

## Elements of Groups 16, 17 and 18

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### EXERCISE [PAGES 163 - 164]

#### Exercise | Q 1.01 | Page 163

Select appropriate answers for the following.

Which of the following has highest electron gain enthalpy?

1. Fluorine
2. Chlorine
3. Bromine
4. Iodine

**Solution:** Chlorine

#### Exercise | Q 1.02 | Page 163

Select appropriate answers for the following.

Hydrides of group 16 are weakly acidic. The correct order of acidity is \_\_\_\_\_.

1.  $\text{H}_2\text{O} > \text{H}_2\text{S} > \text{H}_2\text{Se} > \text{H}_2\text{Te}$
2.  $\text{H}_2\text{Te} > \text{H}_2\text{O} > \text{H}_2\text{S} > \text{H}_2\text{Se}$
3.  $\text{H}_2\text{Te} > \text{H}_2\text{Se} > \text{H}_2\text{S} > \text{H}_2\text{O}$
4.  $\text{H}_2\text{Te} > \text{H}_2\text{Se} > \text{H}_2\text{O} > \text{H}_2\text{S}$

**Solution:**

Hydrides of group 16 are weakly acidic. The correct order of acidity is  $\text{H}_2\text{Te} > \text{H}_2\text{Se} > \text{H}_2\text{S} > \text{H}_2\text{O}$ .

#### Exercise | Q 1.03 | Page 163

Select appropriate answers for the following.

Which of the following element does not show an oxidation state of +4?

1. O
2. S
3. Se
4. Te

**Solution:** O

**Exercise | Q 1.04 | Page 163**

**Select appropriate answers for the following.**

HI acid when heated with conc.  $\text{H}_2\text{SO}_4$  forms \_\_\_\_\_.

1.  $\text{HIO}_3$
2.  $\text{KIO}_3$
3.  $\text{I}_2$
4. KI

**Solution:**

HI acid when heated with conc.  $\text{H}_2\text{SO}_4$  forms  $\text{I}_2$

**Exercise | Q 1.05 | Page 163**

**Select appropriate answers for the following.**

Ozone layer is depleted by \_\_\_\_\_.

1. **NO**
2.  $\text{NO}_2$
3.  $\text{NO}_3$
4.  $\text{N}_2\text{O}_5$

**Solution:**

Ozone layer is depleted by **NO**.

**Exercise | Q 1.06 | Page 163**

**Select appropriate answers for the following.**

Which of the following occurs in a liquid state at room temperature?

1.  $\text{HIO}_3$
2. HBr
3. HCl
4. **HF**

**Solution:** HF

**Exercise | Q 1.07 | Page 163**

**Select appropriate answers for the following.**

In pyrosulfuric acid oxidation state of sulfur is \_\_\_\_\_.

1. only +2
2. only +4
3. +2 and +6
4. **only +6**

**Solution:**

In pyrosulfuric acid oxidation state of sulfur is **only +6**.

**Exercise | Q 1.08 | Page 163**

**Select appropriate answers for the following.**

Stability of interhalogen compounds follows the order \_\_\_\_\_.

1.  $\text{BrF} > \text{IBr} > \text{ICl} > \text{ClF} > \text{BrCl}$
2.  $\text{IBr} > \text{BrF} > \text{ICl} > \text{ClF} > \text{BrCl}$
3.  **$\text{ClF} > \text{ICl} > \text{IBr} > \text{BrCl} > \text{BrF}$**
4.  $\text{ICl} > \text{ClF} > \text{BrCl} > \text{IBr} > \text{BrF}$

**Solution:**

Stability of interhalogen compounds follows the order  **$\text{ClF} > \text{ICl} > \text{IBr} > \text{BrCl} > \text{BrF}$** .

**Exercise | Q 1.09 | Page 163**

**Select appropriate answers for the following.**

$\text{BrCl}$  reacts with water to form \_\_\_\_\_.

1.  $\text{HBr}$
2.  $\text{Br}_2 + \text{Cl}_2$
3.  $\text{HOBr}$
4.  **$\text{HOBr} + \text{HCl}$**

**Solution:**

$\text{BrCl}$  reacts with water to form  **$\text{HOBr} + \text{HCl}$** .

**Exercise | Q 1.1 | Page 163**

**Select appropriate answers for the following.**

Chlorine reacts with excess of fluorine to form \_\_\_\_\_.

1. ClF
2. **ClF<sub>3</sub>**
3. ClF<sub>2</sub>
4. Cl<sub>2</sub>F<sub>3</sub>

**Solution:**

Chlorine reacts with excess of fluorine to form **ClF<sub>3</sub>**.

**Exercise | Q 1.11 | Page 163**

**Select appropriate answers for the following.**

In interhalogen compounds, which of the following halogens is never the central atom?

