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2023, 16, 1783Direct observation of the local microenvironment
in inhomogeneous CO₂ reduction gas diffusion
electrodes *via* versatile pOH imaging†Annette Böhme,^{ab} Justin C. Bui,^{id cd} Aidan Q. Fenwick,^{be} Rohit Bhide,^f
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We report how the micrometer-scale morphology of a carbon dioxide reduction (CO₂R) gas diffusion electrode (GDE) affects the mass transport properties and with it, the local CO₂R performance. We developed a technique to probe the microenvironment in a CO₂R GDE *via* local pOH imaging with time- and three-dimensional spatial, micrometer-scale resolution. The local activity of hydroxide anions (OH[−]), represented by the pOH value, around a GDE in contact with an aqueous electrolyte is a crucial parameter that governs the catalytic activity and CO₂R selectivity. Here, we use fluorescence confocal laser scanning microscopy (CLSM) to create maps of the local pOH around a copper GDE by combining two ratiometric fluorescent dyes, one of which is demonstrated as a pOH sensor for the first time in this work. We observe that the local pOH decreases when current is applied due to the creation of OH[−] as a byproduct of CO₂R. Interestingly, the pOH is lower inside microtrenches compared to the electrode surface and decreases further as trenches become more narrow due to enhanced trapping of OH[−]. We support our experimental results with multiphysics simulations that correlate exceptionally well with measurements. These simulations additionally suggest that the decreased pOH inside microcavities in the surface of a CO₂R GDE leads to locally enhanced selectivity towards multicarbon (C₂₊) products. This study suggests that narrow microstructures on the length scale of 5 μm in a GDE surface serve as local CO₂R hotspots, and thus highlights the importance of a GDE's micromorphology on the CO₂R performance.

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Broader context

Over the past decades, the rising levels of the greenhouse gas CO₂ in the atmosphere have led to a plethora of global environmental hazards. The electrochemical reduction of CO₂ (CO₂R) shows great promise to address this problem as it enables the conversion of CO₂ into useful chemicals and fuels, thereby closing the carbon cycle. The development of gas diffusion electrodes (GDEs) has been a milestone in enhancing the rate of CO₂R because it addresses the problem of mass transport limitations that are encountered in conventional CO₂R setups. Unfortunately, the operating conditions of GDEs are highly complex and difficult to interrogate. They are not yet well enough understood to be commonly applied on an industrial scale. Therefore, insights into the mechanisms that dictate the activity and selectivity of a GDE are urgently needed. In this study, we developed an experimental technique to probe the microenvironment around a CO₂R GDE under operating conditions. We used time-resolved, three-dimensional pOH imaging with a novel sensing mechanism, combined with multiphysics modeling. It was found that the pOH value is locally decreased inside narrow microstructures on the length scale of 5 μm which highlights the influence of the micrometer-scale morphology of a GDE on its CO₂R performance and suggests enhanced selectivity towards multicarbon (C₂₊) products inside microcavities.

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Introduction

It is crucial to store renewable energy effectively for long time periods in order to reach the negative CO₂ emissions required to mitigate the consequences of climate change. Accordingly, the concept of electrochemical CO₂ reduction (CO₂R) is compelling because it enables the storage of renewable energy in the form of chemical bonds. CO₂R establishes CO₂ and water (H₂O) as the primary sustainable feedstocks to form useful chemicals and fuels, thereby closing the carbon cycle.^{1–4}

Though there have been many encouraging advances in the field of electrochemical CO₂R, the process is highly complex and there are still challenges that must be overcome before this technology is energy-efficient and selective enough to be commonly used at an industrial scale.^{1,3,5} One promising approach is the use of gas diffusion electrodes (GDEs).^{6–8} This type of electrode addresses the problem of mass transfer limitations that are encountered in conventional catalytic setups by delivering CO₂ in the gas phase to a catalyst in contact with a liquid electrolyte. Most GDEs consist of a macroporous gas diffusion layer topped with a hydrophobic microporous layer that is then subsequently coated with a catalyst.^{9–11} (Fig. 1(c)). In many cases, the catalyst is wetted by a thin layer of aqueous electrolyte to provide ionic conductivity but not limit CO₂ transport to the surface.^{10,12} This setup allows for increased current densities by more than one order of magnitude over conventional setups: For a conventional setup in which two electrodes are placed in CO₂-saturated electrolyte, mass transport to the cathode limits the rate of CO₂ consumption and hence, the magnitude of the maximum current density does not exceed 30 mA cm^{−2}. On the other hand, tailored GDEs with

current density magnitudes larger than 1 A cm^{−2} have been demonstrated.^{13,14}

Various parameters influence the activity and selectivity of a GDE, and it is the interplay between these operating parameters that dictates the CO₂R performance. It is therefore important to understand the impact of each parameter. One very important parameter that influences the system's performance is the choice of catalyst material.^{15–17} The most commonly studied CO₂R catalyst is copper, because its moderate binding energy to CO as a reaction intermediate allows for its further reduction to desirable higher-order carbon products such as ethanol, propanol, ethylene, and acetate.^{16,18,19} Furthermore, in determining the selectivity and activity of CO₂R, the composition and structure of the GDE, *i.e.* pore size, hydrophobicity, surface depositions or the presence of microstructures, play an ill-understood role.^{9,10,20,21} Other parameters include the applied potential, the properties of the electrolyte (*e.g.* the constituent ions and the viscosity), the cell configuration, and the choice of ion-exchange membranes.^{13,18,19,22–24} In addition to these, the local activities of hydroxides, a_{OH^-} , and protons, a_{H^+} , for which the pOH and pH values are a measure, are critical because both species are involved in the CO₂R process,^{19,25} as per equations (S19)–(S26). The pOH value is defined as $\text{pOH} = -\log_{10}(a_{\text{OH}^-})$ and the pH value as $\text{pH} = -\log_{10}(a_{\text{H}^+})$. In many cases, hydroxide and proton activity can be approximated as the concentration. At equilibrium, $\text{pH} + \text{pOH} = \text{p}K_{\text{w}}$, with $\text{p}K_{\text{w}} = 14$ under ambient conditions, so if the system is in equilibrium, the pH can be directly inferred from the pOH.²⁶

For a GDE at open circuit, CO₂ diffuses through the macroporous and microporous layers into the electrolyte where it

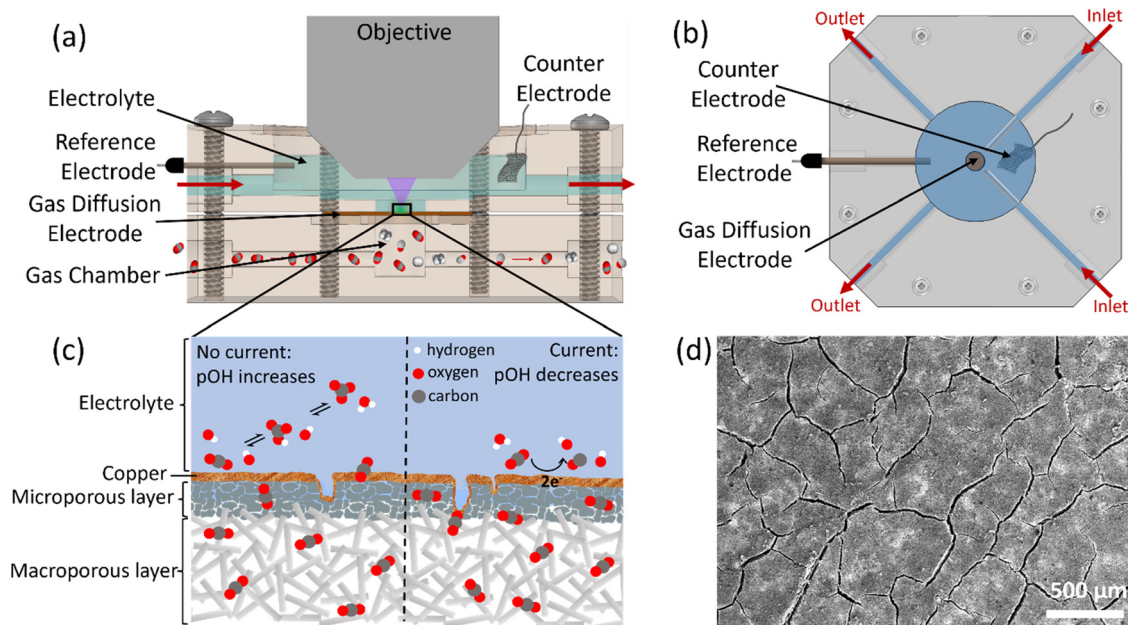


Fig. 1 (a) Cross section of the electrochemical cell with water-immersion objective. The reference electrode is a leakless Ag/AgCl electrode, the counter electrode is a Pt mesh. (b) Top view of the electrochemical cell showing two perpendicular electrolyte inlets and outlets. (c) Schematic cross section of the carbon paper GDE (not to scale), the left hand side depicts the pOH increase if no current is passing, the right hand side shows the pOH decrease with applied current. (d) Top-down SEM image of the GDE surface (Sigracet 22BB carbon paper with 300 nm Cu).

rapidly reacts with OH^- to form bicarbonate and carbonate anions, according to the chemical equilibrium reactions listed in equations (S8)–(S12) (ESI[†]).^{27–29} This decreases the local OH^- concentration and thereby increases the pOH in the electrolyte. This process is schematically depicted in Fig. 1(c) (left side). Due to the participation of CO_2 in these buffer reactions, the ability to map the spatially and time-resolved pOH around a GDE enables us to assess the local concentration of CO_2 in the electrolyte around a GDE.

