

Spectroscopy

1. Objective: to perform two Experiments:

1.1 Basic: The objectives of this experiment are: (1) to study the emission of light from a hydrogen discharge source, (2) to learn the empirical formulas to characterize the pattern of spectral lines from hydrogen, (3) to learn the postulates for developing the Bohr model of the hydrogen atom, (4) to study and develop the Bohr theory of the hydrogen atom, (5) to measure the wavelengths of the Balmer series of visible emission lines from hydrogen, and (6) to learn to analyze the wavelength data to determine the Rydberg constant using the Bohr model formulation.

1.2 Advanced: To determine the line width of several spectral lines including sodium and Mercury.

2. Theory:

2.1 Grating Theory: The principle behind diffraction gratings is exactly the same as in Young's two-slit experiment. The main difference is that we now have thousands of slits, rather than just two. For a given wavelength λ , bright spots will be seen at points of constructive interference, where the path length difference is Δl for rays passing through any two adjacent slits is

$$d \sin \theta = m\lambda, m = 1, 2, 3, \dots (1)$$

as illustrated in Fig. 1. As shown in Fig. 2, the distance x from the central maximum is related to the angle θ by

$$\sin \theta = \frac{x}{\sqrt{x^2 + L^2}} \quad (2)$$

While it is possible to measure the wavelength of light with just two slits (in fact, you've already done it), diffraction gratings have two distinct advantages. First, the large number of slits allows the position of the maxima, and hence the wavelength, to be determined very accurately. Second, the fact that the slits are very close together means that light from different orders will be far apart.

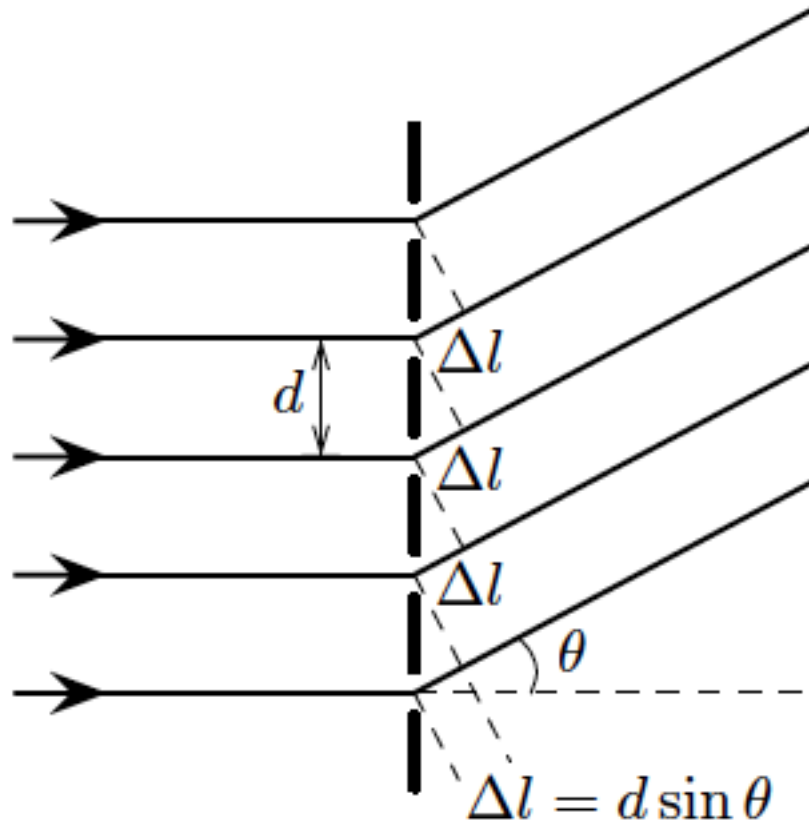


Figure: 1 A Diffraction Grating

Resolution of a Diffraction Grating: It will be useful to know the closest spacing between spectral lines that we expect our system to be able to resolve. That is: if two spectral lines with wavelength near λ are separated by $\Delta\lambda$, when will we be able to tell them apart? To determine the resolution of a diffraction grating, we must first know the angular distance $\Delta\theta$ by which it will split two spectral lines separated by wavelength $\Delta\lambda$.

We define the spreading of wavelengths produced by a grating, known as dispersion, by

$$D \equiv \frac{\Delta\theta}{\Delta\lambda} \quad (3)$$

Regarding θ in Eq. (1) as a function of λ and differentiating both sides gives

$$d \cos \theta \frac{d\theta}{d\lambda} = m \quad (4)$$

or, regarding the differentials as small differences,

$$D = \frac{d\theta}{d\lambda} = \frac{m}{d \cos \theta} \quad (5)$$

The resolving power of a grating is defined to be

$$R = \frac{\lambda}{\Delta\lambda} \quad (6)$$

where λ is the average wavelength of the two spectral lines being considered and $\Delta\lambda$ is the wavelength difference between them. We will say that they are resolved if the angular separation between the peaks is at least

$$\Delta\theta_R = \frac{\lambda}{Nd \cos \theta} \quad (7)$$

where N is the total number of slits the light passes through. (This requirement, called the Rayleigh criterion, essentially states that the peaks must be separated enough that there is a significant dip in intensity between them.) Substituting this $\Delta\theta_R$ for $\Delta\theta$ in Eq. (4), again taking differentials to be differences, we have

$$\frac{\lambda}{N\Delta\lambda} = m \quad (8)$$

$$R = \frac{\lambda}{\Delta\lambda} = Nm \quad (9)$$

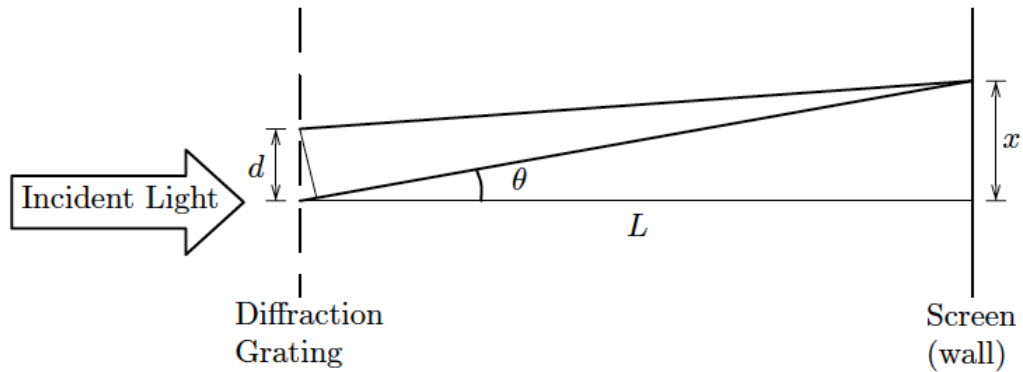


Figure 2: Light passing through the diffraction grating interferes on a screen a distance L away.

2.2 Hydrogen Atomic Spectra: Hydrogen atoms in a discharge lamp emit a series of lines in the visible part of the spectrum. This series is called the Balmer Series after the Swiss teacher Johann Balmer (1825-1898) who, in 1885, found by trial and error a formula to describe the wavelengths of these lines. This formula is given by

$$\frac{1}{\lambda} = R \left[\frac{1}{2^2} - \frac{1}{n^2} \right] \quad (1)$$

where n are integers, 3, 4, 5, ... up to infinity and R is a constant now called the Rydberg constant. In the SI system of units, $R = 1.097 \times 10^7 \text{ m}^{-1}$. Balmer suggested that his formula may be more general and could describe spectra from other elements. Then in 1889, Johannes Robert Rydberg found several series of spectra that would fit a more general relationship, similar to Balmer's empirical formula. This general relationship is known as the Rydberg formula and is given by

$$\frac{1}{\lambda} = R \left[\frac{1}{2^2} - \frac{1}{n^2} \right] \quad n_i > n_f \quad (2)$$

where n_i and n_f are integers, 1, 2, 3, 4, 5, ... up to infinity, with $n_i > n_f$. For the hydrogen atom, $n_i = 2$ corresponds to the Balmer series. There are other series in the hydrogen atom that have been measured. The Lyman series is a set of ultraviolet lines that fit the relationship with $n_i = 1$. A series in the infrared region of the spectrum is the Paschen series that corresponds to $n_i = 3$. The Brackett and Pfund series are two more in the infrared region corresponding to $n_i = 4$ and $n_i = 5$.

The individual lines in the Balmer series are given the names Alpha, Beta, Gamma, and Delta, and each corresponds to a n_i value of 3, 4, 5, and 6 respectively. Wavelengths of these lines are given in Table 1.

