Median dual cell 3D

August 22, 2007

1 Equations

1.1 Mass conservation

$$\frac{\partial c_i}{\partial t} = -\vec{\nabla}.\vec{N}_i + R_i \tag{1}$$

with

$$\vec{N}_i = c_i \vec{v} - \sum_j D_{ij} \vec{\nabla} c_j - w_i c_i \vec{\nabla} U \tag{2}$$

1.2 Poisson's equation

$$\vec{\nabla}^2 U + \frac{F}{\epsilon} \sum_i z_i c_i = 0 \tag{3}$$

1.3 Butler-Volmer kinetics

$$v = k_{ox} \exp \left[\frac{\alpha_{ox} nF}{RT} \left(V - U \right) \right] c_{red} - k_{red} \exp \left[-\frac{\alpha_{red} nF}{RT} \left(V - U \right) \right] c_{ox}$$
 (4)

2 Element matrix

2.1 Convection

2.1.1 Fluctuation in node

$$\Delta c_i^m = \int_{V^m} -\vec{v}.\vec{\nabla}c_i dV$$

2.1.2 Element contribution to fluctuation in node

$$\vec{\mathcal{V}}^1 = 75\vec{v}^1 + 23\vec{v}^2 + 23\vec{v}^3 + 23\vec{v}^4$$

$$\vec{\mathcal{V}}^2 = 23\vec{v}^1 + 75\vec{v}^2 + 23\vec{v}^3 + 23\vec{v}^4$$

$$\vec{\mathcal{V}}^3 = 23\vec{v}^1 + 23\vec{v}^2 + 75\vec{v}^3 + 23\vec{v}^4$$

$$\vec{\mathcal{V}}^4 = 23\vec{v}^1 + 23\vec{v}^2 + 23\vec{v}^3 + 75\vec{v}^4$$

$$\left\{ \begin{array}{l} \Delta c_{i}^{1} \\ \Delta c_{i}^{2} \\ \Delta c_{i}^{3} \\ \Delta c_{i}^{4} \end{array} \right\} = \frac{1}{1728} \left[\begin{array}{ccccc} \vec{\mathcal{V}}^{1}.\vec{n}^{1} & \vec{\mathcal{V}}^{1}.\vec{n}^{2} & \vec{\mathcal{V}}^{1}.\vec{n}^{3} & \vec{\mathcal{V}}^{1}.\vec{n}^{4} \\ \vec{\mathcal{V}}^{2}.\vec{n}^{1} & \vec{\mathcal{V}}^{2}.\vec{n}^{2} & \vec{\mathcal{V}}^{2}.\vec{n}^{3} & \vec{\mathcal{V}}^{2}.\vec{n}^{4} \\ \vec{\mathcal{V}}^{3}.\vec{n}^{1} & \vec{\mathcal{V}}^{3}.\vec{n}^{2} & \vec{\mathcal{V}}^{3}.\vec{n}^{3} & \vec{\mathcal{V}}^{3}.\vec{n}^{4} \\ \vec{\mathcal{V}}^{4}.\vec{n}^{1} & \vec{\mathcal{V}}^{4}.\vec{n}^{2} & \vec{\mathcal{V}}^{4}.\vec{n}^{3} & \vec{\mathcal{V}}^{4}.\vec{n}^{4} \end{array} \right] \left\{ \begin{array}{c} c_{i}^{1} \\ c_{i}^{2} \\ c_{i}^{3} \\ c_{i}^{4} \end{array} \right\}$$

$$(5)$$

2.1.3 Examples: binary electrolyte

with

$$V^{mn} = \frac{\vec{\mathcal{V}}^m . \vec{n}^n}{1728}$$

2.2 Diffusion

2.2.1 Fluctuation in node

$$\Delta c_i^m = \int_{V^m} \vec{\nabla} \cdot \left(\sum_j D_{ij} \vec{\nabla} c_j \right) dV$$
$$= - \int_{\partial V^m} \left(\sum_j D_{ij} \vec{\nabla} c_j \right) . d\vec{n}$$

2.2.2 Element contribution to fluctuation in node

Assume that D_{ij} varies linearly.

$$\vec{\mathcal{D}}_{ij}^{1} = 13\vec{n}^{1}D_{ij}^{1} + \left(7\vec{n}^{1} - 2\vec{n}^{2}\right)D_{ij}^{2} + \left(7\vec{n}^{1} - 2\vec{n}^{3}\right)D_{ij}^{3} + \left(7\vec{n}^{1} - 2\vec{n}^{4}\right)D_{ij}^{4}$$

$$\vec{\mathcal{D}}_{ij}^2 = 13\vec{n}^2D_{ij}^2 + \left(7\vec{n}^2 - 2\vec{n}^3\right)D_{ij}^3 + \left(7\vec{n}^2 - 2\vec{n}^4\right)D_{ij}^4 + \left(7\vec{n}^2 - 2\vec{n}^1\right)D_{ij}^1$$

$$\vec{\mathcal{D}}_{ij}^{3} = 13\vec{n}^{3}D_{ij}^{3} + \left(7\vec{n}^{3} - 2\vec{n}^{4}\right)D_{ij}^{4} + \left(7\vec{n}^{3} - 2\vec{n}^{1}\right)D_{ij}^{1} + \left(7\vec{n}^{3} - 2\vec{n}^{2}\right)D_{ij}^{2}$$

