An ideal gas

An ideal gas has a gas constant $R = 0.3\ kJ/kg \cdot K$ and a constant volume specific heat c_v =0.7 $kJ/kg \cdot K$. If the gas has a temperature change of 100°C, what is the change in enthalpy?

Answer: 100.

Explanation: The change in enthalpy is given by:

$$\Delta h = c_p \cdot \Delta T = (c_v + R) \cdot \Delta T = 100 kJ/kg$$