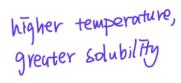
Class Practice 2.3

1.	Which of the following process is NOT used in extracting common salt from sea water?	
	A.	evaporation
	B.	distillation: extract water
	C.	crystallization
	D.	filtration
2.	Who	en a glass of sea water is left on a table under room conditions
	for	10 days, what will be observed?
	A.	Sea water becomes milky.
	B.	Some crystals are formed. (rystallization
	C.	Some white powder is formed.
	D.	Sea water remains unchanged.
3.	When a solution of sodium chloride is concentrated by heating for preparation of crystals, which of the following procedures can be used to check whether the solution is concentrated or not?	
	A.	Taste the solution to see of it is salty enough.
	В.	Observe to see if the solution turns milky.
	(C.)	Dip a glass rod into the solution, take it out to see if small
		crystals appear on the rod.
	D.	Measure the volume of the solution to see if it decreases at
		least by half of the initial volume.
1	16 17	gram of salt to volum of water ratio $\frac{30}{100} = 0.3$
4.		on cm ³ of water can dissolve a maximum of 30 g of salt Y at m temperature, which of the following would produce a > 0.3
		rated solution?
		Saturanted
	(1)	Add 20 g of salt Y to 50 cm ³ water at room temperature.
	(2)	Add 50 g of salt Y to 100 cm ³ water at room temperature. $\frac{50}{100} = 0.5$
	(3)	Add 40 g of salt Y to 200 cm ³ water at room temperature. $\frac{40}{200} = 0.2$



- (1) and (2) only
- B. (1) and (3) only
- C. (2) and (3) only
- D. (1), (2) and (3)
- 5. Which of the following statements concerning a saturated sodium chloride solution is correct?
 - (1) The boiling point of the solution is at 100°C.
 - (2) The amount of sodium chloride dissolved in water is the maximum at that particular temperature. \checkmark
 - (3) The solution is so concentrated that it cannot dissolve any other substances. \checkmark



- A. (1) only
- B. (2) only
- C. (1) and (3) only
- D. (2) and (3) only