1. I
2. Cl
3. Br
4. **F**

**Solution:** F

**Exercise | Q 1.12 | Page 163**

**Select appropriate answers for the following.**

Which of the following has one lone pair of electrons?

1. IF<sub>3</sub>
2. ICl
3. **IF<sub>5</sub>**
4. ClF<sub>3</sub>

**Solution:** IF<sub>5</sub>

**Exercise | Q 1.13 | Page 163**

**Select appropriate answers for the following.**

In which of the following pairs, molecules are paired with their correct shapes?

1. [I<sub>3</sub>]<sup>-</sup> : bent
2. BrF<sub>5</sub> : trigonal bipyramidal
3. ClF<sub>3</sub> : trigonal planar

**4.  $[\text{BrF}_4]^-$  : square planar**

**Solution:**  $[\text{BrF}_4]^-$  : square planar

**Exercise | Q 1.14 | Page 163**

**Select appropriate answers for the following.**

Among the known interhalogen compounds, the maximum number of atoms is

\_\_\_\_\_.

1. 3
2. 6
3. 7
- 4. 8**

**Solution:**

Among the known interhalogen compounds, the maximum number of atoms is 8.

**Exercise | Q 2.01 | Page 163**

**Answer the following.**

Write the order of the thermal stability of the hydrides of group 16 elements.

**Solution:**

The decreasing order of thermal stability:  $\text{H}_2\text{O} > \text{H}_2\text{S} > \text{H}_2\text{Se} > \text{H}_2\text{Te}$

**Exercise | Q 2.02 | Page 163**

**Answer the following.**

What is the oxidation state of Te in  $\text{TeO}_3$ ?

**Solution:**

(Oxidation number of Te) + (Oxidation number of O) = 0

(Oxidation number of Te) +  $3 \times (-2) = 0$

Oxidation number of Te - 6 = 0

Hence, oxidation number of Te in  $\text{TeO}_3$  = +6

**Exercise | Q 2.03 | Page 163**

**Answer the following.**

Name two gases which deplete ozone layer.

**Solution:**

Two gases that deplete ozone layer: Nitric oxide (NO), chlorofluoro carbons (freons).

**Exercise | Q 2.04 | Page 163****Answer the following.**

Give two uses of  $\text{ClO}_2$ .

**Solution:**

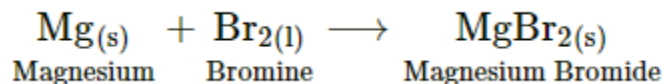
1.  $\text{ClO}_2$  is used as a bleaching agent for paper pulp and textiles.
2. It is also used in the water treatment.

**Exercise | Q 2.05 | Page 163****Answer the following.**

What is the action of bromine on magnesium metal?

**Solution:**

Bromine reacts with magnesium to form magnesium bromide.

**Exercise | Q 2.06 | Page 163****Answer the following.**

Write the names of allotropic forms of selenium.

**Solution:**

Allotropes of Selenium: Red and grey selenium

**Exercise | Q 2.07 | Page 163****Answer the following.**

What is the oxidation state of 'S' in  $\text{H}_2\text{SO}_4$ ?

**Solution:**

(Oxidation number of H) + (Oxidation number of S) + (Oxidation number of O) = 0

$2 \times (+1) + (\text{Oxidation number of S}) + 4 \times (-2) = 0$

Oxidation number of S + 2 - 8 = 0

Hence, oxidation number of 'S' in  $\text{H}_2\text{SO}_4$  = +6

### Exercise | Q 2.08 | Page 163

**Answer the following.**

The  $pK_a$  value of HCl is - 7.0 and that of HI is -10.0. Which is the stronger acid?

**Solution:**

Smaller is the  $pK_a$  value, greater is its acidity. Hence, HI is stronger acid than HCl.

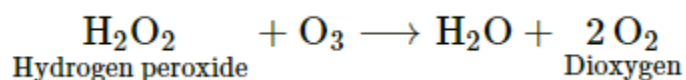
### Exercise | Q 2.09 | Page 163

**Answer the following.**

Give one example showing reducing property of ozone.

**Solution:**

Reaction showing reducing property of ozone is,



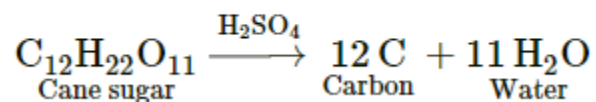
### Exercise | Q 2.1 | Page 164

**Answer the following.**

Write the reaction of conc.  $\text{H}_2\text{SO}_4$  with sugar.

**Solution:**

The reaction is,



### Exercise | Q 2.11 | Page 164

Write any two uses of Chlorine.

**Solution:**

**Uses of chlorine:**

1. Large quantities of chlorine are used industrially for bleaching wood pulp (required for the manufacture of paper and rayon), bleaching cotton and textiles.
2. It is used in the extraction of gold and platinum.
3. It is used for the manufacture of dyes, drugs and organic compounds such as

$\text{CCl}_4$ ,  $\text{CHCl}_3$ , refrigerants, etc.

