

Isotopes and Atomic Mass

Fill in the missing information as outlined in the examples

Standard Atomic Notation	Short Form	Atomic Number	Mass Number	Protons	Neutrons	Electrons
$^{12}_6\text{C}$	C-12	6	12	6	6	6
$^{13}_6\text{C}$	C-13	6	13	6	7	6
He						
Be					4	
Ca			40			
O	O-18			8		
F					8	
P			32			15
Fm					132	
Ne				10	18	
S					17	16
Al	Al-26					
H		1	3			

Calculate the average atomic mass.

- A sample of bromine consists of two isotopes of mass numbers 79 and 81. Isotope 79 have an abundance of 50.69% and Isotope 81 is 49.31% in abundance. What is bromine's average atomic mass?
- Silicon is found in nature with 3 stable isotopes. Si – 28 at 92.23%, Si – 29 at 4.67% and Si – 30 at 3.10%. What is silicon's average atomic mass?
- Calculate the average atomic mass for each element:
 - Li-6 , 7.42% ; Li-7 , 92.58%
 - Ne-20 , 90.92% ; Ne-21 , 0.257% ; Ne-22 , 8.82%
 - Cr-50 , 4.35% ; Cr -52 , 83.79% ; Cr -53 , 9.50% ; Cr-54 , 2.36%