

1.1 exercise 1 - atomic symbols

Deduce the number of protons, neutrons and electrons in the following species:

1. ${}^1_1\text{H}$
2. ${}^{17}_8\text{O}$
3. ${}^4_2\text{He}^{2+}$
4. ${}^{132}_{54}\text{Xe}$
5. ${}^{27}_{13}\text{Al}^{3+}$
6. ${}^{235}_{92}\text{U}$
7. ${}^1_1\text{H}^+$
8. ${}^{45}_{21}\text{Sc}^{3+}$
9. ${}^{37}_{17}\text{Cl}^-$
10. ${}^{14}_6\text{C}$

Use the periodic table to write symbols for the following species:

11. 19 protons, 20 neutrons, 18 electrons
12. 8 protons, 8 neutrons, 10 electrons
13. 1 proton, 2 neutrons, 1 electron
14. 82 protons, 126 neutrons, 80 electrons
15. 53 protons, 74 neutrons, 54 electrons