4. It is used in sterilising drinking water.

5. It is used in preparation of poisonous gases such as phosgene ( $\text{COCl}_2$ ), tear gas ( $\text{CCl}_3\text{NO}_2$ ), mustard gas ( $\text{Cl.C}_2\text{H}_4\text{-S - C}_2\text{H}_4\text{Cl}$ ).

6. It is used in the manufacture of DDT and BHC which are important insecticide.

7. It is used in the manufacture of vinyl chloride which is starting material for polyvinyl chloride plastics.

8. It is used in the manufacture of bleaching powder, aluminium chloride, hydrogen chloride, hypochlorites, chlorates, perchlorates, etc., which are important industrial compounds.

### Exercise | Q 2.12 | Page 164

**Complete the following.**



**Solution:**

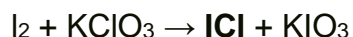


### Exercise | Q 2.12 | Page 164

**Complete the following.**

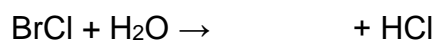


**Solution:**

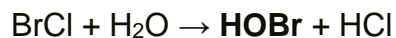


### Exercise | Q 2.12 | Page 164

**Complete the following.**

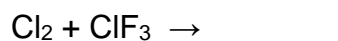


**Solution:**



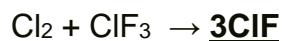
### Exercise | Q 2.12 | Page 164

**Complete the following.**





**Solution:**

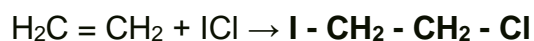


**Exercise | Q 2.12 | Page 164**

**Complete the following.**

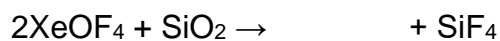


**Solution:**

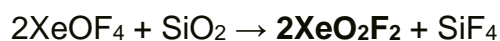


**Exercise | Q 2.12 | Page 164**

**Complete the following.**

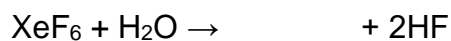


**Solution:**

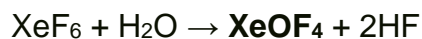


**Exercise | Q 2.12 | Page 164**

**Complete the following.**

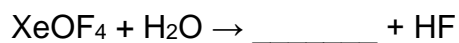


**Solution:**

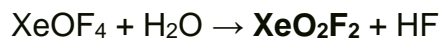


**Exercise | Q 2.12 | Page 164**

**Complete the following.**



**Solution:**



**Exercise | Q 2.13 | Page 164**

**Match the following:**

A	B
$\text{XeOF}_2$	Xenon trioxydifluoride

$\text{XeO}_2\text{F}_2$	Xenon monooxydifluoride
$\text{XeO}_3\text{F}_2$	Xenon dioxytetrafluoride
$\text{XeO}_2\text{F}_4$	Xenon dioxydifluoride

**Solution:**

<b>A</b>	<b>B</b>
$\text{XeOF}_2$	Xenon monooxydifluoride
$\text{XeO}_2\text{F}_2$	Xenon dioxydifluoride
$\text{XeO}_3\text{F}_2$	Xenon trioxydifluoride
$\text{XeO}_2\text{F}_4$	Xenon dioxytetrafluoride

### Exercise | Q 2.14 | Page 164

What is the oxidation state of xenon in the following compounds?

$\text{XeOF}_4$ ,  $\text{XeO}_3$ ,  $\text{XeF}_6$ ,  $\text{XeF}_4$ ,  $\text{XeF}_2$ .

**Solution:**

<b>Compound</b>	<b>Oxidation state of xenon</b>
$\text{XeOF}_4$	+ 6
$\text{XeO}_3$	+ 6
$\text{XeF}_6$	+ 6
$\text{XeF}_4$	+ 4
$\text{XeF}_2$	+ 2

### Exercise | Q 3.01 | Page 164

**Answer the following.**

The first ionization enthalpies of S, Cl, and Ar are 1000, 1256 and 1520  $\text{kJ/mol}^{-1}$ , respectively. Explain the observed trend.

**Solution:**

1. The elements S, Cl, and Ar belong to third period of the periodic table.
2. Across a period, effective nuclear charge increases and atomic size decreases with increase in atomic number. Therefore, the energy required for the removal of an electron from the valence shell (I.E.) increases in the order:  $\text{S} < \text{Cl} < \text{Ar}$ .

### Exercise | Q 3.02 | Page 164

**Answer the following.**

Acidic character of hydrides of group 16 elements increases from  $\text{H}_2\text{O}$  to  $\text{H}_2\text{Te}$ . Explain.