Table 1. Balmer Series – Some Wavelengths in the Visible Spectrum. Name of Line	n_f	n_i	Symbol	Wavelength
Balmer Alpha	2	3	H_α	656.28 nm
Balmer Beta	2	4	H_β	486.13 nm
Balmer Gamma	2	5	H_γ	434.05nm
Balmer Delta	2	6	H_δ	410.17 nm

In 1913 the Danish physicist Niels Bohr was the first to postulate a theory describing the line spectra observed in light emanating from a hydrogen discharge lamp. With Albert Einstein's theory for the photoelectric effect, where a photon has energy proportional to its frequency, Bohr postulated the existence of energy levels in the atom. He assumed that the energy associated with the photons of light were the result of transitions in the atom from one level to another, with the energy of the photon being equal to the difference in the internal energies specific to energy levels involved in the transition from an initial state to a final state, so that

$$\text{Energy of Photon} = - [\text{Energy of Final State} - \text{Energy of Initial State}] \quad (3)$$

or

$$\Delta E = [E_f - E_i] = [E_i - E_f] \quad (4)$$

The negative signs in front of the brackets in Equations (3) and (4) are because the photons are being emitted instead of being absorbed. Also, the levels are bound states and are assigned negative values. In other words, the atom is giving up energy instead of having energy supplied to it.

In 1905 Albert Einstein worked out the theory for the photoelectric effect using a concept that Max Planck had used to describe black body radiation. In this theory, for which Einstein received his Nobel prize, he postulated that light consists of packets of energy called photons or quanta and that each quantum of light has energy proportional to its frequency. He determined that the energy of a photon, E_{photon} is given by

$$E_{\text{photon}} = h\nu \quad (5)$$

where ν is the frequency of the photon and h is a constant of proportionality called Planck's constant. Planck's constant is equal to $6.6260755(40) \times 10^{-34}$ J·s. From the electromagnetic wave theory of light, frequency ν and wavelength λ are related to the speed of light by $c\nu/\lambda =$, so that

$$E_{\text{photon}} = h\nu = h \frac{c}{\lambda} \quad (6)$$

$$h \frac{c}{\lambda} = [E_i - E_f] = \Delta E \quad (7)$$

which relates the wavelength of the emitted light to the difference in energy levels between the final and initial states of the atom.

Bohr postulated that an electron can move about the nucleus of an atom in stable orbits, without emitting radiation and losing energy. Thus its energy would be constant in any single orbit and its energy would change only if it changed orbits and a transition occurred by the electron moving from one stable orbit to another. This postulate was revolutionary in that it contradicted electromagnetic theory, which predicted that the accelerating electron would radiate energy. In order to have stable orbits, Bohr further postulated that the magnitude of the orbiting electron's angular momentum would be quantized and that it must be an integral multiple of the quantity $h/2\pi$. Since the angular momentum L for an electron of mass m_e moving in an orbit of radius r with a speed v is

$$L = m_e vr \quad (8)$$

his postulate then yields

$$m_e v_n r_n = n \frac{h}{2\pi} \quad (9)$$

where $n=1,2,3 \dots$

where v_n is the speed of electron and r_n is its radius as it orbits with an integral number n of the quantity $h/2\pi$. The number n is referred to as the “Principal Quantum number.”

Except for the assumption that the angular momentum is quantized, Bohr’s model for the hydrogen atom was developed using simple classical concepts. He assumed that the electron was small in mass compared to the single proton in the nucleus and that it moved about this proton in a circular orbit. He said that the electron was held in an orbit by the electrical force F_E between the electron and proton, that is given by Coulomb’s Law,

$$F_E = \frac{1}{4\pi\epsilon_0} \frac{e^2}{r_n^2} \quad (10)$$

where e is the magnitude of charges of the electron and proton, r_n is the radius of the orbit, and ϵ_0 is a constant known as the permittivity of free space. The electrical force is the force that supplies the centripetal force, F_C , needed to keep the electron in orbit and is given by

$$F_C = \frac{m_e v_n^2}{r_n} \quad (11)$$

where m_e is the mass of the electron and v_n is the speed of the electron as moves in the circular orbit of radius r_n . Setting Equations (10) and (11) equal gives the relationship

$$\frac{1}{4\pi\epsilon_0} \frac{e^2}{r_n^2} = \frac{m_e v_n^2}{r_n} \quad (12)$$

By solving Equations (9) and (12) simultaneously for r_n and v_n the Bohr model finds expressions for the radii and speeds of the orbiting electron given by

$$r_e = \epsilon_0 \frac{n^2 h^2}{\pi m_e e^2} \quad (13)$$

and

$$V_n = \frac{1}{\epsilon_0} \frac{e^2}{2nh} \quad (14)$$

The minimum radius occurs when $n=1$ and is referred to as the *Bohr radius*, r_1 , and is given by

$$r_1 = \epsilon_0 \frac{h^2}{\pi m_e e^2} \quad (15)$$

In Bohr's model, the total energy E_n of the electron as it orbits the proton is the sum of the kinetic energy KE_n due to its motion and the potential energy PE_n it has because it is electrically bound to the proton so that

$$E_n = KE_n + PE_n \quad (16)$$

The kinetic energy is given by the classical equation for kinetic energy, $KE = mv^2$, and yields

$$KE_n = \frac{1}{2} m_e v_n^2 = \frac{1}{2} \frac{m_e e^2}{\epsilon_0^2 8n^2 h^2} \quad (17)$$

when the orbital speed found in Equation (14) is inserted. The electrical potential energy is given by

$$PE_n = -\frac{1}{4\pi\epsilon_0} \frac{e^2}{r_n} \quad (18)$$

which gives

$$PE_n = -\frac{1}{\epsilon_0^2} \frac{m_e e^4}{4n^2 h^2} \quad (19)$$

when the Bohr radius given by Equation (13) is substituted. The negative sign is a result of the electron being bound to the proton, and the potential is taken to be zero when the electron is infinitely removed from the proton. The total energy is then given by

$$E_n = -\frac{1}{\epsilon_0^2} \frac{m_e e^4}{8n^2 h^2} - \frac{1}{\epsilon_0^2} \frac{m_e e^4}{4n^2 h^2} \quad (20)$$

which reduces simply to

$$E_n = -\frac{1}{\epsilon_0^2} \frac{m_e e^4}{8n^2 h^2} \quad (21)$$

Since the principal quantum number n characterizes the orbit, a change of energy ΔE will occur when it undergoes a transition from an initial value of n_i to a final value n_f so that the energy changes from an initial value of E_i to a final value of E_f . Using Equation (21), ΔE is then given by

$$\Delta E = E_i - E_f = -\frac{1}{\epsilon_0^2} \frac{m_e e^4}{8n_i^2 h^2} + \frac{1}{\epsilon_0^2} \frac{m_e e^4}{8n_f^2 h^2} \quad (22)$$

which reduces to

$$\Delta E = \frac{1}{\epsilon_0^2} \frac{m_e e^4}{8h^2} \left[\frac{1}{n_f^2} - \frac{1}{n_i^2} \right] \quad (23)$$

Substituting Equation (7) gives

$$\frac{1}{\lambda} = \frac{1}{\epsilon_0^3} \frac{m_e e^4}{8h^3 c} \left[\frac{1}{n_f^2} - \frac{1}{n_i^2} \right] \quad (24)$$

and is recognized to be in the form of Equation (2),

$$\frac{1}{\lambda} = R \left[\frac{1}{n_f^2} - \frac{1}{n_i^2} \right] \quad (25)$$

with the Rydberg constant R given by

$$R = \frac{m_e e^4}{8\epsilon_0^2 h^3 c} \quad (26)$$

In this experiment, Equation (25) will be used to determine R from measurements of λ . For the visible, Balmer series n_i is 2 and values of n_f will be matched to the observed spectral pattern.

Advanced Laboratory:

2.3 Mass Ratio of the Deuteron and Proton from the Balmer Spectrum of Hydrogen

Students measure some of the spectral lines of the Balmer series of the D atom using an inexpensive deuterium lamp that contains residual hydrogen. This allows the students to “discover” that the relevant quantity is the reduced mass of the atom and to estimate the ratio of the masses of the two nuclei.

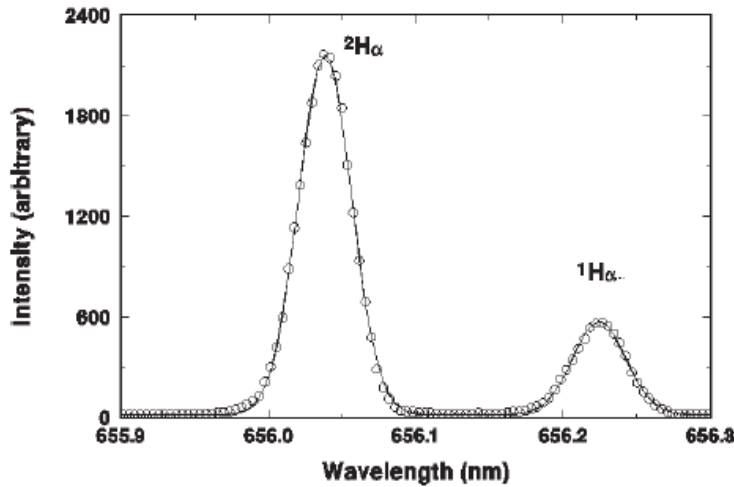


Figure 1: Emission spectrum of a deuterium spectrum tube in the Balmer α region. The open circles are the data points and the solid line is a best fit assuming a gaussian line shape. The fitted full width at half maximum value is $\gg 0.5 \text{ \AA}$.

