1. Define organic chemistry - page 6

Misconception- Supermarket foods

Past - Compounds that can only be synthesized by living organisms

Then...

- Urea, an organic compound, was synthesized by mixing two inorganic compounds

So now...

Organic chemistry is the study of **carbon** based compounds

Exceptions: carbon dioxide, carbonates

1. Properties of Carbon - page 6

- Four valence electrons
  - Four single covalent bonds
  - Double bonds
  - Triple bonds
  - Bond to other carbon atoms
    - Form chains, branches, rings
- Low electronegativity
- Stable - unreactive
- Most compounds are nonpolar
- Functional groups are added to increase the polarity of the molecule

Functional groups section 1.1 page 8

Double or triple bonds

- create sites of reactivity in the molecule
- changes the shapes of the molecule

Adding atoms of high electronegativity

- Example - Oxygen, Nitrogen
- Changes the solubility of the molecule

Naming package - begin hydrocarbon section

The carbon number chant

Alkanes

Branches

Numbering

Read page 1-6 (stop at alkenes)
Complete #5 and #6 page 8