

Definitions

Isolated system: No exchanges

Closed system: Only energy exchange

Open system: Energy & mass exchange

Intensive state variables:

Independent of mass

Extensive state variables:

Proportional to mass

Reservoirs: Infinite/very large system that remains unchanged when in contact with finite system.

Mechanical equilibrium:

No unbalanced forces

Thermal equilibrium:

No temperature differences

Thermodynamic equilibrium:

Intensive state variables of system are constant. Alternatively our system is in mechanical and thermal equilibrium.

Reversible processes:

Every intermediate is an equilibrium state.

Quasi-static processes:

Process sufficiently slow such that only infinitesimal temperature or pressure gradients exist.

Frictionless quasi-static processes are reversible.

Cyclic processes:

$$\Delta U = 0 \text{ and } W = Q$$

For conservative forces:

$$\oint dX = 0$$

where X is a state variable.

Adiabatic processes: $\Delta Q = 0$

Isothermal processes: $\Delta T = 0$

Isobaric processes: $\Delta P = 0$

Density

We define the density of a material as:

$$\rho = \frac{m}{V}.$$

If mass m is constant:

$$\Delta V = m \left(\frac{1}{\rho_f} - \frac{1}{\rho_i} \right)$$

assuming homogeneous material.

Zeroth law

If A is in thermal equilibrium with B and C separately then B and C are also in thermal equilibrium.

Ideal gas state equation

Given n moles of gas at temperature T :

$$\begin{aligned} PV &= nRT \\ &= Nk_B T \end{aligned}$$

where $R = N_A k_B = 8.314 \text{ JK}^{-1} \text{ mol}^{-1}$ and N the number of molecules.

Calculus identities

$$\begin{aligned} 1. \left(\frac{\partial Z}{\partial Y} \right)_X &= \left[\left(\frac{\partial Y}{\partial Z} \right)_X \right]^{-1} \\ 2. \left(\frac{\partial X}{\partial Y} \right)_Z \left(\frac{\partial Y}{\partial Z} \right)_X \left(\frac{\partial Z}{\partial X} \right)_Y &= -1 \end{aligned}$$

First law

Total energy E is conserved and:

$$\begin{aligned} \Delta U &= Q - W \\ dU &= dQ - dW \\ \dot{U} &= \dot{Q} - \dot{W} \end{aligned}$$

where U is internal energy and $E \geq U$.

Note $Q > 0$ represents energy transferred into system. When system does work on surroundings we have $W > 0$.

The work done by a fluid in **reversible** processes is:

$$dW = PdV$$

and has units Joules (J).

Isothermal expansion

Let $P_1 > P_2$ where P_1 and P_2 denote system and external pressure respectively. Only mechanical work is exchanged via a piston. By applying a force such that there exists pressure difference dP , our expansion becomes reversible and hence:

$$W = nRT \int_{V_1}^{V_2} \frac{dV}{V}.$$

Note that for isothermal processes under ideal gas assumption, $\Delta U = 0$.

Isochoric heat capacity

$$C_V(T) = \left(\frac{dQ}{dT} \right)_V = \left(\frac{\partial U}{\partial T} \right)_V$$

Isobaric heat capacity

$$\begin{aligned} C_P &= \left(\frac{dQ}{dT} \right)_P \\ &= C_V + \left[P + \left(\frac{\partial U}{\partial V} \right)_T \right] \left(\frac{\partial V}{\partial T} \right)_P \end{aligned}$$

Heat capacity has units JK^{-1} .

For ideal gases we have that:

$$C_P - C_V = nR.$$

Adiabatic expansion

Under reversible adiabatic processes:

$$TV^{\gamma-1} = \text{constant}$$

$$PV^{\gamma} = \text{constant}$$

$$T^{\frac{1}{\gamma-1}} V = \text{constant}$$

where γ is the adiabatic exponent:

$$\gamma = \frac{C_P}{C_V} = \frac{f+2}{f}$$

$$U = \frac{f}{2} nRT$$

regardless of degrees of freedom f .

Enthalpy

The state function enthalpy simplifies the description of heat transfer.

Enthalpy has units J and is defined as:

$$H = U + PV$$

$$\begin{aligned} dH &= dU + VdP + PdV \\ &= dQ + VdP. \end{aligned}$$

Under isobaric reversible conditions:

$$dH = dQ_P.$$

$$\therefore C_P = \left(\frac{\partial H}{\partial T} \right)_P$$

Latent heats

Latent heat is heat needed for sample to undergo a phase transition:

$$\Delta U_m = Q_m - P\Delta V_m$$

$$L_m = Q_m = \Delta H_m.$$

Chemical reactions

Since $Q = \Delta U + P_0 \Delta V = \Delta H$:

- $Q < 0$: exothermic (heat is released)
- $Q > 0$: endothermic (heat is absorbed)

at constant pressure P_0 .

