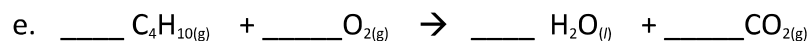
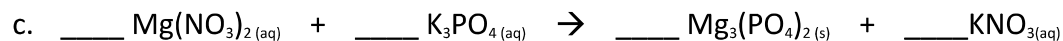
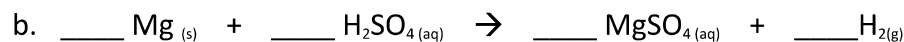
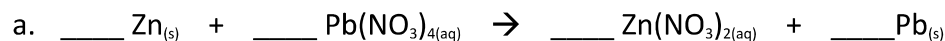


1. Balance the following chemical reactions: (5 x 2 marks each - 10 marks)



2. For each of the following reactions NEATLY write the: (4 x 4 marks each – 16 marks)

i) Balanced Chemical Equation (including states) ii) Type of Reaction

a. calcium metal + dihydrogen monoxide \rightarrow aqueous calcium hydroxide + hydrogen gas

Type: _____

b. iron metal + oxygen gas \rightarrow solid iron(III) oxide

Type: _____

c. solid calcium carbonate \rightarrow solid calcium oxide + carbon dioxide gas

Type: _____

d. calcium chloride solution + silver nitrate solution \rightarrow solid silver chloride + aqueous calcium nitrate

Type: _____

3. For each of the following reactions, write the word equation (no states are required here) (2 x 2 marks each – 4 marks)

