CHAPTER 10 SUMMARY

MAKE A SUMMARY

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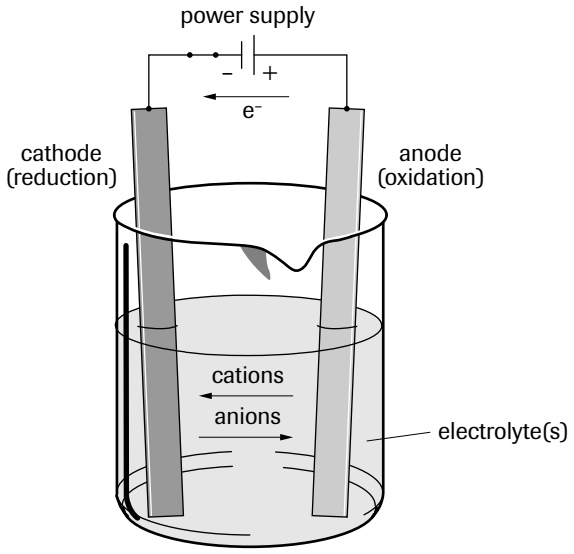

- The minimum electrolytic cell potential is determined by subtracting the $E_r°$ (reduction potential) of the strongest reducing agent present from the $E_r°$ (reduction potential) of the strongest oxidizing agent present, using a standard table of reduction potentials.

  Note: These tabulated values are established for standard half-cells, using aqueous electrolytes, at SATP. Any variation from these conditions will make the predicted value less accurate; and most electrolytic cells are not operated under these conditions. For electrolytic cells using molten electrolytes, these tabled values are essentially meaningless.

Category	Description	Examples
rechargeable batteries	reversing a secondary cell to charge it	batteries for phones and laptop computers
element production	producing elements from naturally occurring compounds	aluminum from aluminum oxide to make aluminum cans
metal refining	producing pure metals from impure metals	copper for electrical wiring
electroplating	covering an object with a layer of metal	chromium plating of faucets

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