

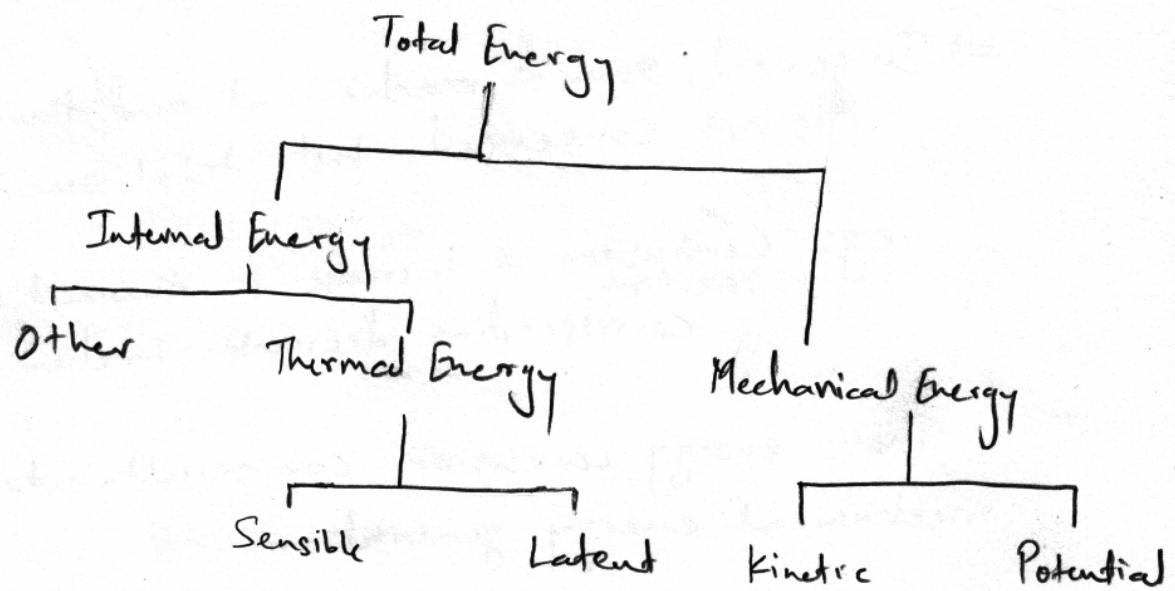
Forms of Energy

First law of thermodynamics

"total energy of a system is conserved"

"energy can neither be created nor destroyed during a process; it can only change forms"

Total Energy is a sum of numerous forms of energy such as thermal, mechanical, kinetic, potential, chemical, nuclear, magnetic, sensible, latent, Coulombic, etc.



Thermal Energy \equiv heat

Sensible heat \equiv kinetic energy of the molecules
(Changes in temperature) (translation, vibration, rotation)

Latent heat \equiv intermolecular forces
(Change in phase)

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(1)

Energy Balance

First law of Thermodynamics

Closed System, over a time interval

$$\text{Change in total energy} = \text{Net heat transferred to the system} - \text{Work done by the system}$$

Control volume, at an instant (or over a time interval Δt)

$$\text{Change in total energy} = \text{Energy entering the control volume} - \text{Energy leaving the control volume}$$

⇒ In general, sum of mechanical and thermal energy is not conserved, but total energy is.

e.g.: Combustion reaction ⇒ increase in thermal energy with corresponding decrease in chemical energy.

Thus, energy conversion can result into thermal or mechanical energy generation

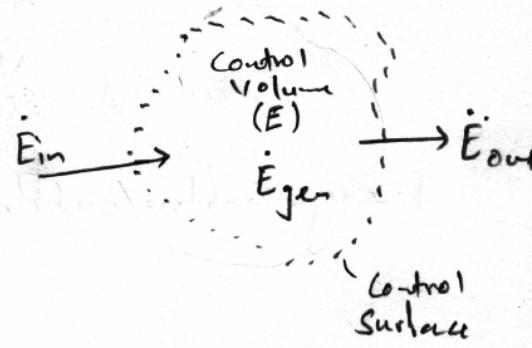
$$\text{Change in thermal + mechanical energy} = \text{Thermal & mechanical energy (entering - leaving) the control volume} + \text{Thermal & mechanical energy generated}$$

* generation can be positive or negative

⇒ at steady state, entering-leaving + generation = 0

(15)

Denoting thermal + mechanical energy by the symbol E , its
rate of change is given by

$$\dot{E} = \frac{dE}{dt} = \dot{E}_{in} - \dot{E}_{out} + \dot{E}_{gen}$$


Note:

Including \dot{E}_{gen} does not violate the first law of thermodynamics.