$$\vec{\mathcal{D}}_{ij}^{4} = 13\vec{n}^{3}D_{ij}^{4} + \left(7\vec{n}^{4} - 2\vec{n}^{1}\right)D_{ij}^{1} + \left(7\vec{n}^{4} - 2\vec{n}^{2}\right)D_{ij}^{2} + \left(7\vec{n}^{4} - 2\vec{n}^{3}\right)D_{ij}^{3}$$

$$\left\{ \begin{array}{l} \Delta c_{i}^{1} \\ \Delta c_{i}^{2} \\ \Delta c_{i}^{3} \\ \Delta c_{i}^{4} \end{array} \right\} = -\frac{1}{324V} \sum_{j} \left[\begin{array}{cccc} \vec{\mathcal{D}}_{ij}^{1}.\vec{n}^{1} & \vec{\mathcal{D}}_{ij}^{1}.\vec{n}^{2} & \vec{\mathcal{D}}_{ij}^{1}.\vec{n}^{3} & \vec{\mathcal{D}}_{ij}^{1}.\vec{n}^{4} \\ \vec{\mathcal{D}}_{ij}^{2}.\vec{n}^{1} & \vec{\mathcal{D}}_{ij}^{2}.\vec{n}^{2} & \vec{\mathcal{D}}_{ij}^{2}.\vec{n}^{3} & \vec{\mathcal{D}}_{ij}^{2}.\vec{n}^{4} \\ \vec{\mathcal{D}}_{ij}^{3}.\vec{n}^{1} & \vec{\mathcal{D}}_{ij}^{3}.\vec{n}^{2} & \vec{\mathcal{D}}_{ij}^{3}.\vec{n}^{3} & \vec{\mathcal{D}}_{ij}^{3}.\vec{n}^{4} \\ \vec{\mathcal{D}}_{ij}^{4}.\vec{n}^{1} & \vec{\mathcal{D}}_{ij}^{4}.\vec{n}^{2} & \vec{\mathcal{D}}_{ij}^{4}.\vec{n}^{3} & \vec{\mathcal{D}}_{ij}^{4}.\vec{n}^{4} \end{array} \right] \left\{ \begin{array}{c} c_{i}^{1} \\ c_{i}^{2} \\ c_{j}^{3} \\ c_{j}^{4} \end{array} \right\}$$

2.2.3 Example: binary electrolyte

with

$$D_{ij}^{mn} = -\frac{\vec{\mathcal{D}}_{ij}^m.\vec{n}^n}{324V}$$

2.3 Migration

2.3.1 Fluctuation in node

$$\begin{array}{rcl} \Delta c_i^m & = & \int_{V^m} \vec{\nabla} . \left(w_i c_i \vec{\nabla} U \right) dV \\ & = & - \int_{\partial V^m} \left(w_i c_i \vec{\nabla} U \right) . d\vec{n} \end{array}$$

2.3.2 Element contribution to fluctuation in node

Assume that w_i varies linearly.

$$\vec{\mathcal{M}}_{i}^{1} = \vec{\mathcal{W}}_{i}^{11}c_{i}^{1} + \vec{\mathcal{W}}_{i}^{12}c_{i}^{2} + \vec{\mathcal{W}}_{i}^{13}c_{i}^{3} + \vec{\mathcal{W}}_{i}^{14}c_{i}^{4}$$

$$\vec{\mathcal{M}}_{i}^{2} = \vec{\mathcal{W}}_{i}^{21}c_{i}^{1} + \vec{\mathcal{W}}_{i}^{22}c_{i}^{2} + \vec{\mathcal{W}}_{i}^{23}c_{i}^{3} + \vec{\mathcal{W}}_{i}^{24}c_{i}^{4}$$