The isotopic shifts of the lines in the Balmer series start at ca. 1.8 \AA for the α line and decrease with decreasing wavelength. The value of the nuclear mass ratio, m_D/m_H , can be obtained from equation 27.

$$\frac{m_D}{m_H} = \left[1 - \frac{m_H}{m_e} \left(1 + \frac{m_e}{m_H} \right) \cdot \frac{\Delta^{(n)}}{\lambda_H^{(n)}} \right] \quad (27)$$

$\Delta^{(n)}$ is the isotope shift of the nth line in the Balmer series, $\lambda_H^{(n)}$ is the wavelength of the nth line in the Balmer series of the H atom, m_H/m_e is the ratio of the rest masses of a proton and an electron, and m_D is the mass of the deuterium nucleus. Measurement of the isotopic shifts of several lines in the series results in better precision.

Table 1. Relevant Theoretical and Experimental Values^a

	μ_D/μ_H	m_D/m_H
Theoretical value	0.999728	1.9990
Value assuming no mass defect	1.00027	2.0014
Typical experimental values	0.999721 ± 0.000005	2.05 ± 0.12

^aThe Rydberg constant for the H atom has a value of $109737.3177 \text{ cm}^{-1}$ measured in vacuum. Theoretical values were calculated using data from refs 3 and 5.

The accuracy of the ratio of nuclear masses depends critically on the accuracy with which the separation between the two lines can be determined. Using the H line as an internal standard, we are able to measure this difference to less than 0.05 \AA . This corresponds to a precision of better than 10% for m_D/m_H .

2.4 Experimental study of line broadening in Plasmas for lightening:

Line broadening plays an important role in high- pressure, industrially important lighting plasmas. In this section of the experiment a joint experimental/theoretical study of line broadening in sodium related plasmas would be conducted. These studies not only impact our understanding of commercially important lighting sources but are also needed to model other plasma processes where radiative transport is important.

3. Appratus:

3.1 Appratus (Basic): The apparatus is shown in Figure 1 and consists of: (1) a Pasco Model SP-9268 precision student spectrometer, (2) a grating with either 300 lines/mm or 600 lines/mm, (3) a hydrogen discharge lamp and power supply, (4) a magnifying glass, (5) a small night light, and (6) a black cloth to block out stray light. The spectrometer is a precision instrument and is shown in Figure 2 with the important parts labeled.

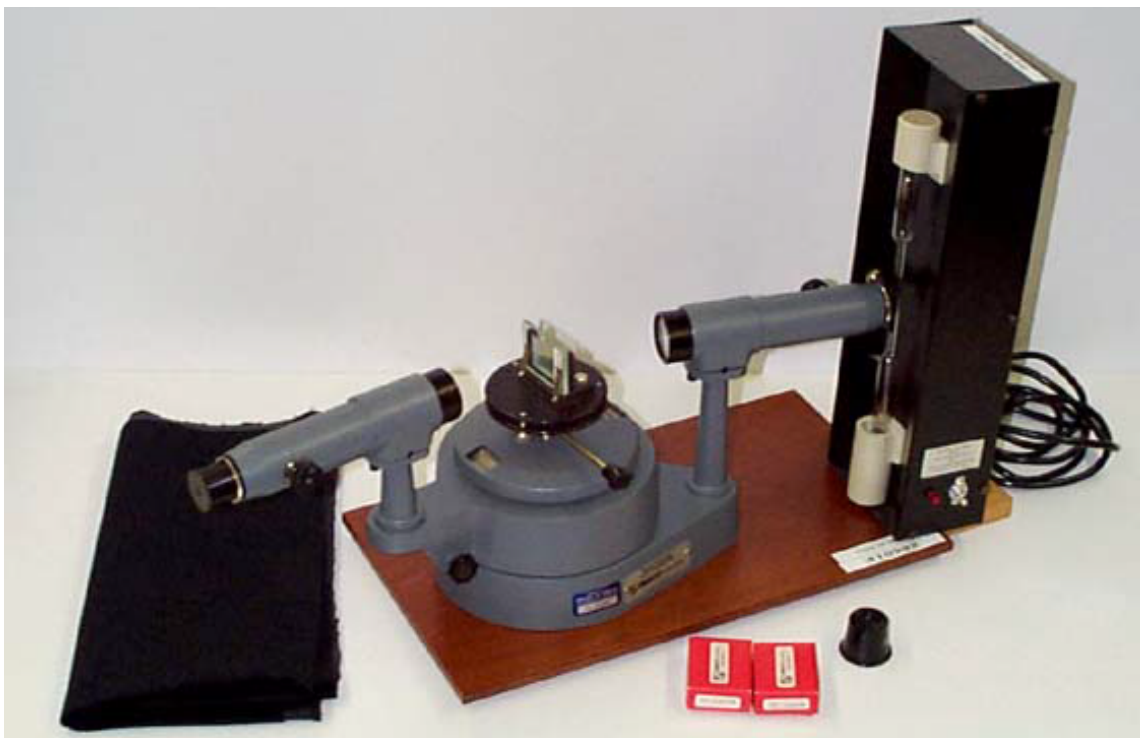


Figure 1. Apparatus and setup for the Balmer series experiment.

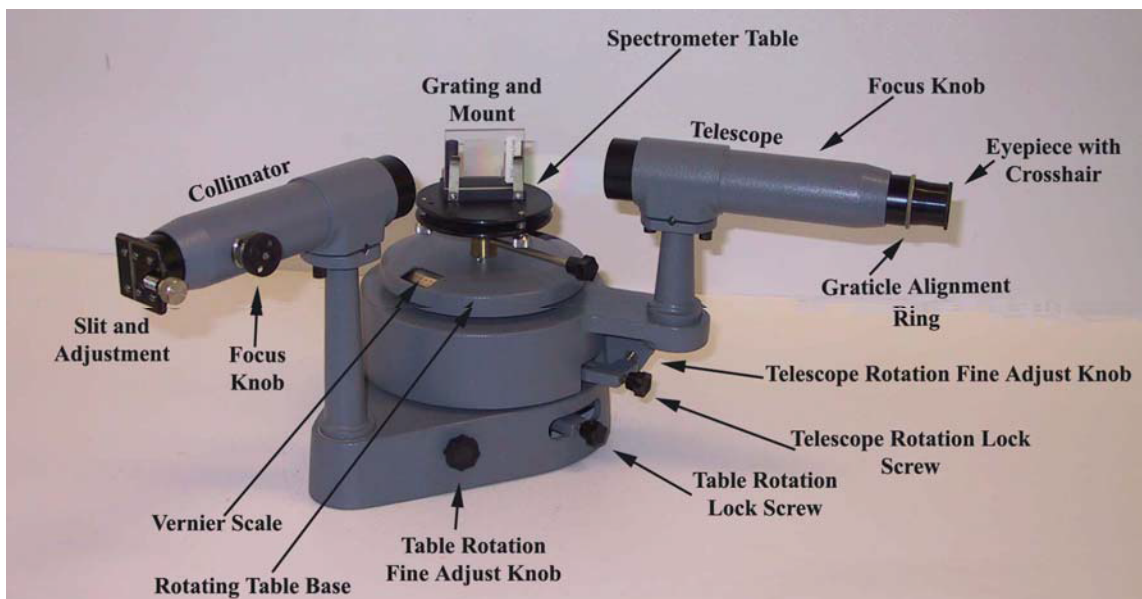


Figure 2. Pasco precision student spectrometer.

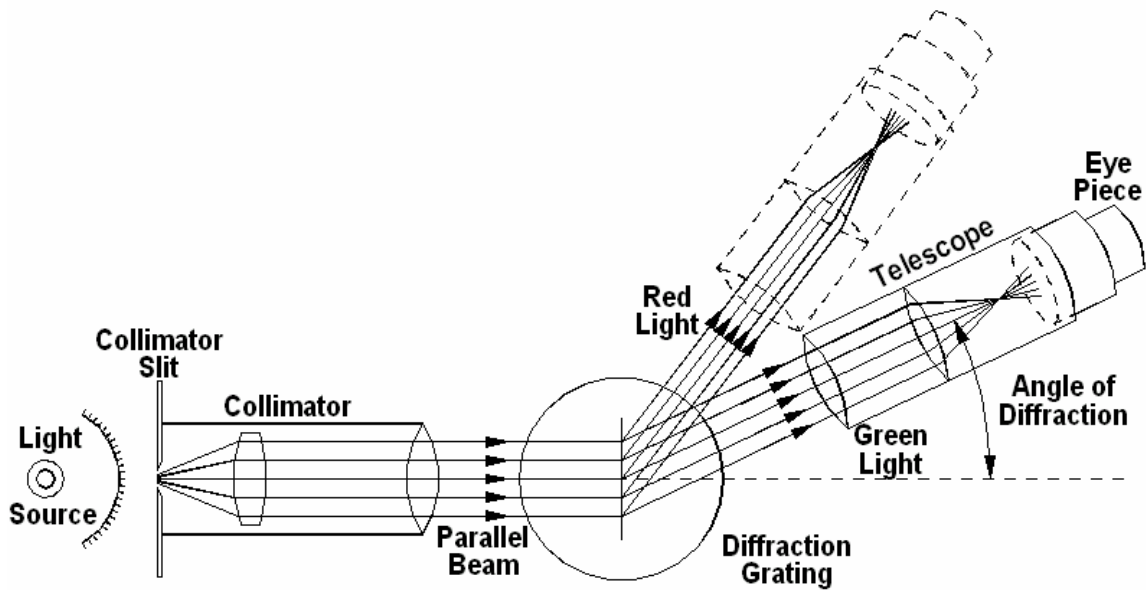


Figure 3. Schematic diagram of student spectrometer.