Reminder:

$E \equiv$ Thermal plus mechanical energy

Mechanical \equiv kinetic and potential energy. (often unchanged)

Thermal \equiv sensible and latent heat

No phase change \Rightarrow No latent heat change

For incompressible substances

$$\frac{\text{change in sensible heat}}{\text{mass}} \equiv m C_v \Delta T$$

$\uparrow \quad \uparrow$

mass change in temperature

$C_v \equiv$ specific heat at constant volume

Specific heat \equiv energy required to change the temperature of a unit mass by one degree

$C_v \equiv$ specific heat at constant volume

$C_p \equiv$ specific heat at constant pressure

For a control volume, C_v is the relevant specific heat for incompressible liquids $C_p = C_v$

for an ideal gas $C_p - C_v = R$

Energy generation: Conversion of some other form of energy to thermal and/or mechanical energy
Volumetric phenomena

Energy inflow/outflow: Surface phenomena

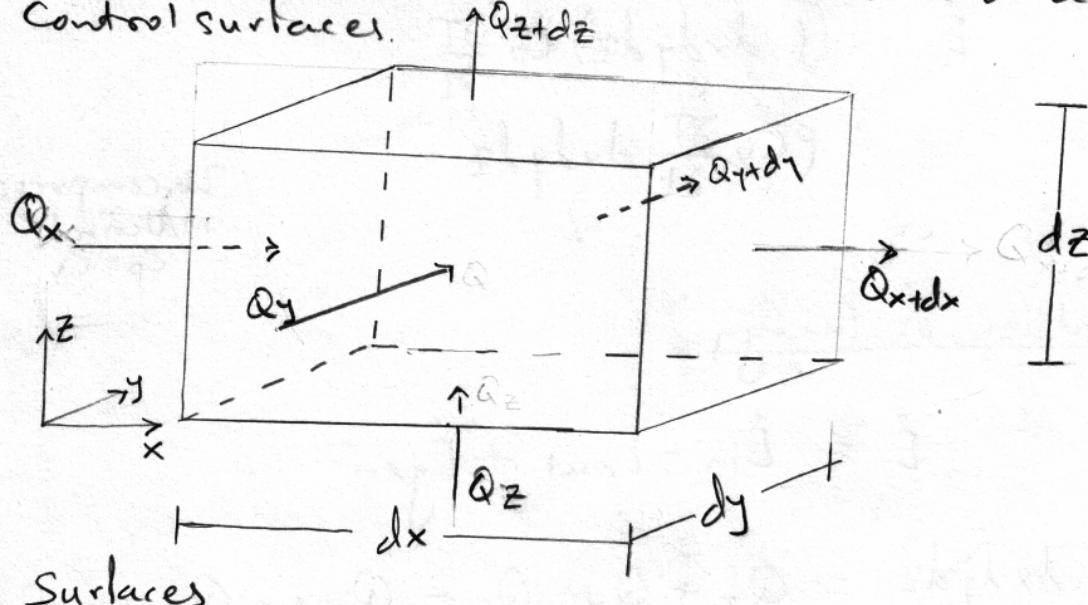
thermal + mechanical energy transferred at the boundaries (control surface enclosing the control volume)

Heat Conduction Equation

Major objective in many problems \rightarrow knowing the temperature field.

- Consider a homogeneous medium. e.g. a metal object
- Assume that the medium is incompressible and there is no bulk motion
- Incompressible \Rightarrow "constant density"
- No bulk motion \Rightarrow no changes in mechanical energy

Defining an infinitesimally small volume as our control volume ($dx \cdot dy \cdot dz$) enclosed by a control surface as shown in the figure, we only consider thermal energy in the control volume and across the control surfaces.



Surfaces

A temperature gradient in the x -direction will give rise to conduction in the x -direction. Considering the rates of heat transfer at surfaces located at x and $x+dx$, as Q_x and Q_{x+dx} , Taylor Series expansion (truncated at first order) gives

$$Q_{x+dx} = Q_x + \frac{\partial Q_x}{\partial x} dx$$

Similarly, in the y - and z -directions

$$Q_{y+dy} = Q_y + \frac{\partial Q_y}{\partial y} dy$$

$$Q_{z+dz} = Q_z + \frac{\partial Q_z}{\partial z} dz$$

Control volume

If there is an energy source associated with the rate of thermal energy generation \dot{E}_{gen}

$$\dot{E}_{gen} = \dot{e}_{gen} dx dy dz$$

$$\dot{e}_{gen} = \dot{E}_{gen}/\text{Volume}$$

In absence of phase change, the rate of thermal energy change is equal to the rate of change of sensible heat

$$\begin{aligned}\dot{E} &= (\rho dx dy dz) C_p \frac{\partial T}{\partial t} \\ &= \rho C_v \frac{\partial T}{\partial t} dx dy dz\end{aligned}$$