$$\begin{split} \vec{\mathcal{M}}_{i}^{4} &= \vec{\mathcal{W}}_{i}^{41} c_{i}^{1} + \vec{\mathcal{W}}_{i}^{42} c_{i}^{2} + \vec{\mathcal{W}}_{i}^{43} c_{i}^{3} + \vec{\mathcal{W}}_{i}^{44} c_{i}^{4} \\ \vec{\mathcal{W}}_{i}^{11} &= 230\vec{n}^{1} w_{i}^{1} + (119\vec{n}^{1} - 37\vec{n}^{2}) w_{i}^{2} + (119\vec{n}^{1} - 37\vec{n}^{3}) w_{i}^{3} + (119\vec{n}^{1} - 37\vec{n}^{4}) w_{i}^{4} \\ \vec{\mathcal{W}}_{i}^{12} &= (119\vec{n}^{1} - 37\vec{n}^{2}) w_{i}^{1} + (92\vec{n}^{1} - 46\vec{n}^{2}) w_{i}^{2} + (69\vec{n}^{1} + 13\vec{n}^{4}) w_{i}^{3} + (69\vec{n}^{1} + 13\vec{n}^{3}) w_{i}^{4} \\ \vec{\mathcal{W}}_{i}^{13} &= (119\vec{n}^{1} - 37\vec{n}^{3}) w_{i}^{1} + (69\vec{n}^{1} + 13\vec{n}^{3}) w_{i}^{2} + (69\vec{n}^{1} + 13\vec{n}^{2}) w_{i}^{3} + (69\vec{n}^{1} + 13\vec{n}^{2}) w_{i}^{4} \\ \vec{\mathcal{W}}_{i}^{14} &= (119\vec{n}^{1} - 37\vec{n}^{4}) w_{i}^{1} + (69\vec{n}^{1} + 13\vec{n}^{3}) w_{i}^{2} + (69\vec{n}^{1} + 13\vec{n}^{2}) w_{i}^{3} + (69\vec{n}^{2} + 13\vec{n}^{3}) w_{i}^{4} \\ \vec{\mathcal{W}}_{i}^{21} &= (92\vec{n}^{2} - 46\vec{n}^{1}) w_{i}^{1} + (119\vec{n}^{2} - 37\vec{n}^{1}) w_{i}^{2} + (69\vec{n}^{2} + 13\vec{n}^{3}) w_{i}^{3} + (19\vec{n}^{2} - 37\vec{n}^{4}) w_{i}^{4} \\ \vec{\mathcal{W}}_{i}^{22} &= (119\vec{n}^{2} - 37\vec{n}^{1}) w_{i}^{1} + 230\vec{n}^{2}w_{i}^{2} + (119\vec{n}^{2} - 37\vec{n}^{3}) w_{i}^{3} + (119\vec{n}^{2} - 37\vec{n}^{4}) w_{i}^{4} \\ \vec{\mathcal{W}}_{i}^{23} &= (69\vec{n}^{2} + 13\vec{n}^{3}) w_{i}^{1} + (119\vec{n}^{2} - 37\vec{n}^{3}) w_{i}^{2} + (69\vec{n}^{2} + 13\vec{n}^{3}) w_{i}^{3} + (69\vec{n}^{2} + 13\vec{n}^{3}) w_{i}^{4} \\ \vec{\mathcal{W}}_{i}^{31} &= (69\vec{n}^{2} + 13\vec{n}^{3}) w_{i}^{1} + (19\vec{n}^{2} - 37\vec{n}^{3}) w_{i}^{2} + (119\vec{n}^{3} - 37\vec{n}^{2}) w_{i}^{3} + (69\vec{n}^{3} + 13\vec{n}^{2}) w_{i}^{4} \\ \vec{\mathcal{W}}_{i}^{33} &= (69\vec{n}^{3} + 13\vec{n}^{3}) w_{i}^{1} + (92\vec{n}^{3} - 46\vec{n}^{2}) w_{i}^{2} + (119\vec{n}^{3} - 37\vec{n}^{2}) w_{i}^{3} + (69\vec{n}^{3} + 13\vec{n}^{2}) w_{i}^{4} \\ \vec{\mathcal{W}}_{i}^{33} &= (69\vec{n}^{3} + 13\vec{n}^{3}) w_{i}^{1} + (69\vec{n}^{3} + 13\vec{n}^{1}) w_{i}^{2} + (119\vec{n}^{3} - 37\vec{n}^{2}) w_{i}^{3} + (69\vec{n}^{3} + 13\vec{n}^{2}) w_{i}^{4} \\ \vec{\mathcal{W}}_{i}^{33} &= (69\vec{n}^{3} + 13\vec{n}^{2}) w_{i}^{1} + (69\vec{n}^{3} + 13\vec{n}^{1}) w_{i}^{2} + (119\vec{n}^{3} - 37\vec{n}^{2}) w_{i}^{3} + (119\vec{n}^{3} - 37\vec{n}^{2}) w_{i}^{4} \\ \vec{\mathcal{W}}_{i}^{43} &= (69\vec{n}^$$

 $\vec{\mathcal{M}}_{i}^{3} = \vec{\mathcal{W}}_{i}^{31}c_{i}^{1} + \vec{\mathcal{W}}_{i}^{32}c_{i}^{2} + \vec{\mathcal{W}}_{i}^{33}c_{i}^{3} + \vec{\mathcal{W}}_{i}^{34}c_{i}^{4}$

2.3.3 Example: binary electrolyte

with

$$M_i^{mn} = -\frac{\vec{\mathcal{M}}_A^m.\vec{n}^n}{15552V}$$

2.4 Homogeneous reactions

2.4.1 Fluctuation in node

$$R_i = \sum_r s_{i,r} v_r$$

• Monomolecular

$$v = kc_j$$

$$\Delta c_i^m = \int_{V^m} k c_j dV$$

• Bimolecular

$$v = kc_j c_k$$

$$\Delta c_i^m = \int_{V^m} k c_j c_k dV$$

2.4.2 Element contribution to fluctuation in node

Assume that k varies linearly.

