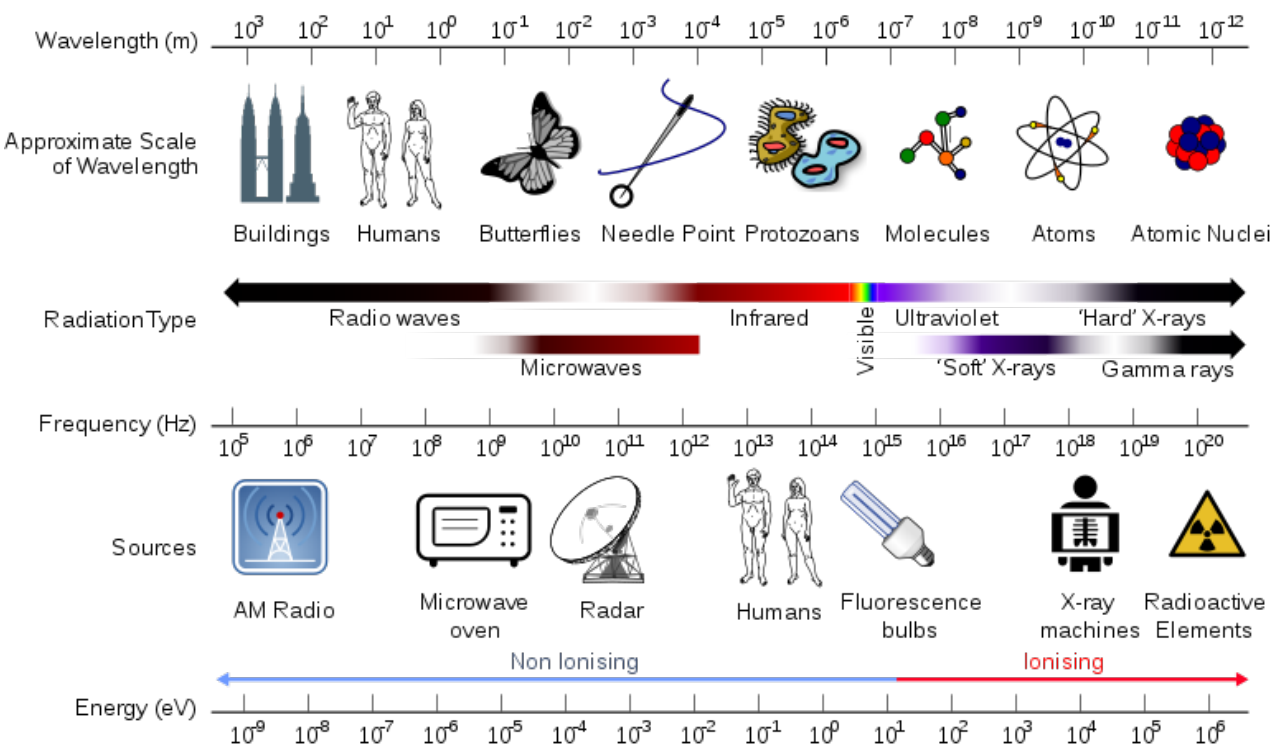
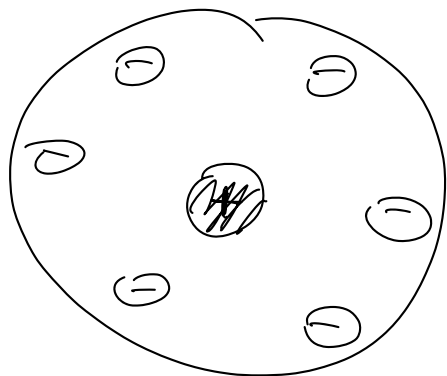