**Solution:**

The acidic character of the hydrides increases from  $\text{H}_2\text{O}$  to  $\text{H}_2\text{Te}$  because the bond dissociation enthalpy of the  $\text{H-E}$  bond decreases down the group. (where,  $\text{E} = \text{O}, \text{S}, \text{Se}, \text{Te}$ )

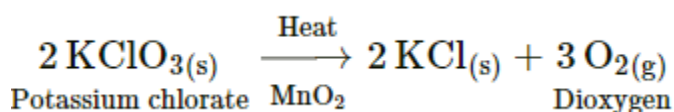
### Exercise | Q 3.03 | Page 164

**Answer the following.**

How is dioxygen prepared in the laboratory from  $\text{KClO}_3$ ?

**Solution:**

Dioxygen is prepared by heating potassium chlorate ( $\text{KClO}_3$ ) in presence of the catalyst,  $\text{MnO}_2$ .



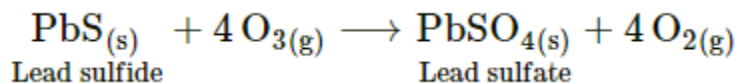
### Exercise | Q 3.04 | Page 164

**Answer the following.**

What happens when lead sulfide reacts with ozone ( $\text{O}_3$ )?

**Solution:**

Ozone oxidizes lead sulfide to lead sulfate



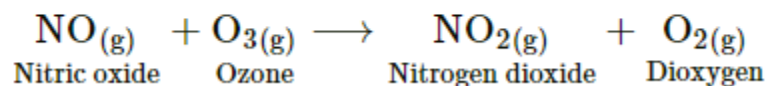
### Exercise | Q 3.04 | Page 164

**Answer the following.**

What happens when nitric oxide reacts with ozone.

**Solution:**

Ozone oxidizes nitric oxide and gives nitrogen dioxide.



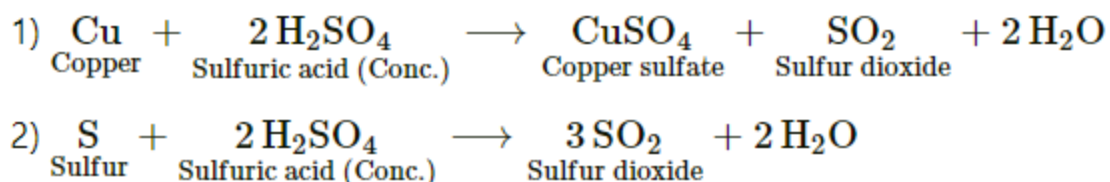
### Exercise | Q 3.05 | Page 164

**Answer the following.**

Give two reactions showing the oxidising property of concentrated  $\text{H}_2\text{SO}_4$ .

**Solution:**

Metals and nonmetals both are oxidized by hot, concentrated sulfuric acid which itself gets reduced to  $\text{SO}_2$ .



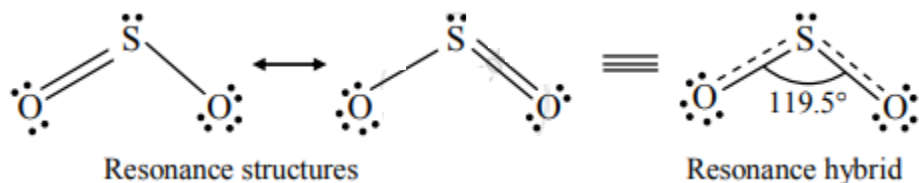
**Exercise | Q 3.06 | Page 164**

**Answer the following.**

Discuss the structure of sulfur dioxide.

**Solution:**

- 1) Sulfur dioxide is angular with  $\text{O} - \text{S} - \text{O}$  bond angle of  $119.5^\circ$ .
- 2) The  $\text{S} - \text{O}$  double bond arises from  $d\pi-p\pi$  bonding.
- 3) It is a resonance hybrid of two canonical forms as shown below:



**Exercise | Q 3.07 | Page 164**

**Answer the following.**

Fluorine shows only  $-1$  oxidation state while other halogens show  $-1$ ,  $+1$ ,  $+3$ ,  $+5$  and  $+7$  oxidation states. Explain.

**Solution:**

- 1) The fluorine atom has no d-orbitals in its valence shell and therefore, cannot expand its octet. Thus, fluorine is the most electronegative exhibit  $-1$  oxidation state only.
- 2)  $\text{Cl}$ ,  $\text{Br}$ , and  $\text{I}$  exhibit  $-1$ ,  $+1$ ,  $+3$ ,  $+5$ , and  $+7$  oxidation states. This is because they are less electronegative than  $\text{F}$  and possess empty d-orbitals in the valence shell and therefore, can expand the octet.

