

8.5 Acid-Base Reactions

Definitions

- Titration
- Titrant
- End point
- Standard solution

Types of Reactions

- Metal + acid \rightarrow hydrogen gas + ionic compound
- E.g.
- Acid + carbonate \rightarrow carbonic acid \rightarrow carbon dioxide + water + ionic compound
- E.g.
- Strong acid + strong base \rightarrow water + ionic compound
- E.g.
- Acid + ionic compound \rightarrow precipitate + acid
- E.g.

Acid Base Titration

- Titration is a lab technique used to determine concentration.
- You add a known volume of a known concentration to determine the unknown.
- The known in the buret is called the titrant.
- You add titrant until it reaches the endpoint (does not have to be pH 7)
- The titrant is a standard solution.
- Know the summary on page 395.

Sample Problems

- We will work through the examples on the chalkboard.

Homework

- Practice Q's: 1-10
- Section Q's: 1-3
- Prepare for titration lab