

Applications of Acids and Bases

While reading over chapter 5.1+5.2 - Determine which acid or base is being described:

Clue	Acid or Base	Other Name
carbon dioxide gas dissolves in the water in your blood		
responsible for passing on inherited characteristics		
toxic to humans, so your body converts it to urea ((NH ₂) ₂ CO)		
helps your stomach to break down your food		
Counteracts stomach acid to protect the rest of the digestive system		
used in producing soft and liquid soaps		
vitamin C		
commonly called baking soda		
used in automobile batteries		
vinegar		
found in drain cleaners and oven cleaners		

ON A SEPARATE PIECE OF PAPER OR CHROMEBOOK Answer the following questions

Read 5.2 Environmental Roles of Acids and Bases

1. Describe several ways that modifying the soil pH can affect plants.
2. How do neutralization reactions aid in the following situations:

Agriculture

Acid Reflux

Bee stings

Processed foods

Eating fish

3. What is Acid Precipitation?
4. How does acid precipitation occur (cause, chemical reactions, sources)?
5. What are three problems that are caused by acid precipitation?
6. Why are some lakes more sensitive to acid precipitation than others?
7. How can we solve the environmental problem of acidified lakes? (Include the chemical reaction)
8. What are two ways technology can reduce the chances of creating acid precipitation?
9. What are heavy metals and give some examples along with their sources?
10. What are some effects of heavy metal accumulation?
11. Explain the problem and solution of acid leaching.