

Atomic number	Element	Group	Electron configuration	Lose or gain electrons to become stable	Charge on ion	Electron configuration of the ion
1	Hydrogen	1	He(2)	Lose 1e ⁻	+1	H ⁺ (0)
2	Helium	2	He(2)	n/a	n/a	n/a
3	Lithium	1	Li(2,1)	Lose 1e ⁻	+1	Li ⁺ (2)
4	Beryllium	2	Be(2,2)	Lose 2e ⁻	+2	Be ²⁺ (2)
5	Boron	13	B(2,3)	Lose 3e ⁻	+3	B3 ⁺ (2)
6	Carbon	14	C(2,4)	n/a	n/a	n/a
7	Nitrogen	15	N(2,5)	Gains 3e ⁻	-3	N ³⁻ (2,8)
8	Oxygen	16	O(2,6)	Gains 2e ⁻	-2	O2 ⁻ (2,8)
9	Fluorine	17	F(2,7)	Gains 1e ⁻	-1	F ⁻ (2,8)
10	Neon	18	Ne(2,8)	n/a	n/a	n/a
11	Sodium	1	Na(2,8,1)	Lose 1e ⁻	+1	Na ⁺ (2,8)
12	Magnesium	2	Mg(2,8,2)	Lose 2e ⁻	+2	Mg ²⁺ (2,8)
13	Aluminium	13	Al(2,8,3)	Lose 3e ⁻	+3	Al ³⁺ (2,8)
14	Silicon	14	Si(2,8,4)	n/a	n/a	n/a
15	Phosphorus	15	P(2,8,5)	Gains 3e ⁻	-3	P ³⁻ (2,8,8)
16	Sulfur	16	S(2,8,6)	Gains 2e ⁻	-2	S ²⁻ (2,8,8)
17	Chlorine	17	Cl(2,8,7)	Gains 1e ⁻	-1	Cl ⁻ (2,8,8)
18	Argon	18	Ar(2,8,8)	n/a	n/a	n/a
19	Potassium	1	K(2,8,8,1)	Lose 1e ⁻	+1	K ⁺ (2,8,8)
20	Calcium	2	Ca(2,8,8,2)	Lose 2e ⁻	+2	Ca ²⁺ (2,8,8)