Atomic number	Element	Group	Electron arrangement	Lose or gain electrons to become stable	Charge on ion	Electron arrangement of the ion
1	Hydrogen	1	He(2)	Lose $1e^-$	+1	$H^+(0)$
2	Helium	2	He(2)	n/a	n/a	n/a
3	Lithium	1	Li(2,1)	Lose $1e^-$	+1	$Li^+(2)$
4	Beryllium	2	Be(2,2)	Lose $2e^-$	+2	$Be^{2+}(2)$
5	Boron	13	B(2,3)	Lose $3e^-$	+3	$B^{3+}(2)$
6	Carbon	14	C(2,4)	n/a	n/a	n/a
7	Nitrogen	15	N(2,5)	Gains $3e^-$	-3	$N^{3-}(2,8)$
8	Oxygen	16	O(2,6)	Gains $2e^-$	-2	$O^{2-}(2,8)$
9	Fluorine	17	F(2,7)	Gains $1e^-$	-1	$F^-(2,8)$
10	Neon	18	Ne(2,8)	n/a	n/a	n/a
11	Sodium	1	Na(2,8,1)	Lose $1e^-$	+1	$Na^+(2,8)$
12	Magnesium	2	Mg(2,8,2)	Lose $2e^-$	+2	$Mg^{2+}(2,8)$
13	Aluminium	13	Al(2,8,3)	Lose $3e^-$	+3	$Al^{3+}(2,8)$
14	Silicon	14	Si(2,8,4)	n/a	n/a	n/a
15	Phosphorus	15	P(2,8,5)	Gains $3e^-$	-3	$P^{3-}(2,8,8)$
16	Sulfur	16	S(2,8,6)	Gains $2e^-$	-2	$S^{2-}(2,8,8)$
17	Chlorine	17	Cl(2,8,7)	Gains $1e^-$	-1	$Cl^-(2,8,8)$
18	Argon	18	Ar(2,8,8)	n/a	n/a	n/a
19	Potassium	1	K(2,8,8,1)	Lose $1e^-$	+1	$K^+(2,8,8)$
20	Calcium	2	Ca(2,8,8,2)	Lose $2e^-$	+2	$Ca^{2+}(2,8,8)$