ullet Monomolecular

$$\mathcal{H}_{j}^{11} = 1245k^{1} + 335k^{2} + 335k^{3} + 335k^{4}$$

$$\mathcal{H}_{j}^{12} = 335k^{1} + 161k^{2} + 97k^{3} + 97k^{4}$$

$$\mathcal{H}_{j}^{13} = 335k^{1} + 97k^{2} + 161k^{3} + 97k^{4}$$

$$\mathcal{H}_{j}^{14} = 335k^{1} + 97k^{2} + 97k^{3} + 161k^{4}$$

$$\mathcal{H}_{j}^{14} = 335k^{1} + 97k^{2} + 97k^{3} + 161k^{4}$$

$$\mathcal{H}_{j}^{22} = 161k^{1} + 335k^{2} + 97k^{3} + 97k^{4}$$

$$\mathcal{H}_{j}^{22} = 335k^{1} + 1245k^{2} + 335k^{3} + 335k^{4}$$

$$\mathcal{H}_{j}^{23} = 97k^{1} + 335k^{2} + 161k^{3} + 97k^{4}$$

$$\mathcal{H}_{j}^{24} = 97k^{1} + 335k^{2} + 97k^{3} + 161k^{4}$$

$$\mathcal{H}_{j}^{31} = 161k^{1} + 97k^{2} + 335k^{3} + 97k^{4}$$

$$\mathcal{H}_{j}^{32} = 97k^{1} + 161k^{2} + 335k^{3} + 97k^{4}$$

$$\mathcal{H}_{j}^{33} = 335k^{1} + 335k^{2} + 1245k^{3} + 335k^{4}$$

$$\mathcal{H}_{j}^{34} = 97k^{1} + 97k^{2} + 335k^{3} + 161k^{4}$$

$$\mathcal{H}_{j}^{41} = 161k^{1} + 97k^{2} + 97k^{3} + 335k^{4}$$

$$\mathcal{H}_{j}^{42} = 97k^{1} + 161k^{2} + 97k^{3} + 335k^{4}$$

$$\mathcal{H}_{j}^{43} = 97k^{1} + 97k^{2} + 161k^{3} + 335k^{4}$$

$$\mathcal{H}_{j}^{44} = 335k^{1} + 335k^{2} + 335k^{3} + 1245k^{4}$$

$$\mathcal{H}_{j}^{44} = 335k^{1} + 335k^{2} + 335k^{3} + 1245k^{4}$$

$$\mathcal{H}_{j}^{44} = 335k^{1} + 35k^{2} + 335k^{3} + 1245k^{4}$$

$$\mathcal{H}_{j}^{44} = 335k^{1} + 37k^{2} + 37k^{2} + 37k^{3} + 37k^{4}$$

$$\mathcal{H}_{j}^{44} = 335k^{1} + 37k^{2} + 37k^{3} + 37k^{4}$$

$$\mathcal{H}_{j}^{44} = 335k^{1} + 37k^{2} + 37k^{3} + 37k^{4}$$

$$\mathcal{H}_{j}^{44} = 37k^{4} + 37k^{2} + 37k^{4} + 37k$$

• Bimolecular

$$\mathcal{K}^{111} = 17535k^1 + 4115k^2 + 4115k^3 + 4115k^4$$

$$\mathcal{K}^{222} = 4115k^1 + 17535k^2 + 4115k^3 + 4115k^4$$

$$\mathcal{K}^{333} = 4115k^1 + 4115k^2 + 17535k^3 + 4115k^4$$

$$\mathcal{K}^{444} = 4115k^1 + 4115k^2 + 4115k^3 + 17535k^4$$

$$\mathcal{K}^{112} = \mathcal{K}^{121} = 4115k^1 + 1815k^2 + 1055k^3 + 1055k^4$$

$$\mathcal{K}^{113} = \mathcal{K}^{131} = 4115k^1 + 1055k^2 + 1815k^3 + 1055k^4$$

$$\mathcal{K}^{114} = \mathcal{K}^{141} = 4115k^1 + 1055k^2 + 1055k^3 + 1055k^4$$

$$\mathcal{K}^{212} = \mathcal{K}^{221} = 1815k^1 + 4115k^2 + 1055k^3 + 1055k^4$$

$$\mathcal{K}^{223} = \mathcal{K}^{232} = 1055k^1 + 4115k^2 + 1055k^3 + 1055k^4$$

$$\mathcal{K}^{224} = \mathcal{K}^{242} = 1055k^1 + 4115k^2 + 1055k^3 + 1815k^4$$

$$\mathcal{K}^{331} = \mathcal{K}^{313} = 1815k^1 + 1055k^2 + 4115k^3 + 1055k^4$$

$$\mathcal{K}^{332} = \mathcal{K}^{323} = 1055k^1 + 1815k^2 + 4115k^3 + 1055k^4$$

$$\mathcal{K}^{334} = \mathcal{K}^{343} = 1055k^1 + 1055k^2 + 4115k^3 + 1815k^4$$

$$\mathcal{K}^{441} = \mathcal{K}^{414} = 1815k^1 + 1055k^2 + 1055k^3 + 4115k^4$$

$$\mathcal{K}^{442} = \mathcal{K}^{424} = 1055k^1 + 1815k^2 + 1055k^3 + 4115k^4$$

$$\mathcal{K}^{442} = \mathcal{K}^{424} = 1055k^1 + 1815k^2 + 1055k^3 + 4115k^4$$

$$\mathcal{K}^{443} = \mathcal{K}^{434} = 1055k^1 + 1055k^2 + 1815k^3 + 4115k^4$$

$$\mathcal{K}^{443} = \mathcal{K}^{434} = 1055k^1 + 1055k^2 + 1815k^3 + 4115k^4$$

$$\mathcal{K}^{443} = \mathcal{K}^{434} = 1055k^1 + 1055k^2 + 1815k^3 + 4115k^4$$

$$\mathcal{K}^{443} = \mathcal{K}^{434} = 1055k^1 + 1055k^2 + 1815k^3 + 4115k^4$$