At a non-zero current, a portion of the CO_2 molecules is reduced to form products such as carbon monoxide, formic acid, methane or ethylene, among others,^{11,18,19,30–32} as per equations (S19)–(S26) (ESI[†]). During this non-equilibrium process, one OH^- is created, or one buffer species is deprotonated for each electron involved.³³ The same holds true for the competing hydrogen evolution reaction (HER) that must be suppressed to maximize the CO_2R yield. As a result, at sufficiently high current density magnitudes, we expect that the pOH will decrease, as is depicted schematically in Fig. 1(c) (right side). This effect is in competition with an increase in pOH caused by unreduced CO_2 molecules that undergo reactions to form bicarbonate and carbonate anions. As a result, a low observed pOH indicates the presence of CO_2 reduction activity.

In addition, it has been shown that the pH and pOH themselves have a substantial impact on the reactivity and selectivity of CO_2R . It has been demonstrated that high pH (corresponding to low pOH) suppresses the parasitic HER^{25,34} and shifts the CO_2R selectivity towards C_{2+} products.^{8,25,32,35} The reason for the latter is that OH^- actively suppresses the creation of single carbon products (C_1) and hydrogen molecules (H_2), while it does not affect the C_{2+} current density. This leads to an increased faradaic efficiency (FE) towards C_{2+} products at a given current density.^{35,36}

As a result, knowledge of the pOH in the vicinity of an operating GDE is critical and can aid researchers in better optimizing the performance of a GDE. Notably, the pOH is expected to vary as a function of distance from the electrode surface. In addition, it also varies in the plane parallel to the surface if there are any inhomogeneities such as microcavities present on the surface. For these reasons, a measuring technique that allows one to map to pOH around an operating GDE with high resolution in three dimensions is desired.

Previous efforts towards this end have included theoretical studies of the pH around GDEs performing CO_2R .^{10,37,38} Experimental approaches have involved scanning electrochemical microscopy (SECM),^{39,40} surface-enhanced Raman (SERS),⁴¹ electrochemical atomic force microscopy (EC-AFM)⁴² and surface-enhanced infrared spectroscopy (SEIRAS).³⁴ While all these techniques are powerful and can reach spatial resolutions on the nanometer scale, they do not have the ability to map the operando pH/pOH of an entire macroscopic sample in three spatial dimensions. Fluorescent confocal laser scanning microscopy (CLSM), a well-known technique in the context of electrochemistry,^{43–48} is our method of choice to address these concerns in the context of CO_2R because it enables time-resolved measurements in three spatial dimensions. The spatial resolution

of this technique can reach 250 nm under ideal conditions. The time resolution is, depending on the size and spatial resolution of the frame of interest, microseconds to several seconds. The combination of time resolution with sub-micrometer spatial resolution in three spatial dimensions with high accuracy is what makes this pOH-mapping technique unique. Importantly, it enables the mapping of the local pOH value under operating conditions over a wide current density scale, from 0 to at least 200 mA cm^{-2} in magnitude. To map the operando pOH in three dimensions with high resolution is important because it allows us to probe the pOH within microstructures on the surface of a sample of interest such as an active GDE. No other technique allows one to probe the operando pOH within cavities in the surface of a GDE. This capability presents the unique opportunity to correlate the microstructure geometry of a GDE with its CO_2R performance. Understanding the local chemical microenvironment near a CO_2R catalyst is key for understanding selectivity and performance of these devices, as well as to validate continuum level theories of CO_2R in porous electrodes.⁵⁸ Furthermore, our novel technique might also prove to be useful for the investigation of other electrochemical applications beyond CO_2R .

In previous experiments,⁴⁹ we used the commercially available ratiometric fluorescent photoacid dye 6,8-dihydroxypyrene-1,3-disulfonic acid disodium salt (DHPDS), which is sensitive to pH values between 6 and 10. The term ‘ratiometric fluorescent dye’ refers to a probe with a spectrum that exhibits at least two distinct peaks with intensities changing in opposite directions upon varying the parameter to which the probe is sensitive. By evaluating the ratio between the peak intensities, the observed signal becomes independent of the local dye concentration.^{50–52} The sensing mechanism of the photoacid DHPDS involves proton-transfer reactions in its electronic ground state to perturb its absorption spectrum. This is the mechanism used by nearly all fluorescent pH indicators in biological studies at near-neutral-pH conditions.⁵³ With DHPDS, we observed that the local pH at the GDE surface increased from pH 6.8 to greater than 10 as the magnitude of the current density was increased from 0 to 28 mA cm^{-2} in 100 mM KHCO_3 electrolyte. We showed that the pH increased locally inside trenches 5–20 μm wide in the GDE surface. Furthermore, the pH increased as the trench width diminished, which indicated that narrow trenches exhibit a higher CO_2R activity than wider trenches and planar surfaces.⁴⁹ While these results were insightful, they were limited by the pH range in which DHPDS is sensitive. We observed that the DHPDS signal saturated at current density magnitudes larger than 20 mA cm^{-2} in 100 mM KHCO_3 . However, there is an interest in operating GDEs at high current density magnitudes to increase the CO_2R rate up to industrially relevant-levels ($|j| > 100 \text{ mA cm}^{-2}$).^{54,55} Therefore, for this study, we developed a new sensing protocol based on our recent discovery of an excited-state proton-transfer mechanism for weak photoacids, such as 8-aminopyrene-1,3,6-trisulfonic acid trisodium salt (APTS). Unlike DHPDS, at near-neutral-pH and alkaline-pH conditions, APTS does not undergo proton-transfer reactions in its electronic ground state. Instead, the fluorescent signal of APTS is altered *via* quenching of its thermally equilibrated electronic excited

state by direct proton transfer to aqueous OH^- .⁵⁶ We also show herein that aqueous dissolved inorganic carbon species do not interfere with this process (Fig. S1, ESI†). Owing to our newly characterized mechanism, the aromatic amine form of APTS can be used as a probe for the pOH value and is sensitive to pOH values between 0 and 2.8, compared to pOH 4–8 for DHPDS. It must be emphasized that this is the first demonstration of APTS as a sensor for the local pOH. APTS exhibits a novel sensing mechanism and greatly enhances the capabilities of our technique to map the local pOH value with CLSM. By combining DHPDS with APTS, we are now able to cover a pOH range from 0 to 8 (with a gap between 2.8 and 4) and investigate operating GDEs under current densities as large in magnitude as 200 mA cm^{-2} . Additionally, we explore the influence of different bicarbonate concentrations in the electrolyte and of different microstructure geometries. All told, the work crucially enables a correlation between the bulk electrolyte properties and electrode structure with the pOH present at the electrochemical active site across a broad range of operating current densities. We also evaluate the CO_2R performance of our GDEs with gas chromatography measurements and supplement our experimental analysis with continuum modeling to enable a better understanding of the correlation between local pOH and CO_2R performance.

Experimental

CLSM with the two fluorescent dyes DHPDS and APTS was used to map the pOH around an operating GDE performing CO_2R over a wide pOH range. An electrochemical cell compatible with CLSM was developed, depicted schematically in Fig. 1(a) as a cross section and Fig. 1(b) as a top view. The electrochemical cell comprises of an electrolyte chamber with two perpendicular electrolyte inlets and outlets and is open at the top which allows the water immersion objective of the confocal microscope to be immersed into the electrolyte. For experiments at constant current densities, a leakless Ag/AgCl reference electrode and a Pt mesh counter electrode are submerged into the electrolyte. This cell enables us to study a GDE that is fed with gaseous CO_2 through a gas chamber. The GDE is made of a Sigracet 22 BB carbon paper substrate covered with 300 nm Cu as well as a carbon black, graphite and Nafion coating. The substrate's surface is covered by an irregular pattern of trenches 5–30 μm wide which is shown schematically as a cross section in Fig. 1(c) as well as in a top-down scanning electron microscopy (SEM) image in Fig. 1(d). The trenches cover approximately 6% of the sample surface. Energy dispersive X-ray spectroscopy (EDS) measurements reveal that copper covers not only the planar carbon paper surface but also the trench walls and bottoms (Fig. S2, ESI†). The maximum spatial resolution of the confocal microscope is 250 nm in the x-y plane and 500 nm in the z-direction, however, electrolyte flow introduces noise, so the resolution is estimated to be on the order of one micron. We investigate features that are no smaller than 5 μm , so this resolution is satisfactory.

We first used the dye APTS in the absence of electrical current to investigate the diffusion of CO_2 through a GDE.

For this, 1 M KOH electrolyte (pOH 0) with 100 μM APTS was used. The fluorescence signal was mapped in the plane perpendicular to the electrode surface as a function of time with one frame captured approximately every four seconds. After one minute of continuous measurements, a 10 standard cubic centimeters per minute (SCCM) CO_2 gas stream was fed into the gas chamber of the electrochemical cell and the change in pOH in the electrolyte was observed. The measurements were performed both with and without circulating the electrolyte at a rate of 6 mL min^{-1} .

