

Review of the Atom

Protons - positively charged particles with a mass of 1

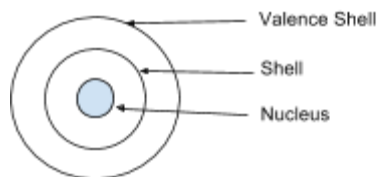
Neutrons - no charge particle with a mass of 1

Electrons - negatively charged particle with a mass of approximately zero

Nucleus - has protons and neutrons

Shells - location of the electrons (circles drawn in Bohr Rutherford Diagrams)

Valence Shell - outer shell of electrons - involved in bonding



Review of Periodic Table

Rows - also called periods

- identifies the number of shells the atom contains

Columns - also called groups

- identifies the number of valence electrons found on the valence shell

Atomic Number - identifies the number of protons and the total number of electrons

Atomic Mass or Mass Number - average mass of all the atoms

- Round the atomic mass to a whole number
- identifies the number of protons and neutrons in the nucleus
- Neutrons = atomic mass - atomic number

Four Special Groups

1. Alkali Metals - group 1
2. Alkaline Earth Metals - group 2
3. Halogens - group 17
4. Noble Gases - group 18

Two main classes of Elements

1. Metals
 - a. Left side of table
 - b. conduct electricity
 - c. malleable
 - d. shiny
2. Non Metals
 - a. right side of table
 - b. do not conduct electricity
 - c. brittle
 - d. dull