

Chemical Reaction

Full marks: 45 Time :1 hr 15 mins

1.	Fill in the Blanks		1×10=10
	a.	CaO + H ₂ O+ Heat	
		Sn0 + 2 Na0H + H ₂ 0	
	C.	Zn +CuSO ₄ + Cu	
		$V_2O_5 + SO_2 + SO_3$	
		+ 3HCl 2AlCl ₃ +3 H ₂ O	
		ZnSO ₄ + 2NH ₄ OH Zn(OH) ₂ +	
		Sn0 + SnCl ₂ + H ₂ 0	
	_	Cl ₂ + 2KCl + Br ₂	
		ZnO +HNO ₃ +H ₂ O	
		+ 2KI Pbl ₂ + 2KNO ₃	
	,		
2.	Answer all the questions		2×5=10
		What are the differences between Endothermic and Exothermic Read	tions?
	b.	What is Double Displacement Reaction? And their types	
		What is Displacement Reaction?	
		What are called Indicators?	
	e.	What are Thermal Decomposition Reactions?	
		•	
3.	Write short notes on: $2 \frac{1}{2} \times 4 = 1$		2 ½×4=10
	a.	Oxides and their types	
	b.	Catalysts	
		Activity Series of Metals	
		Characteristics of Chemical Reactions	
,	14/:	to with Observation I Departies and mounties which types of mounties in that	1 1/10 11

4. Write with Chemical Reaction, and mention which type of reaction is that $1 \% \times 10=15$ What Happen when

- i. Sodium Hydroxide reacts with Dilute Hydrochloric Acid
- ii. Zinc Oxide reacts with Nitric Acid
- iii. Sodium Chloride solution is added with Silver Nitrate Solution
- iv. Ferric Chloride reacts with Sodium Hydroxide Solution
- v. Electric Current is passed through a solution of Sodium Chloride.
- vi. Zinc Sulphate solution reacts with Ammonium Hydroxide solution
- vii. Barium Chloride reacts with Copper Sulphate solution.
- viii. Lead Nitrate Solution reacts with Potassium Iodide Solution
- ix. Electric Current is passed through molten Aluminum Oxide.
- x. Silver Bromide is placed in Day light.