7.3 Reactions in Solution

Definitions

• Total ionic equations

• Spectator

• Net ionic equation

• (aq)

Collision theory

• Solubility tables

Disassociation

- Soap scum (calcium stearate) is difficult to remove. We need a cleaning product that disassociates the molecule into soluble ions that can be rinsed away.
- The production and removal of precipitates is a very important process to chemistry. Soap is a precipitate and many cleaning supplies are used to remove precipitates.

Collision Theory

- A chemical system consists of *particles* (atoms, ions, or molecules) that are in constant random motion at various speeds. The average kinetic energy of the particles is proportional to the temperature of the sample.
- A chemical reaction must involve *collisions of particles* with each other or the walls of the container.
- An *effective collision* is a collision between particles that has sufficient energy and correct orientation (alignment or positioning) of the colliding particles so that bonds can be broken and new bonds formed.
- *Ineffective collisions* involve particles that rebound from the collision, essentially unchanged in nature.

Net Ionic Equations

- Arrhenius theory of dissociation can be used to determine a reaction.
- E.g. balanced chemical equation

$$Pb(NO_3)_{2(aq)} + 2 KI_{(aq)} \rightarrow PbI_{2(s)} + 2 KNO_{3(aq)}$$

total ionic equation

$$Pb^{2+}_{(aq)} + 2NO_{3(aq)}^{-} + 2K^{+}_{(aq)} + 2I^{-}_{(aq)} \rightarrow PbI_{2(s)} + 2K^{+}_{(aq)} + 2NO_{3(aq)}^{-}$$

net ionic equation (spectators from the above reaction are removed)

$$Pb^{2+}_{(aq)} + 2I^{-}_{(aq)} \rightarrow PbI_{2(s)}$$

Summary from Textbook

- Write the balanced chemical equation with full chemical formulas for all reactants and products.
- Using solubility information, such as the table in Appendix C, rewrite the formulas for all high-solubility ionic compounds as dissociated ions, to show the total ionic equation.
- Cancel identical amounts of identical entities appearing on both reactant and product sides.
- Write the net ionic equation, reducing coefficients if necessary.

Homework

- Practice 1-8
- Section 1-3