College of Industrial Technology

King Mongkut's University of Technology North Bangkok

เลขที่นั่งสอบ

Final Examination of Semester 1 **Year:** 2012

Subject: 392151 Chemistry I **Section: 15-16**

Date: 28 September 2012 **Time:** 08.00-10.00

Name: _____ ID: _____ Field of Study: _____

Instructions:

1. Cheating will result in failure of all classes registered for the current semester. Students who are caught cheating will also be denied registering for the following semester.
2. No documents are allowed to be taken out of the examination room.
3. Calculators are allowed in the examination.
4. Dictionaries are NOT allowed.
5. This exam is a closed book examination.
6. No any electronic communication devices allow in the exam room.
7. The examination has 6 pages (including this page), 2 sections and a total score of 60 points.
8. Write all your answers on this exam sheet.

WebElements: the periodic table on the world-wide web

http://www.webelements.com/

1	2		3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
hydrogen 1 H 1.0079																		helium 2 He 4.0026
lithium 3 Li 6.941	beryllium 4 Be 9.0122												boron 5 B 10.811	carbon 6 C 12.011	nitrogen 7 N 14.007	oxygen 8 O 15.999	fluorine 9 F 18.998	neon 10 Ne 20.180
sodium 11 Na 22.990	magnesium 12 Mg 24.305												aluminium 13 Al 26.982	silicon 14 Si 28.086	phosphorus 15 P 30.974	sulfur 16 S 32.065	chlorine 17 Cl 35.453	argon 18 Ar 39.948
potassium 19 K 39.098	calcium 20 Ca 40.078		scandium 21 Sc 44.956	titanium 22 Ti 47.867	vanadium 23 V 50.942	chromium 24 Cr 51.996	manganese 25 Mn 54.938	iron 26 Fe 55.845	cobalt 27 Co 58.933	nickel 28 Ni 58.693	copper 29 Cu 63.546	zinc 30 Zn 65.39	gallium 31 Ga 69.723	germanium 32 Ge 72.61	arsenic 33 As 74.922	selenium 34 Se 78.96	bromine 35 Br 79.904	krypton 36 Kr 83.80
rubidium 37 Rb 85.468	strontium 38 Sr 87.62		yttrium 39 Y 88.906	zirconium 40 Zr 91.224	niobium 41 Nb 92.906	molybdenum 42 Mo 95.94	technetium 43 Tc [98]	ruthenium 44 Ru 101.07	rhodium 45 Rh 102.91	palladium 46 Pd 106.42	silver 47 Ag 107.87	cadmium 48 Cd 112.41	indium 49 In 114.82	tin 50 Sn 118.71	antimony 51 Sb 121.76	tellurium 52 Te 127.60	iodine 53 I 126.90	xenon 54 Xe 131.29
caesium 55 Cs 132.91	barium 56 Ba 137.33	57-70 *	lutetium 71 Lu 174.97	hafnium 72 Hf 178.49	tantalum 73 Ta 180.95	tungsten 74 W 183.84	rhenium 75 Re 186.21	osmium 76 Os 190.23	iridium 77 Ir 192.22	platinum 78 Pt 195.08	gold 79 Au 196.97	mercury 80 Hg 200.59	thallium 81 Tl 204.38	lead 82 Pb 207.2	bismuth 83 Bi 208.98	polonium 84 Po [209]	astatine 85 At [210]	radon 86 Rn [222]
francium 87 Fr [223]	radium 88 Ra [226]	89-102 **	lawrencium 103 Lr [262]	rutherfordium 104 Rf [261]	dubnium 105 Db [262]	seaborgium 106 Sg [266]	bohrium 107 Bh [264]	hassium 108 Hs [269]	meitnerium 109 Mt [268]	ununnilium 110 Uun [271]	unununium 111 Uuu [272]	ununbium 112 Uub [277]		ununquadium 114 Uuq [289]				

Key: element name / atomic number / symbol / atomic weight (mean relative mass)

*lanthanoids	lanthanum 57 La 138.91	cerium 58 Ce 140.12	praseodymium 59 Pr 140.91	neodymium 60 Nd 144.24	promethium 61 Pm [145]	samarium 62 Sm 150.36	europium 63 Eu 151.96	gadolinium 64 Gd 157.25	terbium 65 Tb 158.93	dysprosium 66 Dy 162.50	holmium 67 Ho 164.93	erbium 68 Er 167.26	thulium 69 Tm 168.93	ytterbium 70 Yb 173.04
**actinoids	actinium 89 Ac [227]	thorium 90 Th 232.04	protactinium 91 Pa 231.04	uranium 92 U 238.03	neptunium 93 Np [237]	plutonium 94 Pu [244]	americium 95 Am [243]	curium 96 Cm [247]	berkelium 97 Bk [247]	californium 98 Cf [251]	einsteinium 99 Es [252]	fermium 100 Fm [257]	mendelevium 101 Md [258]	nobelium 102 No [259]

