



CLL371:Project.

Formaldehyde Production
from Methanol.



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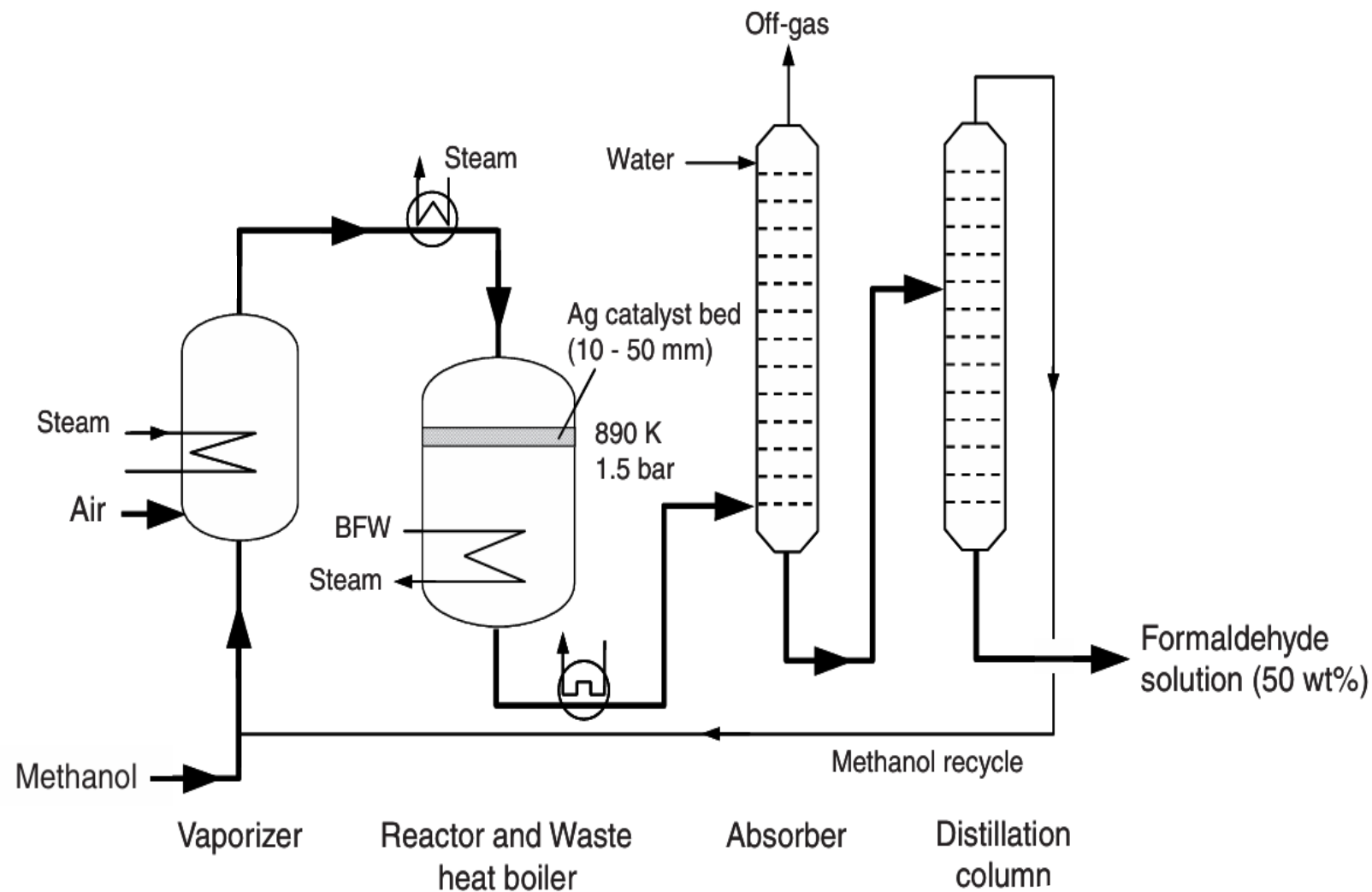
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Production of Formaldehyde using Methanol:

One of the largest applications of methanol, accounting for about 35% of methanol consumption, is the synthesis of formaldehyde, which is based on one of the following reactions:



- formaldehyde can be produced by dehydrogenation (reaction 1) or by partial oxidation (reaction 2). Reaction 1 is endothermic while reaction 2 is exothermic. The overall reaction is highly exothermic, despite some heat consumption by the endothermic dehydrogenation reaction.
- In industrial practice dehydrogenation and partial oxidation are often carried out over a silver-based catalyst in a single reactor in which the heat produced by the exothermic partial oxidation reaction supplies the heat for the endothermic dehydrogenation reaction.
- Fresh and recycle methanol is evaporated in a vaporizer into which air is sparged to generate a feed mixture. The resulting vapor is combined with steam and heated to the reaction temperature. Reaction takes place in a shallow bed of catalyst, after which the product gases are immediately cooled in a waste heat boiler, thereby generating steam.



Process flow diagram of formaldehyde production from Methanol

- After further cooling the gases are fed to an absorber. Subsequently, distillation yields a 40–55 wt% formaldehyde solution in water. Methanol is recycled to the vaporizer.
- A process in which only the partial oxidation reaction over a metal oxide catalyst. In contrast to the silver-catalyzed process, this process operates with excess air. It has the advantage that the distillation column for methanol recovery can be omitted, because in this case methanol conversion is over 99%.

