

Ch14 Winter term 2024

Problem set 4

due May 23, 2024

for Problems 1 and 2: for CO_2 , HCO_3^- , CO_3^{2-} , CaCO_3 at 25°C , use
 $k_{\text{H}} = 0.035 \text{Matm}^{-1}$, $K_{\text{a1}} = 4.25 \times 10^{-7}$, $K_{\text{a2}} = 5.01 \times 10^{-11}$, $K_{\text{sp}} = 4.5 \times 10^{-9}$

1

Calculate the pH of the following solutions. The contribution of H_2CO_3 to the total concentration of dissolved CO_2 -related species can be neglected in this analysis; these samples consist only of a solution, with neither gas phase CO_2 nor solid CaCO_3 .

1.1

20mMNaHCO₃

1.1.1 Answer

This suggests that we have 20 mM of HCO_3^- . Now we have the equilibrium expression for the dissociation of a proton from HCO_3^- :



The equilibrium constant for this reaction is given by:

$$K_a = \frac{[\text{H}^+][\text{CO}_3^{2-}]}{[\text{HCO}_3^-]} \quad (2)$$

We can consider $0.02 - x$ as the concentration of HCO_3^- after dissociation and x as the concentration of H^+ and CO_3^{2-} . Substituting these values into the equilibrium constant expression, we get:

$$5.01 \times 10^{-11} = \frac{x^2}{0.02 - x} \quad (3)$$

Solving this equation, we get the pH of the solution is 6.00.

```

1 import sympy as sp
2
3 # Given constants
4 Ka2 = 5.01e-11
5 initial_HCO3 = 0.02
6 x = sp.symbols('x')
7 final_HCO3 = 0.02 - x
8 final_H = x
9 final_CO3 = x
10 eq = Ka2 - (final_H * final_CO3) / final_HCO3
11 x = sp.solve(eq, x)
12 # now get the pH
13 H = x[0]
14 pH = -sp.log(H, 10)
15 pH.evalf()

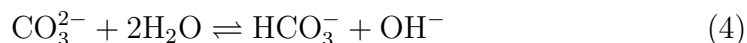
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1.2

20mMNa₂CO₃

1.2.1 Answer

This suggests that we have 20 mM of CO_3^{2-} . We are interested in the hydrolysis of this ionic species in water:



We can use the given K_{a2} to determine the K_b for the hydrolysis of CO_3^{2-} :

$$K_b = \frac{K_w}{K_{a2}} \quad (5)$$

Then the equilibrium expression for this hydrolysis is:

$$K_b = \frac{[\text{HCO}_3^-][\text{OH}^-]}{[\text{CO}_3^{2-}]} \quad (6)$$

This gives a pH of 11.3.

