

will tend to lie in the direction in which the orbital has the maximum value called hybrid orbitals.

ii) Intermolecular force:-

We must remember that the particular method of mentally building molecules that we are learning to use is artificial. It is a purely intellectual process involving imaginary overlap of imaginary orbitals.

— There are other equally artificial ways that use different mental or physical models. Our method is the one that so far has seemed to work out best for the chemist.

But however we arrive at we see the actual structure of a molecule to be the net result of a combination of repulsive & attractive forces. We are related to charge & electron spin.

* Repulsive force:- Electrons tend to stay as far apart as possible because they have the same charge and also if they are unpaired because they have the same spin.
(Pauli exclusion principle)

— The like charged atomic nuclei tend to repel each other