$$\mathcal{K}^{443} = \mathcal{K}^{434} = 1055k^1 + 1055k^2 + 1815k^3 + 4115k^4$$

$$\mathcal{K}^{443} = \mathcal{K}^{434} = 1055k^1 + 1055k^2 + 1815k^3 + 4115k^4$$

$$\mathcal{K}^{122} = 1815k^1 + 1067k^2 + 491k^3 + 491k^4$$

$$\mathcal{K}^{133} = 1815k^1 + 491k^2 + 1067k^3 + 491k^4$$

$$\mathcal{K}^{144} = 1815k^1 + 491k^2 + 491k^3 + 1067k^4$$

$$\mathcal{K}^{211} = 1067k^1 + 1815k^2 + 491k^3 + 491k^4$$

$$\mathcal{K}^{233} = 491k^1 + 1815k^2 + 1067k^3 + 491k^4$$

$$\mathcal{K}^{244} = 491k^1 + 1815k^2 + 491k^3 + 1067k^4$$

$$\mathcal{K}^{311} = 1067k^1 + 491k^2 + 1815k^3 + 491k^4$$

$$\mathcal{K}^{322} = 491k^1 + 1067k^2 + 1815k^3 + 491k^4$$

$$\mathcal{K}^{344} = 491k^1 + 491k^2 + 1815k^3 + 1067k^4$$

$$\mathcal{K}^{411} = 1067k^1 + 491k^2 + 491k^3 + 1815k^4$$

$$\mathcal{K}^{422} = 491k^1 + 1067k^2 + 491k^3 + 1815k^4$$

$$\mathcal{K}^{422} = 491k^1 + 1067k^2 + 491k^3 + 1815k^4$$

$$\mathcal{K}^{433} = 491k^1 + 491k^2 + 1067k^3 + 1815k^4$$

$$\mathcal{K}^{433} = \mathcal{K}^{132} = 1055k^1 + 491k^2 + 491k^3 + 291k^4$$

$$\mathcal{K}^{124} = \mathcal{K}^{142} = 1055k^1 + 491k^2 + 291k^3 + 491k^4$$

$$\mathcal{K}^{134} = \mathcal{K}^{143} = 1055k^1 + 291k^2 + 491k^3 + 491k^4$$

$$\mathcal{K}^{213} = \mathcal{K}^{231} = 491k^1 + 1055k^2 + 491k^3 + 291k^4$$

$$\mathcal{K}^{214} = \mathcal{K}^{241} = 491k^1 + 1055k^2 + 291k^3 + 491k^4$$

$$\mathcal{K}^{214} = \mathcal{K}^{241} = 491k^1 + 1055k^2 + 291k^3 + 491k^4$$

$$\mathcal{K}^{234} = \mathcal{K}^{243} = 291k^1 + 1055k^2 + 491k^3 + 491k^4$$

$$\mathcal{K}^{312} = \mathcal{K}^{321} = 491k^{1} + 491k^{2} + 1055k^{3} + 291k^{4}$$

$$\mathcal{K}^{314} = \mathcal{K}^{341} = 491k^{1} + 291k^{2} + 1055k^{3} + 491k^{4}$$

$$\mathcal{K}^{324} = \mathcal{K}^{342} = 291k^{1} + 491k^{2} + 1055k^{3} + 491k^{4}$$

$$\mathcal{K}^{412} = \mathcal{K}^{421} = 491k^{1} + 491k^{2} + 291k^{3} + 1055k^{4}$$

$$\mathcal{K}^{413} = \mathcal{K}^{431} = 491k^{1} + 291k^{2} + 491k^{3} + 1055k^{4}$$

$$\mathcal{K}^{423} = \mathcal{K}^{432} = 291k^{1} + 491k^{2} + 491k^{3} + 1055k^{4}$$

$$\mathcal{H}^{11}_{j} = \mathcal{K}^{111}c_{k}^{1} + \mathcal{K}^{112}c_{k}^{2} + \mathcal{K}^{113}c_{k}^{3} + \mathcal{K}^{114}c_{k}^{4}$$

$$\mathcal{H}^{12}_{j} = \mathcal{K}^{121}c_{k}^{1} + \mathcal{K}^{122}c_{k}^{2} + \mathcal{K}^{123}c_{k}^{3} + \mathcal{K}^{124}c_{k}^{4}$$

$$\mathcal{H}^{13}_{j} = \mathcal{K}^{131}c_{k}^{1} + \mathcal{K}^{132}c_{k}^{2} + \mathcal{K}^{133}c_{k}^{3} + \mathcal{K}^{134}c_{k}^{4}$$

$$\mathcal{H}^{14}_{j} = \mathcal{K}^{141}c_{k}^{1} + \mathcal{K}^{142}c_{k}^{2} + \mathcal{K}^{143}c_{k}^{3} + \mathcal{K}^{144}c_{k}^{4}$$

$$\mathcal{H}^{21}_{j} = \mathcal{K}^{211}c_{k}^{1} + \mathcal{K}^{212}c_{k}^{2} + \mathcal{K}^{213}c_{k}^{3} + \mathcal{K}^{214}c_{k}^{4}$$

$$\mathcal{H}^{22}_{j} = \mathcal{K}^{221}c_{k}^{1} + \mathcal{K}^{222}c_{k}^{2} + \mathcal{K}^{223}c_{k}^{3} + \mathcal{K}^{224}c_{k}^{4}$$