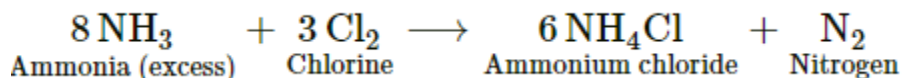
**Exercise | Q 3.08 | Page 164**

**Answer the following.**

What is the action of chlorine on Excess  $\text{NH}_3$ .

**Solution:**

Chlorine when reacted with excess of ammonia gives ammonium chloride and nitrogen.



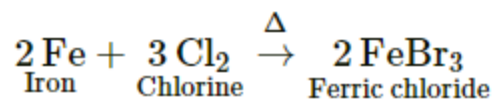
**Exercise | Q 3.08 | Page 164**

**Answer the following.**

What is the action of chlorine on Fe?

**Solution:**

Chlorine reacts with Fe to form ferric chloride.



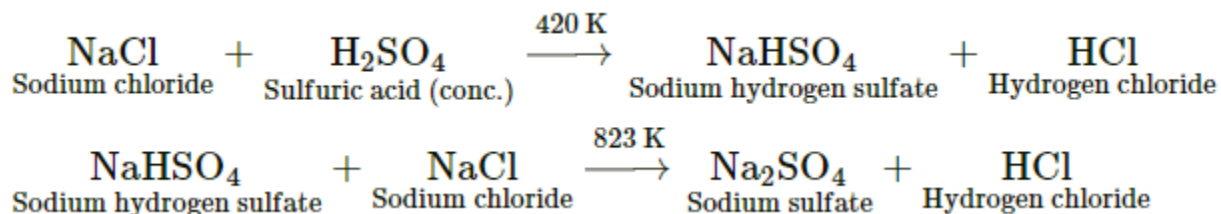
**Exercise | Q 3.09 | Page 164**

**Answer the following.**

How is hydrogen chloride prepared from sodium chloride?

**Solution:**

1) In the laboratory, hydrogen chloride is prepared by heating sodium chloride (common salt) with concentrated sulfuric acid



2) HCl gas can be dried by passing it through concentrated sulfuric acid.

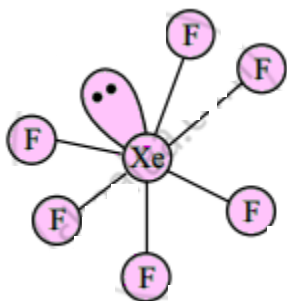
**Exercise | Q 3.1 | Page 164**

**Answer the following.**

Draw structures of  $\text{XeF}_6$

**Solution:**

**structures of  $\text{XeF}_6$  :**



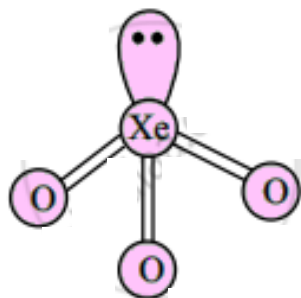
**Exercise | Q 3.1 | Page 164**

**Answer the following.**

Draw structures of  $\text{XeF}_3$

**Solution:**

**structures of  $\text{XeF}_3$  :**



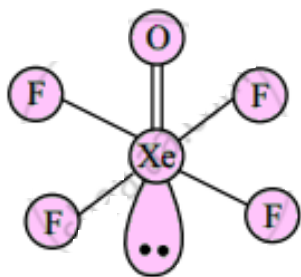
**Exercise | Q 3.1 | Page 164**

**Answer the following.**

Draw structures of  $\text{XeOF}_4$

**Solution:**

**structures of  $\text{XeOF}_4$  :**



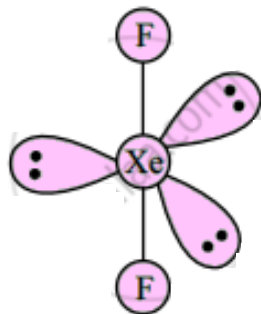
**Exercise | Q 3.1 | Page 164**

**Answer the following.**

Draw structures of  $\text{XeOF}_2$

**Solution:**

structures of  $\text{XeOF}_2$  :



**Exercise | Q 3.11 | Page 164**

**Answer the following.**

What are interhalogen compounds? Give two examples.

**Solution:**

An interhalogen compound is a compound formed by a combination of atoms of different halogens.

e.g.  $\text{ClF}$ ,  $\text{BrF}_3$

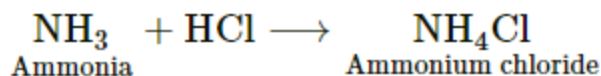
**Exercise | Q 3.12 | Page 164**

**Answer the following.**

What is the action of hydrochloric acid on  $\text{NH}_3$ ?

**Solution:**

Hydrochloric acid reacts with ammonia and gives white fumes of ammonium chloride.