Token: 769 108

Rutherford $r_n \leq 30 \text{ fm}$
 γ, e^-, p, n nuclei



1 H Hydrogen 1.008																	2 He Helium 4.003			
3 Li Lithium 6.94	4 Be Beryllium 9.012													5 B Boron 10.81	6 C Carbon 12.011	7 N Nitrogen 14.007	8 O Oxygen 15.999	9 F Fluorine 18.998	10 Ne Neon 20.180	
11 Na Sodium 22.990	12 Mg Magnesium 24.305													13 Al Aluminum 26.982	14 Si Silicon 28.085	15 P Phosphorus 30.974	16 S Sulfur 32.06	17 Cl Chlorine 35.45	18 Ar Argon 39.948	
19 K Potassium 39.098	20 Ca Calcium 40.078	21 Sc Scandium 44.956	22 Ti Titanium 47.867	23 V Vanadium 50.942	24 Cr Chromium 51.996	25 Mn Manganese 54.938	26 Fe Iron 55.845	27 Co Cobalt 58.933	28 Ni Nickel 58.693	29 Cu Copper 63.546	30 Zn Zinc 65.38	31 Ga Gallium 69.723	32 Ge Germanium 72.630	33 As Arsenic 74.922	34 Se Selenium 78.97	35 Br Bromine 79.904	36 Kr Krypton 83.798			
37 Rb Rubidium 85.468	38 Sr Strontium 87.62	39 Y Yttrium 88.906	40 Zr Zirconium 91.224	41 Nb Niobium 92.906	42 Mo Molybdenum 95.95	43 Tc Technetium [97]	44 Ru Ruthenium 101.07	45 Rh Rhodium 102.906	46 Pd Palladium 106.42	47 Ag Silver 107.868	48 Cd Cadmium 112.414	49 In Indium 114.818	50 Sn Tin 118.710	51 Sb Antimony 121.760	52 Te Tellurium 127.60	53 I Iodine 126.904	54 Xe Xenon 131.293			
55 Cs Cesium 132.905	56 Ba Barium 137.327	* 57 - 70	71 Lu Lutetium 174.967	72 Hf Hafnium 178.49	73 Ta Tantalum 180.948	74 W Tungsten 183.84	75 Re Rhenium 186.207	76 Os Osmium 190.23	77 Ir Iridium 192.217	78 Pt Platinum 195.084	79 Au Gold 196.967	80 Hg Mercury 200.592	81 Tl Thallium 204.38	82 Pb Lead 207.2	83 Bi Bismuth 208.980	84 Po Polonium [209]	85 At Astatine [210]	86 Rn Radon [222]		
87 Fr Francium [223]	88 Ra Radium [226]	** 89 - 102	103 Lr Lawrencium [262]	104 Rf Rutherfordium [267]	105 Db Dubnium [270]	106 Sg Seaborgium [269]	107 Bh Bohrium [270]	108 Hs Hassium [270]	109 Mt Meitnerium [278]	110 Ds Darmstadtium [281]	111 Rg Roentgenium [285]	112 Cn Copernicium [285]	113 Nh Nihonium [286]	114 Fl Flerovium [289]	115 Mc Moscovium [289]	116 Lv Livermorium [293]	117 Ts Tennessine [293]	118 Og Oganesson [294]		
*Lanthanide series			57 La Lanthanum 138.905	58 Ce Cerium 140.116	59 Pr Praseodymium 140.908	60 Nd Neodymium 144.242	61 Pm Promethium [145]	62 Sm Samarium 150.36	63 Eu Europium 151.964	64 Gd Gadolinium 157.25	65 Tb Terbium 158.925	66 Dy Dysprosium 162.500	67 Ho Holmium 164.930	68 Er Erbium 167.259	69 Tm Thulium 168.934	70 Yb Ytterbium 173.045				
**Actinide series			89 Ac Actinium [227]	90 Th Thorium 232.038	91 Pa Protactinium 231.036	92 U Uranium 238.029	93 Np Neptunium [237]	94 Pu Plutonium [244]	95 Am Americium [243]	96 Cm Curium [247]	97 Bk Berkelium [247]	98 Cf Californium [251]	99 Es Einsteinium [252]	100 Fm Fermium [257]	101 Md Mendelevium [258]	102 No Nobelium [259]				

