

UNIVERSITY OF TORONTO

FACULTY OF APPLIED SCIENCE AND ENGINEERING

FINAL EXAMINATION

20 APRIL 2012

First Year

APS 104S

INTRODUCTION TO MATERIALS AND CHEMISTRY

Exam Type B

Examiners: N. P. Kherani, K. Lian, T. Mirkovic

Q1	/15
Q2	/10
Q3	/16
Q4	/9
Q5	/15
Q6	/15
Q7	/20
Total	/100

NAME:

LastFirst

STUDENT NO:

INSTRUCTIONS:

- This is a Type B examination. Only non-programmable calculators are allowed.
- Answer all 7 questions.
- All work is to be done on the pages of this booklet.
- When answering the questions include all the steps of your work on these pages and then fill the answer in the respective boxes where provided. For additional space, you may use the back of the preceding page.
- Do not unstaple this exam booklet.
- A Formula Sheet is attached to the end of this exam booklet; if you wish, you may tear-off this sheet *only*.

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Student No: _____

QUESTION 1 (15 points): Materials Science

Part A. Multiple Choice (circle only one) – 5 points

- (i) The equilibrium interatomic spacing between atoms in a molecule or solid is determined by:
a) Attractive forces; b) Repulsive forces; c) Attractive and repulsive forces; d) none of these.
- (ii) Hydrogen bonding is a form of:
a) Ionic bonding; b) Covalent bonding; c) Dipole bonding; d) Metallic bonding.
- (iii) The cesium chloride (CsCl) crystal structure can be described as having the cation at the cube-centre and the anions at the corners of the cube. The coordination number for both ion types is:
a) 4; b) 6; c) 8; d) 12.
- (iv) The allotropes of carbon include:
a) Carbon nanotubes; b) graphite; c) fullerene; d) diamond; e) all of the preceding.
- (v) Wood, rubber, cotton, wool and silk are:
a) Synthetic polymers; b) polymorphs; c) natural polymers; d) none of the preceding.

Part B. Calculations – 10 points

Gold has an FCC crystal structure with a lattice constant of 407.82×10^{-10} cm and an atomic mass of 196.66 g/mole.

- a) Calculate the density of gold. Note: $N_A = 6.023 \times 10^{23}$ atoms mole⁻¹. (3 points)

- b) Identify the family of planes and family of directions with the highest atomic planar density and atomic linear density, respectively. (2 points)

- c) 24 karat gold is pure gold, while 18 karat gold is 75% pure (i.e., 18 parts pure gold and 6 parts other alloying elements such as silver, copper, and zinc)? From a mechanical perspective, what purpose do the alloying elements serve? (1 point)

- d) Calculate the linear atomic density along the [110] direction? (1 point)

- e) Calculate the atomic planar density for the (110) plane? (3points)

QUESTION 2 (10 points): Materials Science

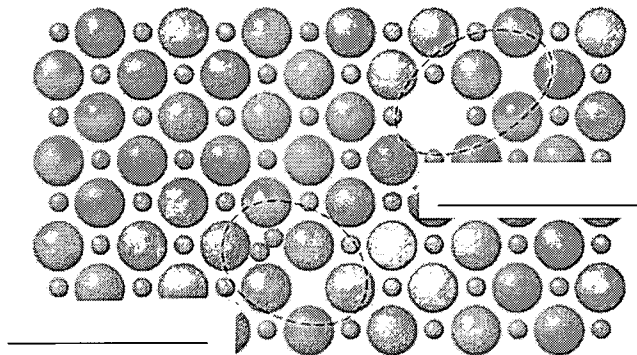
- a) The larger multiplicity of independent combination of slip planes and slip directions in FCC and BCC metals makes it more likely that one of these combinations will be favourably aligned under an applied stress and thus is the basis of these metals exhibiting ductile behaviour. (1 pt)

Circle One :

