

#### Sept 4

1. Welcome
2. Course outline
3. Textbook
4. Resources
5. Grade 10 review

#### Sept 5

1. Concerns of website? Review topics?
2. Safety expectations - understanding
  - A. HHPs, MSDS, WHMIS
  - B. Lab safety - common sense and some new situations
  - C. Safety equipment
  - D. Disposal
3. Textbook safety information
4. Working on review / safety notes

#### Sept 6

1. Matter flowchart
2. Element list - location on periodic table

#### Sept 7

1. Take up "things you should know from grade 10"
2. Atomic structure - Bohr Rutherford and Lewis diagrams
3. Lighting a Bunsen burner
4. Complete safety and element list

#### Sept 10

1. Numbers of the periodic table - relate to electrons, protons, neutrons
  - A. Charged particles vs neutral particles
2. Isotopes and isotopic abundance calculation

#### Sept 11

1. Elements organized video and worksheet
2. Get ready for quiz safety and element

Sept 12

1. Safety and naming quiz
2. Take up safety as a class. Collect element quiz part
3. History of the periodic table
4. Structure of the periodic table
5. Activity looking at figure 8 and 9 page 18+19
  - a. (adding info to blank periodic table)

Sept 13

1. Lab prep for "element or compound"
2. Finish up work from previous days - history, parts of periodic table

Sept 14

1. Element or Compound lab
2. Expectations handout

Sept 17

1. Periodic table review
2. Theory, model, theoretical vs empirical
3. History of atom watching TVO but no video sheets
4. Democritus, Aristotle, Empedocles, alchemists, Dalton, Thomson, Chadwick, Rutherford (making notes after reading pg 22-25)

Sept 18

1. Finish TVO videos up to Bohr
2. Finish notes with video additions

Sept 19

1. Bohr theory - TVO video
2. Spectroscopes

Sept 20

1. Finish history
2. Hint at quantum and electron configuration
3. Quantum number one = period number

Sept 21

PD day no classes

Sept 24

- I. Trends in the periodic table
  - A. Reactivity
  - B. Atomic radius
  - C. Ionic radius
  - D. Ionization energy
  - E. Electron affinity
  - F. Electronegativity
2. Examine images page 50-52

Sept 25

1. Review activity with periodic trends and organization
2. Classifying compounds
  - A. Read 66-71
  - B. Clues to determine the type of compound

Sept 26

- A. Ionic compounds
- B. Properties
- C. Bonding

Sept 27

1. Ionic bonding examples
2. Molecular compounds
3. Covalent bonding

Sept 28

- I. Lewis structures
  - A. Page 79
  - B. Questions page 79-80

Oct 1

Polar, nonpolar, electronegativity

Oct 2

Intermolecular forces

Oct 3

Review period

Oct 4

Unit test

Oct 5

1. Nomenclature
2. IUPAC
3. Binary ionic compounds
  - A. Representative elements
  - B. Oxidation numbers
  - C. Net charge = 0
  - D. Formula crossover
  - E. Name "ide"
  - F. Complete section A

Oct 9

1. Quiz on A
2. Transition metals
  - A. stock vs classical
  - B. Net charge = 0

Oct 10

1. Quiz part A-B
2. Non metals
3. Rest of part I

Oct 11

1. Quiz part A-G
2. Oxy acids
3. Per, ous, hypo-ous
4. Polyatomic ions
5. Per, ite, hypo-ite

Oct 12

1. Quiz part I-J
2. Poly atomic ions
3. Acid poly atomic ions

Oct 15

1. Quiz part I-L5
2. Rest of part 2
3. Hydrates in textbook
4. Hardback unit I test

Oct 16

1. Quiz on M-N
2. Naming review

Oct 17

1. Unit 2 chemical reactions
2. Defining a chemical reaction - overhead notes
3. Reading 3.1

Oct 18

1. The standard format to a chemical reaction
2. Balancing chemical reactions

Oct 19

1. Quiz on part 2
2. Introducing types of chemical reactions
3. S, ox, d, sd, dd
4. Worksheet on balancing and identifying types of reactions

Oct 22

1. Nomenclature test
2. Making notes on 3.2 (little sheet)

Oct 23

Read: 3.3

Questions:

1. Define the single displacement reaction type and write a general reaction equation for it.
2. Describe, in general, what happens in many single displacement reactions.
3. Write 2 specific examples, one involving metals and one involving nonmetals, for this reaction type.
4. How are the metals and nonmetals arranged in an activity series (give 2 ways)?

5. Define alloy, corrosion, and galvanizing. Describe practical applications for each of these ideas.

Additional Questions:

p. 128-129, Q. 1-7; p. 130, Q. 8; p. 134, Q. 10-11; p. 143-144, Q. 3-5

Read: 3.4

Questions:

1. Define the double displacement reaction type and write a general reaction equation for it. Give an example of this reaction type.

2. Define solute, solvent, solubility, precipitate.

3. Describe 2 types of double displacement reactions. Write an example for each reaction type.

Additional Questions:

p. 138, Q. 1-3; p. 141-143, Q. 4-12

Oct 24

Review single and double displacement

Reactions in solution

Activity series

Solubility table

Oct 25

1. Nuclear reactions worksheet

2. Textbook and examples on the board

Oct 26

1. Word to balanced reactions