$$\mathcal{H}^{22}_{j} = \mathcal{K}^{221}c_{k}^{1} + \mathcal{K}^{222}c_{k}^{2} + \mathcal{K}^{233}c_{k}^{3} + \mathcal{K}^{224}c_{k}^{4}$$

$$\mathcal{H}^{23}_{j} = \mathcal{K}^{231}c_{k}^{1} + \mathcal{K}^{232}c_{k}^{2} + \mathcal{K}^{233}c_{k}^{3} + \mathcal{K}^{234}c_{k}^{4}$$

$$\mathcal{H}^{24}_{j} = \mathcal{K}^{241}c_{k}^{1} + \mathcal{K}^{242}c_{k}^{2} + \mathcal{K}^{243}c_{k}^{3} + \mathcal{K}^{234}c_{k}^{4}$$

$$\mathcal{H}^{31}_{j} = \mathcal{K}^{311}c_{k}^{1} + \mathcal{K}^{312}c_{k}^{2} + \mathcal{K}^{313}c_{k}^{3} + \mathcal{K}^{314}c_{k}^{4}$$

$$\mathcal{H}^{31}_{j} = \mathcal{K}^{311}c_{k}^{1} + \mathcal{K}^{312}c_{k}^{2} + \mathcal{K}^{313}c_{k}^{3} + \mathcal{K}^{314}c_{k}^{4}$$

$$\mathcal{H}^{31}_{j} = \mathcal{K}^{311}c_{k}^{1} + \mathcal{K}^{312}c_{k}^{2} + \mathcal{K}^{313}c_{k}^{3} + \mathcal{K}^{314}c_{k}^{4}$$

$$\mathcal{H}^{31}_{j} = \mathcal{K}^{311}c_{k}^{1} + \mathcal{K}^{312}c_{k}^{2} + \mathcal{K}^{313}c_{k}^{3} + \mathcal{K}^{314}c_{k}^{4}$$

$$\mathcal{H}^{31}_{j} = \mathcal{K}^{311}c_{k}^{1} + \mathcal{K}^{312}c_{k}^{2} + \mathcal{K}^{313}c_{k}^{3} + \mathcal{K}^{314}c_{k}^{4}$$

$$\mathcal{H}^{31}_{j} = \mathcal{K}^{311}c_{k}^{1} + \mathcal{K}^{312}c_{k}^{2} + \mathcal{K}^{313}c_{k}^{3} + \mathcal{K}^{314}c_{k}^{4}$$

$$\begin{split} \mathcal{H}_{j}^{33} &= \mathcal{K}^{331}c_{k}^{1} + \mathcal{K}^{332}c_{k}^{2} + \mathcal{K}^{333}c_{k}^{3} + \mathcal{K}^{334}c_{k}^{4} \\ \mathcal{H}_{j}^{34} &= \mathcal{K}^{341}c_{k}^{1} + \mathcal{K}^{342}c_{k}^{2} + \mathcal{K}^{343}c_{k}^{3} + \mathcal{K}^{344}c_{k}^{4} \\ \mathcal{H}_{j}^{41} &= \mathcal{K}^{411}c_{k}^{1} + \mathcal{K}^{412}c_{k}^{2} + \mathcal{K}^{413}c_{k}^{3} + \mathcal{K}^{414}c_{k}^{4} \\ \mathcal{H}_{j}^{42} &= \mathcal{K}^{421}c_{k}^{1} + \mathcal{K}^{422}c_{k}^{2} + \mathcal{K}^{423}c_{k}^{3} + \mathcal{K}^{424}c_{k}^{4} \\ \mathcal{H}_{j}^{43} &= \mathcal{K}^{431}c_{k}^{1} + \mathcal{K}^{432}c_{k}^{2} + \mathcal{K}^{433}c_{k}^{3} + \mathcal{K}^{434}c_{k}^{4} \\ \mathcal{H}_{j}^{44} &= \mathcal{K}^{441}c_{k}^{1} + \mathcal{K}^{442}c_{k}^{2} + \mathcal{K}^{443}c_{k}^{3} + \mathcal{K}^{444}c_{k}^{4} \end{split}$$

2.4.3 Example: monomolecular-monomolecular reversible reaction $A \rightleftharpoons B$

$$v = k_f c_A - k_b c_B$$

$$R_A = -v$$

$$R_B = v$$

with for the forward reaction (replace k by k_f in the formulae!)

$$H_A^{mn} = \frac{V}{17280} \mathcal{H}_A^{mn}$$

and for the backward reaction (replace k by k_b in the formulae!)

$$H_B^{mn} = \frac{V}{17280} \mathcal{H}_B^{mn}$$

2.5 Poisson's equation

2.5.1 Fluctuation in node

$$\begin{array}{lll} \Delta U^m & = & \int_{V^m} \vec{\nabla}^2 U dV + \int_{V^m} \frac{F}{\epsilon} \sum_i z_i c_i dV \\ & = & - \int_{\partial V^m} \vec{\nabla} U . d\vec{n} + \sum_i \frac{z_i F}{\epsilon} \int_{V^m} c_i dV \end{array}$$

