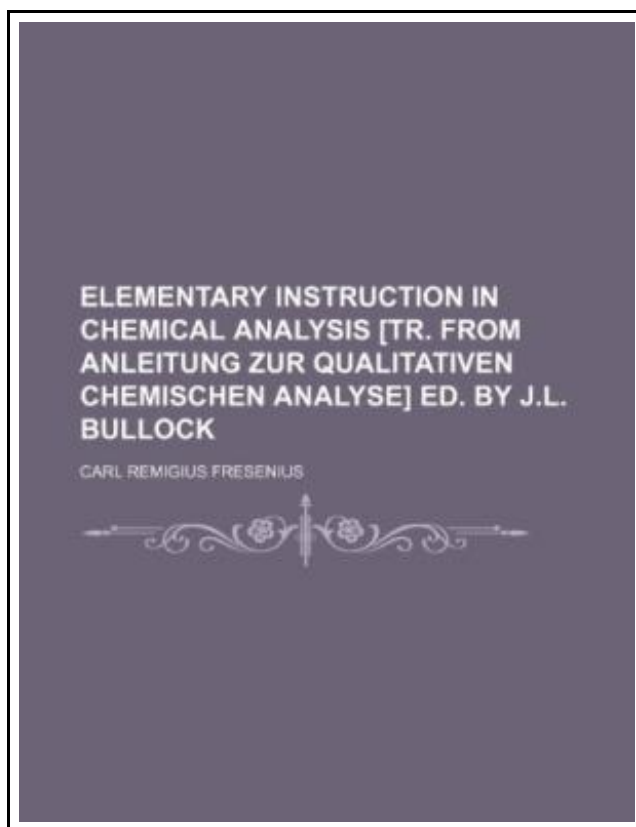


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Rarebooksclub.com, United States, 2012. Paperback. Book Condition: New. 246 x 189 mm. Language: English . Brand New Book ***** Print on Demand *****.This historic book may have numerous typos and missing text. Purchasers can download a free scanned copy of the original book (without typos) from the publisher. Not indexed. Not illustrated. 1864 Excerpt: .are soluble in water; but the greater part of them are insoluble in that menstruum. All the ferrocyanides are decomposed by ignition; where they are not quite anhydrous, hydrocyanic acid, carbonic acid, and ammonia escape, otherwise nitrogen and occasionally cyanogen. In solutions of hydroferrocyanic acid or of soluble ferrocyanides sesquichloride of iron produces a blue precipitate of FerroCyanide Of Iron (Fe_4Cfy ,); sulphate of oxide of copper a brownish-red precipitate of Ferrocyanide Of Copper (Cu_2Cfy); nitrate of silver a white precipitate of Ferrocyanide Of Silver (Ag_2Cfy), which is insoluble in nitric acid and in ammonia, but dissolves in cyanide of potassium. Insoluble ferrocyanides of metals are decomposed by boiling with solution of soda, ferrocyanide of sodium being formed and the oxides thrown down, unless they are soluble in solution of soda. When heated with 3 parts of sulphate and 1 part of nitrate of ammonia, they yield sulphates of the metals contained in them, the whole of the cyanogen volatilizing in form of cyanide of ammonium and the products of its decomposition (bolley). 6. Hydroferricyanic acid (H_2Cfdy). Hydroferricyanic acid and many of the ferricyanides are soluble in water; all ferricyanides are decomposed by ignition in a similar manner as the ferrocyanides. In the aqueous solutions of hydroferricyanic acid and its salts sesquichloride of iron produces no blue precipitate; but sulphate of protoxide of iron produces a blue precipitate of Protoferricyanide Of Iron ($3\text{Fe}_2\text{Cfdy}$); sulphate of copper a yellowish-green precipitate of Ferricyanide Of...



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