True**False**

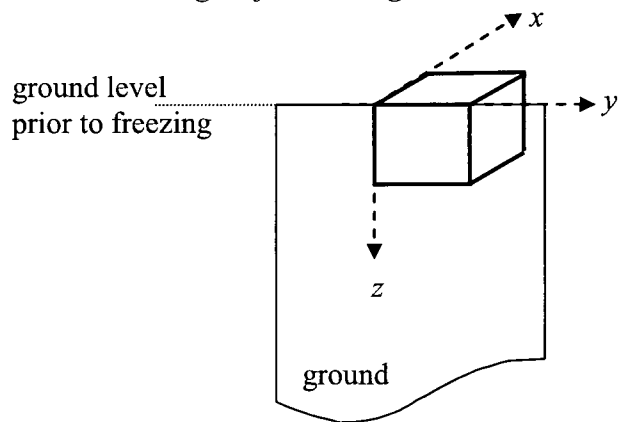
- b) Define resilience (1 point).

- c) Name the two defects in the diagram shown on the right. The larger circles are the anions and the smaller circles are the cations. (2 points)



- d) The northern winters in Canada result in frost damage due to water expansion upon freezing, specifically about 9% increase in volume due to the phase change from liquid to solid. Frost damage occurs to buildings, roads, and equipment every winter. Smaller buildings and roads are particularly affected by ground movements due to frost heave. In southern Ontario the depth of ground freezing varies from 1.0 metre near Windsor to 1.8 metres in the Ottawa area.

- (i) Consider the case where the ground freezes to a depth of 1.8 metres. Assuming that the ground comprises entirely of water, estimate the vertical movement of the ground. Assume that all the movement is upwards, that is the ground does not yield sideways or downwards. *Hint: Start with an unfrozen cube with side length of 1.8 m length, as shown below.* (3 points)



- (ii) If the ground comprises of 50% water, estimate the vertical movement (1 point)
- (iii) Explain why pot-holes get progressively larger during the winter season. (1 point)

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QUESTION 3 (16 points): Thermodynamics

(a) For an isothermal process explain why $|w|_{irrev} > |w|_{rev}$ for compression? Use a diagram to support your answer. (3 marks)

(b) State whether q , w , ΔU , ΔH and ΔS are positive (+), negative (-) or zero (0) for the reversible melting of ice at 1 atm and 0°C. (5 marks)

$$q =$$

$$w =$$

$$\Delta U =$$

$$\Delta H =$$

$$\Delta S =$$

(c) A mass of 30.0 g of H_2O (g) at 373 K (*called A*) flows into 300 g of H_2O (l) at 300 K (*called B*) and 1 atm.

Given: GMW (H_2O) = 18.02 g/mol $C_{P,m}(\text{H}_2\text{O}_{(l)}) = 75.291 \text{ J mol}^{-1} \text{ K}^{-1}$ $\Delta H_{vap}(\text{H}_2\text{O}_{(l)}) = 40.65 \text{ kJ/mol}$

i) Calculate the final temperature of the system once equilibrium has been reached. Assume that $C_{P,m}$ for H_2O is constant at its values for 298 K throughout the temperature range of interest. (3 marks)

ii) Calculate the entropy change of *A*, *B*, and the entire system. (5 marks)

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QUESTION 4: (9 points) Thermodynamics

Consider: $\text{Br}_2(\text{l}) \rightleftharpoons \text{Br}_2(\text{g})$

At 1 bar pressure liquid bromine boils at 58.2°C, and at 9.3°C its vapour pressure is 0.1334 bar.