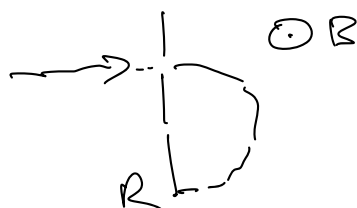
Atomic Number — 6
Symbol — C
Name — Carbon
Average Atomic Mass — 12.011

metals —
nonmetals —
metalloids —

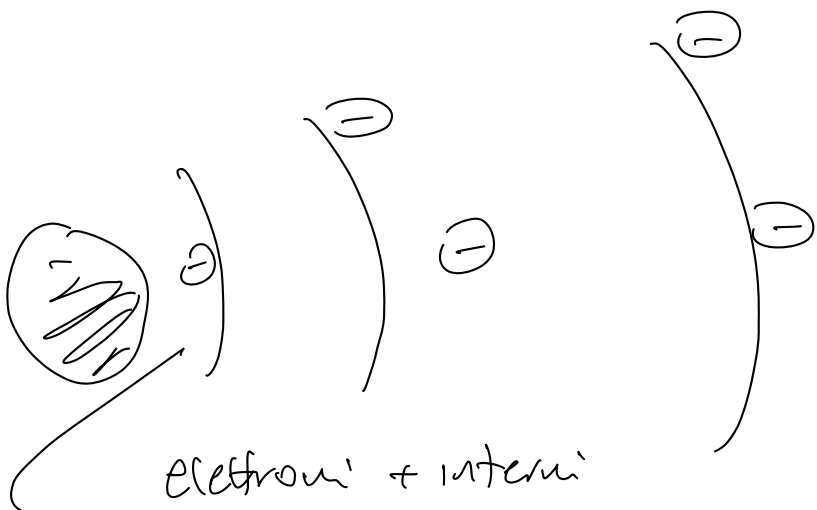
Z_A dei nuclei

nucleo: massa $\propto A$

↳ spettroscopia di massa



Z : carica dei nuclei



legge di Moseley:

$$E_e = \frac{3}{4} R_y (Z-1)^2 \approx 10 \text{ eV } (Z-1)^2$$

misura sperimentale di Z, A dei nuclei ${}^A_Z N$

nucleidi (Z, A)

isotopi: ${}^{12}_6 C$ 0.989 abbondanza in natura

${}^{13}_6 C$ 0.011

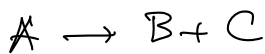
isotopi: $Z = \text{cost.}$ A variabile. (lo stesso chimico lo stesso elemento)

$$A = Z + N$$

↓ ↘
proton # neutron

isobari: $A = \text{cost}$ $Z = \text{diversi}$. Sono elementi diversi, nucleidi con la stessa massa.

nucleidi $\begin{cases} \text{stabili} \\ \text{instabili} \Rightarrow \text{strutture prodotti di decadimento.} \end{cases}$



$$E_i = E_f$$

$$E = m + T \rightarrow \text{en. cinetica}$$

Più solido con A: $M_A = E_B + E_C = m_B + m_C + T_B + T_C$

$$Q = m_A - m_B - m_C = T_B + T_C = T_f$$

en. cinetica nello stato finale.

Q valore del decadimento

$$E_A = E_B + E_C$$

Se $T_A \neq 0$

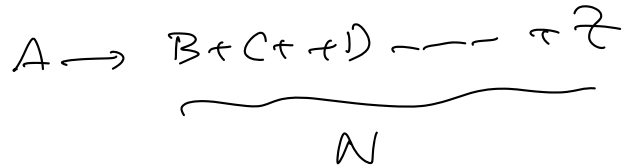
$$M_A + T_A = m_B + T_B + m_C + T_C$$

$$Q = \underbrace{m_A + m_B - m_C}_{\text{mass difference}} = T_B' + T_C' - T_A$$

Energie si conserva $\Rightarrow Q > 0$ nei decadimenti

$$\Gamma(i \rightarrow f) = 2\pi |A|^2 \rho(E) \Big|_{E_f=E_i}$$

$$Q = M_{in} - \sum_f^N m_f$$



Nuclide (A, Z) $M(Z, A)$ masse del nuclide.