In compressible material
 $C_p = C_v$

Conservation of Energy

$$\dot{E} = \dot{E}_{in} - \dot{E}_{out} + \dot{E}_{gen}$$

$$\begin{aligned}\rho C_p \frac{\partial T}{\partial t} dx dy dz &= Q_x + Q_y + Q_z - Q_{x+dx} - Q_{y+dy} - Q_{z+dz} \\ &\quad + \dot{e}_{gen} dx dy dz \\ &= -\frac{\partial Q_x}{\partial x} dx - \frac{\partial Q_y}{\partial y} dy - \frac{\partial Q_z}{\partial z} dz \\ &\quad + \dot{e}_{gen} dx dy dz\end{aligned}$$

Thus,

$$\rho c_p \frac{\partial T}{\partial t} = -\frac{1}{A_x} \frac{\partial Q_x}{\partial x} - \frac{1}{A_y} \frac{\partial Q_y}{\partial y} - \frac{1}{A_z} \frac{\partial Q_z}{\partial z} + \dot{e}_{gen}$$

when $A_x = dy dz$ = area of each of the two surfaces along x-direction

$$A_y = dx dz$$

$$A_z = dx dy$$

Fourier's law of heat conduction

$$q_x = -k A_x \frac{\partial T}{\partial x}$$

$$q_y = -k A_y \frac{\partial T}{\partial y}$$

$$q_z = -k A_z \frac{\partial T}{\partial z}$$

Substituting in the energy balance,

$$\rho c_p \frac{\partial T}{\partial t} = \frac{\partial}{\partial x} (k \frac{\partial T}{\partial x}) + \frac{\partial}{\partial y} (k \frac{\partial T}{\partial y}) + \frac{\partial}{\partial z} (k \frac{\partial T}{\partial z}) + \dot{e}_{gen}$$

— Heat conduction equation in Cartesian coordinates

Notes:

- Energy balance + Fourier's law of conduction
- Solution: $T(x, y, z, t)$
- $k \frac{\partial T}{\partial x} = q_x$ = flux. Thus, at steady state, with no energy source/generation term, heat conduction equation $\Rightarrow q_x$ = constant in the x-direction

In many common problems, $k = \text{constant}$. The conduction equation reduces to (20)

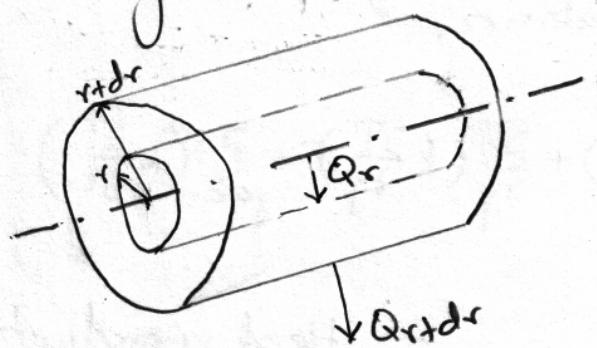
$$\nabla^2 T + \dot{\epsilon}_{gen} = \frac{1}{\alpha} \frac{\partial T}{\partial t}$$

where

$\alpha = \frac{k}{\rho C_p}$ = thermal diffusivity
 (a material property)
 (how fast heat propagates)
 through the material

Cylindrical Coordinates

For the case of 1-dimensional heat conduction (in the radial direction), a control volume corresponding to a thin cylindrical shell can be defined as shown in the figure.



Surface

$$Q_{r+dr} = Q_r + \frac{\partial Q_r}{\partial r} dr$$

Control volume

$$\dot{\epsilon}_{gen} = \dot{\epsilon}_{gen} 2\pi r L dr$$

Area of both surfaces can be approximated as $2\pi r L$, since $dr \rightarrow 0$

$$A_r = 2\pi r L$$

$$\dot{E} = \rho C_p \frac{\partial T}{\partial t} A_r dr$$

Conservation of Energy

$$\rho C_p \frac{\partial T}{\partial t} = - \frac{1}{A_r} \frac{\partial Q_r}{\partial r} + \dot{\epsilon}_{gen}$$

Fourier law of heat conduction

$$Q_r = - k A_r \frac{\partial T}{\partial r}$$

$$\rho C_p \frac{\partial T}{\partial t} = \frac{1}{r} \frac{\partial}{\partial r} \left(r k \frac{\partial T}{\partial r} \right) + \dot{e}_{gen} \quad (21)$$

- Heat conduction equation
in cylindrical coordinates,
(1-D; radial)

Three-Dimensional Heat Conduction equation in
Cylindrical coordinates

$$\rho C_p \frac{\partial T}{\partial t} = \frac{1}{r} \frac{\partial}{\partial r} \left(r k \frac{\partial T}{\partial r} \right) + \frac{1}{r^2} \frac{\partial}{\partial \phi} \left(k \frac{\partial T}{\partial \phi} \right) + \frac{\partial}{\partial z} \left(k \frac{\partial T}{\partial z} \right) + \dot{e}_{gen}$$