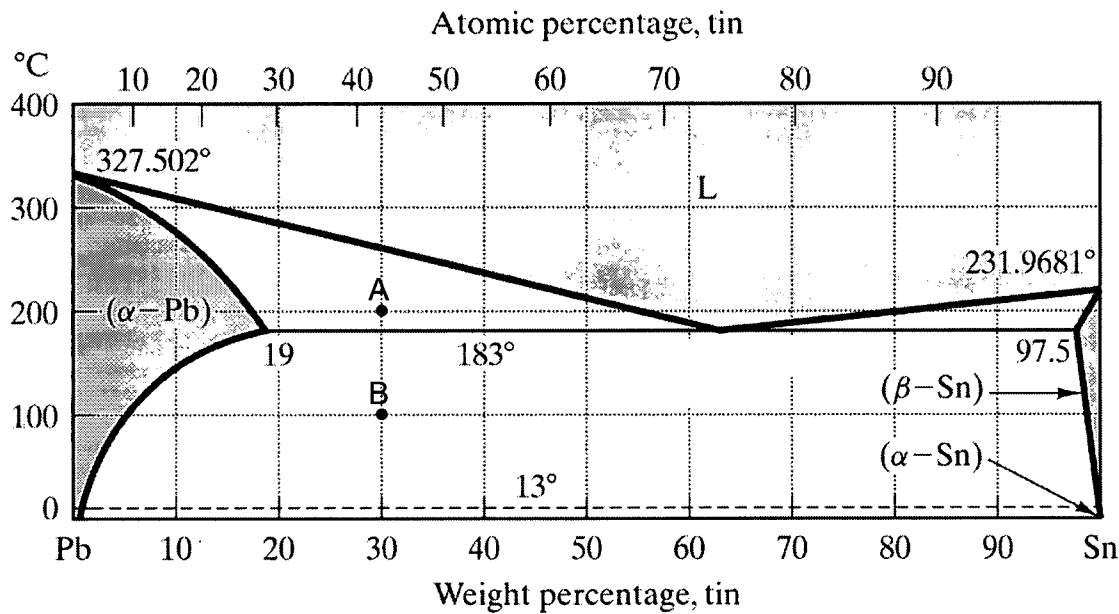
- (a) Write the expression for K_p for the above reaction. (1 mark)
- (b) Calculate ΔG° for the above reaction at 58.2°C and at 9.3°C. (2 marks)
- (c) Assuming ΔS° and ΔH° to be temperature independent, calculate their values. (3 marks)
- (d) Calculate the vapour pressure and ΔG° at 25°C. (3 marks)

QUESTION 5 (15 points): Phase Diagrams

A) Identify “True” or “False” for the following statements (2 marks)

	True	False
(i.) At an invariant point, there are two phases at the equilibrium	[]	[]
(ii.) Eutectoid reaction is the transformation from one solid to two other solids.	[]	[]

B) Based on the phase diagram below, answer the following questions. (7 marks)



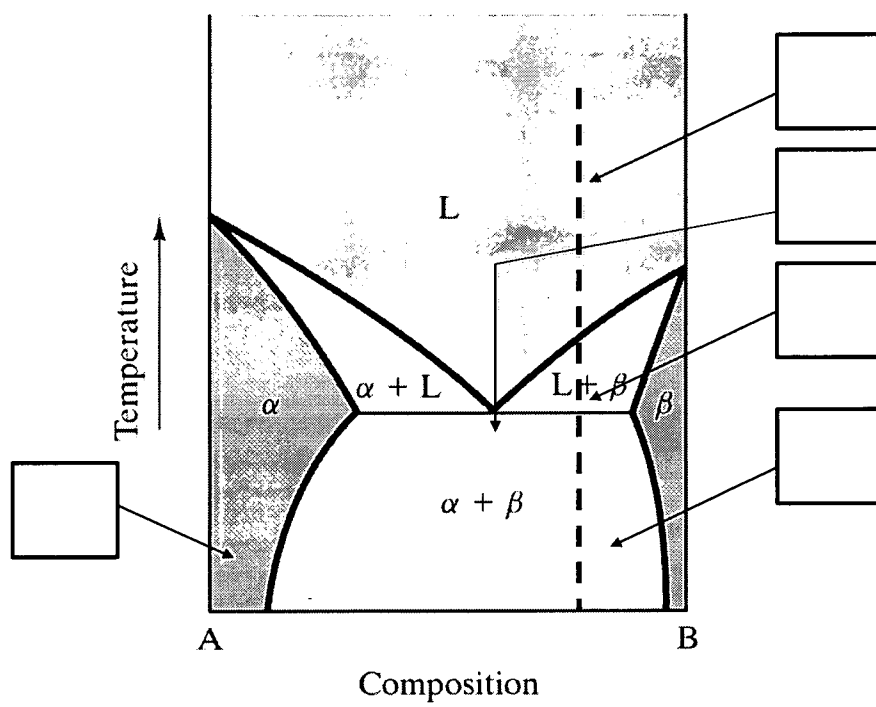
- i) The melting point of Sn is _____.
- ii) The eutectic temperature is _____ and the eutectic composition is _____.
- iii) For point A on the phase diagram,
 - a. Determine the phase compositions.
 - b. The relative fraction of each phase.
- iv) For Point B on the phase diagram,
 - a. Determine the relative fraction of each phase in terms of α and the eutectic composition.
 - b. What is the maximum solubility of Sn in Pb at this temperature?

