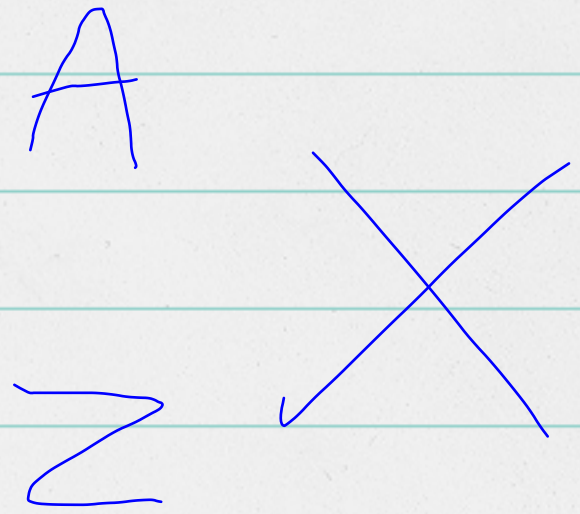
Date March 7

1. Math Basics
   - Significant digits
   - Scientific Notation

2. Calculating Atomic Mass
   - protons
   - neutrons
   - electrons

3. Standard Atomic Notation



Date March 8

1. Atomic Mass
   - not always the periodic table values
2. Isotopes
   - Define
   - calculating

Date March 9

1. Compounds + Mass