Symbols and names: the symbols and names of the elements, and their spellings are those recommended by the International Union of Pure and Applied Chemistry (IUPAC - http://www.iupac.org/). Names have yet to be proposed for the most recently discovered elements 110–112 and 114 so those used here are IUPAC's temporary systematic names. In the USA and some other countries, the spellings aluminum and cesium are normal while in the UK and elsewhere the common spelling is sulphur.
Group labels: the numeric system (1–18) used here is the current IUPAC convention.
Atomic weights (mean relative masses): Apart from the heaviest elements, these are the IUPAC 2001 values and given to 5 significant figures. Elements for which the atomic weight is given within square brackets have no stable nuclides and are represented by the element's longest lived isotope.

Part 1. Multiple Choice (20 points)

Identify the letter of the choice that best completes the statement or answers the question.

___ 1. Calculate the mass of 1.00 mol of SCl_2.

a. 51 amu b. 51 g c. 103 amu d. 103 g

___2. How many moles of hydrogen atoms are presented in 1.25 mol of CH_4?

a. 4 mol b. 5 mol c. 7 mol d. 8 mol

___3. What is the empirical formula of C_5H_{10}.

a. CH b. C_2H c. CH_2 d. C_5H_{10}

____4. Calculate the number of molecules of SO_2 in 41 g of SO_2.

a. 3.86×10^{23} b. 4.01×10^{23} c. 6.02×10^{23} d. 5

____5. Which of the following has the greatest number of atoms?

a. 1 atom b. 1 mole of atom ÷ (6.02×10^{23})

c. 1 mole of He d. 1 mole of NH_3

____6. For the reaction $C + 2H_2 \rightarrow CH_4$, how many moles of hydrogen are required to produce 10 mol of methane, CH_4?

a. 2 mol b. 4 mol c. 10 mol d. 20 mol

____7. The coefficients in a chemical equation represent the

a. masses, in grams, of all reactants and products.

b. relative numbers of moles of reactants and products.

c. number of atoms in each compound in a reaction.

d. number of valence electrons involved in the reaction.

____8. In the reaction $Al_2O_3 \rightarrow Al + O_2$ (unbalanced), what is the mole ratio of aluminum to oxygen?

a. 10:6 b. 2:3 c. 1:1 d. 4:3

___9. Determine percentage of water in $Ba(OH)_2{\bullet}4H_2O$.

a. 10% b. 20% c. 30% d. 40%

___10. A hydrocarbon sample has an empirical formula of CH and the molar mass of this compound is 78 g/mol. What is the chemical formula of this compound?

a. CH_4 b. C_5H_{10} c. C_2H_2 d. C_6H_6

Part 2 Answer (40 points)

1. Calculate the atomic mass of lithium from the following data: (5 points)

Isotope	Natural Relative Abundance (%)	Mass (amu)
^{6}Li	7.5	6.0151
^{7}Li	92.5	7.0160

2. How many moles of H_2O_2 in 34 g of H_2O_2? How many moles of H atoms are there? (5 points)

3. A sample of NH_3 is pumped from a 1.50-L vessel at STP, what is its mass? How many atom of N and H in this gas? (7 points)

4. A forensic scientist analyzes a drug and finds that it contains 80.22% carbon and 9.62% hydrogen. Could the drug be pure tetrahydrocannabinol ($C_{21}H_{30}O_2$)? (6 points)

5. A sample of a hydrocarbon was found to contain 7.2 g of carbon and 1.5 g of hydrogen. The molar mass of this compound was determined to be 58 g mol^{-1}. What are the empirical formula and molecular formula of the compound? (7 points)

6. Balance the following reactions. (4 points)

6.1 $Au(CN)_3 + NaBr \rightarrow AuBr_3 + NaCN$

6.2 $CH_4 + O_2 \rightarrow CO_2 + H_2O$

6.3 $Cl_2 + KBr \rightarrow KCl + Br_2$

6.4 $Cu_2O + CH_4 \rightarrow H2O + Cu + CO$

7. Given the following equation, calculate the mass of O_2 needed to react completely with 7.4 g NO. $2NO + O_2 \rightarrow 2NO_2$ (6 points)