In addition, we used both dyes, DHPDS and APTS, to map the pOH around an operating GDE performing CO_2R at current density magnitudes between 0 mA cm^{-2} and 200 mA cm^{-2} . Measurements were performed in the plane parallel to the electrode surface, at 20 μm above the surface, at the surface, and at 20 μm below the surface inside a trench. A CO_2 -saturated aqueous KHCO_3 solution with KHCO_3 at concentrations of 100 mM, 200 mM and 400 mM spiked with 100 μM DHPDS or 200 μM APTS was used as the electrolyte and pumped through two perpendicular inlets at a rate of 6 mL min^{-1} . The CO_2 gas flow through the gas chamber was set to 10 SCCM. The electrolyte with dilute dye was removed from the cell and replaced after each measurement. Every measurement was performed at least three times.

We further evaluated the CO_2R performance of equivalent copper on carbon paper GDEs with gas chromatography. We used an electrochemical cell that was tailored for use with gas chromatography. More details about the sample fabrication and characterization, dye calibration, and measuring procedure can be found in the ESI,† Section S1 and Fig. S2–S4.

The development of the multiphysics model, including the physics, parameters, and assumptions employed, is described in the ESI,† Section S2 and Fig. S11.

Results and discussion

First, we investigated how the local pOH in the electrolyte surrounding a GDE changed upon exposure to CO_2 at open circuit. This allowed us to draw conclusions about the diffusion of CO_2 molecules through a GDE. Our developed measuring technique is ideal for this study as it combines time resolution with the capability to spatially resolve the local pOH inside inhomogeneities in the surface of a GDE. CO_2 reacts with OH^- and water molecules in the electrolyte to form bicarbonate and carbonate anions. This increases the pOH which becomes an indicator for the CO_2 diffusion pattern. Fig. 2(a) shows color-coded maps of the local pOH in the plane perpendicular to the electrode surface as a cross-section through a trench as well as the above electrolyte for different times without electrolyte flow. We define $t = 0 \text{ s}$ as the time when the CO_2 flow through the gas chamber of the electrochemical cell was turned on. A distinct change in pOH is immediately visible in the first frame after starting the CO_2 flow but it is restricted to the trench. The pOH changes at the electrode surface and in the bulk electrolyte as time proceeds. In Fig. 2(c), the average pOH at three different

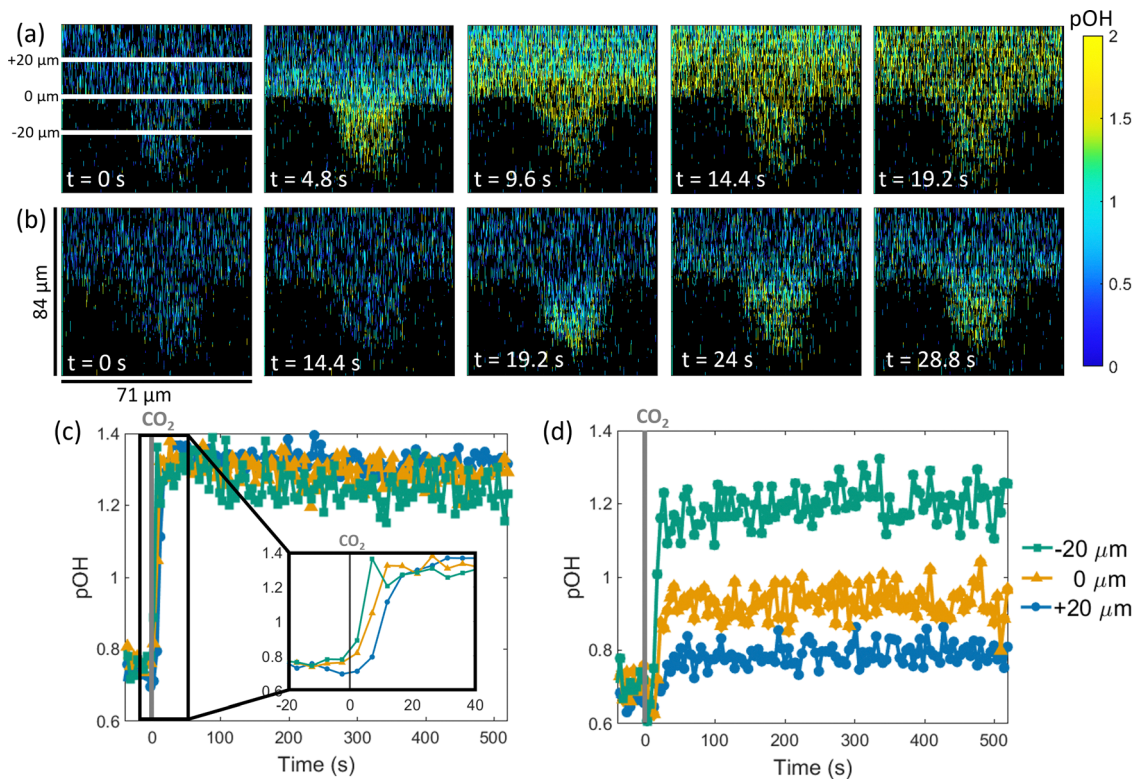


Fig. 2 Changing pOH in the electrolyte due to CO_2 diffusion around a trench in a carbon paper GDE at open circuit, without electrolyte flow (a), (c) and with electrolyte flow (b), (d). Measurements performed with 1 M KOH electrolyte with 100 μM APTS, 10 SCCM CO_2 turned on at the backside of the electrode at $t = 0$ s. Frames in (a) and (b) show pOH maps as a cross section through a trench in the plane perpendicular to the electrode surface for different times. The scale bars apply to both (a) and (b). Panels (c) and (d) track the average pOH at locations 20 μm above the surface (+20 μm), at the surface (0 μm) and 20 μm below the surface in the trench (−20 μm), as specified in the first frame of (a), as a function of time. The vertical gray lines indicate the start of the CO_2 flow.

positions is tracked as a function of time. To this end, we draw a line 20 μm above the electrode surface (+20 μm), at the electrode surface (0 μm), and 20 μm below the electrode surface inside the trench (−20 μm) as indicated in the first panel of Fig. 2(a). The pOH is averaged along these lines as a function of time. The vertical gray line in Fig. 2(c) indicates the time when the CO_2 gas supply was turned on. This data confirms that the pOH first changes inside the trench, then at the GDE surface and approximately five seconds later at 20 μm above the surface, which is consistent with the gaseous CO_2 feed coming from the backside of the GDE. After 20 s, the pOH values converge at all positions monitored and remain constant for the duration of the experiment. An analogous measurement was performed in a region of the electrode void of trenches (Fig. S5, ESI†) and we observed that the pOH increases at a slower rate in comparison to measurements at a location with a trench. This demonstrates that microstructures such as trenches promote faster CO_2 transport through a GDE substrate.

The experiment displayed in Fig. 2(a) was repeated under an electrolyte flowrate of 6 mL min^{-1} (Fig. 2(b)). This has a dramatic effect on the pOH around the GDE: the first change in pOH is only visible after 19 s, compared to less than 5 s without electrolyte flow. A significant increase in pOH caused by the diffusion of CO_2 is solely observed inside the trench. Fig. 2(d) tracks the pOH over time in a similar manner as

Fig. 2(c). This confirms that the pOH increases little at the GDE surface as well as 20 μm above the surface, and even within the trench, the pOH remains below the value obtained without electrolyte flow. A steady state, with constant pOH up to 20 μm above the electrode surface, is reached after 40 s. However, the pOH at the three locations never converges to a single value, in contrast to the measurements performed with stationary electrolyte. Electrolyte flow results in a well-defined boundary layer around the GDE surface which leads to a CO_2 concentration gradient. The boundary layer is taken into consideration in the multiphysics model. In the absence of electrolyte recirculation, diffusion is the dominant transport mechanism for CO_2 and subsequent CO_2R products. When the electrolyte is circulated however, convection dominates the mass transfer and causes OH^- to be more quickly removed from the GDE surface. This prevents the pOH from rising as much as it would without electrolyte flow. As the electrolyte is more stationary inside trenches, a larger pOH increase can be observed in these confined spaces.

An additional study was conducted to compare the CO_2 diffusion through homogeneous GDE substrates made of laminated polytetrafluoroethylene (PTFE) with pore sizes that are specified by the manufacturer as 0.1 to 0.2 μm and 0.45 μm , respectively. Fig. S6 (ESI†) shows a schematic representation of the structure of both a carbon paper GDE and a PTFE GDE

together with SEM images. Fig. S7 (ESI†) compares the flow patterns through PTFE GDEs with different pore sizes. We observe that, as expected, the pOH in the electrolyte above the surface increases faster for PTFE with a larger pore size. After 45 seconds it converges to the same value for both substrates. This result gives evidence that, not only microcavities in a GDE's surface, but also larger pore size can promote faster CO₂ transport. However, there is a critical pore size that should not be exceeded, otherwise the pores will be flooded with liquid electrolyte during an experiment. We observed this for a laminated PTFE substrate with pore sizes of 5 μm or greater. Flooding prevents effective CO₂ transport and inhibits CO₂ diffusion to catalytic sites.^{10,12} These results suggest that consideration of the influence of microstructures, even in commercially available GDE substrates, is a critical parameter to understand device performance.

