Common Polyatomic Ions

Cations

- $\bullet \ \ H_3O^+ \to \ \ \ \text{Hydronium}$
- $\bullet \ \ NH_4^+ \ \to \ \ \ \text{Ammonium}$
- $PH_4^+ \rightarrow Phosphonium$

Anions

Group 3

```
BO3^{3-} \rightarrow Borate
```

Group 4

Inorganic

- $\bullet \ CN^- \to \ \text{Cyanide}$
- $CO_3^- \rightarrow Carbonate$
- $\begin{array}{ccc} \bullet & HCO_3^- \rightarrow & Bicarbonate \\ \bullet & SiO_3^{2-} \rightarrow & Silicate/Metasilicate \\ \bullet & SnO_3^{2-} \rightarrow & Stannate \\ \end{array}$

Organic

- $\bullet \ \ C_2O_4^{2\,-} \to \ \ \text{Oxalate}$
- $C_2H_3O_2^- \rightarrow$ Acetate
- $C_3^2H_5^2O_3^2 \rightarrow$ Lactate/Propionate
- $C_6H_8O_7^{3-} \rightarrow \text{ Citrate}$
- $C_7^{\dagger}H_5^{\dagger}O_2^{-} \rightarrow \text{Benzoate}$
- $C_7H_5O_3^- \rightarrow \text{Salicylate}$

Page 1 of 3 https://schoolnotes.xyz

Group 5

```
\begin{array}{lll} \bullet & N_3^- \to & \text{Azide} \\ \bullet & NO_3^- \to & \text{Nitrate} \\ \bullet & NO_2^- \to & \text{Nitrite} \\ \bullet & PO_4^{3-} \to & \text{Phosphate} \\ \bullet & HPO_4^{2-} \to & \text{Hydrogen Phosphate} \\ \bullet & H_2PO_4^- \to & \text{Dihydrogen Phosphate} \\ \bullet & AsO_3^{3-} \to & \text{Arsenate} \\ \bullet & AsO_3^{3-} \to & \text{Arsenite} \\ \end{array}
```

Group 6

```
 \begin{array}{lll} \bullet & \operatorname{O_2^2}^- & \to & \operatorname{Peroxide} \\ \bullet & \operatorname{OH}^- & \to & \operatorname{Hydroxide} \\ \bullet & \operatorname{SO}_4^{-2} & \to & \operatorname{Sulfate/Sulphate} \\ \bullet & \operatorname{SO}_3^{-2} & \to & \operatorname{Sulfite/Sulphite} \\ \bullet & \operatorname{HSO}_4^- & \to & \operatorname{Bisulphate/Bisulfate} \\ \bullet & \operatorname{HSO}_3^- & \to & \operatorname{Bisulphite/Bisulfite} \\ \bullet & \operatorname{SCN}^- & \to & \operatorname{Thiocyanate} \\ \bullet & \operatorname{S}_2\operatorname{O}_3^{2}^- & \to & \operatorname{Thiosulfate/Thiosulphate} \\ \bullet & \operatorname{S}_2\operatorname{O}_3^{2}^- & \to & \operatorname{Peroxydisulfate/Peroxydisulphate} \\ \bullet & \operatorname{SeO}_4^{2}^- & \to & \operatorname{Selenate} \\ \end{array}
```

Group 7

```
• FO_3^- \rightarrow Fluorate

• ClO_4^- \rightarrow Perchlorate

• ClO_3^- \rightarrow Chlorate

• ClO_2^- \rightarrow Chlorite

• ClO^- \rightarrow Hypochlorite

• BrO_3^- \rightarrow Bromate

• BrO_2^- \rightarrow Bromite

• IO_3^- \rightarrow Iodate

• IO_2^- \rightarrow Iodite

• AtO_3^- \rightarrow Astatate
```

Transitional Metal Anions

```
• CrO_4^{2-} \rightarrow Chromate

• Cr_2O_7^{2-} \rightarrow Dichromate

• MnO_4^{-} \rightarrow Permaganate
```

https://schoolnotes.xyz Page 2 of 3

Common Polyatomic Ions

- $\begin{array}{ll} \bullet & MnO_4^{2\,-} \to & \text{Maganate} \\ \bullet & WO_4^{2\,-} \to & \text{Tungstate} \\ \bullet & MoO_4^{2\,-} \to & \text{Molybdate} \\ \bullet & Fe(CN)_6^{4\,-} \to & \text{Ferrocyanide} \\ \bullet & VO_3^{-} \to & \text{Vanadate} \\ \end{array}$

Page 3 of 3 https://schoolnotes.xyz