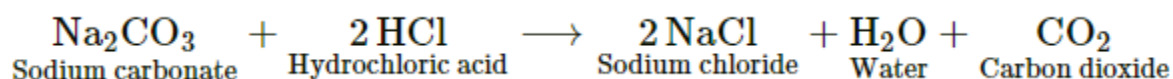
### Exercise | Q 3.12 | Page 164

**Answer the following.**

What is the action of hydrochloric acid on  $\text{Na}_2\text{CO}_3$ ?

**Solution:**

Hydrochloric acid reacts with sodium carbonate to give sodium chloride with the release of  $\text{CO}_2$  gas.



### Exercise | Q 3.13 | Page 164

**Answer the following.**

Give two uses of HCl.

**Solution:**

1. It is used in the manufacture of chlorine and chlorides such as ammonium chloride.
2. It is used in the preparation of aqua regia which is used to dissolve noble metals like gold.

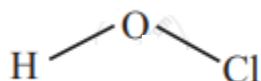
### Exercise | Q 3.14 | Page 164

**Answer the following.**

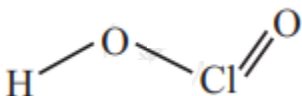
Write the names and structural formulae of oxoacids of chlorine.

**Solution:**

- 1) Hypochlorous acid, HOCl

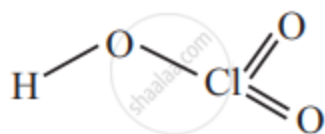


- 2) Chlorous acid, HOClO or  $\text{HClO}_2$

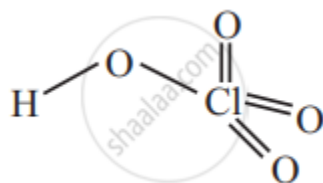


- 3) Chloric acid,  $\text{HClO}_3$





4) Perchloric acid,  $\text{HClO}_4$



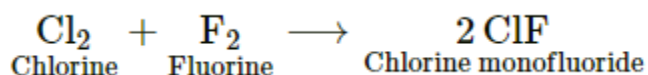
### Exercise | Q 3.15 | Page 164

**Answer the following.**

What happens when  $\text{Cl}_2$  reacts with  $\text{F}_2$  in equal volume at 437 K.

**Solution:**

When  $\text{Cl}_2$  reacts with  $\text{F}_2$  in equal volume at 437 K, chlorine monofluoride is formed.



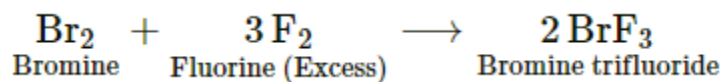
### Exercise | Q 3.15 | Page 164

**Answer the following.**

What happens when  $\text{Br}_2$  reacts with excess of  $\text{F}_2$ .

**Solution:**

When  $\text{Br}_2$  reacts with excess of  $\text{F}_2$ , bromine trifluoride is formed.



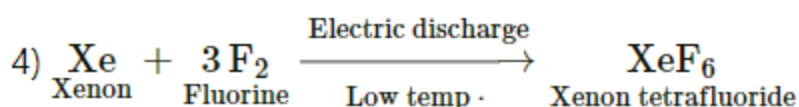
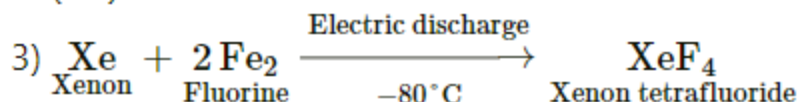
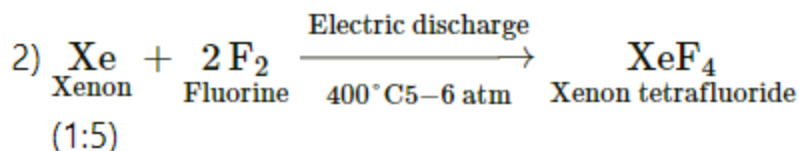
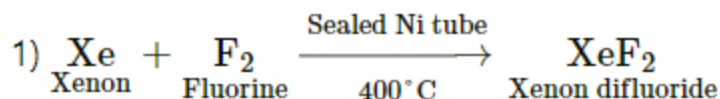
### Exercise | Q 3.16 | Page 164

**Answer the following.**

How are xenon fluorides  $\text{XeF}_2$ ,  $\text{XeF}_4$  and  $\text{XeF}_6$  obtained? Give suitable reactions.

**Solution:**

Xenon fluorides are generally prepared by direct reaction of xenon and fluorine in different ratios and conditions, such as temperature, electric discharge and photochemical reaction.



### Exercise | Q 3.17 | Page 164

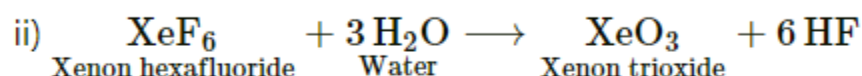
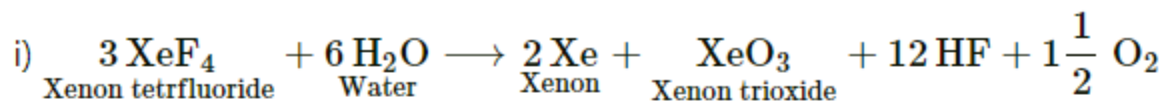
Answer the following.