A schematic diagram of the student spectrometer is shown in Figure 3, which illustrates the principles of its operation. The student spectrometer consists of three basic components: a collimator, a diffraction grating, and a telescope. The light to be analyzed enters the spectrometer through an adjustable slit, which forms a narrow, well-defined line source of light. This slit is located at the focal point of the collimator, which transforms the light into a parallel, collimated beam of light. The beam uniformly illuminates the grating so that all light rays strike the grating at the same angle of incidence. A parallel beam of light is necessary to illuminate the grating so that a sharp image of the slit can be formed when viewed with the telescope.

The diffraction grating disperses the light so that the diffraction angle, θ , depends on the wavelength of light, λ , and the grating spacing, d , according to the grating equation,

$$d \sin \theta = m\lambda \quad (28)$$

where n is an integer given by

$$m = 0, 1, 2, 3, \text{-----} \quad (29)$$

The integer, m , is referred to as the diffraction order. Light of a given color will be diffracted at a specific angle for a given order. All light rays of that color will be parallel if the incident light rays are parallel. Light of a different color is diffracted at a different angle, but its light rays will also be parallel.

When the telescope is focused at infinity, parallel light of a specific color entering the telescope will form an image of the slit of that color in the focal plane of the objective lens of the telescope. Thus a green image of the slit will be formed at one angle and a red

image of the slit will be formed at another, larger angle. The telescope is attached to a rotating arm whose angle of rotation can be precisely and accurately measured. This device, often referred to as a goniometer, measures the diffraction angles of the spectra of colors of light. By rotating the telescope arm, images of the slit of various colors can be viewed and magnified with the eyepiece and their respective angles of diffraction measured. The eyepiece has a graticule with an orthogonal set of cross-hairs, which help align and reference the viewed images of the slit. From the grating spacing and measurements of the diffraction angles, the wavelength of the observed light can be determined from the grating equation, Equation (1).

3.2 Procedure (Basic)

1. The student spectrometer will provide the best and most accurate results if it is properly aligned and focused. The light source should be located about 1 cm from the collimator slit. A block of wood is provided so that the discharge light source can rest partially on the spectrometer platform. The slit should be examined and adjusted so that it is just barely open. Note that only one side of the slit moves while the other side remains fixed as adjustments are made.
2. Look through the eyepiece of the telescope and view the cross-hairs of the graticule. (This should be done preferably without your glasses if you wear them. The eyepiece is adjustable so that you can correct for your eyesight.) Slide the eyepiece in and out until the cross-hairs come into sharp focus. By rotating the graticule alignment ring, adjust the orientation of the cross-hairs so that they are aligned vertically and horizontally. You may need to refocus the cross-hairs.
3. The telescope should be focused for an object located at infinity. Loosen the telescope rotation lock-screw (see Figure 2) and rotate the telescope arm to view a distant object on a wall across the room, or a building outside. Adjust the telescope focus knob so that the viewed image of the distant object is clearly in focus.
4. Rotate the telescope arm so that it is directly across and aligned with the collimator. The telescope focus should remain unchanged, kept focused at infinity, during this step and at all times following.
5. Turn the light source on and view the slit of the collimator. The grating should not be in its mount. The telescope arm may need to be rotated slightly to see the red light and outline of the slit. The light source may need to be moved slightly to give the best, most intense illumination of the slit. You may loosen and re-tighten the telescope rotation lock-screw as needed to facilitate fixing and moving the arm position.

CAUTION: The hydrogen discharge lamp is powered by high voltage and the tube gets hot. DO NOT touch the tube anywhere especially near the ends where the electrical contacts are made.

Note: Simple hydrogen discharge tubes can lose their hydrogen by reacting with

impurities inside the tube and by small leaks from the outside. A good tube will have a bright red color near its center, and a poor tube will be more pinkish near the center. For best results, it is important to have a good tube.

6. Use the focus knob to adjust the collimator focus for the clearest, sharpest image of the slit. If needed, re-adjust the light source to provide the best illumination of the slit and to give the sharpest images.
7. Gratings with either 300 lines/mm or 600 lines/mm are provided for this experiment, and the grating spacings d are either $1/300000$ m or $1/600000$ m. The number of lines per mm should be labeled on the grating, but if not, ask your instructor for this value. The grating is a replica grating approximately 1-inch square that is mounted in the center, on one side of a 1."x2"x5/32" glass plate. The grating itself should not be touched and the glass mount should be handled only by its edges. The diffraction grating should now be placed in the grating mount with the grating side of the glass mount against the vertical posts. Loosen the spectrometer table lock-screw and align the plane of the grating so that it is perpendicular to the optical axis formed by the collimator and telescope when they are directly opposite one another. There is a line inscribed on the spectrometer table to assist with the orientation. This adjustment doesn't have to be precise, but in your best judgment, it should appear to be perpendicular.
8. The image you see with the telescope directly in line with the collimator is the undiffracted, central image. Now with the black cloth over the apparatus to shield out stray light, rotate the telescope arm to the left and right to survey the different lines and colors of diffracted light. Moving away from the central image, you should see the first order diffraction lines of violet, turquoise, and red light followed by the second order diffraction pattern of lines of the same colors. These visible lines are three of the Balmer lines corresponding to $n_f = 5, 4$, and 3 . There is a fourth one corresponding to $n_f = 6$, but it is weak, and sometimes so extremely weak that it cannot be seen. The room may have to be completely dark to see this line, if at all possible.

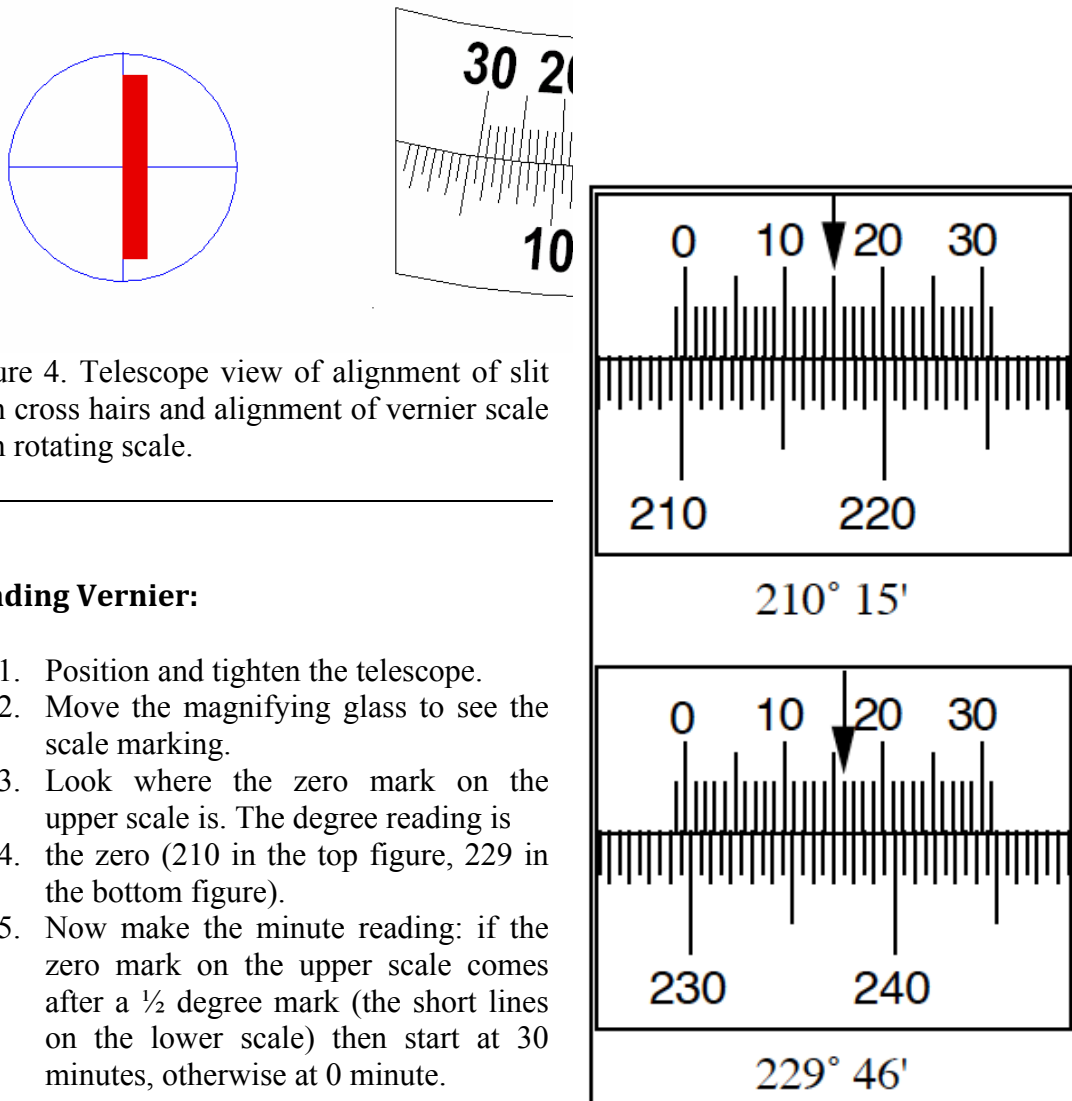
Note: There are bands of very weak transitions due to molecular hydrogen that can sometimes be observed between the blue, green, and red lines. The emissions from these bands grow stronger as a tube ages, and at the same time, the hydrogen lines grow weaker.

