

II. OXY ACIDS			
Write Formulas		Write Names	
1.	chloric acid	9.	HClO ₃
2.	phosphoric acid	10.	H ₃ PO ₄
3.	iodic acid	11.	H ₂ SO ₄
4.	nitric acid	12.	HFO ₃
5.	sulfuric acid	13.	HBrO ₃
6.	bromic acid	14.	HIO ₃
7.	carbonic acid	15.	HNO ₃
8.	fluoric acid	16.	H ₂ CO ₃

JJ. OXY-ACID DERIVATIVES			
Write Formulas		Write Names	
1.	nitric acid	26.	H_2CO_3
2.	nitrous acid	27.	H_2CO_4
3.	persulfuric acid	28.	HClO
4.	sulfuric acid	29.	HClO_4
5.	pernitric acid	30.	H_2CO
6.	sulfurous acid	31.	H_3PO_2
7.	hyponitrous acid	32.	HFO_4
8.	carbonous acid	33.	HClO_2
9.	hypoiodous acid	34.	H_3PO_4
10.	hyposulfurous acid	35.	HBrO_4
11.	perphosphoric acid	36.	HFO_3
12.	bromic acid	37.	HIO_4
13.	bromous acid	38.	HFO_2
14.	iodic acid	39.	HFO
15.	hypobromous acid	40.	HIO_2
16.	carbonic acid	41.	HNO_3
17.	hypochlorous acid	42.	H_2SO_5
18.	hypocarbonous acid	43.	HNO_4
19.	perfluoric acid	44.	HNO
20.	phosphoric acid	45.	HIO
21.	fluoric acid	46.	H_3PO_5
22.	hypofluorous acid	47.	HBrO_3
23.	iodous acid	48.	HBrO
24.	periodic acid	49.	H_2SO_3
25.	percarbonic acid	50.	H_3PO_3

KL. POLYATOMIC ANIONS - salts			
Write Formulas		Write Names	
1.	potassium bromate	26.	NaClO ₃
2.	silver iodate	27.	Li ₃ PO ₄
3.	zinc phosphate	28.	CaSO ₄
4.	cupric carbonate	29.	B(FO ₃) ₃
5.	lead(IV) sulfate	30.	Fe(BrO ₃) ₃
6.	beryllium chlorate	31.	Pb(IO ₃) ₄
7.	manganous bromate	32.	KNO ₃
8.	ferrous nitrate	33.	Sb ₂ (SO ₄) ₅
9.	tin(IV) iodate	34.	Sn(ClO ₃) ₂
10.	copper(II) carbonate	35.	HgSO ₄
11.	magnesium fluorate	36.	Ag ₂ CO ₃
12.	lithium phosphate	37.	Be(NO ₃) ₂
13.	aluminium sulfate	38.	SnSO ₄
14.	cobalt(II) phosphate	39.	Cu ₃ PO ₄
15.	ferric phosphate	40.	KIO ₃
16.	chromium(III) sulfate	41.	NaNO ₃
17.	arsenous carbonate	42.	Si(ClO ₃) ₄
18.	ammonium nitrate	43.	Ag ₃ PO ₄
19.	cesium chlorate	44.	NaFO ₃
20.	lead(II) phosphate	45.	SrCO ₃
21.	lithium fluorate	46.	Sn(SO ₄) ₂
22.	magnesium bromate	47.	Zn(BrO ₃) ₂
23.	mercuric sulfate	48.	AlPO ₄
24.	nickel(III) iodate	49.	Sb(IO ₃) ₅
25.	phosphorous carbonate	50.	Ba(NO ₃) ₂

LL. DERIVATIVE ANIONS - salts			
Write Formulas		Write Names	
1.	ferrous sulfite	26.	Na_2SO_4
2.	nickel(III) carbonate	27.	$\text{Fe}_3(\text{PO}_2)_2$
3.	mercuric phosphite	28.	KClO_2
4.	calcium perchlorate	29.	AgBrO
5.	radium nitrite	30.	NaClO
6.	cobaltous sulfate	31.	ZnCO_3
7.	manganese(IV) periodate	32.	$\text{Ba}(\text{NO}_2)_2$
8.	iron(II) hyponitrite	33.	Li_2SO_3
9.	lithium bromate	34.	$\text{Mg}_3(\text{PO}_5)_2$
10.	cesium iodite	35.	PbCO_3
11.	magnesium percarbonate	36.	$\text{Ca}(\text{ClO}_3)_2$
12.	barium hypophosphite	37.	FePO_4
13.	manganese(II) nitrite	38.	$\text{Sn}(\text{SO}_4)_2$
14.	tin(IV) sulfite	39.	AlPO_2
15.	magnesium periodate	40.	$\text{Pb}(\text{IO}_3)_2$
16.	aluminium hyposulfite	41.	Cu_3PO_3
17.	plumbic perphosphate	42.	HgSO_4
18.	stannous sulfate	43.	$\text{Sb}_2(\text{SO}_4)_5$
19.	antimonic hyponitrite	44.	CaCO_3
20.	copper(II) chlorate	45.	FeCO_2
21.	aluminum phosphite	46.	Cs_3PO_5
22.	antimony(III) persulfate	47.	CoSO_2
23.	barium hyponitrite	48.	CuNO_4
24.	beryllium carbonate	49.	$\text{Fe}(\text{ClO}_2)_2$
25.	calcium chlorite	50.	$\text{Zn}(\text{IO}_3)_2$