The focus of our study is to investigate the local pOH around an operating GDE performing CO₂R over a wide range of current densities. Our newly developed technique allows us to perform pOH measurements with superior three-dimensional resolution inside cavities in an electrode under operating conditions. To ensure that our GDE samples (300 nm Cu on carbon paper) perform CO₂R, we evaluated their CO₂R performance with gas chromatography. We performed chronopotentiometry experiments for different current densities. Our results are consistent with copper GDEs previously reported in literature.^{15,18,57} H₂ production dominates for low current density magnitudes. As the magnitude of the current density increases, we observe more CO₂R products, and the selectivity shifts towards C₂₊ products. We observe ethylene for current densities of 50 mA cm⁻² or higher in magnitude and ethanol for 200 mA cm⁻² (Fig. S8, ESI†). We can therefore assume that the analogous electrodes we use for our pOH imaging experiments behave as one would expect from literature.

By combining the two pOH-dependent fluorescent ratiometric dyes, DHPDS and APTS, we are able to cover the pOH range from 0 to 8. In the context of CO₂R, a low local pOH under operation indicates high CO₂R activity and is desirable

because it has been shown that high pH can help suppress the parasitic HER and favor the formation of C₂₊ products.^{8,25,32,35,36} Fig. 3 displays pOH maps in the plane parallel to the electrode surface in 100 mM KHCO₃ electrolyte for different current density magnitudes between 0 mA cm⁻² and 100 mA cm⁻² and at three different z-positions: 20 μm above the electrode surface, at the electrode surface, and 20 μm below the electrode surface inside a trench. The panels on the left-hand side were captured using DHPDS, and the ones on the right-hand side were captured using APTS.

We observed that APTS degrades and slowly loses its fluorescence for high current density magnitudes. We hypothesize that this is connected to the reduction of APTS at the electrode surface. A more detailed analysis of this effect can be found in the ESI,† Section S1. Despite the slow degradation of APTS, the pOH measurements obtained with APTS are reliable because APTS is a ratiometric dye. The pOH value is calculated from the ratio between two signals that are captured independently in two different wavelength intervals. As some dye molecules degrade, both signals become weaker. This leads to a decreased signal-to-noise ratio, however, the ratio between both signals remains unchanged (see Fig. S9, ESI†). The dye degradation effect could be mitigated by using an APTS concentration of 200 μM (compared to DHPDS, where 100 μM is sufficient), removal and replacement of the electrolyte after each measurement, and the introduction of two perpendicular electrolyte inlets to ensure the transportation of fresh APTS to the GDE surface. Some of the panels under high current density magnitudes still appear relatively dark and noisy. This is especially evident inside the trench as the electrolyte is more stationary in confined spaces and this encumbers the transport of fresh, undegraded dye. However, since APTS is a ratiometric dye, the pOH it predicts does not depend on concentration. At magnitudes of the current density as high as 200 mA cm⁻², we were still able to collect enough fluorescence signal to determine the pOH.

For a current density of $J = 0$ mA cm⁻², the pOH equals 7.2 everywhere (pOH of CO₂-saturated 100 mM KHCO₃). When

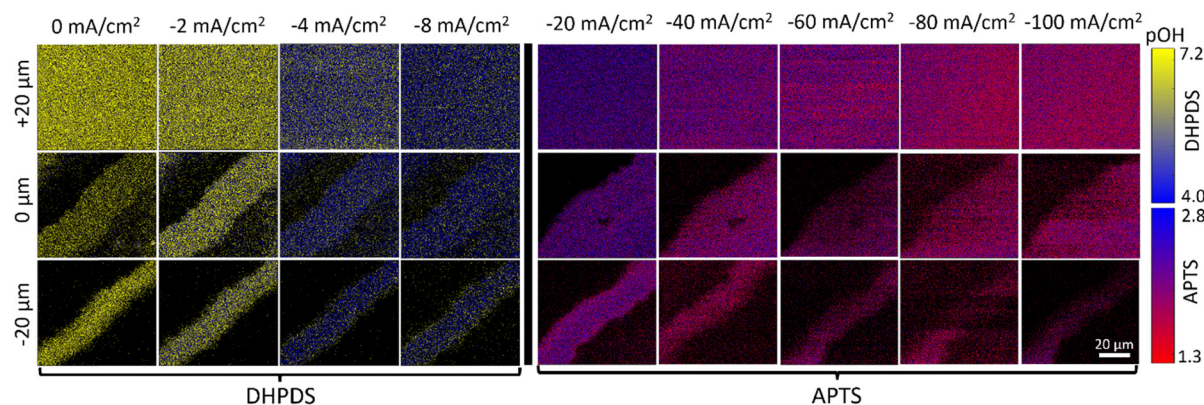


Fig. 3 pOH maps in the plane parallel to the electrode surface 20 μm above the surface, at the surface and 20 μm below the surface inside a trench in 100 mM KHCO₃ electrolyte for different current densities between 0 mA cm⁻² and -100 mA cm⁻², obtained with 100 μM DHPDS (left) and 200 μM APTS (right). Note the different color legends for the panels obtained with DHPDS and APTS.

the current is non-zero, the local pOH decreases because OH^- is created as a byproduct of CO_2R . The pOH is lower at the electrode surface than $20\text{ }\mu\text{m}$ above the surface because CO_2 is reduced at the electrode surface and for the pOH to decrease at $+20\text{ }\mu\text{m}$, OH^- has to diffuse away from the surface. Due to electrolyte flow, a concentration gradient is created. Furthermore, the pOH inside the trench is lower than at the surface. This can be seen especially well in the panels for current densities of -2 mA cm^{-2} and -20 mA cm^{-2} .

The surface morphology of copper does not change during CO_2R experiments. We performed SEM as well as EDS measurements for samples before and after CO_2R . We observed potassium deposits on the sample after CO_2R that originate from KHCO_3 molecules in the electrolyte, but the copper catalyst doesn't change in appearance, neither in the trenches nor on the planar electrode surface (see Fig. S2, ESI†).

Fig. 4 shows the average pOH as a function of current density, both from experiments and from multiphysics simulations. For the experimental values, the pOH is averaged over at least three independent measurements. The data for low current density magnitudes was obtained with DHPDS, while the data for higher current density magnitudes was obtained with APTS. The red shaded areas indicate the pOH range that is inaccessible by either dye. The insets are a zoom-in on the experimental data for low current density magnitudes.

Simulations were performed to better understand the variations of pOH with respect to position, electrolyte concentration and trench geometry. A 2-dimensional continuum model of a periodic trench in a GDE was developed. The multiphysics simulations solve for the concentration profiles of all species involved, along with the liquid electrolyte potential in a single, periodic trench of a given width and depth. Concentration-dependent Tafel kinetics are employed to simulate the electrochemical reduction of CO_2 to various products on the Cu electrode surface.^{35,36} Greater detail regarding the physics, parameters, and assumptions employed in the continuum model (shown schematically in Fig. S11, ESI†) can be found

in the ESI,† Section S2. The model replicates the experimental results with outstanding accuracy. This is true both for the pOH values observed in all panels in Fig. 4 as well as for the product distribution that was measured with gas chromatography. The agreement between model and experiment gives us confidence that the model is a good representation of the real world and predictions beyond what can be observed in experiments can be trusted.

Fig. 4(a) compares the pOH in 200 mM KHCO_3 electrolyte at positions around a GDE analogous to that shown in Fig. 3. As the magnitude of the current density increases, the pOH decreases accordingly, due to the generation of OH^- from the CO_2R reactions. At first, the slope of the pOH decrease is steep, then gradually tapers as the current density magnitude increases. This can be explained by the logarithmic nature of the pOH scale. A decrease in pOH from 4 to 3 requires the OH^- concentration to increase from 0.1 mM to 1 mM , while a decrease in the pOH from 2 to 1 requires the OH^- concentration to rise from 10 mM to 100 mM , a much larger absolute change in concentration. The pOH in trenches is lower than at the surface, which again is lower than $20\text{ }\mu\text{m}$ above the surface for all current densities; this confirms our previously described results. The simulations agree well with the CLSM experiments, particularly for data at the electrode surface ($z = 0\text{ }\mu\text{m}$) and in the trench ($z = -20\text{ }\mu\text{m}$). Agreement between experiments and simulations is similarly strong for bulk electrolyte concentrations of 100 mM and 400 mM (Fig. S10a–c, ESI†). The greatest discrepancy between experiments and simulations is apparent in the $z = +20\text{ }\mu\text{m}$ data, which more closely represents the pOH of the bulk electrolyte. The discrepancy is likely due to the static nature of the boundary layer in the simulations, which is set to a constant value of $150\text{ }\mu\text{m}$, regardless of the applied current density. Previous experimental studies have shown that the bulk electrolyte pH, and thus, the boundary layer thickness is a function of current density at higher current density magnitudes.^{25,34} However, implementing an adaptive, 2-D mesh that grows with current would be very computationally

