Worksheet 1

Name Class Date

**1** This question is about the ammonium chloride reaction previously seen. The reaction is thermal decomposition.



**a** Explain how you can tell the reaction is reversible.

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**b** The reaction needs to be heated. Looking at the forward reaction, what would you observe during the reaction?

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**c** When the reaction reaches equilibrium, which substance(s) would be present?

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**d** At equilibrium under certain conditions, the proportion of reactants to products is 70% to 30%. Describe where the position of equilibrium is found.

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**e** Altering the conditions causes the proportion of reactants to products to change to 60% to 40%. Describe what has now happened to the position of equilibrium.

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