
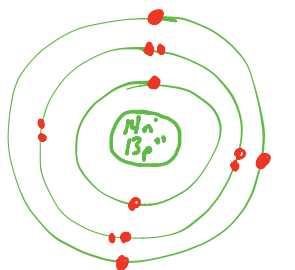



Modelling the atomic structure:

	Bohr	Bohr-Rutherford	Lewis
Protons	13	13	13
Neutrons	13	14	13
Electrons	13	13	13 → 3
Example			

Section 4.2

1. Why do elements react?
2. What does it mean to become stable?
3. What are three ways that elements can become "stable"?

Ionic Bonding

4. What is an ion?
5. Cation?
6. Anion?
7. What types of elements form ionic bonds?
8. Describe how an ionic compound is formed. (Example: NaCl)
9. What are the properties of an ionic compound?

Covalent Bonding

10. What is a covalent bond?
11. What types of elements form covalent bonds?
12. What is a diatomic molecule?
13. Describe how a covalent bond is formed. (Example: H₂O)
14. What are the properties of molecular compounds?