MN. ACID POLYATOMIC ANIONS - salts			
Write Formulas		Write Names	
1.	chromium(II) monohydrogen hypophosphite	26.	Ba(HCO ₃) ₂
2.	lithium dihydrogen phosphate	27.	Fe ₂ (HPO ₄) ₃
3.	manganese(II) monohydrogen phosphite	28.	FeHPO ₃
4.	tin(IV) hydrogen hypocarbonite	29.	Sn(HCO ₃) ₄
5.	magnesium hydrogen sulfate	30.	Si(HCO ₃) ₄
6.	copper(II) dihydrogen perphosphate	31.	Ca(HCO ₂) ₂
7.	aluminium hydrogen carbonite	32.	Fe(HSO ₄) ₃
8.	silver hydrogen percarbonate	33.	Fe ₂ (SO ₄) ₃
9.	zinc dihydrogen phosphite	34.	Sn(HPO ₂) ₂
10.	iron(III) hydrogen persulfate	35.	Be(HCO ₃) ₂
11.	barium monohydrogen perphosphate	36.	NaH ₂ PO ₅
12.	stannic hydrogen carbonate	37.	LiHSO ₂
13.	manganous hydrogen sulfite	38.	NH ₄ HSO ₃
14.	rubidium dihydrogen hypophosphite	39.	Pb(HCO) ₄
15.	arsenic(V) hydrogen hyposulfite	40.	KHSO ₅
16.	sodium bicarbonate	41.	NH ₄ H ₂ PO ₅
17.	calcium biphosphate	42.	(NH ₄) ₂ HPO ₂
18.	ferrous bisulfate	43.	P(HCO ₄) ₃
19.	lithium bisulfite	44.	Ni(HCO ₂) ₂
20.	nickel(II) biphosphite	45.	CaHPO ₄
21.	strontium bicarbonite	46.	Na ₂ HPO ₄
22.	beryllium biphosphite	47.	Sn(HSO ₂) ₂
23.	plumbous biphosphate	48.	Sb(HSO ₂) ₅
24.	mercury(I) bisulfite	49.	As(H ₂ PO ₃) ₃
25.	cesium bicarbonite	50.	NH ₄ HCO ₃

OO1. MISCELLANEOUS POLYATOMIC IONS			
Write Formulas		Write Names	
1.	sodium hydroxide	9.	Al(OH) ₃
2.	copper(II) hydroxide	10.	NH ₄ F
3.	ammonium sulfide	11.	(NH ₄) ₂ O
4.	ammonium phosphide	12.	NH ₄ OH
5.	stannous hydroxide	13.	Zn(OH) ₂
6.	ammonium hydroxide	14.	Pb(OH) ₄
7.	calcium hydroxide	15.	Fe(OH) ₂
8.	ammonium chloride	16.	LiOH

OO2. POLYATOMIC IONS - others examples			
Write Formulas		Write Names	
1.	aluminium thiocyanate	16.	Sn(MnO ₄) ₄
2.	potassium permanganate	17.	Mg(SCN) ₂
3.	ferric acetate	18.	BeCrO ₄
4.	calcium chromate	19.	Cu(CH ₃ CO ₂) ₂
5.	lead(IV) dichromate	20.	As(CN) ₃
6.	sodium cyanide	21.	Sr(OCN) ₂
7.	mercuric cyanate	22.	LiMnO ₄
8.	ammonium thiocyanate	23.	Ni(CH ₃ CO ₂) ₃
9.	zinc permanganate	24.	AgSCN
10.	nickelic acetate	25.	FeCrO ₄
11.	arsenous chromate	26.	Si(Cr ₂ O ₇) ₂
12.	manganese(II) dichromate	27.	NH ₄ OCN
13.	potassium cyanide	28.	B(SCN) ₃
14.	phosphorus(V) cyanate	29.	Ca(MnO ₄) ₂
15.	ammonium permanganate	30.	NH ₄ SCN

ZZ. FINAL REVIEW			
Write Formulas		Write Names	
1.	sodium fluoride	26.	CaO
2.	phosphorous oxide	27.	HgBr
3.	barium bicarbonate	28.	SnBr ₂
4.	nitrogen dioxide	29.	SO ₂
5.	hydrochloric acid	30.	H ₂ S (l)
6.	chlorine gas	31.	Xe
7.	zinc biphosphite	32.	NH ₄ F
8.	ferrous hydroxide	33.	Sn(OH) ₂
9.	chloric acid	34.	CuH ₂ PO ₂
10.	silver nitrate	35.	Sn(CO ₃) ₂
11.	persulfuric acid	36.	H ₃ PO ₂
12.	nickel(III) carbonite	37.	Pb(HSO ₄) ₂
13.	zinc dihydrogen hypophosphite	38.	Sn(HCO ₄) ₂
14.	sodium chromate	39.	Zn(MnO ₄) ₂
15.	ammonium sulfide	40.	HNO ₃
16.	mercury(II) sulfide	41.	LiNO ₂
17.	cobaltous bisulfite	42.	CO ₂
18.	ammonium hydroxide	43.	Na ₂ CO ₂
19.	arsenic(III) hydrogen persulfate	44.	NH ₄ NO ₃
20.	phosphorous periodate	45.	ZnCO ₄
21.	nitrous acid	46.	HFO
22.	beryllium bromate	47.	Pb(IO ₃) ₂
23.	iodic acid	48.	H ₂ CO
24.	sulfur dioxide	49.	H ₂ O
25.	iodine vapour	50.	H ₂ S (g)