General form for first law

Given system with m conjugate pairs (x_i, X_i) that represent multiple modes of work exchange:

$$dU = dQ + \sum_{i=1}^m x_i dX_i$$

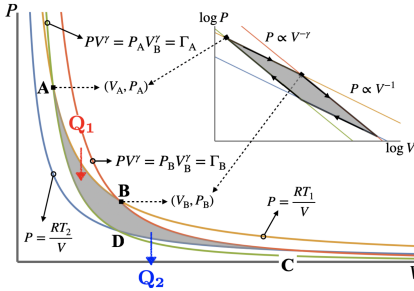
for each generalised force $\{x_i\}$ drives generalised differential displacement $\{X_i\}$.

Carnot's theorem

Peak efficiency of a cyclic heat engine:

$$\eta = \frac{\dot{W}}{\dot{Q}_H} = 1 - \frac{Q_C}{Q_H} = 1 - \frac{T_C}{T_H}$$

and is either in terms of rate or value, with units J or Js⁻¹. This ideal cycle is known as the Carnot cycle:



where AB,CD are isothermal processes and BC, DA are adiabatic processes.

Entropy

The state function entropy is a measure of disorder defined as:

$$S = \frac{Q}{T}$$

where Q is heat received from a reservoir at temperature T and units JK⁻¹.

Then for a reversible cyclic heat engine:

$$dU = TdS - PdV.$$

For all processes the following holds:

$$dH = TdS + VdP.$$

Entropy of mixing

$$\Delta S = n_A R \ln \frac{V_A + V_B}{V_A} + n_B R \ln \frac{V_A + V_B}{V_B}$$

$$\Delta s_{mix} = -R(x_A \ln x_A + x_B \ln x_B)$$

$$x_A = \frac{n_A}{n_A + n_B} \text{ and } x_B = \frac{n_B}{n_A + n_B}$$

Second law

$$\Delta S_{total} = \Delta S_{system} + \Delta S_{reservoir} \geq 0$$

$$dS \geq \frac{dQ}{T}$$

Helmholtz free energy

$$F = U - TS$$

$$dF = -SdT - PdV$$

Gibbs free energy

$$G = H - TS$$

$$dG = -SdT + VdP$$

Chemical reactions are spontaneous if:

$$\Delta G = \Delta H - T\Delta S < 0.$$

Maxwell relations

$$\left(\frac{\partial T}{\partial V}\right)_S = -\left(\frac{\partial P}{\partial S}\right)_V$$

$$\left(\frac{\partial T}{\partial P}\right)_S = \left(\frac{\partial V}{\partial S}\right)_P$$

$$\left(\frac{\partial S}{\partial V}\right)_T = \left(\frac{\partial P}{\partial T}\right)_V$$

$$-\left(\frac{\partial S}{\partial P}\right)_T = \left(\frac{\partial V}{\partial T}\right)_P$$

The isobaric expansivity is defined as:

$$\beta = \frac{1}{V} \left(\frac{\partial V}{\partial T}\right)_P$$

and the isothermal compressibility:

$$\kappa_T = -\frac{1}{V} \left(\frac{\partial V}{\partial P}\right)_T.$$

Throttling

Throttling is the adiabatic reduction in gas pressure and is an isenthalpic process. At constant volume we define:

$$\begin{aligned} \mu_{JK} &= \left(\frac{\partial T}{\partial P}\right)_H \\ &= \frac{V(T, P)}{C_P}(\beta T - 1) \end{aligned}$$

as the slope of a P-T plot known as the Joule-Kelvin coefficient.

Clausius-Clapeyron equation

The slope of any phase boundary is:

$$\frac{dP}{dT} = \frac{\Delta S}{\Delta V} = \frac{\Delta H}{T\Delta V}$$

since constant pressure at boundaries.

Van der Waals state equation

$$\left(P + \frac{an^2}{V^2}\right)(V - nb) = nRT$$

Chemical potentials

The Euler equation for a 1-component **open** system with N particles is:

$$U = TS - PV + \mu N$$

with modified first law statement:

$$dU = TdS - PdV + \mu dN.$$

This gives the Gibbs-Duhem relation:

$$SdT - VdP + Nd\mu = 0$$

where μ is the chemical potential:

$$\mu = \frac{G}{N}$$

since G is extensive. At constant T with ideal gas assumptions:

$$\mu(P, T) = RT \ln \frac{P}{P_0} + \mu_0(P, T).$$

Chemical potential μ has units J.

Third law

$$S = 0 \text{ at } T = 0\text{K}.$$