

Physical and Chemical Properties of Hydrocarbons

① Physical Properties pg 22

Intermolecular Forces van der Waals

(London Dispersion Force)

e^- of 1 molecule attracted to nuclei of another molecule

mp./b.p - low because of intermolecular forces

- as the molecules get bigger the boiling points increase because the number of e^- increases

Solubility - Non polar

- do not dissolve in water

- good solvent for non polar chemicals