

IONIC COMPOUNDS: Names and Formulas

1. Write the formulas for the following compounds.

a. magnesium oxide _____

b. aluminum nitride _____

c. potassium sulfide _____

d. calcium bromide _____

e. aluminum sulfide _____

f. beryllium oxide _____

g. strontium phosphide _____

h. sodium fluoride _____

i. lithium nitride _____

j. barium oxide _____

k. tin (II) fluoride _____

l. lead (IV) nitride _____

m. iron (III) chloride _____

n. copper (I) oxide _____

o. gold (III) sulfide _____

p. mercury (II) oxide _____

q. tin (IV) iodide _____

r. copper (II) phosphide _____

s. cobalt (II) sulphide _____

t. tin (IV) sulphide _____

2. Write the names for the following compounds.

a. Li_2O _____

b. AlCl_3 _____

c. MgS _____

d. CaF_2 _____

e. Al_2O_3 _____

f. BeF_2 _____

g. K_3P _____

h. Mg_3P_2 _____

i. CaO _____

j. Ag_2S _____

k. PbS _____

l. SnO_2 _____

m. NiO _____

n. CuI_2 _____

o. PbCl_4 _____

p. FeP _____

q. AuBr_3 _____

r. Hg_2S _____

s. SbF_3 _____

t. MnO_2 _____

POLYATOMIC COMPOUNDS: Names and Formulas

3. Write the formulas for the following compounds.

a. magnesium carbonate _____

b. aluminum nitrate _____

c. potassium sulfate _____

d. calcium hydroxide _____

e. aluminum sulfate _____

f. sodium carbonate _____

g. strontium phosphate _____

h. sodium hydrogen carbonate _____

i. lithium nitrate _____

j. aluminum hydroxide _____

k. tin (II) hydrogen carbonate _____

l. lead (IV) nitrate _____

m. iron (III) carbonate _____

n. copper (II) hydroxide _____

o. lead (II) nitrate _____

p. mercury (II) hydrogen sulfate _____

q. tin (IV) phosphate _____

r. lead (IV) hydroxide _____

s. potassium nitrate _____

t. tin (IV) sulphate _____

4. Write the names for the following compounds.

a. Li_2SO_4 _____

b. $\text{Al}(\text{HCO}_3)_3$ _____

c. MgSO_4 _____

d. K_2CO_3 _____

e. Na_2SO_4 _____

f. AgNO_3 _____

g. K_3PO_4 _____

h. $\text{Sr}(\text{HSO}_4)_2$ _____

i. RbOH _____

j. NaHCO_3 _____

k. PbSO_4 _____

l. AuOH _____

m. GaPO_4 _____

n. $\text{Cu}(\text{NO}_3)$ _____

o. $\text{Pb}(\text{HSO}_4)_4$ _____

p. $\text{Fe}(\text{HCO}_3)_3$ _____

q. Au_2CO_3 _____

r. HgOH _____

s. $\text{Sb}_2(\text{SO}_4)_3$ _____

t. MnSO_4 _____

MOLECULAR COMPOUNDS: Names and Formulas

5. Write the formulas for the following compounds.

a. carbon dioxide _____

b. silicon dioxide _____

c. water _____

d. carbon disulphide _____

e. sulphur trioxide _____

f. carbon tetrachloride _____

g. sulphur dioxide _____

h. dinitrogen tetroxide _____

i. nitrogen monoxide _____

j. arsenic tribromide _____

k. diphosphorus trisulphide _____

l. dinitrogen monoxide _____

m. dichlorine monoxide _____

n. bromine gas _____

o. carbon monoxide _____

p. xenon tetrafluoride _____

q. neon gas _____

r. silicon tetrahydride _____

s. iodine heptachloride _____

t. krypton difluoride _____

6. Write the names for the following compounds.

a. CF_4 _____

b. NH_3 _____

c. PBr_3 _____

d. F_2 gas _____

e. CS_2 _____

f. CO _____

g. SiC _____

h. N_2O_4 _____

i. P_2O_5 _____

j. SF_4 _____

k. NF_3 _____

l. P_2S_5 _____

m. PF_5 _____

n. ICl _____

o. SeCl_2 _____

p. Cl_2O _____

q. AsBr_3 _____

r. H_2S _____

s. B_2H_8 _____

t. TeCl_2 _____

PUTTING IT ALL TOGETHER:Names and Formulas

7. Write the formulas for the following compounds.

a. calcium fluoride _____

b. carbon disulfide _____

c. nitrogen triiodide _____

d. sodium phosphide _____

e. dichlorine monoxide _____

f. iron (III) carbonate _____

g. manganese (IV) sulphate _____

h. diphosphorus pentasulphide _____

i. tin (IV) chloride _____

j. magnesium hydrogen carbonate _____

k. potassium sulphate _____

l. barium nitride _____

m. aluminum hydroxide _____

n. fluorine gas _____

o. silicon dioxide _____

p. calcium hydroxide _____

q. xenon gas _____

r. gold (I) nitrate _____

s. sulphur trioxide _____

t. iron (II) phosphate _____

8. Write the names for the following compounds.

a. CCl_4 _____

b. $\text{Mg}(\text{HSO}_4)_2$ _____

c. PBr_3 _____

d. H_2 gas _____

e. PbS_2 _____

f. $\text{Al}_2(\text{CO}_3)_3$ _____

g. Na_2SO_4 _____

h. Na_2O _____

i. $\text{Al}_2(\text{SO}_4)_3$ _____

j. Li_2SO_4 _____

k. NaNO_3 _____

l. PCl_5 _____

m. BiF_5 _____

n. $\text{Al}(\text{HCO}_3)_3$ _____

o. FeCl_2 _____

p. N_2O _____

q. CuHSO_4 _____

r. Li_3PO_4 _____

s. SnO _____

t. SeCl_2 _____