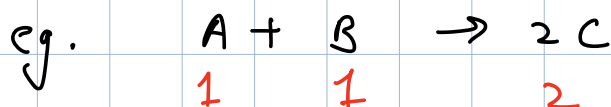


Few Points for the exam.

1. (a) Constant density systems:

→ if the system is in aqueous phase
or

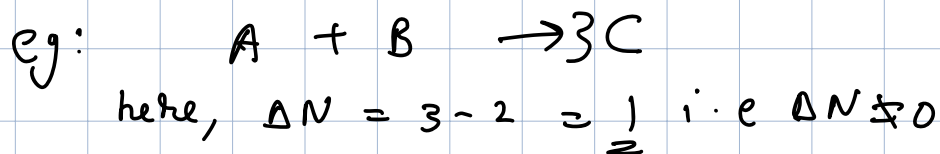
→ if the system is in gas phase with
 $\Delta N = 0$



$\Delta N =$ no. of moles on product side
— no. of moles on reactant side.
 $= 0$

(b) variable density systems:

→ if the system is in gas phase
with $\Delta N \neq 0$



2. If the order of the reaction is not mentioned
look at the units of reaction constant (k)
on the pre exponent factor (k_0)

→ For first order rxn: $[k]$ or $[k_0] = \text{min}^{-1}$ or sec^{-1}

→ For second order rxn: $[k]$ or $[k_0] = \text{L/mol} \cdot \text{sec}$