Comparing Strong Acid Strong Base titration to Weak Acid Strong Base titration

Ex. In a titration, 20.00 mL of 0.300 mol/L HCl(aq) is titrated with standardized 0.300 mol/L NaOH(aq). What is the amount of unreacted HCl(aq) and the pH of the solution: (a) before titration begins (b) after 10.00 mL of NaOH is added (c) after 20 mL of NaOH is added.	E.g. In a titration of 20.00 mL of 0.300 mol/L $HC_2H_3O_2$(aq) (acetic acid) with standardized 0.300 mol/L NaOH(aq), what is the amount of unreacted $HC_2H_3O_2$(aq) and the pH of the solution: (a) before titration begins (b) after 10.00 mL of NaOH is added (c) after 20 mL of NaOH is added.