9. The following steps involve aligning and setting the reference value for making the precise angular measurements of the diffracted light. It is important to do these steps carefully, since all measurements will be made relative to this setting.
10. Rotate the telescope arm back to the position where it is directly opposite and in line with the optical axis of the collimator and where the central, undiffracted image can be observed. Adjust the arm so that the vertical cross-hair is near or on the central image and tighten the telescope rotation lock-screw. Use the fine adjust knob near and

just to the right of the lock-screw to align the vertical cross-hair exactly with the left, stationary slit, the left edge of the central line image. See Figure 2.

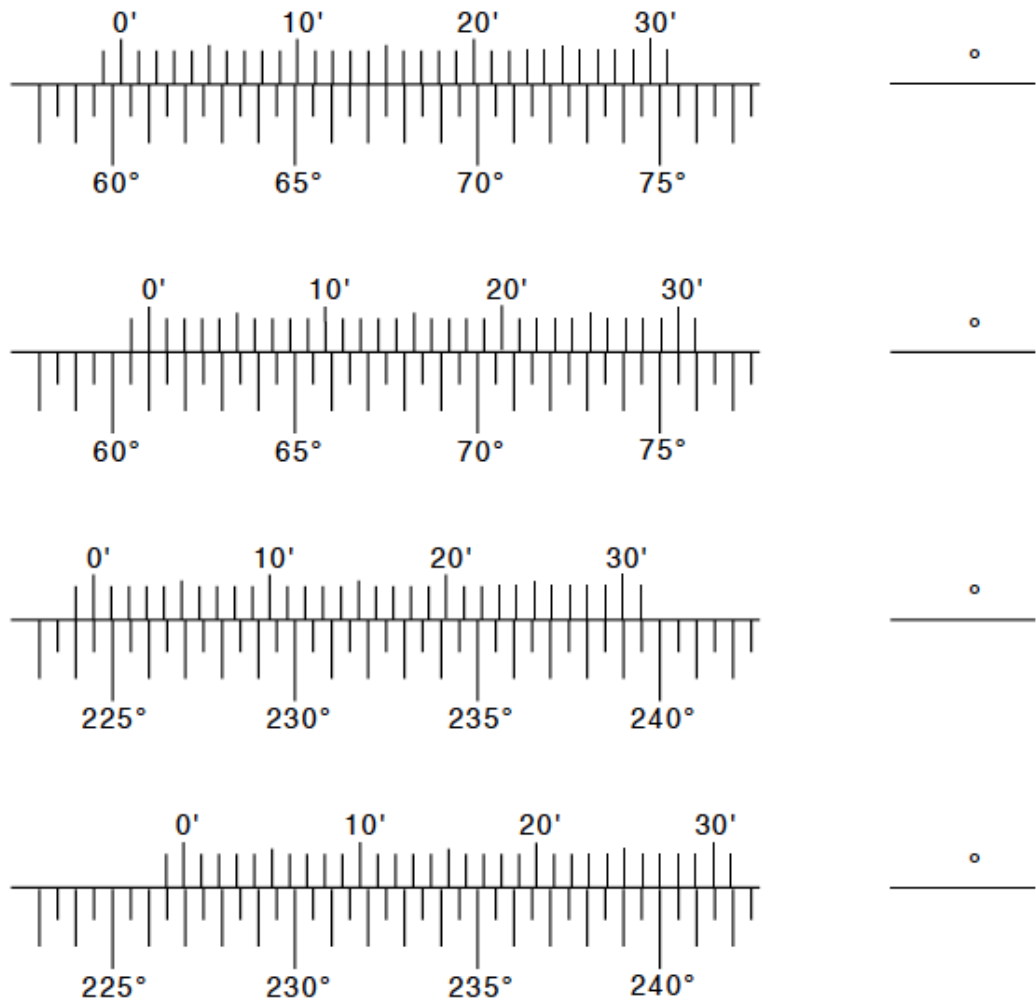
Note the exact position of the vertical cross-hair and left edge as this will be the relative position of these for all subsequent alignments.

11. Angles are measured with the 0-360° scale engraved into the rotating telescope base and are read relative to the zero fiducial ruling on the vernier scale that is fixed to the rotating table base. Set the zero reference angle by loosening the table rotation lock-screw and rotating the table base until the zero fiducial mark on the vernier scale is closely aligned with the zero mark on the 0-360° main scale. Tighten the lock-screw and use the table rotation fine adjust knob to position the two zero marks so that they are exactly aligned with each other. Use the magnifying glass to better make this adjustment. As a check, the 30' ruling on the vernier scale also should be aligned exactly with the 14.5° ruling on the main scale. See Figure 4.



6. Next look for the two lines which matchup the best between the two scales. Read the number of the appropriate line on the upper scale and add it to the number of minutes you started with. In the figure the zero mark is just past the $\frac{1}{2}$ degree mark so we begin with 30 minute. The 15 minute mark is the one that lines up best so we get 210 degree 15 minutes for the top reading, and 229 degree 46 minute for the bottom reading.
7. Now convert the decimal degree, in this case $15/60 = 0.25$, so $\theta = 2210.25$ degree.
8. You might try the following examples.

Vernier Problems:



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12. At this point the following conditions should exist: (1) the telescope should be aligned with the optical axis of the collimator, (2) the vertical cross-hair in the eyepiece should be aligned with the left edge of the central image, (3) the zero marks of the vernier and main scales should be aligned exactly, (4) the plane of the grating should be perpendicular to the optical axis, and (5) all adjustments should be fixed with their respective lock-screws. The spectrometer should be ready for

measurements.

13. The general procedure for measuring the angles of the various diffraction lines is to un-tighten the telescope arm rotation lock-screw, rotate the telescope to view the diffraction line of interest, center the image of the line near the vertical cross-hair, tighten the lock-screw, use the telescope rotation fine adjust knob to align the vertical cross-hair with the left edge of the slit image, and then read the angular scale.
14. The angles for each of the diffracted lines can be read most simply by avoiding the vernier scale feature and estimating the angles to the nearest 0.1° on the scale on the rotating table base. The precision and accuracy provided by this estimate will be more than adequate for obtaining excellent results in this experiment. The angle can be read more accurately by using the magnifier. The night-light will be needed in the darkened room.
15. The zero on the vernier scale is the reference mark used to indicate the exact angular position on the main scale. Remember that the smallest division on the main scale is 0.5° and that the estimate only has to involve 5 equal imaginary increments of 0.1° in each of those.
16. Open up an Excel spreadsheet and label the columns and rows as shown in Figure 5.

	A	B	C	D	E	F	G	H	I	J	K	L
1		Grating Line Spacing			d=							
2												
3	Color	Color	n_i	m	$\theta_R(^{\circ})$ Reading	$\theta_L(^{\circ})$ Reading	$\theta_L(^{\circ})$ Corrected	$\theta_R(\text{rad})$	$\theta_L(\text{rad})$	λ	$1/\lambda$	$(1/4 - 1/n_i^2)$
4	v		6	1								
5	bv		5	1								
6	bg		4	1								
7	r		3	1								
8	v		6	2								
9	bv		5	2								
10	bg		4	2								
11	r		3	2								
12												

Figure 5. Example of Excel spreadsheet for hydrogen Balmer series experiment.