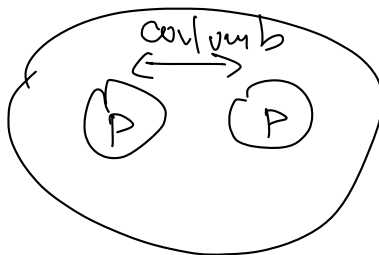
$$M(Z, A) = Z \cdot m_p + (A - Z) m_n$$

Esperimenti di diffusione sui nuclei

- distrib. uniforme di massa e carica.
- distrib. sferica uniforme (buona approssim.)
- $q = Ze$ $Z \in \mathbb{N}$

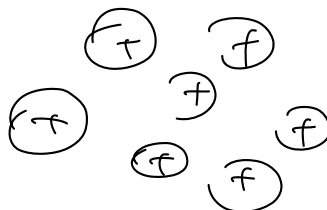
\hookrightarrow le stesse cariche elettriche di e^-

$$Z = 2$$



$$F \propto \frac{1}{r^2}$$

Z cariche positive



$$\frac{Z(Z-1)}{2}$$

termini di repulsione

\Rightarrow esistere una forza più forte di Coulomb

\Rightarrow forze nucleari

sperimentalmente

$$m(Z, A) < Z m_p + (A - Z) m_n$$

Definizione

$$m(Z, A) = Z m_p + (A - Z) m_n - B(Z, A) \quad B \geq 0$$

energia di legame nucleare

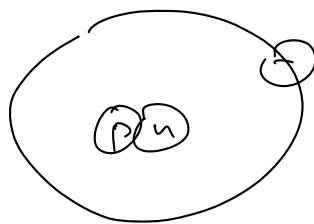
$$\rightarrow m({}_1^1\text{H}) = m_p + m_e - 13.6 \text{ eV}$$