2.5.2 Element contribution to fluctuation in node

$$\begin{cases}
\Delta U^{1} \\
\Delta U^{2} \\
\Delta U^{3} \\
\Delta U^{4}
\end{cases} = \begin{bmatrix}
\frac{75V}{576} \frac{z_{i}F}{\epsilon} & -\frac{\vec{n}^{1} \cdot \vec{n}^{1}}{12V} & \frac{23V}{576} \frac{z_{i}F}{\epsilon} & -\frac{\vec{n}^{1} \cdot \vec{n}^{2}}{12V} & \frac{23V}{576} \frac{z_{i}F}{\epsilon} & -\frac{\vec{n}^{1} \cdot \vec{n}^{3}}{12V} & \frac{23V}{576} \frac{z_{i}F}{\epsilon} & -\frac{\vec{n}^{2} \cdot \vec{n}^{3}}{12V} & \frac{23V}{576} \frac{z_{i}F}{\epsilon}$$

2.5.3 Example: binary electrolyte

with

$$Z_i = \frac{V}{576} \frac{z_i F}{\epsilon}$$

$$n^{mn} = -\frac{\vec{n}^m \cdot \vec{n}^n}{12V}$$

2.6 Time

2.6.1 Fluctuation in node

$$\Delta c_i^m = \int_{V^m} \frac{\partial c_i}{\partial t} dV$$

2.6.2 Element contribution to fluctuation in node

$$\begin{cases}
\Delta c_i^1 \\
\Delta c_i^2 \\
\Delta c_i^3 \\
\Delta c_i^4
\end{cases} = \frac{V}{576} \begin{bmatrix} 75 & 7 & 7 & 7 \\
7 & 75 & 7 & 7 \\
7 & 7 & 75 & 7 \\
7 & 7 & 7 & 75
\end{bmatrix} \begin{cases}
\frac{\partial c_i^1}{\partial t_i} \\
\frac{\partial c_i^2}{\partial t_i} \\
\frac{\partial c_i^2}{\partial t_i} \\
\frac{\partial c_i^2}{\partial t_i} \\
\frac{\partial c_i^2}{\partial t_i}
\end{cases} \tag{16}$$

3 Boundary element vector

3.1 Electrode reactions

3.1.1 Fluctuation in node

$$R_i = \sum_r s_{i,r} v_r$$

$$v = k_{ox} \exp \left[\frac{\alpha_{ox} nF}{RT} (V - U) \right] c_{red} - k_{red} \exp \left[-\frac{\alpha_{red} nF}{RT} (V - U) \right] c_{ox}$$

$$\Delta c_i^m = \int_{S^m} R_i dS$$

3.1.2 Boundary element contribution to fluctuation in node

Assume that the normal flux varies linearly on the boundary.

$$\left\{ \begin{array}{l} \Delta c_i^1 \\ \Delta c_i^2 \\ \Delta c_i^3 \end{array} \right\} = \frac{S}{108} \left\{ \begin{array}{l} 22R_i^1 + 7R_i^2 + 7R_i^3 \\ 7R_i^1 + 22R_i^2 + 7R_i^3 \\ 7R_i^1 + 7R_i^2 + 22R_i^3 \end{array} \right\}$$
(17)

3.1.3 Example: binary electrolyte, $A \rightleftharpoons B + ne^{-}$

$$v = k_{ox} \exp \left[\frac{\alpha_{ox} nF}{RT} (V - U)\right] c_B - k_{red} \exp \left[-\frac{\alpha_{red} nF}{RT} (V - U)\right] c_A$$

$$R_A = -v$$

$$R_B = v$$

$$\begin{cases}
\Delta c_A^1 \\
\Delta c_B^1 \\
\Delta U^1 \\
\Delta c_A^2 \\
\Delta c_A^2 \\
\Delta U^2 \\
\Delta c_B^3 \\
\Delta U^3
\end{cases} = \begin{cases}
-\frac{S}{108} \left(22N_i^1 + 7N_i^2 + 7N_i^3\right) \\
\frac{S}{108} \left(22N_i^1 + 7N_i^2 + 7N_i^3\right) \\
0 \\
-\frac{S}{108} \left(7N_i^1 + 22N_i^2 + 7N_i^3\right) \\
\frac{S}{108} \left(7N_i^1 + 22N_i^2 + 7N_i^3\right) \\
0 \\
-\frac{S}{108} \left(7N_i^1 + 7N_i^2 + 22N_i^3\right) \\
\frac{S}{108} \left(7N_i^1 + 7N_i^2 + 22N_i^3\right) \\
\frac{S}{108} \left(7N_i^1 + 7N_i^2 + 22N_i^3\right) \\
0
\end{cases}$$
(18)

4 Element jacobian

4.1 Convection

Zero contribution.

4.2 Diffusion

Zero contribution (approximately).