Molecules — molecular mass

Ionic Compounds — formula unit mass

Elements - atomic mass

2. The Mole

- Molar mass

Date March 19

1. Unit Test Returned
   - discuss problems
   - 12 % of course is done

2. Mole + Compounds

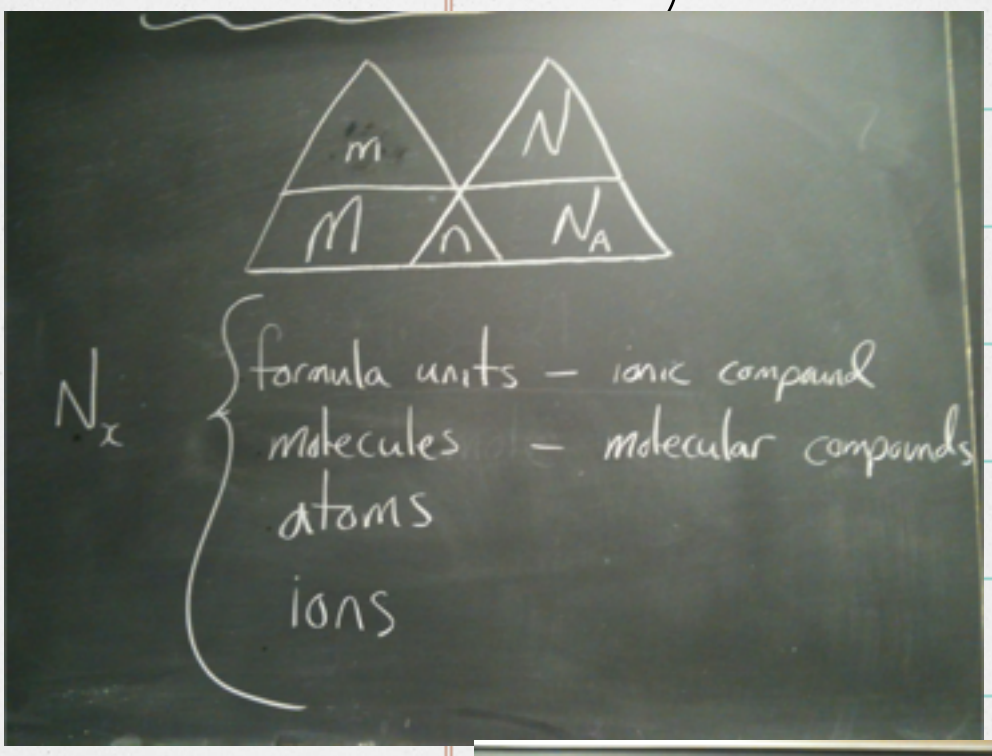
3. Mole calculations

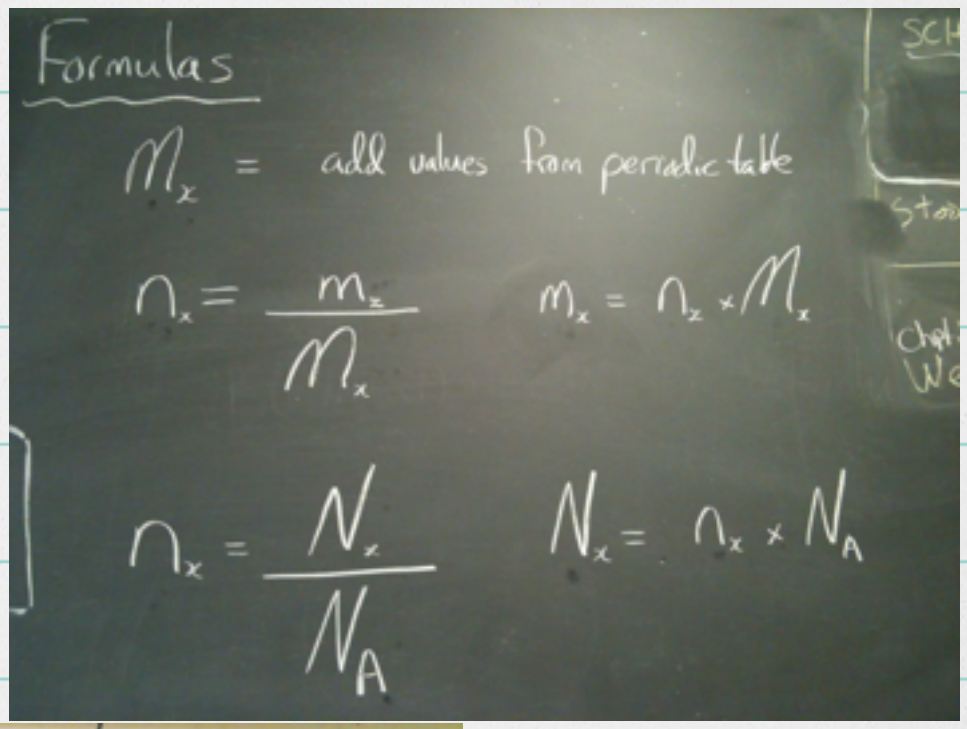
Date March 20

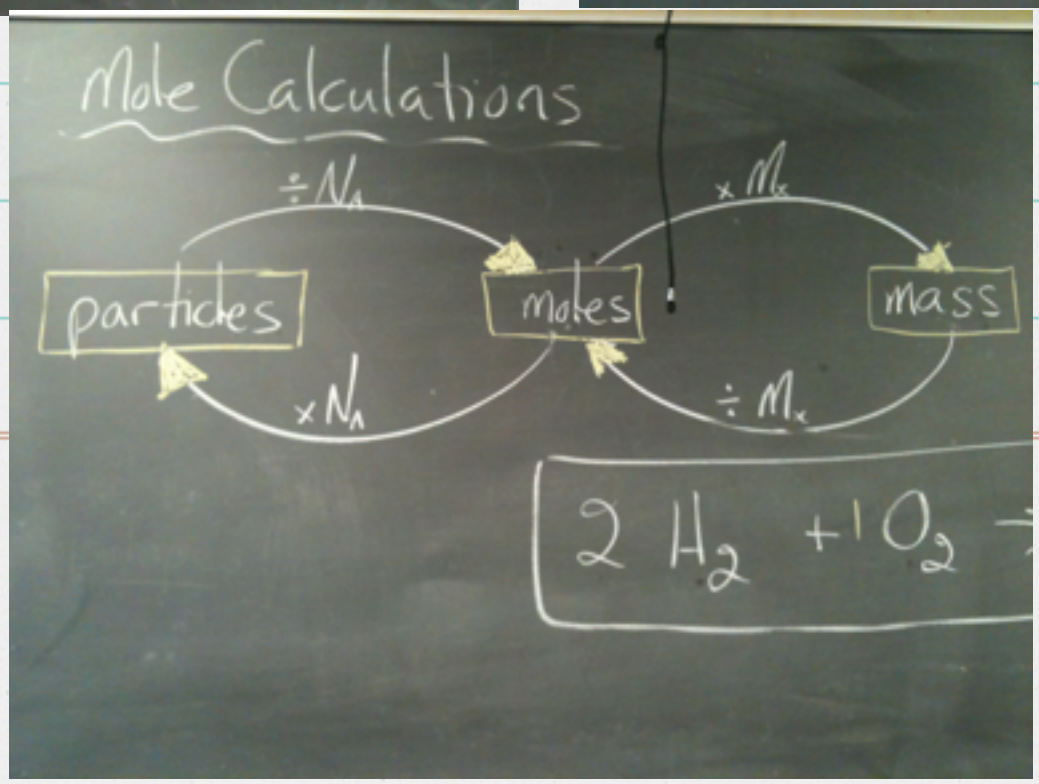