$$= 938.9 \text{ MeV} = 1.008 \text{ uA}$$

$$1 \text{ uA} = \frac{1}{12} m({}_6^{12}\text{C}) = 931.494 \text{ MeV}$$

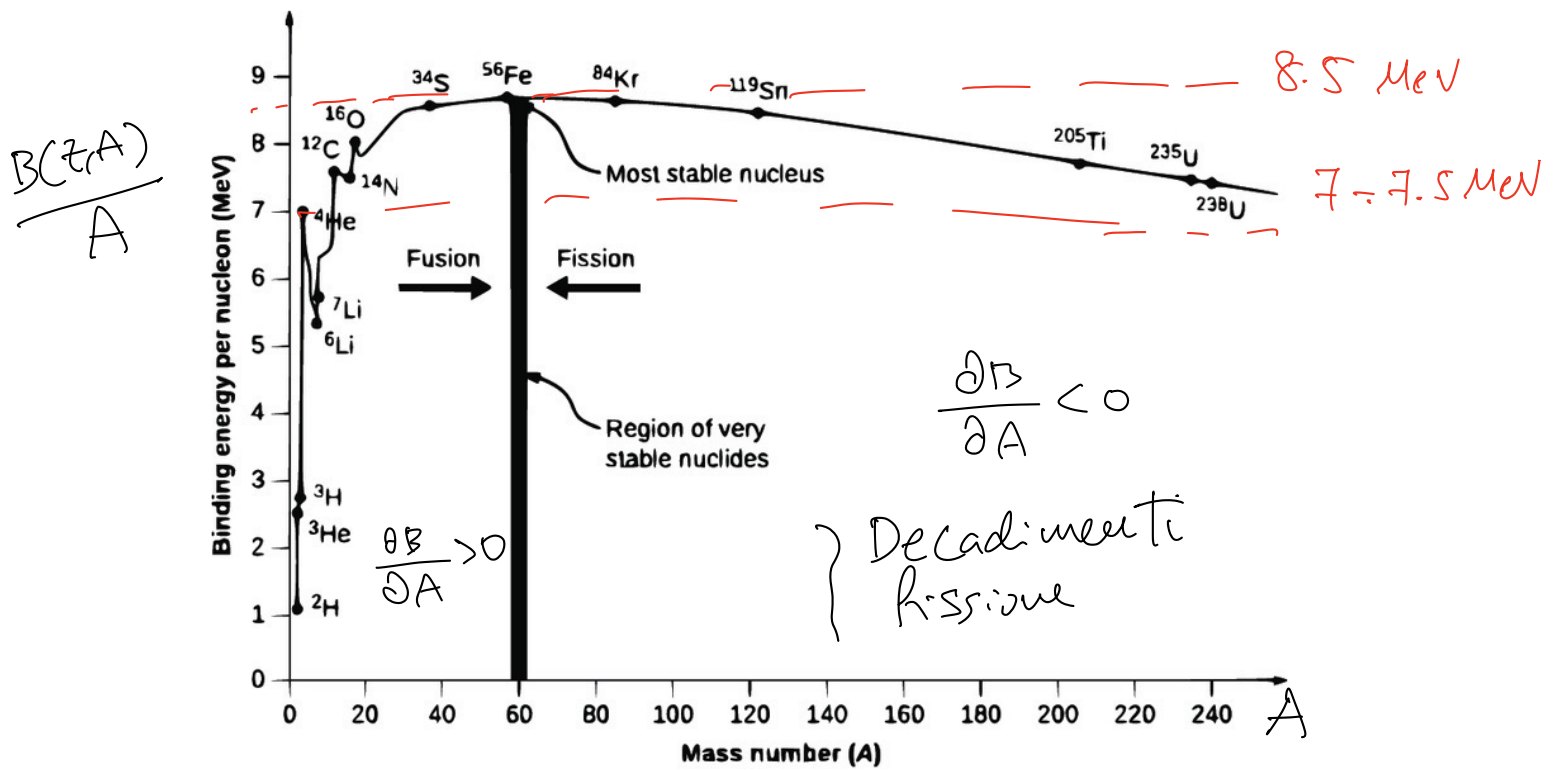
$$\text{H: } B(1, 1) = 0$$

$$\text{Deuterio: } {}_1^2\text{H}$$



deutone: nucleo di Deuterio ${}_1^2\text{H}$

$$B(1, 2) = m({}_1^1\text{H}) + m_n - m({}_1^2\text{H}) = 2.225 \text{ MeV}$$

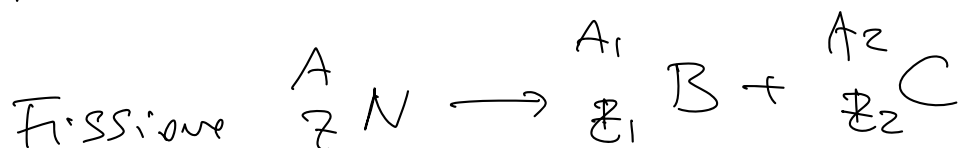


nucleone è protone o neutrone

$$B(^4\text{He}) = B(Z, A) = 7 \text{ MeV} \quad \text{particella } \alpha$$

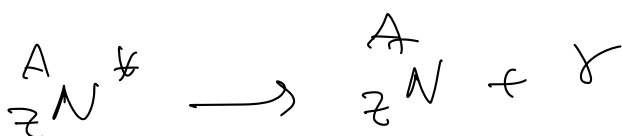
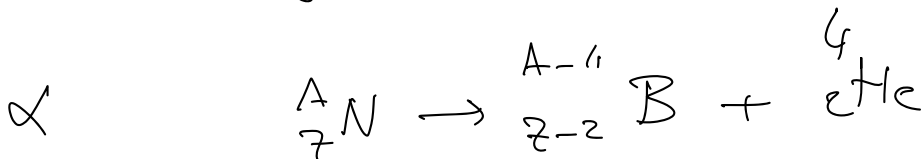
\Rightarrow uno dei nuclei più legati.