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C) In a hypothetical phase diagram given below, please complete the following: (6 marks)

- i) Draw the corresponding structures in the squares
- ii) Write the eutectic reaction and mark the eutectic point on the graph

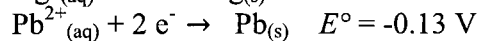
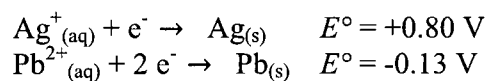
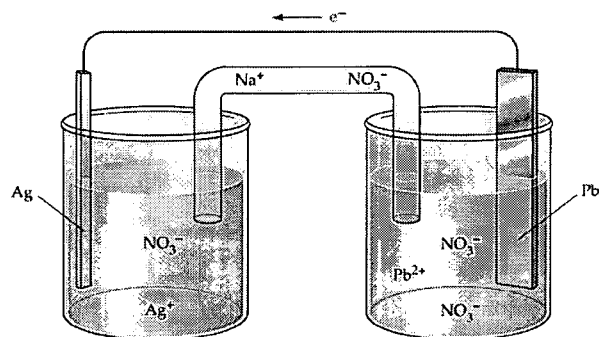


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QUESTION 6 (15 points): Electrochemistry

Consider the galvanic cell depicted below.



(a) Write the balanced reaction equation and shorthand cell expression for this cell. (3 marks)

(b) Assuming standard conditions, calculate E°_{cell} , K and ΔG° for this reaction? (6 marks)(c) What is E_{cell} if the cathode ions are present at 5 M concentration, and anode ions are present at 1 M? (3 marks)

(d) Calculate the mass change (in grams) of the silver electrode after the cell has supplied a constant current of 10.0 A for 25 min? (3 marks)

Given: atomic mass of Ag = 107.9 g/mol

QUESTION 7 (20 points): Electricity & Magnetism

A) Identify “True” or “False” for the following statements (3 marks)

	True	False
i) The energy band gap in intrinsic semiconductors is usually less than 2 eV	[]	[]
ii) Permanent magnets are often composed of hard magnetic materials	[]	[]
iii) At the Curie temperature, antiferromagnetic turns into paramagnetic	[]	[]

B) Classify the following as diamagnetic, paramagnetic, ferromagnetic, antiferromagnetic or ferrimagnetic (4 marks)

- i) Fe₃O₄ _____
- ii) O₂ _____
- iii) Argon _____
- iv) MnO _____

C) Select the best answer for each question (2 marks)

i) Electrical conductivity increases with temperatures in which of the following materials: (*Circle one*)

- a) conductor & semiconductor b) semiconductor & insulator c) conductor & insulator
- d) all of the above

ii) Which statement(s) is(are) correct? (*Circle one or more*)

- a) “Hole” has a positive charge
- b) “Hole” has a negative charge
- c) “Holes” are located in the conduction band
- d) “Holes” are located in the valence band

D) i) Calculate the mass ratio of steel (Fe) and aluminum (Al) wires having the same resistance and same length.

Given $\sigma_{\text{Fe}} = 1.1 \times 10^7 \text{ } (\Omega\text{-m})^{-1}$; $\sigma_{\text{Al}} = 3.8 \times 10^7 \text{ } (\Omega\text{-m})^{-1}$; $d_{\text{Fe}} = 7.9 \text{ Mg/m}^3$; $d_{\text{Al}} = 2.7 \text{ Mg/m}^3$. (4 marks)

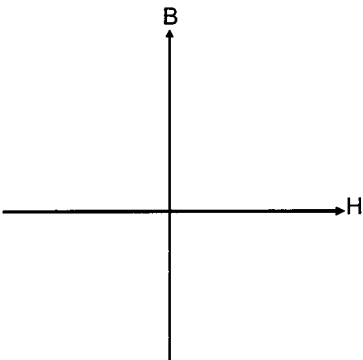
ii) If the price of Al wire is 5 times more than steel wire and based on your result in D) i), which material would you recommend for streetcar cables (1mark)?

iii) In addition to the thermal effect, what other 2 factors influence/affect the resistivity of a material (2 marks)?

