\bar{z}^2 + 4z - 8 = 0$ .... (2)

By (1) – (2)

$\Rightarrow 2\left(z^2 - \bar{z}^2\right) + 4(\bar{z} - z) = 0$

$\Rightarrow \boxed{z + \bar{z} = 2}$ .... (3)

$\Rightarrow |z + \bar{z}| = 2$ .... (4)

Let z = x + iy

 $\boxed{x = 1}$ z = 1 + iy

Put in (1)

$\Rightarrow (1 + y^2)^{3/2} + 2(1 - y^2 + 2iy) + 4(1 - iy) - 8 = 0$

$\Rightarrow (1 + y^2)^{3/2} = 2(1 + y^2)$

$\Rightarrow \sqrt{1 + y^2} = 2 = |z|$

Also $\boxed{y = \pm\sqrt{3}}$

 $z = 1 \pm i\sqrt{3}$

$\Rightarrow z - \bar{z} = \pm 2i\sqrt{3}$

$\Rightarrow |z - \bar{z}| = 2\sqrt{3}$

$\Rightarrow |z - \bar{z}|^2 = 12$

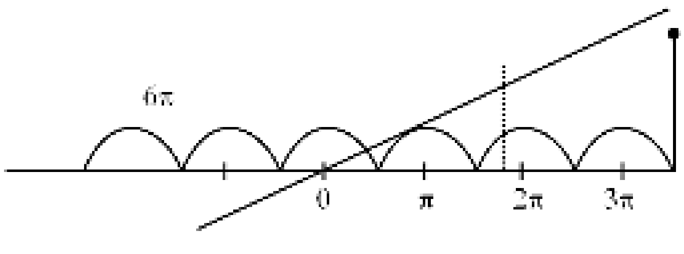
Now $z + 1 = 2 + i\sqrt{3}$

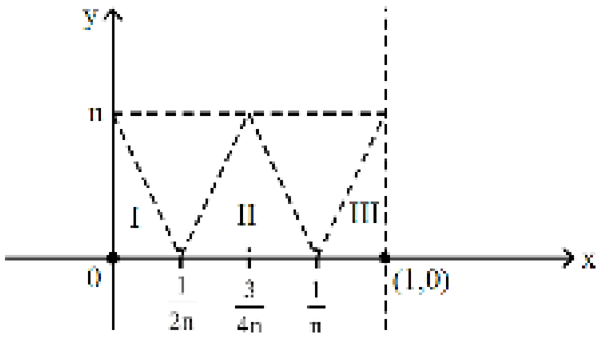
$|z + 1|^2 = 4 + 3 = 7$

 $P \to 2$; $Q \to 1$; $R \to 3$; $S \to 5$

 Option [C] is correct.

46) $|\cos x| = \tan^{-1} \tan\frac{x}{6}$



47) 

Area = Area of (I + II + III) = 4

$= \frac{1}{2} \times \frac{1}{2n} \times n + \frac{1}{2} \times \frac{1}{2n} \times n + \frac{1}{2}\left(1 - \frac{1}{n}\right) \times n$

$= \frac{1}{4} + \frac{1}{4} + \frac{n-1}{2} = 6$

$\boxed{n = 12}$

48) S = 77 + 757 + 7557 + ... + $7\underbrace{5.....5}_{100}7$

10S = 770 + 7570 + ... + 75 ... 570 + 755 570

9S = -77 + $\underbrace{13 + 13 + ... + 13}_{100 \text{ times}}$ + $7\underbrace{5.....5}_{100}70$

= -77 + 1300 + $7\underbrace{5.....5}_{101}7$ + 13

$S = \frac{7\underbrace{5.....5}_{101}7 + 1236}{9}$

m = 1236

n = 9

m + n = 1245

49) $\frac{133 - i(900\cos\theta + 1686\sin\theta)}{7 - 3i\cos\theta} = 300 - \frac{1967 + 1686i\sin\theta}{7 - 3i\cos\theta}$

$= 300 - \frac{281(7 + 6i\sin\theta)}{7 - 3i\cos\theta} \times \frac{7 + 3i\cos\theta}{7 + 3i\cos\theta}$

