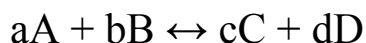


7.2 Equilibrium Law in Chemical Reactions

- Equilibrium describes the relationship between reactants and products
- The equilibrium value does not have a unit
- We compare the how much “product” is formed (right hand side values)
- The coefficients of a balanced equation are used as exponents of the concentration values



$$K_e = \frac{[C]^c [D]^d}{[A]^a [B]^b}$$

(The textbook uses K , however, we will use K_e (equilibrium constant) to avoid confusion with the K (rate constant) value associated with rate.)

- The above equation is used with the concentrations of products and reactants when they reach equilibrium.
Empirical evidence: page 439 table 1 and page 440 table 2
- We can use the same formula to determine if the reaction has reached equilibrium. Provided values can be placed into the expression and then its K value can be compared to the K_e
- The equilibrium law expression and its associated constant, K_e , describes the behaviour of almost all gaseous and aqueous chemical equilibria.
Solids and liquids have values that remain constant.

The Equilibrium Constant and Reaction Kinetics

- Read this section to see how the equilibrium law equation is derived.

Calculating K_e

- Always work from a balanced equation, determine the equilibrium law equation, substitute, solve.
- Make sure values are in mol/L
- See examples. Page 443-445

Heterogeneous Equilibria

- Homogeneous equilibria: all entities are in the same phase.
- Heterogeneous equilibria: reactants and products are in more than one phase. From before we only include (aq) and (g)

E.g. $\text{NaCl}_{(s)} \leftrightarrow \text{Na}^+_{(aq)} + \text{Cl}^-_{(aq)}$

$$K_e = [\text{Na}^+_{(aq)}] [\text{Cl}^-_{(aq)}]$$

- We ignore the material in the solid or liquid state since their concentrations do not change. The term mol/L in the case of solids and liquids is equivalent to their density, which will not change during the reaction.
- Reactions involving ionic compounds – single or double displacement
 - Write equilibrium expressions only for net ionic equations

Limitations of Equilibrium Constants and Percent Reaction Values

- We previously learned that **rate of reaction** constant (K) was affected by many factors such as temperature, pressure, concentration, catalysts, etc.
- The **equilibrium constant** (K_e) is also affected by other factors. The equilibrium is temperature and pressure dependent.
- Also, the K_e value will only tell you the ratio of the amounts of products and reactants in a closed system. It does not tell you how fast the reaction takes place.

Magnitude of K

- $K \gg 1$ products are heavily favoured and reaction nears completion.
- $K \cong 1$ concentrations of products and reactants are approximately equal at equilibrium.
- $K \ll 1$ reactants are heavily favoured and reactants do not tend to react.

Homework

- Practice 1,2,3,4,5,6,7
- Questions 1,2,3,4,5,6