

1.4 EXERCISE 1 – trends in Period 3

1. State and explain the following trends:
 - a) atomic size across Period 3
 - b) first ionisation energy across Period 3
 - c) electronegativity across Period 3
 - d) melting point from sodium to aluminium
 - e) electrical conductivity from sodium to aluminium

2. Explain the following:
 - a) the melting point of silicon is the highest in Period 3
 - b) the melting point of phosphorus is lower than silicon
 - c) the melting point of phosphorus is lower than that of sulphur
 - d) the boiling point of chlorine is greater than that of argon, but lower than that of sulphur
 - e) neither silicon nor sulphur conduct electricity