$$\frac{dB}{dZ} \neq 0 \text{ per } A \leq 60 \Rightarrow \text{fusione favorita}$$



$$A = A_1 + A_2 \quad \text{tipicamente } Z_1 \approx Z_2 \approx \frac{Z}{2}$$

$$Z = Z_1 + Z_2$$



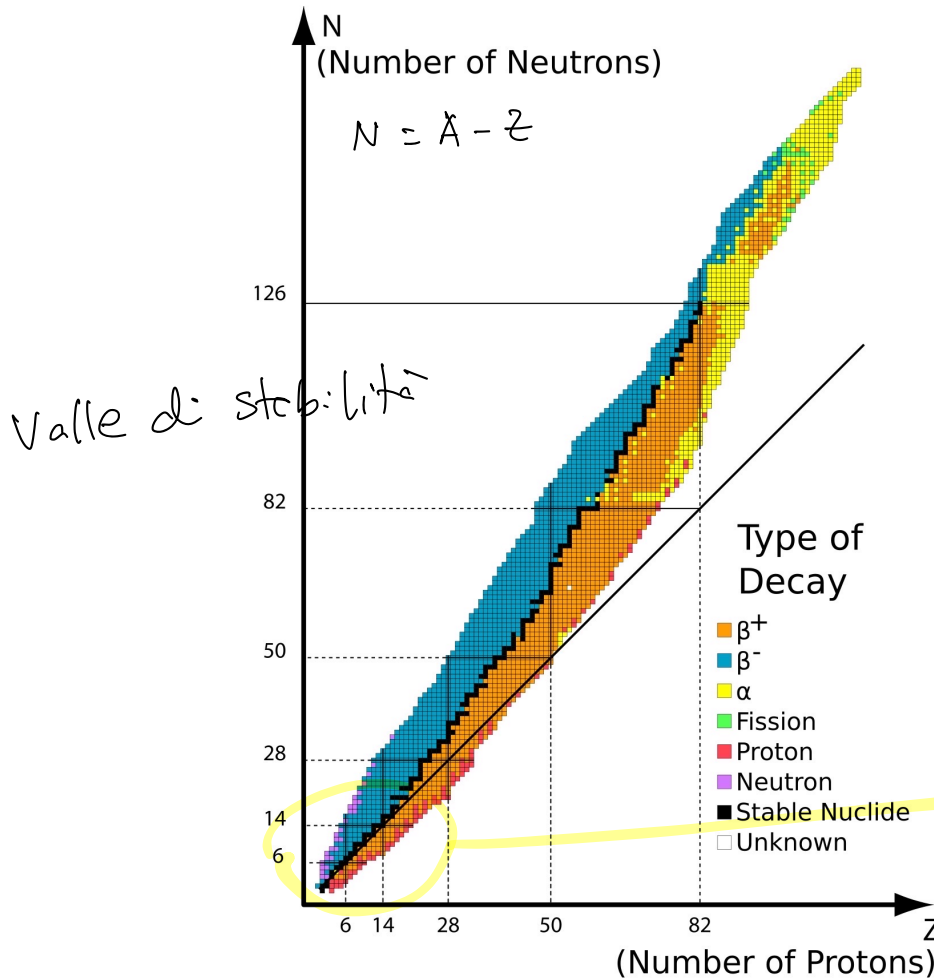
Decadimento γ
 $E_\gamma \approx 1 - 10 \text{ MeV}$

conservazione # barionico

barioni: (qqq) 3 quark.

protoni (uud)
neutroni (udd)

Carta di Segrè



Z grandi
nuclei stabili hanno
un eccesso di neutroni

$$N \approx Z$$

$$A \approx 2Z$$

$$\frac{Z}{A} \approx \frac{1}{2}$$