```
1 import sympy as sp
2
3 # Given constants
4 Ka2 = 5.01e-11
5 # calculate the Kb from the Ka2
6 Kw = 1e-14
7 Kb = Kw / Ka2
8 initial_CO3 = 0.02
9 x = sp.symbols('x')
10 final_CO3 = 0.02 - x
11 final_OH = x
12 final_HCO3 = x
13 eq = Kb - (final_OH * final_HCO3) / final_CO3
14 x = sp.solve(eq, x)
15 # now get the pOH
16 OH = x[0]
17 pOH = -sp.log(OH, 10)
18 # convert this into a pH
19 pH = 14 - pOH
20 pH.evalf()
```

2

This problem is modified from the 2022 midterm, problem 6B "Is this a soluble problem?", which, ironically, was not soluble as it was written...

2.1

What partial pressure (in atm) of gas phase CO_2 exists in equilibrium with a solution saturated with CaCO_3 at $\text{pH} = 7.5$? The problem originally specified that the free Ca^{2+} concentration = 0.001 M, but this constraint is inconsistent with the relevant equilibrium constants.

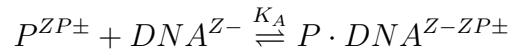
2.1.1 Answer

The solubility product constant for CaCO_3 is given by:

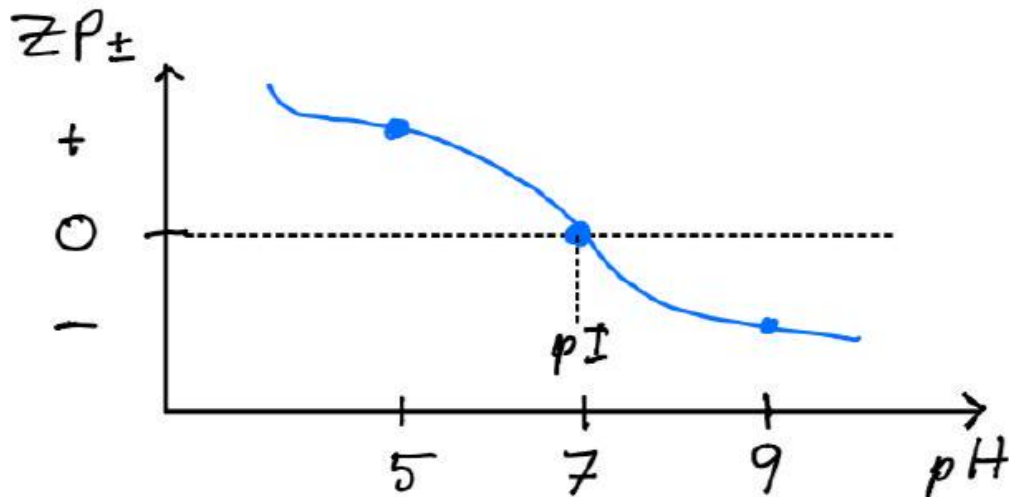
$$K_{sp} = [\text{Ca}^{2+}][\text{CO}_3^{2-}] \quad (7)$$

3

A protein P binds to DNA according to the following reaction



where K_A is the association constant for the binding of P to DNA. P has multiple positively charged amino groups and multiple negatively charged carboxylate groups, with an overall charge of $ZP\pm$. Due to the titration of these ionizable residues, the overall charge $ZP\pm$ is a function of pH as shown below:



At the isoelectric point, pI , the positively charged residues are balanced by the negatively charged residues, so that the net charge is 0. The DNA remains negatively charged at all pHs. The association of P and DNA has a strong electrostatic component based on the net charges of P and DNA; i.e., at pHs lower than the pI , P has a net positive charge, so there is an attractive electrostatic contribution stabilizing complex formation with the negatively charged DNA, while at pHs above the pI , P has a net negative charge so there is a repulsive electrostatic contribution destabilizing complex formation with DNA. At every pH, the charge of the complex is the sum of the charges on the protein at that pH, plus the negative charge on DNA.

From Ch14, we know that equilibrium concentrations involving charged species will depend on ionic strength (I) due to the ionic strength dependence

of the activity coefficients. Based on the ionic strength dependence of the activity coefficients captured in the Debye-Hückel limiting law, answer the following questions and provide a brief (1 sentence) explanation.

3.1

The Debye-Hückel limiting law is given by:

$$\log \gamma_{\pm} = -0.509 z_{\pm}^2 \sqrt{I} \quad (8)$$

3.1.1 Question

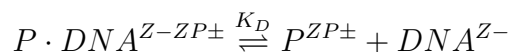
Will the association constant K_A increase or decrease with increasing I at pH5 ?

3.1.2 Answer

As we

- Will the association constant K_A increase or decrease with increasing I at pH5 ?
- Will the association constant K_A increase or decrease with increasing I at pH 9?

This reaction can also be modeled as a dissociation reaction, where K_D is the dissociation constant.

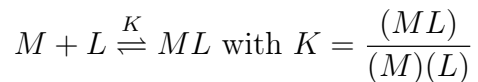


3b. (10 pts)

- Will the dissociation constant K_D increase or decrease with increasing I at pH5 ?
- Will the dissociation constant increase or decrease with increasing I at pH9 ?

Problem 4 (20 points total)

M and L interact to form a complex ML with association constant K .



In our analysis of ligand binding, we typically have assumed that the ligand, L, is in vast excess over the metal, M, so that the binding of ligand to M to form ML does not appreciably impact the concentration of free L. Under these conditions, when $(L) = \frac{1}{K}$

Eq. A

$$\bar{n} = \frac{(ML)}{(M) + (ML)} = \frac{K(L)}{1 + K(L)} = \frac{1}{2}$$

In the case of tight binding systems where the total ligand concentration is comparable to the total metal concentration, however, the formation of ML will influence the amount of free L left in solution. The following example illustrates this point.

Define the total concentrations of the metal and ligand as (M_{tot}) and (L_{tot}) , respectively:

$$\begin{aligned}(M) + (ML) &= (M_{tot}) \\ (L) + (ML) &= (L_{tot})\end{aligned}$$

Taking into account the effect of (ML) on free (L), calculate the numerical value of \bar{n} to 3 significant figures for the case where

$$(M_{tot}) = (L_{tot}) = \frac{1}{K}$$

(continued on next page!)

Hint: incorporate the conservation of mass relations into the association constant expression and solve for (ML) when $(M_{tot}) = (L_{tot}) = \frac{1}{K}$; i.e.

$$K = \frac{(ML)}{(M)(L)} = \frac{(ML)}{[(M_{tot}) - (ML)][(L_{tot}) - (ML)]}$$

Explain briefly why \bar{n} for the tight binding case when $(L) = \frac{1}{K}$ is less than that observed when $(L) = \frac{1}{K}$ for the weak binding case of Eq. A.

Problem 5 (20 points total)

Balance the following reaction,

