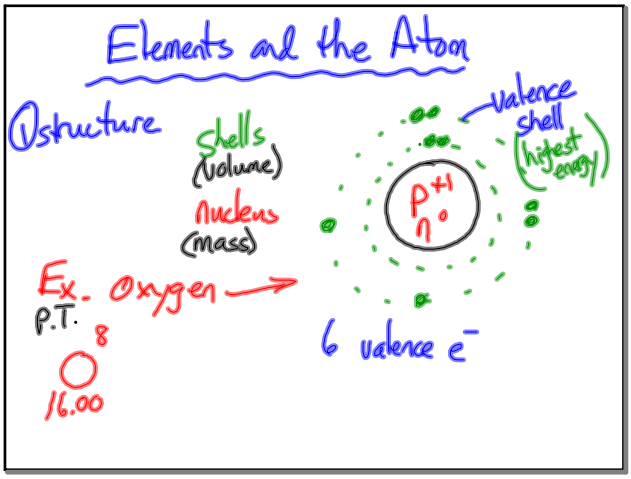
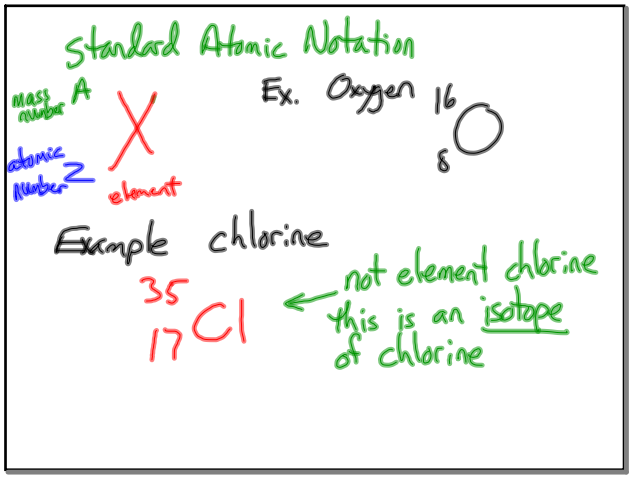
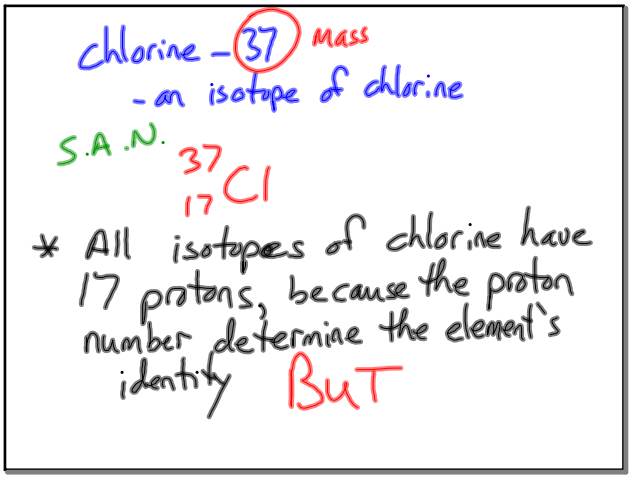


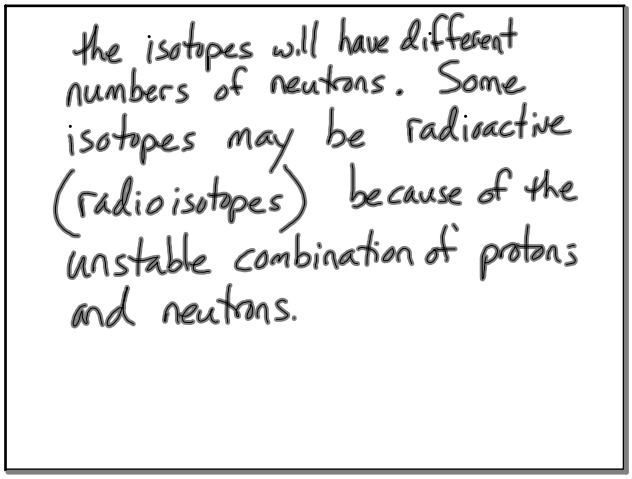
Feb 9-12:19 PM



Feb 9-12:38 PM



Feb 9-12:42 PM



Feb 9-12:47 PM

So the atomic mass is not the mass of an atom, it is the weighted average mass of all the isotopes of that element.

→ The abundance of each isotope is included in the average.

Feb 9-12:56 PM

② Calculating Relative Atomic Mass

Given – isotope and % abundance

Ex. Calculate relative atomic mass of a sample of Beryllium.

Be - 7	15 %
Be - 9	65 %
Be - 11	20 %

Feb 9-1:05 PM

Be - 7	15 %
Be - 9	65 %
Be - 11	20 %

$$A_r = \left(\text{mass isotope A} \times \text{percent decimal}\right) + \left(\ldots\ldots\right.$$

$$A_r = (7 \times 0.15) + (9 \times 0.65) + (11 \times 0.2)$$

$$= 9.1\ u \text{ (atomic mass unit)}$$

Feb 9-1:11 PM