GCSE - Chemistry - Acids

Question 1 (2 marks)

Define an acid in terms of its behaviour in aqueous solution.

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Question 2 (3 marks)

Describe the difference between a strong acid and a weak acid. Give an example of each.

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Question 3 (2 marks)

What is the pH scale, and how does it relate to acidity?

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Question 4 (3 marks)

Write the balanced chemical equation for the reaction between sulfuric acid (H_2SO_4) and sodium hydroxide (NaOH).

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Question 5 (2 marks)

What type of reaction occurs between an acid and a base? What are the products formed?

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Question 6 (3 marks)

Describe a simple experiment to determine the relative strengths of two acids.

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Question 7 (2 marks)

Explain why it is dangerous to taste an unknown acid.

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Question 8 (3 marks)

What is the role of an indicator in a titration? Give an example of an indicator used in acid-base titrations.

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Question 9 (2 marks)

Write the ionic equation for the reaction of any strong acid with any strong alkali.

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Question 10 (3 marks)

A solution of hydrochloric acid has a pH of 2. What is the approximate pH of a solution of ethanoic acid of the same concentration? Explain your reasoning.

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