1. Mole Calculations

2. Worksheet - molar mass + 1 step calculations

3. Page 106 - Homework







Date March 21

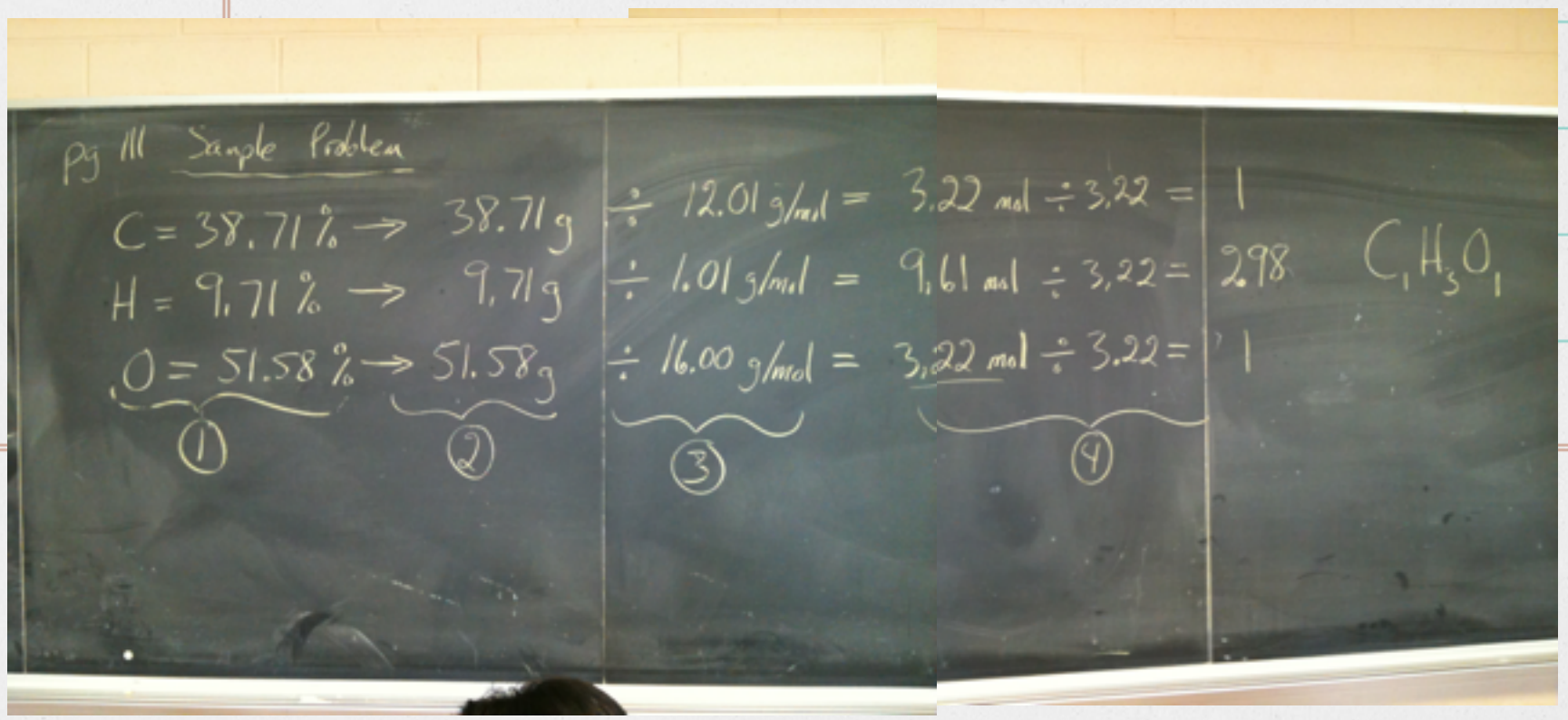
1. Page 106 - concerns

2. Determining Chemical Formulas

(A) Percent Composition

- every compound can be described in terms of their element abundance in percent
- determined by a combustion analyzer