- 1) _____
- 2) _____

E) Draw a hysteresis *B-H* loop (start from origin) for a ferromagnetic material and indicate on it (4 marks)

- (a) the saturation induction B_s ,
- (b) the remanence flux density B_r , and
- (c) the coercive force H_c .



FORMULAE & CONSTANTS (You may tear this sheet off.)

Constants

$R = 8.3145 \text{ J}\cdot\text{K}^{-1} \text{ mol}^{-1} = 0.0820574587 \text{ L}\cdot\text{atm}\cdot\text{K}^{-1} \text{ mol}^{-1} = 0.083145 \text{ L}\cdot\text{bar}\cdot\text{K}^{-1}\cdot\text{mol}^{-1}$			
$1 \text{ atm} = 101.325 \text{ kPa} = 1.01325 \text{ bar} = 14.696 \text{ psi} = 760 \text{ Torr} = 760 \text{ mmHg}$			
$N_A = 6.022 \times 10^{23} \text{ mol}^{-1}$	$k = 8.62 \times 10^{-5} \text{ eV/K}$	$1 \text{ eV} = 1.60 \times 10^{-19} \text{ J}$	
$T(\text{K}) = T(^{\circ}\text{C}) + 273.15 \text{ K}$	$F = 9.6485309 \times 10^4 \text{ C mole}^{-1}$	$e = 1.60217733 \times 10^{-19} \text{ C}$	

Formulae from Callister

$V = \frac{4\pi r^3}{3} \quad \rho = \frac{nA}{V_C N_A} \quad \%IC = \{1 - \exp[-0.25(X_A - X_B)^2]\} \times 100 \quad E = \int F dr \quad \frac{d}{dx}(x^n) = nx^{n-1}$			
$APF = \frac{\text{Total Sphere Volume}}{\text{Total Unit Cell Volume}}$		$LD = \frac{\text{Number Of Atoms Centred On Direction Vector}}{\text{Length Of Direction Vector}}$	
$PD = \frac{\text{Number Of Atoms Centred On A Plane}}{\text{Area Of Plane}}$		$DP = \frac{\overline{M}_n}{m}$	$\overline{M}_n = \sum x_i M_i \quad \overline{M}_w = \sum w_i M_i$
$N_V = N \exp(-\frac{Q_v}{kT}) \quad N_S = N \exp(-\frac{Q_s}{2kT}) \quad N_{fr} = N \exp(-\frac{Q_{fr}}{2kT})$			
$E = \frac{\sigma}{\epsilon} \quad \sigma = \frac{F}{A} \quad \epsilon = \frac{\Delta l}{l} \quad \tau = \frac{F}{A} \quad \tau = G\gamma \quad U_r = \frac{1}{2} \sigma_Y \epsilon_Y \quad \%CW = \left(\frac{A_0 - A_d}{A_0}\right) \times 100 \quad G = E/(2(1+\nu))$			
$\sigma_y = \sigma_0 + k_y d^{-1/2} \quad \nu = -\frac{\epsilon_x}{\epsilon_z} = -\frac{\epsilon_y}{\epsilon_z} \quad \tau_R = \sigma \cos\phi \cos\lambda \quad \sigma_y = \frac{\tau_{crss}}{(\cos\phi \cos\lambda)_{max}} \quad V = IR \quad \rho = \frac{RA}{l}$			
$\sigma = \frac{1}{\rho} \quad J = \sigma E \quad E = \frac{V}{l} \quad v_d = \mu_e E \quad \sigma = n e \mu_e \quad \theta = \cos^{-1} \left(\frac{u_1 u_2 + v_1 v_2 + w_1 w_2}{\sqrt{(u_1^2 + v_1^2 + w_1^2)(u_2^2 + v_2^2 + w_2^2)}} \right)$			