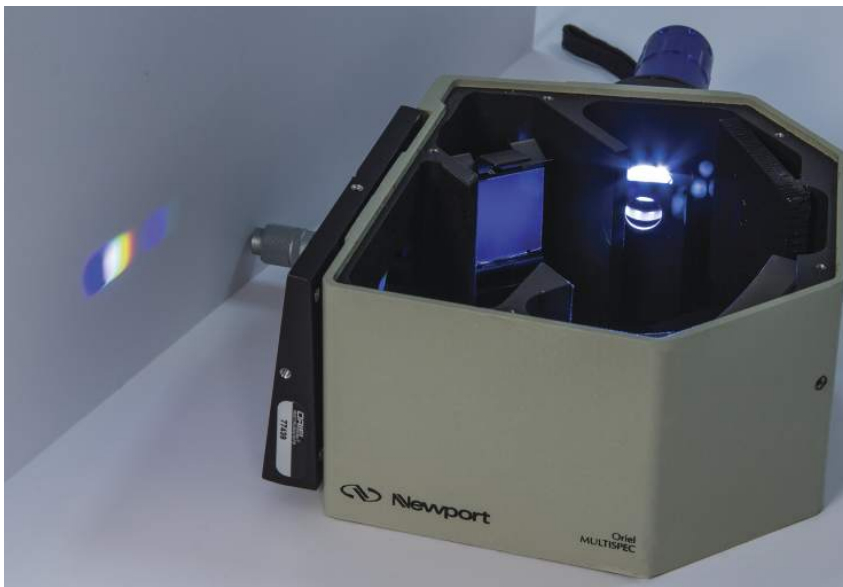
17. In your spreadsheet, measure and record the readings on the main scale for the angles θ_R and θ_L in degrees. This should be done for $n_i = 5, 4$, and 3 , both right and left diffracted Balmer lines, and for both the first and second order patterns, $m = 1$ and $m = 2$. If possible the measurement should also include the $n_i = 6$ line. The readings for θ_R should be in the range of $0-30^\circ$ and the readings for θ_L should be in the range $330-360^\circ$. Of course the values of θ_L should be in the range of $0-30^\circ$, but will be found later in the analysis of the data by subtracting the reading value from the initial value of 360° .

3.3 Analysis (Basic): (You can use Excel or any other software you like. Here guide for Excel is given.

1. In the extra column for the corrected value for the angle θ_L , subtract the reading for θ_L from 360° . These values will be in the range from 0 to 60° , and should be nearly the same values as θ_R . Examine your values of θ_L and θ_R to see if they are approximately the same on both sides. This will allow you to check for any mistakes in the angle measurements. The pairs of values should be within about 0.5° of each other, You should repeat the measurement if there appears to be a mistake.
2. In the columns for θ_L (rad) and θ_R (rad) convert the angles in degrees to radian units. Excel requires the angles for trigonometry functions to be in radians. The conversion factor is $180^\circ = \pi$ radians, therefore multiply the angle in degrees times the value for π and then divide by 180. In **Excel**, the value of π can be obtained with the *Math & Trig* function, *PI ()* with no argument.
3. Use Equation (27) to calculate λ from each of the values for θ_L (rad) and θ_R (rad) and record the values in the column labeled for these values. The value for the sine of the angle in cell H4 is written as $=\text{SIN}(H4)$. Remember that the argument must be in radians and placed inside parentheses.
4. In the columns for $1/\lambda$ and $(1/4 - 1/n_i^2)$, compute these values.
5. Equation (1) suggests that a plot of $1/\lambda$ versus $(1/4 - 1/n_i^2)$ should give a straight line passing through zero with a slope of R . Make a graph of $1/\lambda$ versus $(1/4 - 1/n_i^2)$ and label the axes with proper titles and units.
6. Use the **Add Trendline** feature of Excel to find the slope of the straight line that best fits your data. Force the intercept to be zero.
7. Compare your value for the slope and R with the accepted value of the Rydberg constant. Calculate the percent difference and enter your result in your spreadsheet.
$$\% \text{ Difference} = [|\text{Accepted Value} - \text{Measured Value}| / \text{Accepted Value}] 100\% \quad (29)$$
8. Use your measured value for the Rydberg constant to calculate the wavelengths of the observed emission lines. Compare these values with your measurements of the wavelength and calculate the percent difference using Equation (29).
9. Print your graph and spreadsheet of data.

3.4 Apparatus (Advanced): The apparatus for these experiments consists of several light sources and the spectrograph, which consists of a rotatable grating and two mirrors for directing the light, a solid-state detector, and a computer for recording the data. Here is a

picture of the spectrograph made by Oriel Instruments, MS125J that has been created specifically to be compatible with diode array detectors, and has a flat focal field for use with diode arrays up to 1 inch (25 mm) long. A calibrated micrometer drive for changing wavelength and a wide variety of gratings allow great versatility in the selection of wavelength ranges. Optional fiber optic light guides or camera 35 mm focusing lenses allow you to take instantaneous spectra almost anywhere. The optical arrangement of MS125J was optimized by computer in order to squeeze the highest performance from the smallest possible instrument. This instrument has a low F/number of 3.7, a focal length of 120 mm, variable wavelength ranges, interchangeable gratings, minimal coma, and a flat focal field, all in a neat compact unit which you can hold on your hand!



Spectrograph Specifications

Design: Crossed Czerny-Turner
F/number: 3.7

Focal length: 120 mm
Stray light: 0 - 0.1%
Grating type: Plane, (ruled or holographic)
Grating mount: Clip, Kinematic
Drive: Micrometer, calibrated
Wavelength accuracy: 1 nm over full range (with 1200 l/mm grating)
Wavelength reproducibility: 1 nm with standard micrometer
0.1 nm with digital micrometer (with 1200 l/mm grating)
Focal Plane: 15 mm from, and parallel to face
Optical axis: height 1.78 inch, (45.2 mm)
Dimensions: approx. 6 x 6 x 3 inch (150 x 150 x 80 mm)
Weight: 3.3 lb, (1.5 kg)

3.5 Procedure (Advanced):

3.5.1 Setup: First, read the Spectra-Array software user manual and MS 125 1/8m spectrograph model 77400 documents. Familiarize yourself with the instrument and the LineSpec software. Second, it is important that all of the mechanical connections of the instrument be carefully and securely made so that the parts do not wobble out of alignment. If they seem loose, ask the lab assistant to fix them.

3.5.2 Calibration: Although the exact procedure may vary from instrument to instrument, the calibration process generally involves using the instrument to test samples of one or more known values called calibrants. The calibrants used in these experiments are lines of very well-known wavelengths from a mercury discharge tube. These lines fall onto pixels of the CCD detector. Calibration is essentially the assignment of pixels to known wavelength values. Mercury has a distinctive yellow doublet between approximately 575 and 580 nm and a strong single green line and single violet line as shown below.

Yellow doublet at 578.97 nm and 576.96 nm
Green line 546.074 nm
Violet line 435.833 nm

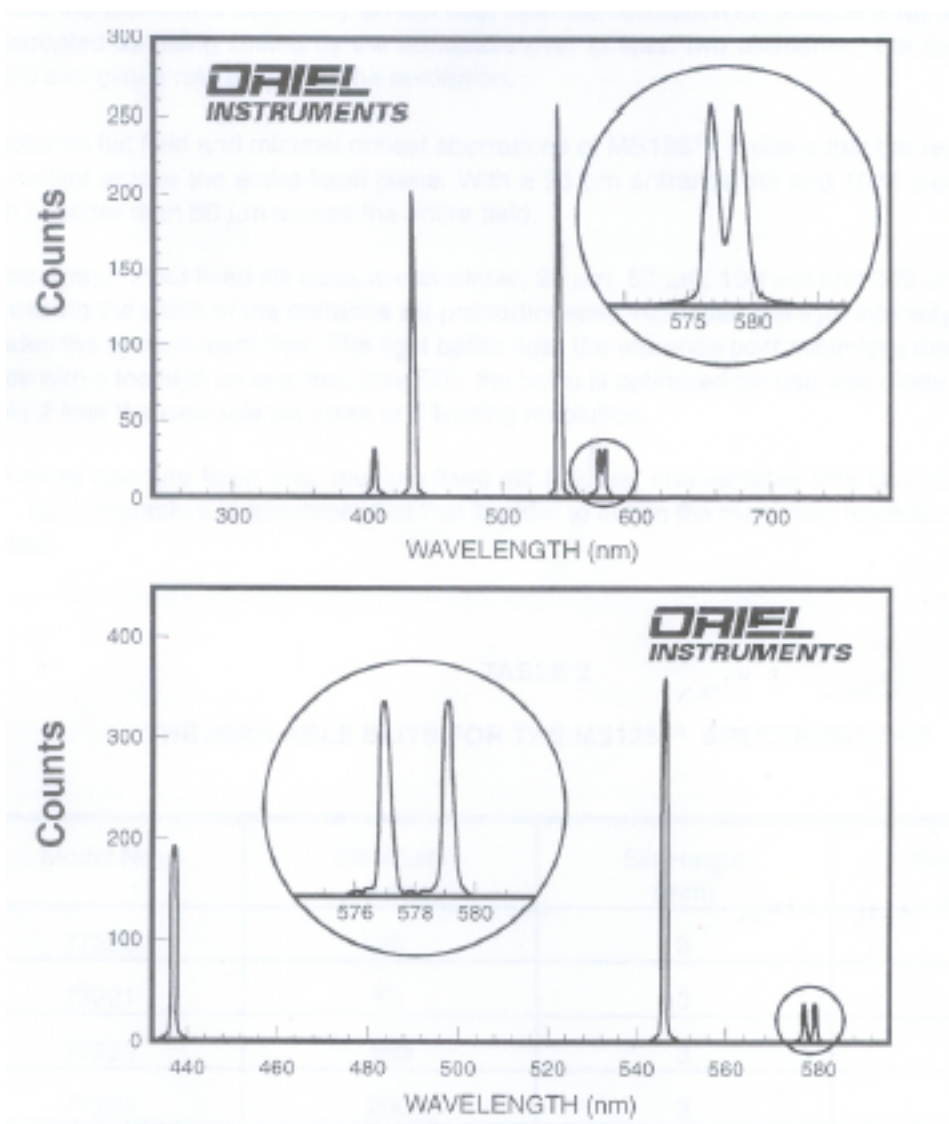


Figure 4. Location of lines of mercury used for calibration.