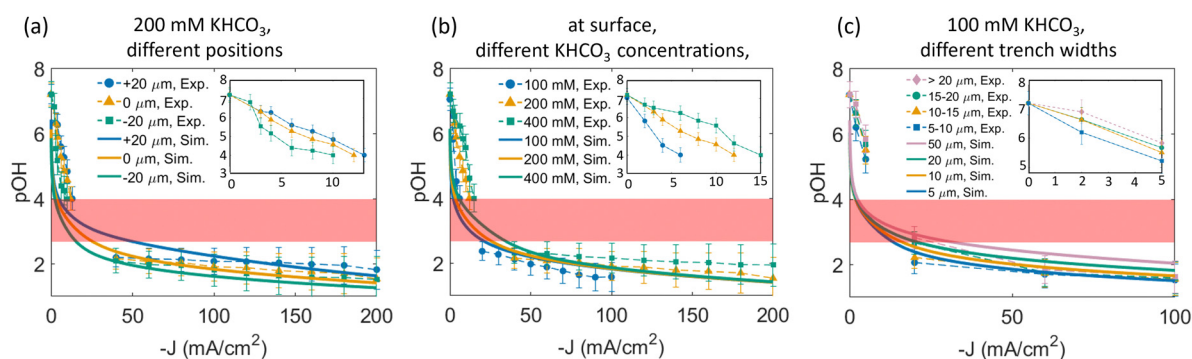


Fig. 4 The pOH as a function of current density at different positions (a), for different KHCO_3 concentrations (b) and for different trench widths (c), in experiments and simulations. (a) Averaged pOH as a function of current density for 200 mM KHCO_3 , obtained with $100\text{ }\mu\text{m DHPDS}$ ($|J| \leq 13\text{ mA cm}^{-2}$)/ $300\text{ }\mu\text{m APTS}$ ($|J| \geq 20\text{ mA cm}^{-2}$) at three different positions around the electrode surface together with simulation results. (b) Averaged pOH as a function of current density at the electrode surface for three different KHCO_3 concentrations between 100 and 400 mM , obtained with $100\text{ }\mu\text{m DHPDS}$ ($|J| \leq 15\text{ mA cm}^{-2}$)/ 200 – $300\text{ }\mu\text{m APTS}$ ($|J| \geq 20\text{ mA cm}^{-2}$) together with simulation results. (c) pOH inside a trench $20\text{ }\mu\text{m}$ below the surface as a function of current density for different trench widths between $5\text{ }\mu\text{m}$ and $30\text{ }\mu\text{m}$. In all panels, the pOH range shaded in red is not accessible by either dye, the inset is a zoom on the data for low current density magnitudes.

intensive, and for the present study, the pOH in the trench, where CO_2R occurs, is more critical to simulate accurately.

Fig. 4(b) displays a study of the influence of KHCO_3 concentration in the electrolyte on the pOH at the GDE surface. The trend observed at the electrode surface is the same as it is $20\ \mu\text{m}$ above the electrode surface and $20\ \mu\text{m}$ below in the trenches (Fig. S10d–f, ESI†). Measurements for higher concentrations could be carried out at higher current density magnitudes due to the enhanced electrical conductivity of the electrolyte. Again, the model generally agrees with observed experimental trends that show that the pOH decreases more slowly as electrolyte concentration is increased. This is because KHCO_3 acts as a buffer and, as the KHCO_3 concentration increases, the buffering capacity is enhanced, which restrains the decrease in pOH. However, at higher current density magnitudes, the simulated pOH becomes more independent of current density whereas the experimental pOH remains impacted by buffer capacity. This discrepancy at higher current density magnitudes can be attributed to a multitude of factors, such as the large uncertainty in the rate constants for the homogeneous-phase buffer reactions used in the simulations,^{28,58,59} and the fact that dilute-solution theory was employed in the simulations whereas concentrated solution theory may be necessary to simulate adequately the behavior at higher current density magnitudes.⁶⁰ Additionally, the simulations assume that there are no mass transport losses of gas phase CO_2 throughout the GDE itself, and thus the CO_2 concentration is at a constant value of $34\ \text{mM}$ at the electrode surface. This assumption is unrealistic at high current density magnitudes and could potentially impact the simulated pOH due to the participation of CO_2 in homogeneous buffer reactions. Nonetheless, the simulations are capable of generally capturing the impact of buffer composition, and the incorporation of these second-order phenomena is beyond the scope of the present study.

Lastly, the influence of the trench geometry on the local pOH was investigated (Fig. 4(c)). Eighteen trenches with different widths in $100\ \text{mM}\ \text{KHCO}_3$ were taken into consideration. The average pOH inside the trenches, *i.e.* $20\ \mu\text{m}$ below the surface, was evaluated at different current densities. The trenches were divided into widths of $5\text{--}10\ \mu\text{m}$, $10\text{--}15\ \mu\text{m}$, $15\text{--}20\ \mu\text{m}$, and $>20\ \mu\text{m}$ and the average pOH was calculated for each. Both measurements and simulations agree that the pOH decreases as trench width diminishes for all current densities investigated. This behavior is due to the increased surface-area-to-volume-ratio of the thinner trenches (Fig. S13, ESI†). The increased surface-area-to-volume ratio means that more OH^- is generated *via* CO_2R on the trench walls relative the electrolyte volume contained within a trench. In addition, the mass transfer out of trenches is poor compared to a planar surface, which leads to trapping of OH^- . Thus, the OH^- concentration is increased for thinner trenches as a result of both the increased surface-area-to-volume-ratio as well as mass transport effects. This result is in line with previously reported electrochemical studies that have demonstrated that smaller pore radii in meso- or micro-structured electrodes enhance local OH^- trapping.^{54,61–63} Interestingly, the measured pOH is less sensitive to trench width at

higher current density magnitudes than the simulated pOH. A possible explanation for this discrepancy could be the lack of bubble-induced convection in the present simulations. At high current density magnitudes, the latter could ameliorate OH^- trapping within GDE trenches and reduce sensitivity to trench geometry. The influence of trench width on the pOH further emphasizes the importance to understand how a GDE's micrometer-scale geometry affects CO_2R .

When we combine the results from Fig. 4(a) and (c), we find that the pOH is decreased inside trenches compared to the planar electrode surface and furthermore, this effect is stronger for more narrow trenches. This observation is caused by the altered mass transport inside narrow trenches. Our experiments therefore suggest that the micrometer-scale morphology of a GDE plays an important role in determining the mass transport properties of a GDE, and with it the CO_2R performance.

Beyond elucidating trends observed in the CLSM experiments, continuum simulations can be used to identify pOH variations at resolutions that are not achievable experimentally. Fig. 5(a) depicts OH^- transport within a simulated periodic GDE trench operating at a current density of $-50\ \text{mA cm}^{-2}$ (OH^- transport for -1 , -10 , and $-100\ \text{mA cm}^{-2}$ is depicted in Fig. S15, ESI†). As shown in the contour plot, the pOH is lowest, and thus the concentration of OH^- is greatest deep within the trench. Examination of the OH^- flux rationalizes the diminished pOH within the trench: OH^- generated *via* CO_2R in the trench exits to the bulk electrolyte, but the presence of the trench increases the surface area available for the electrochemical reduction of CO_2 and, accordingly, the generation of OH^- . The trench also provides a volume where the generated OH^- can accumulate due to the altered mass transport. OH^- generated by CO_2R in the trench is thus effectively trapped within the trench, decreasing the local pOH. Similar OH^- trapping behavior has been observed in prior multi-dimensional continuum studies of mesostructured electrodes^{38,58} and has been shown to be beneficial for CO_2R selectivity.³⁸

To evaluate the influence of the pOH gradient within the trench on local selectivity variations, concentration-dependent Tafel kinetics were employed in the multiphysics model. These concentration dependent Tafel kinetics were fit to experimental data collected on Cu GDEs used in the present study, and are capable of replicating observed partial current densities quite well (Fig. S12, ESI†). The study relates local changes in OH^- activity to changes in current densities for product formation. Fig. 5(b) depicts the change in pH as a function of depth into the trench. Note that Fig. 5(b) shows the pH value while for the most part, this study focuses on the dynamics of OH^- and hence, the pOH. Here, we chose to display the pH instead to emphasize the correlation between Fig. 5(b) and (c). In the model, the relation $\text{pH} + \text{pOH} = 14$ holds true everywhere because we only consider the diffuse transport layer while the electrical double layer is not taken into consideration in the model because its length scale is much smaller than the length scales we study. Therefore, within the frame of this model, pH and pOH are directly connected. Consistent with Fig. 5(a), pH increases (pOH decreases) deeper inside the trench (more negative z) due to trapping of CO_2R -generated OH^- .

