Worksheet 2

Name Class Date

Students only need one set of instructions – but multiple sets are included to reduce printing costs.

Dynamic Equilibrium

Complete the following sentences about reversible reactions.

In a reversible reaction the products can reform the original \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_. At the start:

* the forward reaction happens \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ and gets slower over time
* the reverse or backward reaction does not happen, but it gets \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ in time.

At a certain point, the forward and the reverse reactions occur at the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ speed. At this point \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ equilibrium has been achieved. At equilibrium the forward and \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ reactions are in balance. After this point the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ of reactant and product in the mixture stays the same. It is a dynamic \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ because the forward and reverse reactions are still \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

Dynamic equilibrium can only occur in a \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ system. This is a system where substances cannot enter or \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ the container where the reaction is taking place.

Dynamic Equilibrium

Complete the following sentences about reversible reactions.

In a reversible reaction the products can reform the original \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_. At the start:

* the forward reaction happens \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ and gets slower over time
* the reverse or backward reaction does not happen, but it gets \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ in time.

At a certain point, the forward and the reverse reactions occur at the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ speed. At this point \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ equilibrium has been achieved. At equilibrium the forward and \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ reactions are in balance. After this point the \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ of reactant and product in the mixture stays the same. It is a dynamic \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ because the forward and reverse reactions are still \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_.

Dynamic equilibrium can only occur in a \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ system. This is a system where substances cannot enter or \_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_\_ the container where the reaction is taking place.