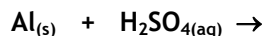


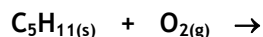
Stoichiometry Worksheet #2**ON YOUR OWN SHEET OF PAPER!!!**

1) Balance the chemical equation and use it to solve for the following problems:



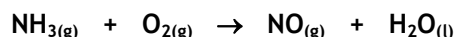
- If 58.83 g of Aluminum is completely reacted, how many grams of Hydrogen sulfate are needed? (320.3g H_2SO_4)
- If 45.60 L of Hydrogen gas forms at STP, what mass of Hydrogen sulfate was used? (199.5g H_2SO_4)
- How many grams of Aluminum were needed to produce 107.72 g of Aluminum sulfate? (17.008g Al)
- What volume of hydrogen gas will be produced @ STP from the complete rxn of 36.1g of Aluminum? (44.9 L H_2)

2) Balance the chemical equation and use it to solve for the following complete combustion:



- If 180.3 g of C_5H_{11} is burned how many Liters of CO_2 are produced? (284.4 L CO_2)
- If 63.4 Liters of carbon dioxide gas forms at STP, how many Liters of oxygen were used? (98.3 L O_2)
- How many grams of oxygen gas were needed to produce 24.5 grams of water vapor? (61.4 g O_2)
- What mass of C_5H_{11} when burned will produce 74 grams of H_2O vapor? (53 g C_5H_{11})

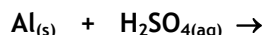
3) Balance the chemical equation and use it to solve for the following problems:



- How many grams of NO were needed to produce 30.2 g of water? (33.6 g NO)
- What volume of O_2 gas is required to produce 34.0 Liters of NO? (42.5 L O_2)
- If 115.3 g of water were produced, how many grams of Ammonia were used? (72.60 g NH_3)
- We want to collect 5 Liters of water, what volume of Ammonia will we need to start with? (4000 L NH_3)

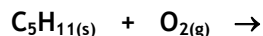
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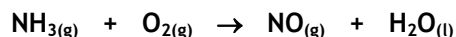
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