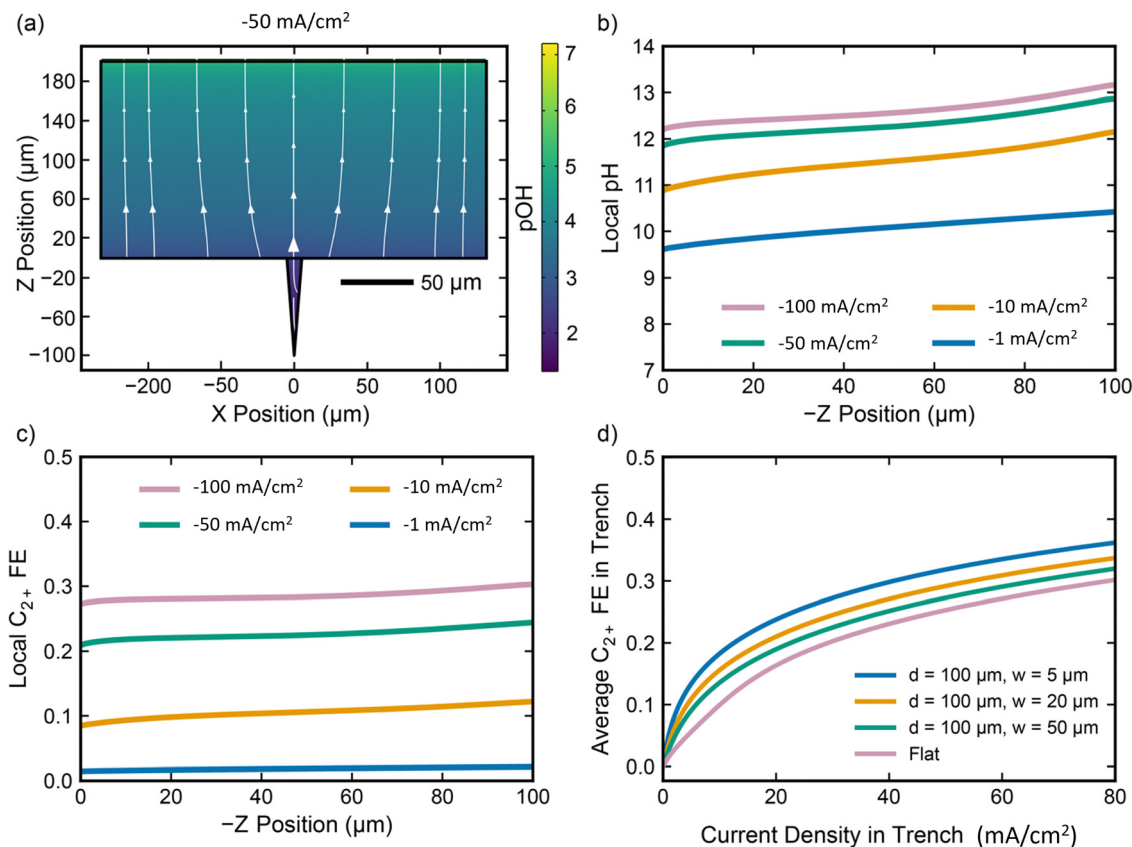


Fig. 5 (a) pOH contour plot and OH^- flux (white arrows) throughout the simulated domain for a periodic GDE trench (width of 20 μm and depth of $-100 \mu\text{m}$) operating at -50 mA cm^{-2} of geometric current density (see Section S4, ESI† for a description of how geometric current density is calculated). (b) Simulated electrode surface pH ($14 - \text{pOH}$) as a function of depth into the trench for different current densities. (c) Simulated C_{2+} FE as a function of depth into the trench for different current densities. (d) Average C_{2+} FE on trench walls as a function of the current density passed solely through the trench. $Z = 0 \mu\text{m}$ represents the top of the trench, and $Z = -100 \mu\text{m}$ represents the bottom of the trench. d and w are the trench depth and width, respectively.

The simulations exhibit a direct, albeit modest, relationship between pH and C_{2+} FE (Fig. 5(c)). As the pH increases or the pOH decreases within the trench, the C_{2+} FE correspondingly increases (variations of C_1 , H_2 , and local current density within the trench are shown in Fig. S15, ESI†). This relationship between C_{2+} FE and pH has been widely reported, and is connected to the suppression of HER and C_1 products.^{9,19,35,36,63} The disproportionate suppression of H_2 , CH_4 , and HCOOH partial current density within the trench can be observed in Fig. S16 (ESI†). This suppression of C_1 and H_2 in the trench results in an approximately 5% increase in C_{2+} FE across the depth of the trench, regardless of current density (Fig. 5(c)). The enhancement of the performance in the trench is modest, as the change in pH is only about 1 pH unit across the trench depth. Nonetheless, the simulations serve to understand local selectivity trends and provide better understanding of variations in performance within the GDE heterogeneities.

Lastly, performance of CO_2R within trenches was compared to the performance on a flat surface without trenches (Fig. 5(d)). The simulations demonstrate that the C_{2+} FE at a given current density is higher within a trench compared to that on a flat surface due to the improved OH^- trapping. In addition, as the width of the trench decreases, and with it the pOH (Fig. 4(c) and

Fig. S18, ESI†), the C_{2+} FE increases. This is a result of the changed mass transport properties in the confined space as well as the enhanced trench surface-area-to-volume ratio (Fig. S13, ESI†), leading to a lower pOH value. The simulations indicate that the micrometer-scale morphology on a GDE surface has an impact on the product distribution with narrow trenches on the order of 5 μm favoring the creation of C_{2+} products across the entire current density range (Fig. 5(d)) by approximately 5% compared to the flat Cu case. Within the scope of this work, the connection between surface morphology and CO_2R selectivity can only be demonstrated *via* multiphysics simulations because we have no means to locally resolve the product distribution. However, the accuracy of our model has been demonstrated as there is an excellent agreement with the experimentally measured CO_2R performance as well as the local pOH. The simulations demonstrate the potential for improved C_{2+} FE within these microcavities and experimental demonstration of enhanced C_{2+} FE inside microcavities in the surface of an operating GDE should be the subject of a future study.

The effect of the trench depth on CO_2R was also evaluated (Fig. S19, ESI†), which demonstrates that the ability of a trench to enhance C_{2+} FE increases as the trench deepens. This is due to the increased surface area available for electrochemical OH^-

generation for deeper trenches. The surface-area-to-volume ratio is far less sensitive to trench depth than to trench width (Fig. S18, ESI†) which indicates that compared to the surface area, the volume plays a minor role in determining the C_{2+} FE. However, enhancements in C_{2+} FE become marginal beyond trench depths of 50 μm . Finally, the fraction of the surface covered by the trenches was varied in the simulations, to further evaluate the efficacy of trenches. Fig. S20 (ESI†) depicts the pOH and C_{2+} FE as a function of trench coverage, evaluated at current densities of -20 mA cm^{-2} and -50 mA cm^{-2} (normalized now to the electrochemical active surface area (ECSA)). This figure aims to deconvolute the effects of roughness from the effects of changes in the microenvironment. As the trench surface coverage increases, the pOH decreases, and the C_{2+} FE increases, due to the OH^- trapping ability of the trenches. While the increase in C_{2+} FE is minor (still approximately 5% from a flat surface to a completely trenched surface), the enhancement with morphology is similarly observed at both current densities and the results reveal that heterogeneities in commercially available GDEs can serve as local hotspots for enhanced C_{2+} FE and benefit CO_2R due to their OH^- trapping capacity. The simulation results serve as a great support to our novel measurements that allow to probe the microenvironment in a GDE under operando conditions. Both experiment and simulations highlight the importance of the micromorphology of a GDE that influences the mass transport properties. Future work should aim to leverage this insight to develop geometries for which transport is tailored to better trap OH^- . Techniques like laser ablation can be utilized to realize these tailored electrodes.

Conclusions

We demonstrate a novel technique to probe the microenvironment in a CO_2R GDE under operating conditions. We map the local pOH value around a GDE resolved in time and three spatial dimensions with fluorescent confocal laser scanning microscopy (CLSM) with a combination of two ratiometric dyes, DHPDS and APTS, that operate by different sensing mechanisms. The novel application of APTS as a dynamic excited-state pOH probe was demonstrated for the first time. Critically, this work could hold potential applications beyond the investigation of CO_2R . We were able to capture pOH maps in the pOH range between 0 and 8 and at current densities as large in magnitude as 200 mA cm^{-2} .

In a first step, we investigated the diffusion patterns of CO_2 through a GDE. This study revealed that CO_2 diffuses through microcavities first, and that electrolyte flow leads to a CO_2 concentration gradient with the highest concentration confined to trenches. These microcavities in a GDE surface as well as large pore sizes promote faster CO_2 transport, highlighting the importance to understand the influence of the micron-scale morphology of a CO_2R GDE.

We mapped the local pOH around an operating GDE under applied current. This revealed that the pOH in the vicinity of a

GDE decreases as the magnitude of the current density is increased. For all current densities covered, the pOH inside trenches is lower than on the planar surface with the lowest pOH observed inside narrow trenches that are approximately 5 μm wide. This decrease in pOH can be connected to the trapping of OH^- in narrow trenches caused by altered mass transport properties.

Multiphysics simulations confirmed all experimental trends, providing pOH maps consistent with measured pOH and also reproduced product distributions with high accuracy. These simulations indicate that there is a correlation between the pOH and the faradaic efficiency towards C_{2+} products. Narrow trenches exhibit a decreased pOH and simulations suggest that this results in a locally increased faradaic efficiency towards C_{2+} products.

We conclude that we successfully developed a superior technique to probe the microenvironment around a CO_2R GDE. The micrometer-scale morphology of a GDE has a strong effect on the mass transport properties and with it, on the CO_2R performance. Microcavities on the order of 5 μm on the surface of a copper GDE serve as local hotspots with decreased pOH. We expect that this will inform the design of future CO_2R devices.

Author contributions

AB performed dye calibration and all pOH imaging experiments as well as data analysis and wrote the original draft of the manuscript. JCB and AJK developed the multiphysics model with oversight by AZW. JCB ran simulations and performed data analysis on results from the multiphysics model. AQF conducted the gas chromatography measurements. RB contributed to conceptualization of the choice of APTS photoacid. CNF investigated possible interference in signal from APTS due to dissolved inorganic carbon species. AJW developed the experimental setup and designed the electrochemical cell. AB, AQF, RB, AJW, SA and HAA contributed to the conceptualization of the study. JCB, AQF, RB, AJW, ATB, AZW, SA and HAA reviewed and edited the manuscript. ATB, AZW, SA and HAA supervised the project. All authors discussed the results and contributed to the preparation of the manuscript.