How are  $\text{XeF}_2$  and  $\text{XeOF}_4$  prepared?

**Solution:**

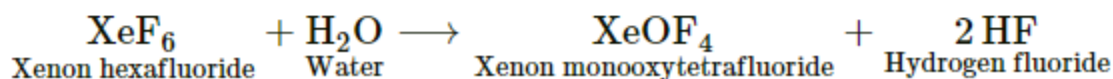
#### 1) Preparation of $\text{XeO}_3$ :

Fluorides of xenon react with water to form  $\text{XeO}_3$ .



#### 2) Preparation of $\text{XeOF}_4$ :

Partial hydrolysis of  $\text{XeF}_6$  yields  $\text{XeOF}_4$ .



### Exercise | Q 3.18 | Page 164

Answer the following.

List the uses of Neon and argon gases.

### Solution1:

#### Uses of neon gas:

- (i) It is mixed with helium to protect electrical equipments from high voltage.
- (ii) It is filled in discharge tubes with characteristic colours.
- (iii) It is used in beacon lights.

#### Uses of Argon gas:

- (i) Argon along with nitrogen is used in gas-filled electric lamps. This is because Ar is more inert than N.
- (ii) It is usually used to provide an inert temperature in a high metallurgical process.
- (iii) It is also used in laboratories to handle air-sensitive substances.

### Solution2:

#### Uses of Neon

Neon is used in discharge tubes and fluorescent bulbs for advertisement display purposes. Glow of different colours 'neon signs' can be produced by mixing neon with other gases. Neon bulbs are used in botanical gardens and in green houses.

#### Uses of Argon

Argon is used mainly to provide an inert atmosphere in high temperature metallurgical processes such as arc welding of metals and alloys. In the laboratory, it is used for handling substance which are air sensitive.

It is used in filling incandescent and fluorescent lamps where its presence retards the sublimation of the filament and thus increases the life of the lamp. It is also used in "neon signs" for obtaining lights of different colours.

### Exercise | Q 3.19 | Page 164

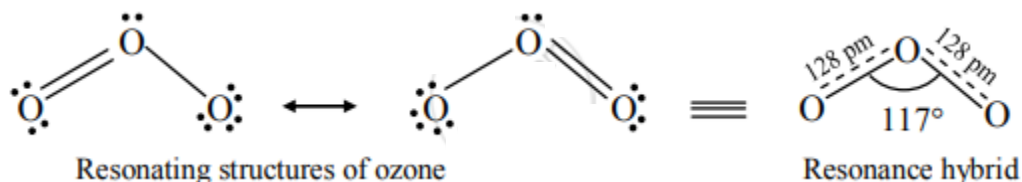
#### Answer the following.

Describe the structure of the Ozone. Give two uses of ozone.

#### Solution:

#### Structure of ozone:

Ozone ( $O_3$ ) is an angular molecule. The two O – O bond lengths in the ozone molecule are identical, 128 pm, and the O – O – O bond angle of about  $117^\circ$ . It is a resonance hybrid of two canonical forms.



### Exercise | Q 3.2 | Page 164

#### Answer the following.

Explain the trend in following atomic properties of group 16 elements:

Atomic radii

#### Solution:

##### Atomic radii:

The atomic radii increase down the group, as a result of an increase in the number of quantum shells.

### Exercise | Q 3.2 | Page 164

#### Answer the following.

Explain the trend in following atomic properties of group 16 elements:

Ionisation enthalpy

#### Solution:

##### Ionisation enthalpy:

The ionisation enthalpy decreases down the group due to an increase in the atomic size. Group 16 elements have less ionization energy than the corresponding group 15 elements.

### Exercise | Q 3.2 | Page 164

#### Answer the following.

Explain the trend in following atomic properties of group 16 elements:

Electronegativity

#### Solution:

##### Electronegativity:

The electronegativity decreases down the group. Oxygen has the highest electronegativity next to fluorine amongst all the elements.

### Exercise | Q 4.1 | Page 164

#### Answer the following.

Distinguish between rhombic sulfur and monoclinic sulfur.

