6.2 Explaining Solutions

Definitions
- Intermolecular force
- Intramolecular force
- Dissociation

Explaining Water Mixtures
- Intermolecular forces need to be used to explain solubility.

Molecular Substances in Water
- Like dissolves like.
- Water slightly polar and capable of hydrogen bonds.

The Effects of Polarity of Hydrogen Bonds
- Dipole-dipole bonds
- Hydrogen bonds
- E.g. Hydrogen bonding and dipole bonding with water (ammonia and water)

Ionic Compounds in Water
- E.g. Sodium chloride and water.

Explaining Non-aqueous Mixtures
- London dispersion forces hold non-polar in solution
- E.g. Two different non-polar molecules.

Water – "The Universal Solvent"
- Water is nearly a perfect solvent.

Homework
- Practice 1-12
- Section Questions 1-5