Conflicts of interest

The authors declare the following competing financial interests: Co-authors are co-inventors on the following pending patent applications that are relevant to this article: US20180065095 and WO2018049061 (SA), US20190217255 and WO2019191326 (SA and RB). SA received sponsored research funding support from Nissan Chemical Corporation for work related to that reported in this article.

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References

- 1 S. Jin, Z. Hao, K. Zhang, Z. Yan and J. Chen, Advances and Challenges for the Electrochemical Reduction of CO₂ to CO: From Fundamentals to Industrialization, *Angew. Chem., Int. Ed.*, 2021, **133**, 2–24.
- 2 S. J. Davis, N. S. Lewis, M. Shaner, S. Aggarwal, D. Arent, I. L. Azevedo, S. M. Benson, T. Bradley, J. Brouwer and Y. Chiang, *et al.*, Net-Zero Emissions Energy Systems, *Science*, 2018, **360**, 6396.
- 3 M.-Y. Lee, K. T. Park, W. Lee, H. Lim, Y. Kwon and S. Kang, Current achievements and the future direction of electrochemical CO₂ reduction: A short review, *Crit. Rev. Environ. Sci. Technol.*, 2020, **50**, 769–815.
- 4 J.-P. Jones, G. K. S. Prakash and G. A. Olah, Electrochemical CO₂ Reduction: Recent Advances and Current Trends, *Isr. J. Chem.*, 2014, **54**, 1451–1466.
- 5 S. Wang, T. Kou, S. E. Baker, E. B. Duoss and Y. Li, Electrochemical Reduction of CO₂ to Alcohols: Current Understanding, Progress, and Challenges, *Adv. Energy Sustainability Res.*, 2022, **3**, 2100131.
- 6 D. Higgins, C. Hahn, X. Chengxiang, T. F. Jaramillo and A. Z. Weber, Gas-Diffusion Electrodes for Carbon Dioxide Reduction: A New Paradigm, *ACS Energy Lett.*, 2019, **4**, 317–324.
- 7 M. Inaba, A. W. Jensen, G. W. Sievers, M. Escudero-Escribano, A. Zana and M. Arenz, Benchmarking High Surface Area Electrocatalysts in a Gas Diffusion Electrode: Measurement of Oxygen Reduction Activities under Realistic Conditions, *Energy Environ. Sci.*, 2018, **11**, 988–994.
- 8 C. Dinh, T. Burdyny, G. Kibria, A. Seifitokaldani, C. M. Gabardo, F. P. G. De Arquer, A. Kiani, J. P. Edwards, P. De Luna and O. S. Bushuyev, *et al.*, CO₂ Electroreduction to Ethylene via Hydroxide-Mediated Copper Catalysis at an Abrupt Interface, *Science*, 2018, **360**, 783–787.
- 9 B. Kim, F. Hillman, M. Ariyoshi, S. Fujikawa and P. J. A. Kenis, Effects of Composition of the Microporous Layer and the Substrate on Performance in the Electrochemical Reduction of CO₂ to CO, *J. Power Sources*, 2016, **312**, 192–198.
- 10 L.-C. Weng, A. T. Bell and A. Z. Weber, Modeling Gas-Diffusion Electrodes for CO₂ Reduction, *Phys. Chem. Chem. Phys.*, 2018, **20**, 16973–16984.
- 11 T. N. Nguyen and C. T. Dinh, Gas Diffusion Electrode Design for Electrochemical Carbon Dioxide Reduction, *Chem. Soc. Rev.*, 2020, **49**, 7488–7504.
- 12 A. Q. Fenwick, A. J. Welch, X. Li, I. Sullivan, J. S. DuChene, C. Xiang and H. A. Atwater, Probing the Catalytically Active Region in a Nanoporous Gold Gas Diffusion Electrode for Highly Selective Carbon Dioxide Reduction, *ACS Energy Lett.*, 2022, **7**(2), 871–879.
- 13 M. R. Singh, E. L. Clark and A. T. Bell, Effects of Electrolyte, Catalyst, and Membrane Composition and Operating Conditions on the Performance of Solar-Driven Electrochemical Reduction of Carbon Dioxide, *Phys. Chem. Chem. Phys.*, 2015, **17**, 18924–18936.
- 14 F. P. García de Arquer, C. T. Dinh, A. Ozden, J. Wicks, C. McCallum, A. R. Kirmani, D. H. Nam, C. Gabardo, A. Seifitokaldani and X. Wang, *et al.*, CO₂ Electrolysis to Multicarbon Products at Activities Greater than 1 A cm⁻², *Science*, 2020, **367**, 661–666.
- 15 Y. Hori, K. Kikuchi and S. Suzuki, Production of CO and CH₄ in the Electrochemical Reduction of CO₂ at Metal Electrodes in Aqueous Hydrogencarbonate Solution, *Chem. Lett.*, 1985, 1695–1698.
- 16 R. Kortlever, J. Shen, K. J. P. Schouten, F. Calle-Vallejo and M. T. M. Koper, Catalysts and Reaction Pathways for the Electrochemical Reduction of Carbon Dioxide, *J. Phys. Chem. Lett.*, 2015, **6**, 4073–4082.
- 17 A. Bagger, W. Ju, A. S. Varela, P. Strasser and J. Rossmeisl, Electrochemical CO₂ Reduction: A Classification Problem, *ChemPhysChem*, 2017, **18**, 3266–3273.
- 18 K. Kuhl, E. R. Cave, D. N. Abram and T. F. Jaramillo, New Insights into the Electrochemical Reduction of Carbon Dioxide on Metallic Copper Surfaces, *Energy Environ. Sci.*, 2012, **5**, 7050–7059.
- 19 S. Nitopi, E. Bertheussen, S. B. Scott, X. Liu, A. K. Engstfeld, S. Horch, B. Seger, I. E. L. Stephens, K. Chan and C. Hahn, *et al.*, Progress and Perspectives of Electrochemical CO₂ Reduction on Copper in Aqueous Electrolyte, *Chem. Rev.*, 2019, **119**, 7610–7672.
- 20 P. Jeanty, C. Scherer, E. Magori, K. Wiesner-Fleischer, O. Hinrichsen and M. Fleischer, Upscaling and Continuous Operation of Electrochemical CO₂ to CO Conversion in Aqueous Solutions on Silver Gas Diffusion Electrodes, *J. CO₂ Util.*, 2018, **24**, 454–462.
- 21 E. R. Cofell, U. O. Nwabara, S. S. Bhargava, D. E. Henckel and P. J. A. Kenis, Investigation of Electrolyte-Dependent Carbonate Formation on Gas Diffusion Electrodes for CO₂ Electrolysis, *ACS Appl. Mater. Interfaces*, 2021, **13**, 15132–15142.
- 22 M. R. Singh, Y. Kwon, Y. Lum, J. W. Ager and A. T. Bell, Hydrolysis of Electrolyte Cations Enhances the Electrochemical Reduction of CO₂ over Ag and Cu, *J. Am. Chem. Soc.*, 2016, **138**, 13006–13012.
- 23 D. M. Weekes, D. A. Salvatore, A. Reyes, A. Huang and C. P. Berlinguette, Electrolytic CO₂ Reduction in a Flow Cell, *Acc. Chem. Res.*, 2018, **51**, 910–918.
- 24 Y. Y. Birdja, E. Pérez-Gallent, M. C. Figueiredo, A. J. Göttle, F. Calle-Vallejo and M. T. M. Koper, Advances and Challenges in Understanding the Electrocatalytic Conversion of Carbon Dioxide to Fuels, *Nat. Energy*, 2019, **4**, 732–745.
- 25 Z. Zhang, L. Melo, R. P. Janssonius, F. Habibzadeh, E. R. Grant and C. P. Berlinguette, pH Matters When Reducing CO₂ in an Electrochemical Flow Cell, *ACS Energy Lett.*, 2020, **5**, 3101–3107.
- 26 D. Kolb, The pH Concept, *J. Chem. Educ.*, 1979, **56**, 49–53.