(B) Empirical Formula vs Molecular Formula

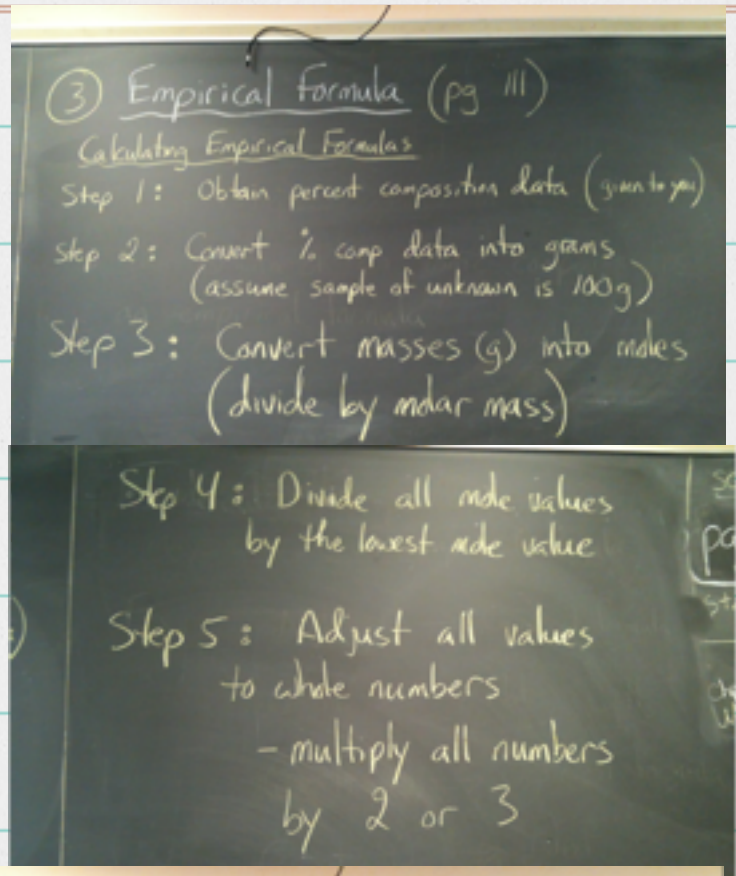


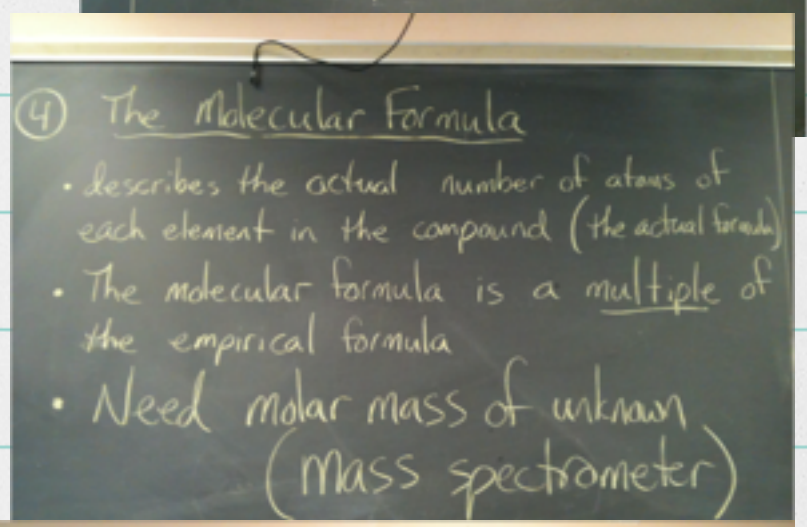
Date March 22

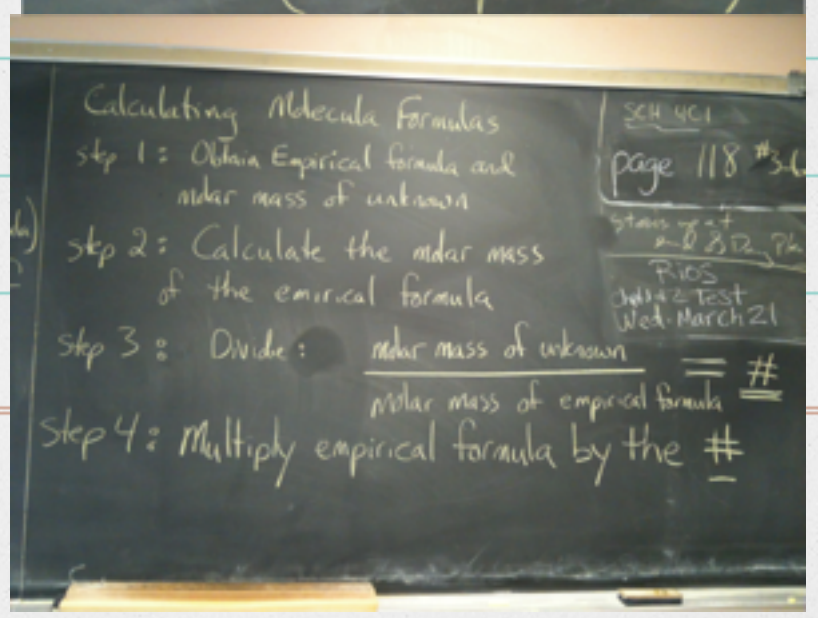
1. Empirical Formula
   - Define
   - Steps for calculation