Formulae from Engel & Reid

Idea gas equation of state: $PV = nRT$ Van der Waals equation of state: $P = \frac{nRT}{(V - nb)} - \frac{an^2}{V^2}$ First law, closed systems $\Delta U = q + w \quad dU = dq + dw$ $dw = -P_{ext} dv$ $dU = nC_{v,m} dT$		$H \equiv U + PV$ $dH = nC_{p,m} dT$ For ideal gases, $C_{p,m} = C_{v,m} + R$ Solids, Liquids, $C_{p,m} = C_{v,m}$ $G \equiv H - TS$ $dS \equiv \frac{dQ_{reversible}}{T}$ For a process at constant temperature $\Delta G = \Delta H - T\Delta S$	
For an isothermal reversible process (ideal gas): $W_{rev} = -\int_{V_1}^{V_2} \frac{nRT}{V} dV = -nRT \ln \frac{V_2}{V_1} = -nRT \ln \frac{P_1}{P_2}$ Adiabatic reversible process (ideal gas): $P_1 V_1^\gamma = P_2 V_2^\gamma$ $T_1 V_1^{(\gamma-1)} = T_2 V_2^{(\gamma-1)} \quad T_1 P_1^{[(1-\gamma)/\gamma]} = T_2 P_2^{[(1-\gamma)/\gamma]} \quad \left(\frac{\overline{C}_p}{\overline{C}_v}\right) = \gamma$		ν_i : stoichiometric coefficient Assuming no phase change, constant C_p $\Delta H^{\circ}_{rxn} = \sum \nu_i \Delta H^{\circ}_{f,i} + \Delta C_p (T - 25^{\circ}\text{C})$ $\Delta C_p = \sum \nu_i C_{p,i}$ Efficiency = $ w /q_{in}$ $\epsilon(\text{ideal}) = 1 - T_C/T_H$	
For solids or liquids: $\Delta S_m = C_m \ln \frac{T_f}{T_i}$ Phase transition $\Delta S_{trans} = \frac{\Delta H_{trans}}{T_{trans}}$ $\Delta S^{\circ}(T_2) = \Delta S^{\circ}(T_1) + \int_{T_1}^{T_2} \Delta C_p \frac{dT}{T}$ Standard entropy of the reaction $\Delta S^{\circ}_{rxn} = \sum \nu_i S_m^{\circ}$ Standard free energy of a reaction: $\Delta G^{\circ}_{rxn} = \sum \nu_i \Delta G^{\circ}_{f,i}$ or $\Delta G^{\circ}_{rxn} = \Delta H^{\circ}_{rxn} - T\Delta S^{\circ}_{rxn}$ $\Delta G^{\circ} = -RT \ln K$ Total free energy of the reaction $aA + bB \Rightarrow cC + dD$ $\Delta G_{rxn} = \Delta G_{rxn}^{\circ} + RT \ln(Q)$ where $Q = \left[\frac{a_C^c a_D^d}{a_A^a a_B^b} \right]$ where a = activity (~ concentration)		$\Delta S = nR \ln \frac{V_f}{V_i}$ (isothermal) $\Delta S = nC_v \ln \frac{T_f}{T_i}$ (change in T at const V) $\Delta S = nC_p \ln \frac{T_f}{T_i}$ (change in T at const P) $\Delta G_{rxn} = \Delta G_{rxn}^{\circ} + RT \ln(Q_p)$, where $Q_p = \left(\frac{P_C}{P^{\circ}} \right)^c \left(\frac{P_D}{P^{\circ}} \right)^d \left(\frac{P_A}{P^{\circ}} \right)^a \left(\frac{P_B}{P^{\circ}} \right)^b$ $\Delta G = -nFE \quad E = E^{\circ} - \frac{RT}{nF} \ln Q \quad E = E^{\circ} - \frac{0.0592}{n} \log Q \quad \text{at } 25^{\circ}\text{C} \quad l = \frac{nC}{t}$	