The first step is to set the grating of the spectrometer to 546 nm. When the micrometer is set to 4 mm the grating is approximately centered at 400 nm; 5 mm corresponds to 500 nm and so forth. So 5.46 on the micrometer will move the grating so that it is approximately centered at 546 nm. Now you must be sure that all of the four lines needed for calibration show in the sample window. Select the Spectrum item from the Mode menu. Check the Sample box. Click the button in the upper left of the display with the circular arrow icon. If you see the four lines shown in the lower part of Fig. 4, no more adjustment is needed. If you do not see all four lines adjust the micrometer until the doublet is near the right edge of the sample window and the 435 nm line is near the left edge. Clear the spectrum by using the drop down menu File and select Delete Spectrum.

Now select the third button from the left (Scan with Averaging) near the top of the of the LineSpec window. Enter 100 in the number of scans box and choose OK. The spectrum will appear in the Sample window of the Dump window. You may narrow the area

around a peak by using the mouse to draw a selection box around it.

- **To determine the calibration coefficients using a spectral calibration lamp:**

1. Record the emission spectrum of a spectral calibration lamp for the appropriate detector (Master, Slave 1, Slave 2 or Slave 3. We only have a master.). An example of an emission spectrum corresponding to a spectral calibration lamp is shown in Figure 5.

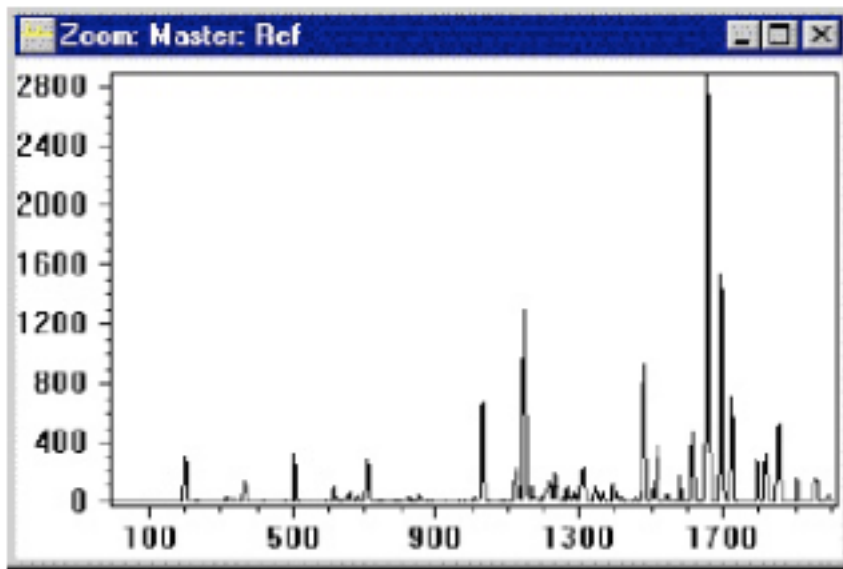


Figure: 5

2. Select Wavelength Calibration from the Setup pull-down menu to display the Wavelength calibration control window. Figure 6.

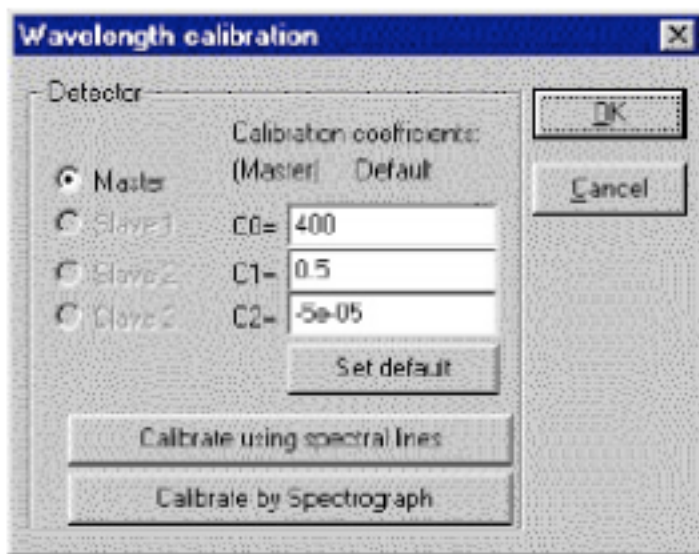


Figure: 6

3. Click the Calibrate using spectral lines button to display the Find calibration for a Master detector control window as shown in Figure 7.



Figure: 7

4. Find the position of a known spectral line using the mouse pointer. Note the position (pixel units) of the spectral line displayed in the status bar (see Figure 8 below):

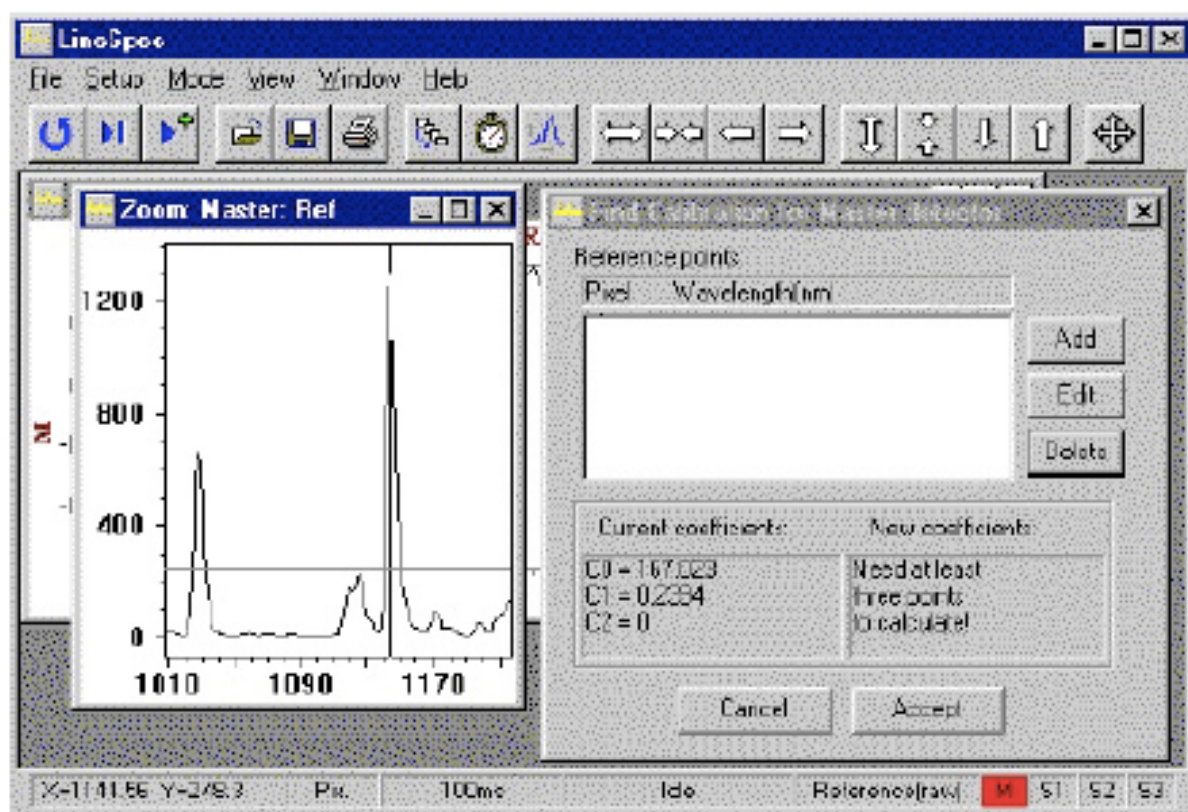


Figure: 8

5. Click the Add button in the Find calibration control window to display the following window:

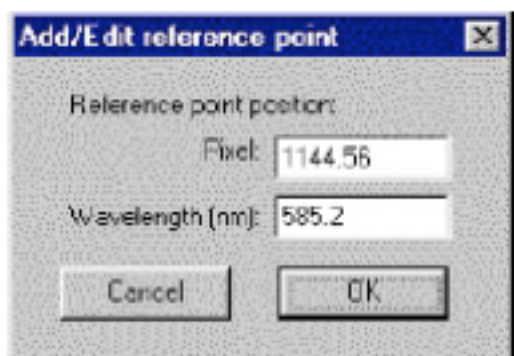


Figure: 9

Enter the pixel position of the spectral line and the known position in nanometer units.