2. Molecular Formula
   - Define
   - steps for calculation

3. Textbook Questions
   - page 114
   - page 118
   - page 120

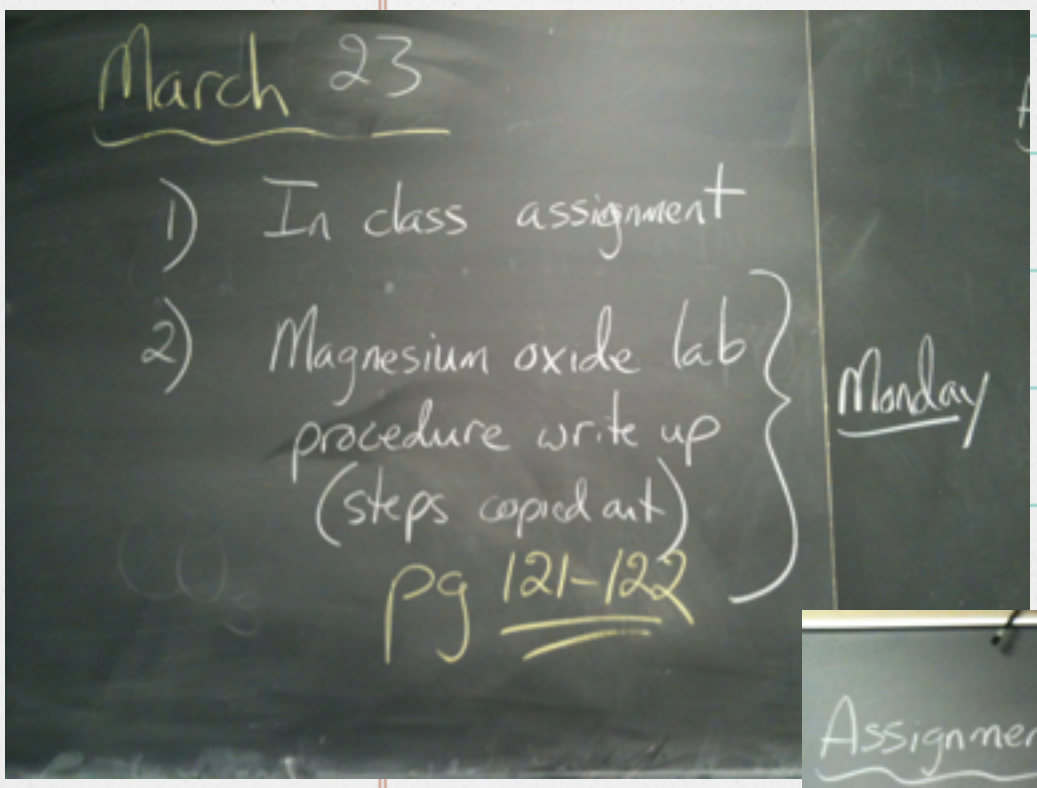


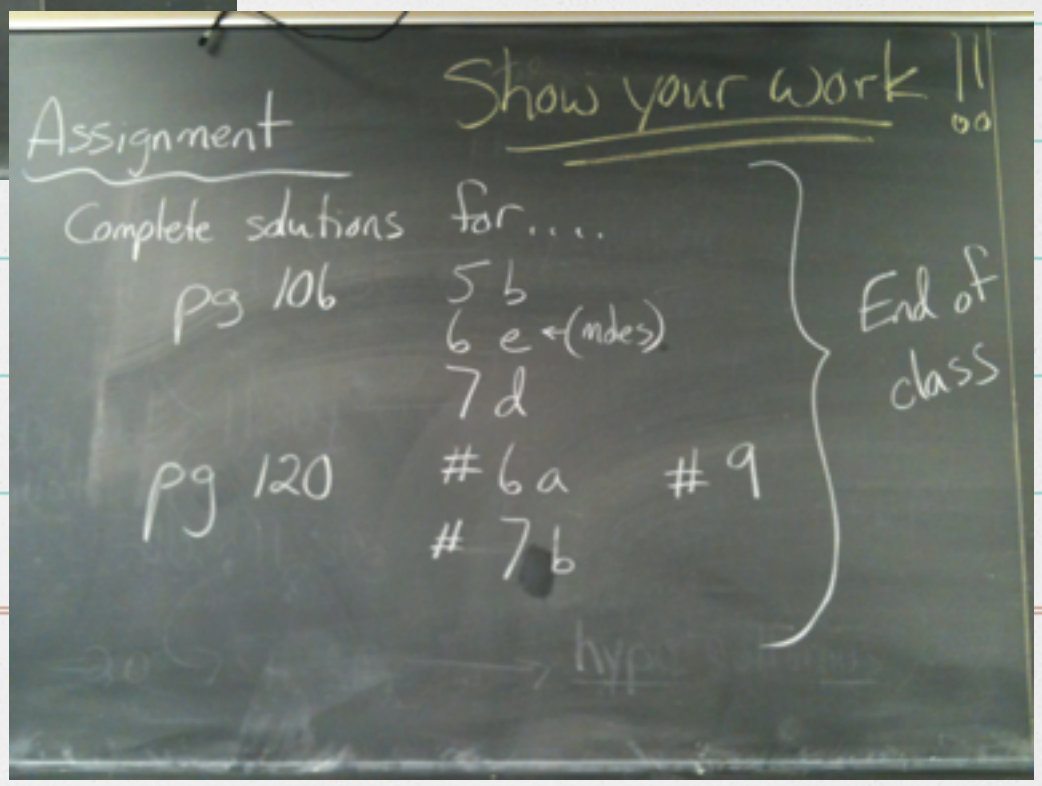




Date March 23

1. In class assignment
   - page 120 #6-8, page 106 7d, 6e, 5b
2. Read MgO lab procedure
   - write out steps





Date March 26

1. Bunsen Burner Instructions

2. MgO procedure review / MgO Lab

Write up: pg 121 Friday !!

1. Title page Name, Lab Title, Due Date

2. Purpose, hypothesis (prediction - a), Periodic Table
procedure(copied with changes),
observations (qualitative and quantitative), ← TABLES
calculations (analysis part d, e),
analysis (b, c, h),
evaluation (f, g, i, j, percent error calculation)
pg 107

3. Procedure (if copied before lab)

4. Lab sheet

Date March 27

1. Discus Lab?

2. Concentration Calculations

-page 123

- Define, equation (general)

$$C = \frac{\text{Solute}}{\text{Solution}}$$

3. Specific Equations

%Conc PPM/PPB/PPT Molar

- Summary page 13

4. Go through Textbook section complete 7-12 questions

Date March 28

pg 123-133 Textbook questions

1. Concentration
   - % Concentration
   - ppm
   - Molar Concentration

work period + work on lab

2. Dilutions
   - more solvent, same solute
   - mole initial = mole final

Friday

$C \times V = n_i$ $\quad$ $C \times V = n_f$

$C_i V_i = C_f V_f$

Date March 28 + March 30th

1. Hand in Lab

2. Dilutions

3. Worksheet → Textbook Questions

4. Read over Lab $CuSO_4$ Spectrophotometer

Nope

→ Making Solutions

1) From a solid

$n = c \times v$, $m = n \times M_x$

2) From dilutions

$C_1V_1 = C_2V_2$