Spherical Coordinates

$$\rho C_p \frac{\partial T}{\partial t} = \frac{1}{r^2} \frac{\partial}{\partial r} \left(k r^2 \frac{\partial T}{\partial r} \right) + \frac{1}{r^2 \sin^2 \theta} \frac{\partial}{\partial \phi} \left(k \frac{\partial T}{\partial \phi} \right) + \frac{1}{r^2 \sin \theta} \frac{\partial}{\partial \theta} \left(k \sin \theta \frac{\partial T}{\partial \theta} \right) + \dot{e}_{gen}$$

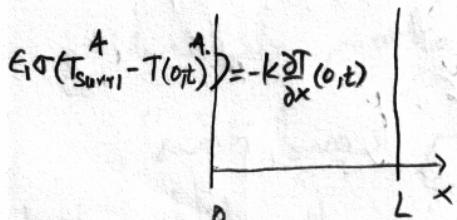
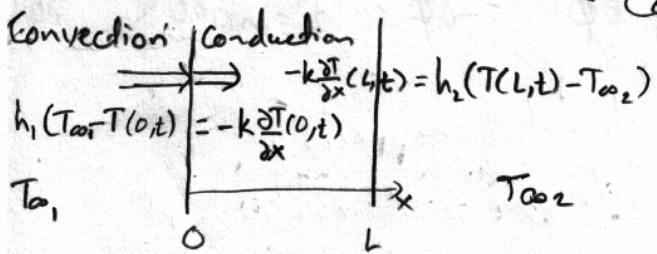
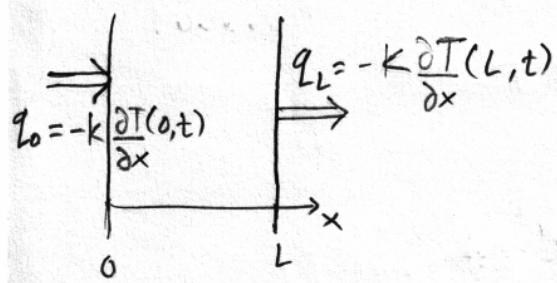
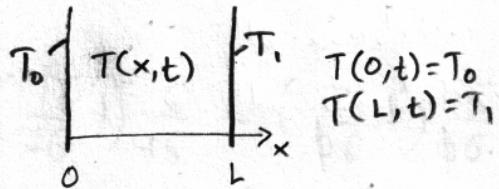
After choosing the appropriate coordinate system
and simplifying the heat conduction equation, depending
on the problem, a number of "boundary conditions"
and an "initial condition" are required to solve the differential
equation.

Initial Condition

$T(x, r, z, 0)$ specified

Boundary Conditions

- Thermal conditions at the boundary
- Required to integrate/solve the governing heat conduction equation
- Commonly encountered boundary conditions



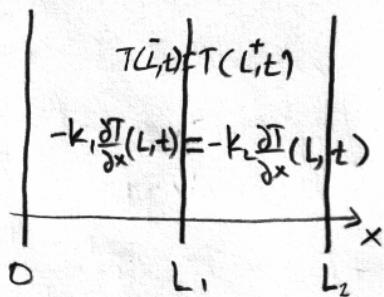
• Convection boundary condition

- most common
- Conduction at surface = convection at surface in the same direction

• Radiation boundary condition

- less common
- Conducting material surrounded by vacuum \Rightarrow no convection
- typically negligible compared to convection, especially forced convection.

• Interface boundary Condition



- two layers in contact

- Same temperature at contact surface

- matching flux at contact surface

• Combined/Generalized boundary conditions

- heat transfer to heat transfer from the surface in = the surface in all modes all modes

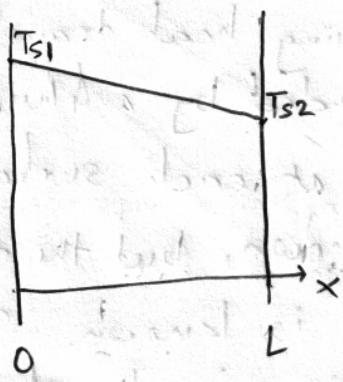
Steady State Conduction:

One Dimensional

Consider steady heat transfer through a wall. If temperature at different locations of the same surface of the wall is a constant, heat transfer can be assumed to be 1-D (one dimensional) in the direction perpendicular to the surface, because heat transfer is governed by a temperature gradient.

Choice: Cartesian coordinate system
Assumption:

- Steady state
- No heat source or sink
- One dimensional heat conduction



Simplified Heat conduction equation:

$$0 = \frac{d}{dx} \left(k \frac{dT}{dx} \right)$$