#### Solution:

The following are the points of difference between rhombic ( $\alpha$ -sulfur) and monoclinic ( $\beta$  - sulfur):

<b>Rhombic sulfur (<math>\alpha</math>-sulfur)</b>	<b>Monoclinic sulfur (<math>\beta</math>- sulfur)</b>
1. It is a pale yellow coloured solid.	1. It is bright yellow solid
2. It forms orthorhombic crystals	2. It forms needle-shaped monoclinic crystals
3. Its melting point is 385.8 K.	3. Its melting point is 393 K.
4. Its density is 2.06 g/cm <sup>3</sup>	4. Its density is 1.98 g/cm <sup>3</sup>
5. It is insoluble in water and soluble in CS <sub>2</sub>	5. Soluble in CS <sub>2</sub>
6. It is stable below 369 K and transforms to $\beta$ -sulfur above this temperature	6. It is stable above 369 K and transforms into $\alpha$ -sulfur below this temperature.
7. It is prepared by the evaporation of rolls sulfur in CS <sub>2</sub> .	7. It is prepared from rhombic sulfur.

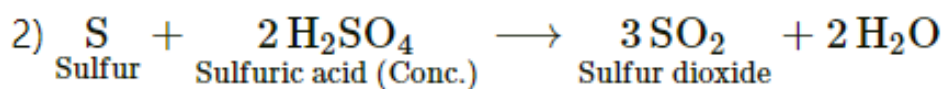
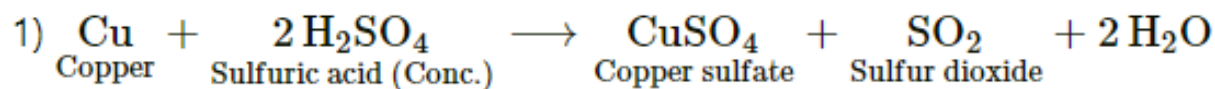
### Exercise | Q 4.2 | Page 164

**Answer the following.**

Give two reactions showing the oxidising property of concentrated H<sub>2</sub>SO<sub>4</sub>.

**Solution:**

Metals and nonmetals both are oxidised by hot, concentrated sulfuric acid which itself gets reduced to SO<sub>2</sub>.



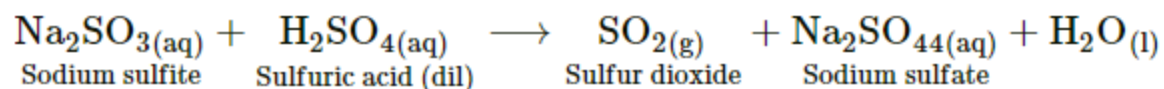
### Exercise | Q 4.3 | Page 164

**Answer the following.**

How is SO<sub>2</sub> prepared in the laboratory from sodium sulfite?

**Solution:**

In laboratory, sulfur dioxide is prepared by treating sodium sulfite with dilute sulfuric acid.



### Exercise | Q 4.4 | Page 164

**Answer the following.**

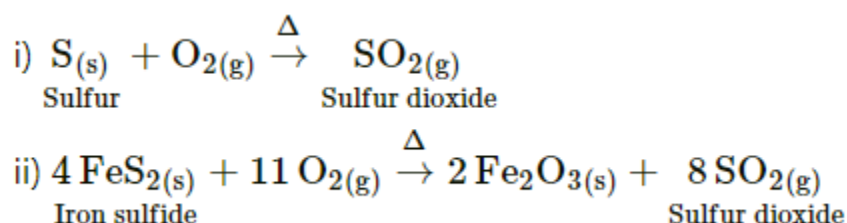
Describe the manufacturing of  $\text{H}_2\text{SO}_4$  by the contact process.

**Solution:**

Sulfuric acid is manufactured by the contact process, which involves the following three steps.

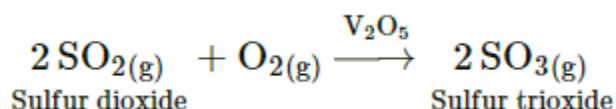
#### 1) Roasting in air:

Sulfur or sulfide ore (iron pyrites) on burning or roasting in air produces sulfur dioxide.



#### 2) Catalytic oxidation of sulfur dioxide:

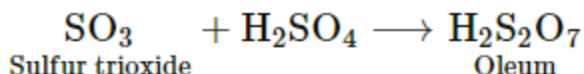
i) Sulfur dioxide is oxidised catalytically with oxygen to sulfur trioxide, in the presence of  $\text{V}_2\text{O}_5$  catalyst.



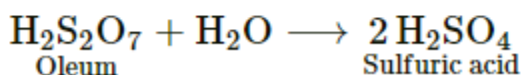
ii) The reaction is exothermic and reversible and the forward reaction leads to decrease in volume. Therefore, low temperature (720K) and high pressure (2 bar) are favourable conditions for maximum yield of  $\text{SO}_3$ .

#### 3) Absorption, followed by dilution of sulfur trioxide gas:

i) Sulfur trioxide gas (from the catalytic converter) is absorbed in concentrated  $\text{H}_2\text{SO}_4$  to produce oleum.



ii) Dilution of oleum with water gives sulfuric acid of desired concentration.



iii) The sulfuric acid obtained by contact process is 96-98 % pure.