- 27 J. A. Rabinowitz and M. W. Kanan, The Future of Low-Temperature Carbon Dioxide Electrolysis Depends on Solving One Basic Problem, *Nat. Commun.*, 2020, **11**, 5231.
- 28 K. G. Schulz, U. Riebesell, B. Rost, S. Thoms and R. E. Zeebe, Determination of the Rate Constants for the Carbon Dioxide to Bicarbonate Inter-Conversion in pH-Buffered Seawater Systems, *Mar. Geochem.*, 2006, **100**, 53–65.
- 29 G. A. Mills and H. C. Urey, The Kinetics of Isotropic Exchange between Carbon Dioxide, Bicarbonate Ion, Carbonate Ion and Water, *J. Am. Chem. Soc.*, 1940, **62**, 1019–1026.
- 30 A. J. Garza, A. T. Bell and M. Head-Gordon, Mechanism of CO₂ Reduction at Copper Surfaces: Pathways to C₂ Products, *ACS Catal.*, 2018, **8**, 1490–1499.
- 31 X. Zhao, L. Du and Y. Sun, Integrated Design for electrocatalytic carbon dioxide reduction, *Catal. Sci. Technol.*, 2020, **10**, 2711–2720.
- 32 X. Liu, P. Schlexer, J. Xiao, Y. Ji, L. Wang, R. B. Sandberg, M. Tang, K. S. Brown, H. Peng, S. Ringe, C. Hahn, T. F. Jaramillo, J. K. Nørskov and K. Chan, pH Effects on the Electrochemical Reduction of CO₂ towards C₂ Products on Stepped Copper, *Nat. Commun.*, 2019, **10**, 32.
- 33 J. A. Gauthier, Z. Lin, M. Head-Gordon and A. T. Bell, Pathways for the Formation of C₂₊ Products under Alkaline Conditions during the Electrochemical Reduction of CO₂, *ACS Energy Lett.*, 2022, **5**, 1679–1686.
- 34 K. Yang, R. Kas and W. A. Smith, In Situ Infrared Spectroscopy Reveals Persistent Alkalinity near Electrode Surfaces during CO₂ Electroreduction, *J. Am. Chem. Soc.*, 2019, **141**, 15891–15900.
- 35 L. Wang, S. A. Nitopi, E. Bertheussen, M. Orazov, C. G. Morales-Guio, X. Liu, D. C. Higgins, K. Chan, J. K. Nørskov, C. Hahn and T. F. Jaramillo, Electrochemical Carbon Monoxide Reduction on Polycrystalline Copper: Effects of Potential, Pressure, and pH on Selectivity toward Multicarbon and Oxygenated Products, *ACS Catal.*, 2018, **8**, 7445–7454.
- 36 J. C. Bui, C. Kim, A. Z. Weber and A. T. Bell, Dynamic Boundary Layer Simulation of Pulsed CO₂ Electrolysis on a Copper Catalyst, *ACS Energy Lett.*, 2021, **6**, 1181–1188.
- 37 N. Nesbitt and W. Smith, Water Activity Regulates CO₂ Reduction in Gas-Diffusion Electrodes, *J. Phys. Chem. C*, 2021, **125**(24), 13085–13095.
- 38 S. Suter and S. Haussener, Optimizing Mesostructured Silver Catalysts for Selective Carbon Dioxide Conversion into Fuels, *Energy Environ. Sci.*, 2019, **12**, 1668–1678.
- 39 A. Botz, J. Clausmeyer, D. Öhl, T. Tarnev, D. Franzen, T. Turek and W. Schuhmann, Local Activities of Hydroxide and Water Determine the Operation of Silver-Based Oxygen Depolarized Cathodes, *Angew. Chem., Int. Ed.*, 2018, **57**, 12285–12289.
- 40 S. Dieckhöfer, D. Öhl, J. R. C. Junqueira, T. Quast, T. Turek and W. Schuhmann, Probing the Local Reaction Environment During High Turnover Carbon Dioxide Reduction with Ag-Based Gas Diffusion Electrodes, *Chem. – Eur. J.*, 2021, **27**, 5906–5912.
- 41 X. Lu, C. Zhu, Z. Wu, J. Xuan, J. S. Francisco and H. Wang, In Situ Observation of the pH Gradient near the Gas Diffusion Electrode of CO₂ Reduction in Alkaline Electrolyte, *J. Am. Chem. Soc.*, 2020, **142**, 15438–15444.
- 42 N. Nesbitt and W. A. Smith, Operando Topography and Mechanical Property Mapping of CO₂ Reduction Gas-Diffusion Electrodes Operating at High Current Densities, *J. Electrochem. Soc.*, 2021, **168**, 044505.
- 43 M. C. O. Monteiro and M. T. M. Koper, Measuring Local pH in Electrochemistry, *Curr. Opin. Electrochem.*, 2021, **25**, 100649.
- 44 N. C. Rudd, S. Cannan, E. Bitziou, I. Ciani, A. Whitworth and P. R. Unwin, Fluorescence Confocal Laser Scanning Microscopy as a Probe of pH Gradients in Electrode Reactions and Surface Activity, *Anal. Chem.*, 2005, **77**, 6205–6217.
- 45 L. Bouffier and T. Doneux, Coupling Electrochemistry with In Situ Fluorescence (Confocal) Microscopy, *Curr. Opin. Electrochem.*, 2017, **6**, 31–37.
- 46 N. Pande, S. K. Chandrasekar, D. Lohse, G. Mul, J. A. Wood, B. T. Mei and D. Krug, Electrochemically Induced pH Change: Time-Resolved Confocal Fluorescence Microscopy Measurements and Comparison with Numerical Model, *J. Phys. Chem. Lett.*, 2020, **11**, 7042–7048.
- 47 J. E. Vitt and R. C. Engstrom, Imaging of Oxygen Evolution and Oxide Formation Using Quinine Fluorescence, *Anal. Chem.*, 1997, **69**, 1070–1076.
- 48 W. J. Bowyer, J. Xie and R. C. Engstrom, Fluorescence Imaging of the Heterogeneous Reduction of Oxygen, *Anal. Chem.*, 1996, **68**, 2005–2009.
- 49 A. J. Welch, A. Q. Fenwick, A. Boehme, H.-Y. Chen, I. Sullivan, X. Li, J. DuChene, C. Xiang and H. A. Atwater, Operando Local pH Measurement within Gas Diffusion Electrodes Performing Electrochemical Carbon Dioxide Reduction, *J. Phys. Chem. C*, 2021, **125**, 20896–20904.
- 50 R. Gui, H. Jin, X. Bu, Y. Fu, Z. Wang and Q. Liu, Recent advances in dual-emission ratiometric fluorescent probes for chemo/biosensing and bioimaging of biomarkers, *Coord. Chem. Rev.*, 2019, **383**, 82–103.
- 51 M. H. Lee, J. S. Kim and J. L. Sessler, Small molecule-based ratiometric fluorescent probes for cations, anions, and biomolecules, *Chem. Soc. Rev.*, 2015, **44**, 4185.
- 52 S.-H. Park, N. Kwon, J.-H. Lee, J. Yoon and I. Shin, Synthetic ratiometric fluorescent probes for detection of ions, *Chem. Soc. Rev.*, 2020, **49**, 143–179.
- 53 A. Hakonen and S. A. Hulth, High-Performance Fluoresensor for pH Measurements between 6 and 9, *Talanta*, 2010, **80**, 1964–1969.
- 54 J. M. Spurgeon and B. Kumar, A comparative techno-economic analysis of pathways for commercial electrochemical CO₂ reduction to liquid products, *Energy Environ. Sci.*, 2018, **11**, 1536.
- 55 M. G. Kibria, J. P. Edwards, C. M. Gabardo, C.-T. Dinh, A. Seifitokaldani, D. Sinton and E. H. Sargent, Electrochemical CO₂ Reduction into Chemical Feedstocks: From Mechanistic Electrocatalysis Models to System Design, *Adv. Mater.*, 2019, **31**, 1807166.
- 56 R. Bhide, C. N. Feltenberger, G. S. Phun, G. Barton, D. Fishman and S. Ardo, Quantification of Excited-State Brønsted–Lowry Acidity of Weak Photoacids Using Steady-

- State Photoluminescence Spectroscopy and a Driving-Force-Dependent Kinetic Theory, *J. Am. Chem. Soc.*, 2022, **144**, 14477–14488.
- 57 M. Sassenburg, R. de Rooji, N. T. Nessbitt, R. Kas, S. Chandrashekar, N. J. Firet, K. Yang, K. Liu, M. A. Blommaert, M. Kolen, D. Ripepi, W. A. Smith and T. Burdyny, Characterizing CO₂ Reduction Catalysts on Gas Diffusion Electrodes: Comparing Activity, Selectivity and Stability of Transition Metal Catalysts, *ACS Appl. Energy Mater.*, 2022, **5**, 5983–5994.
 - 58 J. C. Bui, E. W. Lees, L. M. Pant, I. V. Zenyuk, A. T. Bell and A. Z. Weber, Continuum Modeling of Porous Electrodes for Electrochemical Synthesis, *Chem. Rev.*, 2022, **122**(12), 11022–11084.
 - 59 M. Eigen, Proton Transfer, Acid–Base Catalysis, and Enzymatic Hydrolysis, *Angew. Chem., Int. Ed. Engl.*, 1964, **3**(1), 1–19.
 - 60 V. M. Ehlinger, A. R. Crothers, A. Kusoglu and A. Z. Weber, Modeling Proton-Exchange-Membrane Fuel Cell Performance/Degradation Tradeoffs with Chemical Scavengers, *J. Phys. Energy*, 2020, **2**(4), 044006.
 - 61 A. S. Hall, Y. Yoon, A. Wuttig and Y. Surendranath, Mesostructure-Induced Selectivity in CO₂ Reduction Catalysis, *J. Am. Chem. Soc.*, 2015, **137**, 14834–14837.
 - 62 A. J. Welch, J. S. Duchene, G. Tagliabue, A. Davoyan, W. H. Cheng and H. A. Atwater, Nanoporous Gold as a Highly Selective and Active Carbon Dioxide Reduction Catalyst, *ACS Appl. Energy Mater.*, 2019, **2**, 164–170.
 - 63 C. Kim, J. C. Bui, X. Luo, J. K. Cooper, A. Kusoglu, A. Z. Weber and A. T. Bell, Tailored Catalyst Microenvironments for CO₂ Electroreduction to Multicarbon Products on Copper Using Bilayer Ionomer Coatings, *Nat. Energy*, 2021, **6**(11), 1026–1034.