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L1-CHEMR





## Level 1 Chemistry, 2019

9.30 a.m. Monday 18 November 2019

### RESOURCE BOOKLET

Refer to this booklet to answer the questions in your Question and Answer Booklets.

Check that this booklet has pages 2–4 in the correct order and that none of these pages is blank.

YOU MAY KEEP THIS BOOKLET AT THE END OF THE EXAMINATION.

### **Activity series**

Ca Mg Al Zn Fe Pb (H) Cu Ag

### Colours of selected ions and solids

Colourless ions	chloride, iodide, sulfate, hydroxide, carbonate, calcium, magnesium, zinc, lead, barium, silver
Blue ions	copper
Pale green ions	iron(II)
White solids	calcium sulfate, calcium hydroxide, calcium carbonate, magnesium hydroxide, magnesium carbonate, zinc carbonate, lead chloride, lead sulfate, lead carbonate, barium sulfate, barium hydroxide, barium carbonate, silver chloride
Green solid	iron(II) hydroxide, iron(II) carbonate
Blue solid	copper hydroxide
Yellow solid	lead iodide
Cream solid	silver iodide

### Solubility rules

nitrates	All soluble
chlorides	All soluble except silver chloride, lead chloride
iodides	All soluble except silver iodide, lead iodide
sulfates	All <b>soluble</b> except barium sulfate, lead sulfate, calcium sulfate
hydroxides	All insoluble except potassium hydroxide, sodium hydroxide
carbonates	All <b>insoluble</b> except potassium carbonate, sodium carbonate

### Table of ions

+1	+2	+3	-3	-2	-1
NH <sub>4</sub> <sup>+</sup>	Ca <sup>2+</sup>	A1 <sup>3+</sup>		O <sup>2-</sup>	OH_
Na <sup>+</sup>	Mg <sup>2+</sup>	Fe <sup>3+</sup>		S <sup>2-</sup>	Cl <sup>-</sup>
K <sup>+</sup>	Cu <sup>2+</sup>			CO <sub>3</sub> <sup>2-</sup>	Г
Ag <sup>+</sup>	Pb <sup>2+</sup>			SO <sub>4</sub> <sup>2-</sup>	NO <sub>3</sub>
$H^{+}$	Fe <sup>2+</sup>				HCO <sub>3</sub>
Li <sup>+</sup>	Ba <sup>2+</sup>				F-
	Zn <sup>2+</sup>				

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# PERIODIC TABLE OF THE ELEMENTS

18 2 <b>He</b>	10 Ne	18 <b>Ar</b>	36 <b>Kr</b>	54 Xe	86 <b>Rn</b>	0g
17	9 F	17 CI	35 Br	53 I	85 At	117 Ts
91	© 8	16 S	34 Se	52 Te	84 <b>Po</b>	116 Lv
15	Z	15 <b>P</b>	33 As	51 <b>Sb</b>	83 <b>Bi</b>	115 Mc
14	) 9	14 <b>Si</b>	32 Ge	50 <b>Sn</b>	82 <b>Pb</b>	114 F1
13	5 <b>B</b>	13 Al	31 <b>Ga</b>	49 In	81 TI	113 Nh
		12	30 <b>Zn</b>	48 Cd	80 Hg	112 Cn
		II	29 Cu	47 <b>Ag</b>	97 <b>Au</b>	Rg
		10	28 <b>Z</b> i	46 <b>Pd</b>	78 <b>Pt</b>	110 <b>Ds</b>
		6	27 Co	45 Rh	77 Ir	109 <b>Mt</b>
	1	8	26 Fe	44 <b>Ru</b>	92 Os	108 <b>Hs</b>
H H		7	25 Mn	43 Te	75 Re	107 <b>Bh</b>
Atomic number		9	24 C <b>r</b>	42 Mo	74 <b>W</b>	106 S <b>g</b>
Atomic		8	23 V	41 Nb	73 Ta	105 <b>Db</b>
		4	22 <b>Ti</b>	40 <b>Zr</b>	72 Hf	104 <b>Rf</b>
		S.	21 Sc	39 Y	71 <b>Lu</b>	103 <b>Lr</b>
2	4 <b>Be</b>	12 Mg	20 Ca	38 Sr	56 <b>Ba</b>	88 <b>Ra</b>
<i>F</i>	3 Li	11 Na	19 <b>K</b>	37 Rb	55 Cs	87 Fr

57	58	59	<b>PN</b>	61	62	63	64	65	66	67	68	69	70
La	Ce	Pr		<b>Pm</b>	Sm	Eu	<b>Gd</b>	Tb	<b>Dy</b>	<b>Ho</b>	Er	Tm	<b>Yb</b>
89	90	91	92	93	94	95	96	97	98	99	100	101	102
<b>Ac</b>	Th	<b>Pa</b>	U	<b>Np</b>	<b>Pu</b>	Am	Cm	<b>Bk</b>	Cf	Es	Fm	<b>Md</b>	<b>No</b>