$= 300 - \frac{281(49 - 18\sin\theta\cos\theta + i(21\cos\theta + 42\sin\theta))}{49 + 9\cos^2\theta}$

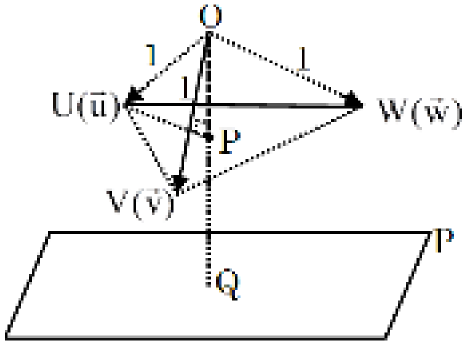
for positive integer

Im(A) = 0

$21\cos\theta + 42\sin\theta = 0$

$\tan\theta = \frac{-1}{2}$; $\sin 2\theta = \frac{-4}{5}$; $\cos^2\theta = \frac{4}{5}$

$\text{Re}(A) = 300 - \frac{281(49 - 9\sin 2\theta)}{49 + 9\cos^2\theta}$

50) 

Given $|\vec{u}-\vec{v}| = |\vec{v}-\vec{w}| = |\vec{w}-\vec{u}|$

$\Rightarrow$ ΔUVW is an equilateral Δ

Now distances of U, V, W from $P = \dfrac{7}{2}$

$\Rightarrow PQ = \dfrac{7}{2}$

Also, Distance of plane P from origin

$\Rightarrow$ OQ = 4

 OP = OQ – PQ $\Rightarrow OP = \dfrac{1}{2}$

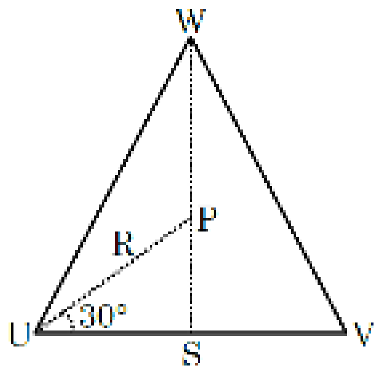
Hence, $PU = \sqrt{OU^2 - OP^2} \Rightarrow PU = \dfrac{\sqrt{3}}{2} = R$

Also, for ΔUVW, P is circumcenter

 for ΔUVW : US = Rcos30°

$\Rightarrow$ UV = 2Rcos30°

$\Rightarrow UV = \dfrac{3}{2}$



 $Ar(\Delta UVW) = \dfrac{\sqrt{3}}{4}\left(\dfrac{3}{2}\right)^2 = \dfrac{9\sqrt{3}}{16}$

 Volume of tetrahedron with coterminous edges $\vec{u}, \vec{v}, \vec{w}$

$= \dfrac{1}{3}(Ar.\Delta UVW) \times OP = \dfrac{1}{3} \times \dfrac{9\sqrt{3}}{16} \times \dfrac{1}{2} = \dfrac{3\sqrt{3}}{32}$

 parallelopiped with coterminous edges

$\vec{u}, \vec{v}, \vec{w} = 6 \times \dfrac{3\sqrt{3}}{32} = \dfrac{9\sqrt{3}}{16} = V$

 $\dfrac{16}{\sqrt{3}}V = 9$

51) $T_{r+1} = {}^4C_r\left(a.x^3\right)^{4-r}.\left(\dfrac{70}{27bx^2}\right)^r$

$= {}^4C_r.a^{4-r}.\frac{70^r}{(27b)^r}.x^{12-5r}$

here 12 – 5r = 7

⇒ r = 1

∴ coeff. = $4.a^3.\frac{70}{27b}$

$T_{r+1} = {}^7C_r(ax^2)^{7-r}\left(\frac{-1}{bx^3}\right)^r$

$= {}^7C_r.a^{7-r}\left(\frac{-1}{b}\right)^r.x^{14-5r}$

14 – 5r = -6 ⇒ r = 4

coeff. : ${}^7C_4.a^3.\left(\frac{-1}{b}\right)^4 = \frac{35a^3}{b^4}$

now $\frac{35a^3}{b^4} = \frac{280a^3}{27b}$

$b^3 = \frac{35 \times 27}{280} = b = \frac{3}{2}$ ⇒ 2b = 3