	Date
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	Aim:-
	To identity the
	To identify the extra element and the functional group present in the given sample (salicytic acid).
	Reference:
	Requirement:
	Ignition tube, Pain of tongs, test tube holder, China dish
	Ignition tube, Pain of tongs, test tube holder, China dish tripod stand, wwise gauze, funnel and filter paper.
	Theory :-
	Sodium bicarbonate fest:
	Place a little of the substance in a test tube and
	The acids from their corresponding radium salts with
	the liberation of earbon dioxide.
	COONA
	ОН
	+ NOHCO3 -> + H2O + CO2
	Solium Solium salt Carbon diorride Solium salt of soliutio aud
	Salicylic acid bisarborate of salicylic acid  Teacher's Signature
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Ester test:	
Heat about 0.5 g of the sample with	In of methyl
alt alcohol and a few drops of concentra	ted Hosog in a
dry test lube for about a minute. Cool	and bown the
contents to a few ml of water contain	ed in a small beaker.
Salighic acid can be identified by the	e colour of
wintergreen oil.	
	снэ
	THE PARTY AND ADDRESS OF THE PARTY OF THE PA
+ CH20H Conc. H2504	V <sup>0H</sup>
+ CH20H Conc. H250a	H20
Salicylie Methyl Methyl : acid alihol odow of i	salicylate
acid alihal odown of i	wintergreen oil
Concentrated H2509 checks the revers	ibility of the
reaction and acts as a dehydrating	
Acriflavine test:	
Appearance of yellow or brown yellow	precipitate confirms
Appearance of yellow or brown yellow the presence of COOH growp.	
Ferric chloride test:	
This fest is carried out in a control new	to al solution. Put
about 0.5, of the substance in a boiling	test the and gdd
dilute ammoria solution with the contexts.	are just alkaline
Teacher's Sign	nature

S.No	Experiment	Observation	Inference
1	Sodum bicarbonate test - Place a little		Verilla State
1.	of the substance in a test tube and add	TO LEAD OF	Alle Me
	2 ml of 51 agueous sodium bicarbonate sol .	CALL AND A	
Annie 1	Ester formation - Head about 0.5g of	V-104-10-13	
2.	the sample with 1 ml of methyl alcohol	ok ans. Nos	
	and add a few derops of conc. H2504	Na.	
	in a day test tube for about a minute.		
	Good and power the contents to a few ml		
	of water contained in a small beaker.		<u> </u>
	Acoiflavine test To a mixture of 0.2 1.	HALHER	War Janes
3.	acriflovine and I'm potassium chromate		
	solution odd a few drops of eguerns	Dell's M	Taylor Server
-	solution of the given sample.	Y ANNA SAN	, , ,
	Ferrie Chloride Test - Put about 0.59 of the		
4.	substance in a boiling test tube and add distribe		Jardway
	ammonia solution until the contents are just		whate.
	alkaline to litmus paper. Boil the solution		
	gently until the odown of ammonio is not there. To cold solution odd few strops of		and the A
	ferric shloode solution.		
	Benzylthiousonium test - Dissolve about		
5.	0.5g of the sample in minimum volume		
	of water. Add 5 % sodium hydroxide sol2		
	until the solution is just alkaline to methy!		
3.8	orange. Add a drop of dilute HCI. In another		
1000	test tube dissolve 0.8 g of benzylthiouxonium		
	chloride in 2 m of water. Mix the 2.		

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to litmus paper. Boil the solution gently odown of ammonia is that not there. To add few doops of ferric chloride solution. colour indicates the presence of salighing Benzyl thiouronium Test	until the the cold solution A violet acid.
Benzylthiousonium salt of the acid step Echemistry same as previous experiment).	
6 + Fe Cl 3 3H FO	e + HCI
Salicylic Ferric Violet coloured  acid Chloride Complex	Hydrochloric acid.
Result:-	
The given organic compound has been strain the laboratory.	rdied successfully
Teacher's Signatur	re