4.3 Migration

4.3.1 Element contribution to fluctuation in node

$$\vec{\nabla} U = \frac{\vec{n}^1 U^1 + \vec{n}^2 U^2 + \vec{n}^3 U^3 + \vec{n}^4 U^4}{3V}$$

$$\left\{ \begin{array}{c} \Delta c_i^1 \\ \Delta c_i^2 \\ \Delta c_i^3 \\ \Delta c_i^4 \end{array} \right\} = -\frac{\vec{\nabla} U}{5184} \cdot \left[\begin{array}{cccc} \vec{W}_i^{11} & \vec{W}_i^{12} & \vec{W}_i^{13} & \vec{W}_i^{14} \\ \vec{W}_i^{21} & \vec{W}_i^{22} & \vec{W}_i^{23} & \vec{W}_i^{24} \\ \vec{W}_i^{31} & \vec{W}_i^{32} & \vec{W}_i^{33} & \vec{W}_i^{34} \\ \vec{W}_i^{41} & \vec{W}_i^{42} & \vec{W}_i^{43} & \vec{W}_i^{44} \end{array} \right] \left\{ \begin{array}{c} c_i^1 \\ c_i^2 \\ c_i^3 \\ c_i^4 \end{array} \right\}$$
(19)

4.3.2 Example: binary electrolyte

with

$$\tilde{M}_i^{mn} = -\vec{\nabla}U.\frac{\vec{\mathcal{W}}_i^{mn}}{5184}$$

4.4 Homogeneous reactions

4.4.1 Element contribution to fluctuation in node

• Monomolecular

Zero contribution (approximately).

 \bullet Bimolecular

Because of the symmetry it is the same contribution.

4.5 Poisson's equation

Zero contribution.

4.6 Time

Zero contribution.

5 Boundary element jacobian

5.1 Electrode reactions

5.1.1 Partial derivatives

$$\frac{\partial v}{\partial U} = -\frac{\alpha_{ox}nF}{RT}k_{ox}\exp\left[\frac{\alpha_{ox}nF}{RT}\left(V - U\right)\right]c_{red} - \frac{\alpha_{red}nF}{RT}k_{red}\exp\left[-\frac{\alpha_{red}nF}{RT}\left(V - U\right)\right]c_{ox}$$

$$\frac{\partial v}{\partial c_{ox}} = -k_{red}\exp\left[-\frac{\alpha_{red}nF}{RT}\left(V - U\right)\right]$$

$$\frac{\partial v}{\partial c_{red}} = k_{ox}\exp\left[\frac{\alpha_{ox}nF}{RT}\left(V - U\right)\right]$$

$$\left\{\begin{array}{c} \Delta c_{i}^{1} \\ \Delta c_{i}^{2} \\ \Delta c_{i}^{3} \end{array}\right\} = \frac{S}{108}\begin{bmatrix} 22\frac{\partial R_{i}^{1}}{\partial c_{i}^{1}} & 22\frac{\partial R_{i}^{1}}{\partial U^{1}} & 7\frac{\partial R_{i}^{2}}{\partial c_{i}^{2}} & 7\frac{\partial R_{i}^{2}}{\partial U^{2}} & 7\frac{\partial R_{i}^{3}}{\partial c_{i}^{3}} & 7\frac{\partial R_{i}^{3}}{\partial U^{3}} \\ 7\frac{\partial R_{i}^{1}}{\partial c_{i}^{1}} & 7\frac{\partial R_{i}^{1}}{\partial U^{1}} & 22\frac{\partial R_{i}^{2}}{\partial c_{i}^{2}} & 22\frac{\partial R_{i}^{2}}{\partial U^{2}} & 7\frac{\partial R_{i}^{3}}{\partial c_{i}^{3}} & 7\frac{\partial R_{i}^{3}}{\partial U^{3}} \\ 7\frac{\partial R_{i}^{1}}{\partial c_{i}^{1}} & 7\frac{\partial R_{i}^{1}}{\partial U^{1}} & 7\frac{\partial R_{i}^{2}}{\partial c_{j}^{2}} & 7\frac{\partial R_{i}^{2}}{\partial U^{2}} & 22\frac{\partial R_{i}^{3}}{\partial C_{i}^{3}} & 22\frac{\partial R_{i}^{3}}{\partial U^{3}} \\ 7\frac{\partial R_{i}^{1}}{\partial c_{i}^{1}} & 7\frac{\partial R_{i}^{1}}{\partial U^{1}} & 7\frac{\partial R_{i}^{2}}{\partial c_{j}^{2}} & 7\frac{\partial R_{i}^{2}}{\partial U^{2}} & 22\frac{\partial R_{i}^{3}}{\partial C_{j}^{3}} & 22\frac{\partial R_{i}^{3}}{\partial U^{3}} \end{bmatrix} \begin{cases} c_{i}^{1} \\ U^{1} \\ c_{j}^{2} \\ U^{2} \\ c_{j}^{3} \\ U^{3} \end{cases}$$

5.1.2 Example: binary electrolyte, $A \rightleftharpoons B + ne^-$

$$\frac{\partial v}{\partial U} = -\frac{\alpha_{ox}nF}{RT}k_{ox}\exp\left[\frac{\alpha_{ox}nF}{RT}\left(V - U\right)\right]c_B - \frac{\alpha_{red}nF}{RT}k_{red}\exp\left[-\frac{\alpha_{red}nF}{RT}\left(V - U\right)\right]c_A$$

$$\frac{\partial v}{\partial c_A} = -k_{red}\exp\left[-\frac{\alpha_{red}nF}{RT}\left(V - U\right)\right]$$

$$\frac{\partial v}{\partial c_B} = k_{ox}\exp\left[\frac{\alpha_{ox}nF}{RT}\left(V - U\right)\right]$$

with

$$\tilde{C}_i^m = \frac{S}{108} \frac{\partial v^m}{\partial c_i^m}$$

$$\tilde{U}^m = \frac{S}{108} \frac{\partial v^m}{\partial U^m}$$