6. Repeat steps 4 to 5 identifying the four spectral lines that span the detection region of interest.



Figure: 10

7. Once the features of interest have been identified and assigned to known spectral lines, the program automatically calculates the wavelength calibration coefficients as shown in Figure 10:

The Edit button allows the user to modify any reference points within the table, while the Delete button allows points to be removed.

8. Click the Accept button to automatically update the wavelength calibration coefficients or Cancel to retain the original settings.

9. Confirm the accuracy of your calibration by testing it against a helium light source. Clear the calibration spectrum, record a spectrum from the helium lamp, and view the lines in nanometers. If your helium lines are all within one nanometer of the established values shown in Appendix A, you may proceed with the rest of the experiments.

3.5.3 Data Recording

Take several measurements of each spectrum and use the mean as your result.

- Observe the spectra from the incandescent and the fluorescent (overhead) light sources using the hand held spectroscope. Describe these spectra qualitatively.
- Measure and record spectral lines from the florescent and incandescent light sources using the Oriel spectrograph.
- Measure and record spectral lines from the white, red, and blue LEDs.
- Measure and record spectra from the H, 2H (deuterium), and Na.

3.6 Analysis (Advanced):

When writing your report (2 pages) include all instrument parameters such as the grating constant, slit width, resolution, focal lengths, etc. Compare all of your results to currently

accepted standard values and do error analyses.

- Determine Rydberg's constant for hydrogen and deuterium.
- Use your data to compute Planck's constant.
- Find the ratio of the mass of deuterium to hydrogen.
- Determine the ground state energy of hydrogen by using the Bohr model and the measured wavelengths of the lines in the Balmer series.
- Fit a curve to the observed spectrum of Hydrogen, Deuterium and Sodium lamps. determine the width of the line.

4. Questions:

Spectral Line:

1. Show that the SI units for Planck's constant h are equal to the SI units of angular momentum.
2. Calculate the magnitude of the Bohr radius.
3. Calculate the speed of the electron moving in an orbit whose radius is equivalent to the Bohr radius.
4. Using your value for the Rydberg constant, calculate the wavelength of an ultraviolet transition in the Lyman series from the $n = 2$ level and to the $n = 1$ level.
5. Using your value for the Rydberg constant, calculate the wavelength of an infrared transition in the Paschen series from the $n = 4$ level and to the $n = 3$ level.
6. How much energy does it take to ionize a hydrogen atom in its ground state? That is how much energy has to be supplied to move an electron from the $n = 1$ level to the $n = \infty$ level?
7. In the analysis of your results, suppose you had plotted $1/\lambda$ versus $1/n_i^2$ instead of $(1/4 - 1/n_i^2)$. Would your data give a straight line and would its slope give the same results for R ? How would the intercept change?
8. What is the energy of a red photon of wavelength 600 nm? Of a blue photon of wavelength 400 nm? Express your answers in both Joules and electron-volts. (1 eV = 1.60×10^{-19} J)
9. How is the wavelength of the light to be measured in this experiment? Describe each of the terms in Equation (27).
10. Of the following three transitions: from $n = 2$ to $n = 1$, from $n = 3$ to $n = 2$, from $n = 4$ to $n = 3$, which produces photons with the shortest wavelength? The longest? Which is in the visible part of the spectrum? Which is in the ultraviolet? Which is in the infrared?

11. Use your measured value for the Rydberg constant to calculate the wavelengths of the observed emission lines. Compare these values with your measurements of the wavelength and calculate the percent difference using Equation (29). Print your graph and spreadsheet of data.

13. Construct as complete an energy level diagram of atomic hydrogen as you can, and show the transitions that give rise to the Balmer lines. (Incidentally, just exactly what is a spectral “line”?)

14. Define the following terms and, where applicable, calculate them for the two optical systems used in this experiment: {it grating equation, diffraction orders, angular dispersion, linear dispersion, resolving power, spectral resolution, bandpass, focal length, $f/\#$, and free spectral range}. (See [1, 2])

Spectrograph:

15. Suppose perfectly monochromatic light of wavelength 500 nm enters a reflection grating spectrograph (see description of the instrument) with the slit width set at 10.0 μm , and suppose the beam fully illuminates the concave grating which is a 3cm x 3cm square with 1200 lines/mm. Make a plot with properly scaled axes of the light intensity versus angular displacement in the focal plane of the spectrograph in the first order diffracted image of the slit. Show quantitatively the salient features of the multiple slit diffraction pattern. (see [2])

17. Throughout this experiment the tabulated wavelengths of the mercury spectrum will be used as the calibration reference. Suppose, however, you had to start from scratch with no reference wavelengths. How would you establish an absolute scale of wavelengths?

18. Predict the isotope shift in wavelength (i.e., the difference in wavelength $\Delta\lambda = \lambda_H - \lambda_D$ of the first 3 Balmer lines of hydrogen and deuterium. (see [3])

19. Visible light is incident perpendicularly on a grating with 315 rulings/mm. What is the longest wavelength that can be seen in the fifth-order diffraction?

20. A grating has 400 lines/mm. How many orders of the entire visible spectrum (400-700 nm) can it produce in a diffraction experiment, in addition to the $m = 0$ order?

21. A diffraction grating 20.0 mm wide has 600 rulings. Light of wavelength 589 nm is incident perpendicularly on the grating. What are the (a) largest, (b) second largest, (c) third largest value of θ at which maxima appear on a distant viewing screen.

22. A diffraction grating having 180 lines/mm is illuminated with a light signal containing only two wavelengths $\lambda_1 = 400\text{nm}$ and $\lambda_2 = 500\text{nm}$. The signal is incident perpendicular on the grating. (a) What is the angular separation between the second order maxima of these two wavelengths? (b) What is the smallest angle at which two of the

resulting maxima are superposed? (c) What is the highest order for which maxima for both wavelengths are present in the diffraction pattern?

33. A diffraction grating is made up of slits of width 300 nm with separation 900 nm. The grating illuminated by monochromatic plane waves of wavelength $\lambda = 600$ nm at normal incident. (a) How many maxima are there in the full diffraction pattern? (b) What is the angular width of a spectral line observed in the first order if the grating has 1000 slits?

24. Light of wavelength 600 nm is incident normally on a diffraction grating. Two adjacent maxima occur at angles given by $\sin\theta = 0.2$ and $\sin\theta = 0.3$. The fourth order maxima are missing. (a) What is separation between adjacent slits? (b) What is the smallest slit width this grating can have? For that slit width what are the (c) largest, (b) second largest, and (e) third largest values of the order number m of the maxima produced by the grating?

25. With light from a gaseous discharge tube incident normally on a grating with slit separation $1.73 \mu\text{m}$, sharp maximum of green light are experimentally found at $\theta = \pm 17.6^\circ, 37.3^\circ, -37.1^\circ$, and -65.0° . Compute the wavelength of the green light that best fits these data.

26. Assume that the limits of the visible spectrum are arbitrarily chosen at 430 and 680 nm. Calculate the number of rulings per millimeter of a grating that will spread the first order spectrum through an angle of 20.0° .

27. A beam of light consisting of wavelengths from 460 nm to 640 nm is directed perpendicularly onto the diffraction grating with 160 lines/mm. (a) What is the lowest order that is overlapped by another order? (b) What is the highest order for which the complete wavelength range of beam is present? In that highest order, at what angle does the light at wavelength (c) 460.0 nm and (d) 640.0 nm appear? (e) What is the greatest angle at which the light at wavelength 560.0 nm appear? Show that in a spectrum of white light obtained with a grating, the red ($\lambda_r = 700$ nm) of the second order overlaps the violet ($\lambda_v = 400$ nm) of the third order.

28. The D line in the spectrum of sodium is a doublet with wavelengths 589.0 nm and 589.6 nm. Calculate the minimum number of lines needed in a grating that will resolve this doublet in the second order spectrum.

APPENDIX 1

National Institute of Standards and Technology Physics
Laboratory
Basic Atomic Spectroscopy Data
Mercury (Hg)

Strong Lines of Mercury (Hg)

Intensity	Air Wavelength (Å)	Spectrum	Reference
1000 P ,c	3983.931	Hg II	SR01
400 P	4046.563	Hg I	BAL50
60	4339.223	Hg I	BAL50
100	4347.494	Hg I	BAL50
1000 P	4358.328	Hg I	BAL50
12 c	5128.442	Hg II	SR01
15	5204.768	Hg II	SR01
80 P	5425.253	Hg II	SR01
500 P	5460.735	Hg I	BAL50
200 P	5677.105	Hg II	SR01
50	5769.598	Hg I	BAL50
60	5790.663	Hg I	BAL50
12	5871.279	Hg II	SR01
20 c	5888.939	Hg II	SR01
15	6146.435	Hg II	SR01
250 P ,c	6149.475	Hg II	SR01
25	7081.90	Hg I	F54

Helium wavelengths (nm)

438.793	w
443.755	w
447.148	s blue-violet
471.314	m blue-green
492.193	m green
501.567	s green
504.774	w
587.562	s yellow
667.815	m red

s=strong, m=med, w=weak